

(FILE 'HOME' ENTERED AT 14:57:29 ON 14 JUN 2003)

FILE 'CAPLUS, WPIDS, JAPIO' ENTERED AT 14:58:13 ON 14 JUN 2003

L1 2187 S (HYDROGEN PEROXIDE OR H2O2 OR HYDROGEN DIOXIDE) (50A) (((PHOS
L2 163 S L1 AND ((HYDROGEN PEROXIDE OR H2O2 OR HYDROGEN DIOXIDE) (25A)
L3 158 DUP REM L2 (5 DUPLICATES REMOVED)

=> d que

L1 2187 SEA (HYDROGEN PEROXIDE OR H2O2 OR HYDROGEN DIOXIDE) (50A)
(((PHOSPHORUS OR PHOSPHOROUS) (3A) ACID#) OR PHOSPHORIC ACID
OR PHOSPHONIC ACID# OR PHOSPHONATE# OR DIPHOSPHONIC OR
DIPHOSPHONATE# OR DEQUEST OR BRIQUEST)
L2 163 SEA L1 AND ((HYDROGEN PEROXIDE OR H2O2 OR HYDROGEN DIOXIDE)
(25A) (SULFONIC OR SULFONATE# OR SULFONATED OR SULFATE#))
L3 158 DUP REM L2 (5 DUPLICATES REMOVED)

=>

10/067, 809

All Reviewed
6/14/03

=> d 1-158 bib ab kwic

L3 ANSWER 1 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2003:173391 CAPLUS
DN 138:209911
TI Hair care compositions containing hydrogen peroxide
IN Murakami, Atsushi; Matsumoto, Heiichiro; Ishimura, Tomoyuki; Tajiri, Miki;
Sano, Mitsuo; Tsujino, Yoshio
PA Mandom Corporation, Japan
SO PCT Int. Appl., 17 pp.
CODEN: PIXXD2
DT Patent
LA Japanese
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2003017954	A1	20030306	WO 2002-JP8608	20020827
	W:	AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, OM, PH, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, YU, ZA, ZM, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM			
	RW:	GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW, AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, SK, TR, BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG			

PRAI JP 2001-256052 A 20010827

AB Disclosed are hair care compns. contg. hydrogen peroxide which are to be used as hair bleachings, oxidative hair dyes, perming agents and so on. These compns. are characterized by contg. hydrogen peroxide together with polyoxyethylene lanolin, an oily component, a sequestrant and water. If desired, these compns. may further contain an anionic surfactant. These compns. exhibit excellent emulsion stability without showing any sepn. in case of processed into emulsified compns. A hair bleach compn. contg. **hydrogen peroxide** 10, polyoxyethylene lanolin 1.5, polyoxyethylene lauryl ether **sulfate** sodium salt 2, cetyl alc. 2, **phosphonic acid** 0.05, phenacetin 0.2, and water balance to 100 % was prepd.

RE.CNT 4 THERE ARE 4 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB Disclosed are hair care compns. contg. hydrogen peroxide which are to be used as hair bleachings, oxidative hair dyes, perming agents and so on. These compns. are characterized by contg. hydrogen peroxide together with polyoxyethylene lanolin, an oily component, a sequestrant and water. If desired, these compns. may further contain an anionic surfactant. These compns. exhibit excellent emulsion stability without showing any sepn. in case of processed into emulsified compns. A hair bleach compn. contg. **hydrogen peroxide** 10, polyoxyethylene lanolin 1.5, polyoxyethylene lauryl ether **sulfate** sodium salt 2, cetyl alc. 2, **phosphonic acid** 0.05, phenacetin 0.2, and water balance to 100 % was prepd.

IT 9004-82-4, Polyoxyethylene lauryl ether **sulfate** sodium salt
RL: COS (Cosmetic use); BIOL (Biological study); USES (Uses)
(hair care compns. contg. **hydrogen peroxide**, polyoxyethylene lanolin oily component, sequestering agents, and anionic surfactants)

L3 ANSWER 2 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2003:133188 CAPLUS
DN 138:172784
TI Synthesis of hydrogen peroxide from oxygen and hydrogen using a catalyst

IN Papparatto, Giuseppe; De Alberti, Giordano; D'Aloisio, Rino; Buzzoni, Roberto
 PA Eni S.p.A., Italy; Polimeri Europa S.p.A.
 SO PCT Int. Appl., 31 pp.
 CODEN: PIXXD2
 DT Patent
 LA English
 FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2003014014	A2	20030220	WO 2002-EP8546	20020730
	W: AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, OM, PH, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VN, YU, ZA, ZM, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM RW: GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW, AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, SK, TR, BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG				

PRAI IT 2001-MI1688 A 20010802

AB A catalyst useful for the synthesis of hydrogen peroxide starting from hydrogen and oxygen consists of at least one metal of the platinum group as active component, a polyolefin, and a carrier. The catalyst contains 0.05-2 wt.% of Pd, 0.005-0.5 wt.% of Pt with an at. ratio of Pt/Pd of (1-30)/(70-99), and optionally Ru, Rh, Ir, and Au. The polyolefins can be rubbers and copolymers of butadiene-styrene (synthetic rubber; GRS, SBR); ethylenepropylene (EPM, EPR) copolymers, ethylenepropylenediene copolymers (EPDM rubbers), styrene-butadiene-styrene (SBR thermoplastic rubbers); isobutylene isoprene rubber (butylrubbers). The carrier can be silica, alumina, silica-alumina, zeolites, and preferably activated carbon or activated carbon functionalized with sulfonic groups with a surface area of > 600 m²/g. The catalyst is prepd. by dispersing the precursors of the single metal components on an inert carrier which can be pretreated with a polyolefin by pptn. or impregnation. The reaction solvent contains a halogenated promoter, such as HBr, NaBr, KBr, or NH₄Br, and/or an acid promoter, such as sulfuric, phosphoric, nitric acid or sulfonic acids. The solvent consists of at least one alc. or a mixt. of alc.-water optionally contg. an aliph. ether and/or one or more C₅-32 hydrocarbons. The alc. can be ethanol, tert. butanol, or preferably methanol. The ether is Me tert-Bu ether. The hydrocarbon is selected from paraffins, such as n-hexane, n-heptane, n-octane, and n-decane, cyclo-paraffinic hydrocarbons, such as cyclohexane, decalin, methylcyclohexane, ethylcyclohexane and dimethylcyclohexane, arom. hydrocarbons, such as benzene, naphthalene, toluene, xylenes, ethylbenzene, cumene, and alkyl-naphthalenes. The reaction is carried out at 20-40.degree.C, 30-100 bars, and in the presence of an inert gas, such as N₂, He, or Ar. The produced hydrogen peroxide soln. can be directly used in an oxidn. process of a substrate, e.g. olefins, arom. hydrocarbons, ammonia, and carbonyl compds., catalyzed by titanium silicalite.

IT Sulfonic acids, uses

RL: NUU (Other use, unclassified); USES (Uses)
 (acid promoter; synthesis of **hydrogen peroxide** from oxygen and hydrogen using catalyst)

IT 7664-38-2, **Phosphoric acid**, uses 7664-93-9, Sulfuric acid, uses 7697-37-2, Nitric acid, uses

RL: NUU (Other use, unclassified); USES (Uses)
 (acid promoter; synthesis of **hydrogen peroxide** from oxygen and hydrogen using catalyst)

IT 1344-28-1, Alumina, uses 7440-44-0, Carbon, uses 7440-44-0D, Carbon, functionalized with **sulfonic** groups 7631-86-9, Silica, uses 159995-97-8, Aluminum silicon oxide

RL: CAT (Catalyst use); USES (Uses)
(catalyst support; synthesis of **hydrogen peroxide**
from oxygen and hydrogen using catalyst)

L3 ANSWER 3 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2003:334694 CAPLUS

DN 138:340423

TI Direct synthesis of hydrogen peroxide and integration of the process into
oxidation processes

IN Haas, Thomas; Stochniol, Guido; Rollmann, Jurgen

PA Germany

SO U.S. Pat. Appl. Publ., 6 pp.

CODEN: USXXCO

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 2003083510	A1	20030501	US 2002-254746	20020925
	DE 10153546	A1	20030522	DE 2001-10153546	20011030
	EP 1308416	A1	20030507	EP 2002-19947	20020905
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO, MK, CY, AL, TR, BG, CZ, EE, SK				

PRAI DE 2001-10153546 A 20011030

AB Aq.-org. or org. hydrogen peroxide solns. can be produced by direct
synthesis in the presence of a catalyst fixed bed and an org. solvent.
According to the invention, a non-explosive gas mixt. comprising H2 and O2
is employed and a liq. reaction medium comprising org. solvent and bromide
and/or iodide is passed over the fixed bed with a cross-section loading of
at least 0.3 m/h. In the case of a trickle bed procedure with a
cross-section loading of 0.3 m/h to 2 m/h, 4 to 10 wt.% methanolic H2O2
solns. can be prepd. with a high productivity.

IT Alkenes, processes

Aromatic hydrocarbons, processes

Sulfonic acids, processes

RL: CPS (Chemical process); PEP (Physical, engineering or chemical
process); PROC (Process)

(direct synthesis of **hydrogen peroxide** and
integration of process into oxidn. processes)

IT 1333-74-0, Hydrogen, processes 7647-01-0, Hydrochloric acid, processes
7782-44-7, Oxygen, processes 13598-36-2, **Phosphonic**
acid 24959-67-9, Bromide, processes

RL: CPS (Chemical process); PEP (Physical, engineering or chemical
process); PROC (Process)

(direct synthesis of **hydrogen peroxide** and
integration of process into oxidn. processes)

L3 ANSWER 4 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2003:172898 CAPLUS

DN 138:209910

TI Method for simultaneous dyeing and permanent waving of hair

IN Schonert, Dieter; Schmidt-Hoerr, Anette; Lenz, Uwe

PA Wella Aktiengesellschaft, Germany

SO Eur. Pat. Appl., 12 pp.

CODEN: EPXXDW

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 1287812	A2	20030305	EP 2002-13122	20020614
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO, MK, CY, AL, TR				
	DE 10143293	A1	20030320	DE 2001-10143293	20010904

US 2003084518 A1 20030508 US 2002-198548 20020718
PRAI DE 2001-10143293 A 20010904
OS MARPAT 138:209910

AB The invention concerns the simultaneous dyeing and permanent waving of hair by using two components that are mixed upon use; component (A) includes a keratin-reducing agent, component (B) contains oxidative precursor and coupling dyes; after the treatment with the mixt. an oxidative fixation agent is applied. The oxidative dyes are selected from the group of pyrazole derivs. and p-aminophenol derivs. Thus components included (g): (A) ammoniumthioglycolate (70% aq. soln.) 16.0; ammonia (25% aq. soln.) 1.0; ammonium hydrogen carbonate 4.0; oxoethylated castor oil (35 mol EO) 1.0; perfume oil 0.5; water to 100; (B) ammonium thioglycolate (70% aq. soln.) 10.00; oxoethylated castor oil (35 mol EO) 1.0;; sodium sulfite 0.20; perfume oil 0.1; 3-aminophenol 0.10; amino-4-[(2-hydroxyethyl)amino]-anisole 0.04; resorcin 0.60; p-toluylenediamino **sulfate** 0.40; 2-amino-5-methyl-phenol 0.03; 4-amino-3-methylphenol 0.50; ethanol 2.00; water to 100; oxidative fixation agent: **hydrogen peroxide** (50%) 4.00; salicylic acid 0.10; disodium hydrogen phosphate 0.20; o-**phosphoric acid** 0.15; oxoethylated castor oil (35 mol EO) 1.0; vinylpyrrolidone-styrene copolymer 0.10; perfume oil 0.10; water to 100.

AB The invention concerns the simultaneous dyeing and permanent waving of hair by using two components that are mixed upon use; component (A) includes a keratin-reducing agent, component (B) contains oxidative precursor and coupling dyes; after the treatment with the mixt. an oxidative fixation agent is applied. The oxidative dyes are selected from the group of pyrazole derivs. and p-aminophenol derivs. Thus components included (g): (A) ammoniumthioglycolate (70% aq. soln.) 16.0; ammonia (25% aq. soln.) 1.0; ammonium hydrogen carbonate 4.0; oxoethylated castor oil (35 mol EO) 1.0; perfume oil 0.5; water to 100; (B) ammonium thioglycolate (70% aq. soln.) 10.00; oxoethylated castor oil (35 mol EO) 1.0;; sodium sulfite 0.20; perfume oil 0.1; 3-aminophenol 0.10; amino-4-[(2-hydroxyethyl)amino]-anisole 0.04; resorcin 0.60; p-toluylenediamino **sulfate** 0.40; 2-amino-5-methyl-phenol 0.03; 4-amino-3-methylphenol 0.50; ethanol 2.00; water to 100; oxidative fixation agent: **hydrogen peroxide** (50%) 4.00; salicylic acid 0.10; disodium hydrogen phosphate 0.20; o-**phosphoric acid** 0.15; oxoethylated castor oil (35 mol EO) 1.0; vinylpyrrolidone-styrene copolymer 0.10; perfume oil 0.10; water to 100.

L3 ANSWER 5 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2003:40194 CAPLUS
DN 138:91852
TI Two-agent type liquid bleaching compositions
IN Ozaki, Kazuyoshi; Maki, Masataka; Ogura, Nobuyuki; Muneo, Aoyagi
PA Kao Corporation, Japan
SO Eur. Pat. Appl., 29 pp.
CODEN: EPXXDW

DT Patent
LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 1275708	A1	20030115	EP 2002-14962	20020709
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO, MK, CY, AL, TR, BG, CZ, EE, SK				
	JP 2003020498	A2	20030124	JP 2001-209555	20010710
	JP 2003041295	A2	20030213	JP 2001-231687	20010731
	JP 2003041296	A2	20030213	JP 2001-231688	20010731
	CN 1396252	A	20030212	CN 2002-140923	20020710
PRAI	JP 2001-209555	A	20010710		
	JP 2001-231687	A	20010731		
	JP 2001-231688	A	20010731		
OS	MARPAT 138:91852				

AB To provide 2-agent type liq. bleaching compns. having excellent bleaching effect even if a mixing ratio of the 2 agents varies, great usability, and no problem in storage stability, 2-agent type liq. bleaching compns. contain an agent A and an agent B filled and held in sep. chambers of a container and the agent A is made of 0.1-10% H2O2 and H2O provided with certain buffering capacity, and the agent B is made of an alkali agent and H2O provided with certain buffering capacity. Agent A and agent B meet the following conditions, resp. : (I) pH of agent A 1-6.5 at 20.degree. and a vol. of aq. 0.1N NaOH soln. required to adjust 1000 mL agent A to pH 7 at 20.degree. is 50-1000 mL and (II) pH of agent B 9-12 at 20.degree. and a vol. of aq. 1N H2SO4 soln. required to adjust 1000 mL agent B to pH 7 at 20.degree. is 450-2000 mL. Thus, an alk. (pH 10.5) bleaching detergent contained 2/1 ratio A/B of **hydrogen peroxide** 5, citric acid 1.5, polyoxyethylene lauryl ether 2, ethylene oxide-propylene oxide copolymer monolauryl ether 30, LAS 1, sodium salt of polyoxyethylene lauryl ether **sulfate** 2, alkyl(C12-15)benzenesulfonic acid sodium salt, 0.5, N-tetradecyl-N,N,N-trimethylammonium chloride 1, p-methoxyphenol 0.3, N-lauryl-N,N-dimethyl-N-(2-hydroxy-1-sulfopropyl)ammonium sulfobetaine 1, lauroxyloxybenzenesulfonic acid sodium salt 1 parts, and the balance H2O, in combination with KCO3 6, NaHCO3 0.3, above sulfobetaine 4, LAS 3 parts, and the balance H2O, showing 86% bleaching efficiency (reflectance).

RE.CNT 9 THERE ARE 9 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB To provide 2-agent type liq. bleaching compns. having excellent bleaching effect even if a mixing ratio of the 2 agents varies, great usability, and no problem in storage stability, 2-agent type liq. bleaching compns. contain an agent A and an agent B filled and held in sep. chambers of a container and the agent A is made of 0.1-10% H2O2 and H2O provided with certain buffering capacity, and the agent B is made of an alkali agent and H2O provided with certain buffering capacity. Agent A and agent B meet the following conditions, resp. : (I) pH of agent A 1-6.5 at 20.degree. and a vol. of aq. 0.1N NaOH soln. required to adjust 1000 mL agent A to pH 7 at 20.degree. is 50-1000 mL and (II) pH of agent B 9-12 at 20.degree. and a vol. of aq. 1N H2SO4 soln. required to adjust 1000 mL agent B to pH 7 at 20.degree. is 450-2000 mL. Thus, an alk. (pH 10.5) bleaching detergent contained 2/1 ratio A/B of **hydrogen peroxide** 5, citric acid 1.5, polyoxyethylene lauryl ether 2, ethylene oxide-propylene oxide copolymer monolauryl ether 30, LAS 1, sodium salt of polyoxyethylene lauryl ether **sulfate** 2, alkyl(C12-15)benzenesulfonic acid sodium salt, 0.5, N-tetradecyl-N,N,N-trimethylammonium chloride 1, p-methoxyphenol 0.3, N-lauryl-N,N-dimethyl-N-(2-hydroxy-1-sulfopropyl)ammonium sulfobetaine 1, lauroxyloxybenzenesulfonic acid sodium salt 1 parts, and the balance H2O, in combination with KCO3 6, NaHCO3 0.3, above sulfobetaine 4, LAS 3 parts, and the balance H2O, showing 86% bleaching efficiency (reflectance).

IT 60-00-4, Ethylenediaminetetraacetic acid, uses 77-92-9, Citric acid, uses 102-71-6, Triethanolamine, uses 111-42-2, Diethanolamine, uses 141-43-5, Monoethanolamine, uses 144-55-8, Sodium hydrogen carbonate, uses 497-19-8, Sodium carbonate, uses 515-42-4D, Benzenesulfonic acid sodium salt, C12-15 alkyl derivs. 584-08-7, Potassium carbonate 657-84-1, p-Toluenesulfonic acid sodium salt 1310-58-3, Potassium hydroxide, uses 1310-73-2, Sodium hydroxide, uses 1330-43-4, Sodium tetraborate 1643-20-5, Lauryldimethylamine oxide 2809-21-4, 1-Hydroxyethylidene-1,1-diphosphonic acid 4574-04-3 4615-13-8 7558-79-4, Disodium phosphate 7601-54-9, Trisodium phosphate 7664-38-2, Orthophosphoric acid, uses 7722-84-1, Hydrogen peroxide, uses 9002-92-0, Polyoxyethylene lauryl ether 9003-04-7, Polyacrylic acid sodium salt 9004-82-4, Sodium polyoxyethylene lauryl ether **sulfate** 13197-76-7 37311-00-5, Ethylene oxide-propylene oxide copolymer monolauryl ether 88380-00-1, Lauroxyloxybenzenesulfonic acid sodium salt

RL: TEM (Technical or engineered material use); USES (Uses)

(two-agent-type liq. bleaching compns. contg. acidic **hydrogen**

peroxide soln. combination with alkali soln. for laundering of fabrics)

L3 ANSWER 6 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2003:4757 CAPLUS

DN 138:44447

TI Hair bleaching agents containing ubiquinones

IN Noecker, Bernd; Wilz, Ruediger; Ghiasi, Fariba; Garbe, Barbara

PA Goldwell G.m.b.H., Germany

SO Eur. Pat. Appl., 8 pp.

CODEN: EPXXDW

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 1269976	A2	20030102	EP 2002-14196	20020626
	EP 1269976	A3	20030423		
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO, MK, CY, AL, TR				
	DE 10131540	A1	20030116	DE 2001-10131540	20010629
PRAI	DE 2001-10131540	A	20010629		

OS MARPAT 138:44447

AB The invention concerns hair bleaching agents that contain ubiquinones and bleaching agents; bleaching agents are peroxides and ammonium salts. Thus a two component bleaching agent was composed of a bleaching powder and hydrogen peroxide lotion. The bleaching powder included (wt./wt.%): silica 16.27; ammonium chloride 14.70; sodium carbonate 10.00; sodium silicate 2.30; phthalimidoperoxyhexanoic acid 41.50; sodium persulfate 15.20; Coenzyme Q6 0.03. The **hydrogen peroxide** lotion contained (wt./wt.%): **hydrogen peroxide** 6.00; cetylstearylalc. 1.70; **phosphoric acid** 0.50; sodium lauryl **sulfate** 0.20; Coenzyme Q6 0.05; disodium hydrogen phosphate 0.10; salicylic acid 0.10; water to 100.00.

AB The invention concerns hair bleaching agents that contain ubiquinones and bleaching agents; bleaching agents are peroxides and ammonium salts. Thus a two component bleaching agent was composed of a bleaching powder and hydrogen peroxide lotion. The bleaching powder included (wt./wt.%): silica 16.27; ammonium chloride 14.70; sodium carbonate 10.00; sodium silicate 2.30; phthalimidoperoxyhexanoic acid 41.50; sodium persulfate 15.20; Coenzyme Q6 0.03. The **hydrogen peroxide** lotion contained (wt./wt.%): **hydrogen peroxide** 6.00; cetylstearylalc. 1.70; **phosphoric acid** 0.50; sodium lauryl **sulfate** 0.20; Coenzyme Q6 0.05; disodium hydrogen phosphate 0.10; salicylic acid 0.10; water to 100.00.

L3 ANSWER 7 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2003-248242 [24] WPIDS

DNN N2003-197191 DNC C2003-064090

TI Disinfecting and cleaning system for medical device, e.g. contact lens, comprises first composition containing peroxidase enzyme and iodide salt, and second composition containing hydrogen peroxide source in aqueous solution.

DC D22 E19 P34

IN MOWREY-MCKEE, M F; SILLS, M A

PA (NOVS) NOVARTIS AG; (NOVS) NOVARTIS PHARMA GMBH

CYC 88

PI WO 2003013621 A1 20030220 (200324)* EN 28p

RW: AT BE BG CH CY CZ DE DK EA EE ES FI FR GB GR IE IT LU MC NL PT SE SK TR

W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CO CR CU CZ DE DK DM DZ EC EE ES FI GB GD GE GH HR HU ID IL IN IS JP KE KG KP KR KZ LC LK LT LU LV MA MD MK MN MX NO NZ OM PH PL PT RO RU SE SG SI SK TJ TM TN TR TT UA US UZ VC VN YU ZA ZW

ADT WO 2003013621 A1 WO 2002-EP8839 20020807
PRAI US 2001-310893P 20010808
AB WO2003013621 A UPAB: 20030410

NOVELTY - A disinfecting and cleaning system consists of a first composition containing peroxidase enzyme and iodide salt, and a second composition containing hydrogen peroxide source in an aqueous solution.

DETAILED DESCRIPTION - An INDEPENDENT CLAIM is included for a method of disinfecting a medical device by combining first and second compositions to provide an activated disinfecting solution, and contacting the medical device with the disinfecting solution for a predetermined time.

USE - For disinfecting and cleaning medical device, e.g. contact lens (claimed).

ADVANTAGE - The inventive disinfecting and cleaning system is compatible with most types of contact lenses while maintaining both high level of antibacterial activity and low order of toxicity to eye tissue. It is effective against wide spectrum of microorganisms including Staphylococcus aureus, Pseudomonas aeruginosa, Serratia marcescens, Candida albicans and Fusarium solani. Contact lenses treated with the inventive system were non-irritating, even when they were not rinsed after disinfection and prior to insertion into the eye.

Dwg.0/0

TECH.

amino acid, sodium sulfite, sodium hydrogen sulfite, sodium thiosulfate, or imidazole. The buffering agent is acetate, phosphate, borate, citrate, nitrate, sulfate, tartrate, lactate, carbonate, bicarbonate, TRIS, BIS-TRIS propane, and/or Goode buffer.

The **hydrogen peroxide** stabilizer is a compound of formula $(H_2PO_3-X)_2N-(X-N(XPO_3H_2))_z-(XPO_3H_2)$ (I) or its water-soluble salt or a compound of formula (II) or its water-soluble salt:

X = 1-4C alkylene;

z = 0-3;

n, m, p, q = 0-4.

The **hydrogen peroxide** stabilizer is diethylene triamine penta(methylene-phosphonic acid) or its water-soluble salt.

TECHNOLOGY FOCUS - INORGANIC CHEMISTRY - Preferred Component: The hydrogen peroxide source is hydrogen peroxide, sodium.

L3 ANSWER 8 OF 158 CAPLUS COPYRIGHT 2003 ACS DUPLICATE 1

AN 2002:31372 CAPLUS

DN 136:73513

TI Magnesia foam panels containing naphthalene sulfonate and **hydrogen peroxide** for boat compartment and cabin walls

IN Vaghini, Alberto; Di Maggio, Rosa

PA Italy

SO PCT Int. Appl., 21 pp.

CODEN: PIXXD2

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2002002477	A1	20020110	WO 2001-IB1591	20010702
	W:	AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, TZ, UA, UG, US, UZ, VN, YU, ZA, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM			
	RW:	GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, TR, BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG			

	AU 2001082396	A5	20020114	AU 2001-82396	20010702
PRAI	IT 2000-PC25	A	20000703		
	WO 2001-IB1591	W	20010702		

AB Magnesia foam panels for use in making cabins and compartments for boats and the like comprise at least one layer of magnesia foam. The panels consist of one layer of material contg. 100 parts MgO, 90 parts compd. B, 0.1-3 % naphthalene **sulfonate** and 0-20 parts H2O2 with compd. B comprising 50 parts H2O, 50 parts MgCl2 and H3PO4 0.5%.

RE.CNT 11 THERE ARE 11 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

TI Magnesia foam panels containing naphthalene **sulfonate** and **hydrogen peroxide** for boat compartment and cabin walls

AB Magnesia foam panels for use in making cabins and compartments for boats and the like comprise at least one layer of magnesia foam. The panels consist of one layer of material contg. 100 parts MgO, 90 parts compd. B, 0.1-3 % naphthalene **sulfonate** and 0-20 parts H2O2 with compd. B comprising 50 parts H2O, 50 parts MgCl2 and H3PO4 0.5%.

ST magnesia naphthalene **sulfonate hydrogen peroxide** foaming agent panel; strength elasticity magnesia naphthalene **sulfonate hydrogen peroxide** foam panel; magnesium chloride **phosphoric acid** magnesia **hydrogen peroxide** foam panel

IT Ships
(boats; magnesia foam panels contg. naphthalene **sulfonate** and **hydrogen peroxide** for boat compartment and cabin walls)

IT Controlled atmospheres
(carbon dioxide steam-curing; magnesia foam panels contg. naphthalene **sulfonate** and **hydrogen peroxide** for boat compartment and cabin walls)

IT Sawdust
(filler; magnesia foam panels contg. naphthalene **sulfonate** and **hydrogen peroxide** for boat compartment and cabin walls)

IT Thermal insulators
(foamed magnesia panels; magnesia foam panels contg. naphthalene **sulfonate** and **hydrogen peroxide** for boat compartment and cabin walls)

IT Foaming agents
Strength
Young's modulus
(magnesia foam panels contg. naphthalene **sulfonate** and **hydrogen peroxide** for boat compartment and cabin walls)

IT Foams
(magnesia; magnesia foam panels contg. naphthalene **sulfonate** and **hydrogen peroxide** for boat compartment and cabin walls)

IT Balloons
Microspheres
(microballoons, silica; magnesia foam panels contg. naphthalene **sulfonate** and **hydrogen peroxide** for boat compartment and cabin walls)

IT Construction materials
(panels; magnesia foam panels contg. naphthalene **sulfonate** and **hydrogen peroxide** for boat compartment and cabin walls)

IT Hydration, chemical
(steam curing; magnesia foam panels contg. naphthalene **sulfonate** and **hydrogen peroxide** for boat compartment and cabin walls)

IT 9003-53-6, Polystyrene
RL: PEP (Physical, engineering or chemical process); PYP (Physical process); PROC (Process)

(foam, filler; magnesia foam panels contg. naphthalene
sulfonate and hydrogen peroxide for boat
 compartment and cabin walls)

IT 1309-48-4, Magnesium oxide (MgO), processes
 RL: PEP (Physical, engineering or chemical process); PYP (Physical
 process); TEM (Technical or engineered material use); PROC (Process); USES
 (Uses)
 (foamed panels; magnesia foam panels contg. naphthalene
sulfonate and hydrogen peroxide for boat
 compartment and cabin walls)

IT 7722-84-1, **Hydrogen peroxide (H2O2)**,
 processes
 RL: PEP (Physical, engineering or chemical process); PYP (Physical
 process); PROC (Process)
 (foaming agents; magnesia foam panels contg. naphthalene
sulfonate and hydrogen peroxide for boat
 compartment and cabin walls)

IT 1309-42-8, Magnesium hydroxide
 RL: FMU (Formation, unclassified); PEP (Physical, engineering or chemical
 process); PYP (Physical process); FORM (Formation, nonpreparative); PROC
 (Process)
 (in foamed panels; magnesia foam panels contg. naphthalene
sulfonate and hydrogen peroxide for boat
 compartment and cabin walls)

IT 7664-38-2, **Phosphoric acid**, processes 7786-30-3,
 Magnesium chloride (MgCl2), processes 50852-11-4, Naphthalene
sulfonate
 RL: PEP (Physical, engineering or chemical process); PYP (Physical
 process); PROC (Process)
 (in foamed panels; magnesia foam panels contg. naphthalene
sulfonate and hydrogen peroxide for boat
 compartment and cabin walls)

IT 373643-90-4D, Magnesium carbonate chloride hydroxide (Mg2(CO3)Cl(OH)),
 hydrate
 RL: FMU (Formation, unclassified); PEP (Physical, engineering or chemical
 process); PRP (Properties); PYP (Physical process); FORM (Formation,
 nonpreparative); PROC (Process)
 (strengthening agent; magnesia foam panels contg. naphthalene
sulfonate and hydrogen peroxide for boat
 compartment and cabin walls)

L3 ANSWER 9 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2002:868698 CAPLUS

DN 137:357893

TI Method and compositions for coloring hair using hydrogen peroxide-based
 developers

IN Duffer, Dalal Ibrahim Esber; Vena, Louann C.; Tian, Minmin; Singh,
 Shailendra K.

PA Revlon Consumer Products Corporation, USA

SO PCT Int. Appl., 33 pp.

CODEN: PIXXD2

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2002089748	A2	20021114	WO 2002-US14760	20020509
	WO 2002089748	A3	20030227		
	W:	AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN,			
		CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH,			
		GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR,			
		LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, OM, PH,			
		PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TN, TR, TT, TZ,			
		UA, UG, UZ, VN, YU, ZA, ZM, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM			

RW: GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW, AT, BE, CH,
CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, TR,
BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG

US 2003028979 A1 20030213 US 2001-852982 20010510

PRAI US 2001-852982 A 20010510

AB An aq. compn. for oxidatively coloring hair in 12 min or less is prepd. by combining a dye mixt. and a developer mixt. The compn. comprises (i) 0.01-2.0% (combined wt.) of primary intermediates and couplers, (ii) a free alky. concn. of 0.20-0.75 meq/g, and (iii) a hydrogen peroxide concn. of 4-6% by wt. A hair coloring kit contg. sep. containers of an aq. dye mixt. and an aq. developing mixt. is also described. For example, a hair dye mixt. in the level 5, medium brown, shade was prepd. contg. (by wt.) p-phenylenediamine 0.63% (primary intermediate), N,N-bis(2-hydroxyethyl)-p-phenylenediamine sulfate 0.11 (primary intermediate), resorcinol 0.50 (coupler), 1-naphthol 0.04 (coupler), m-aminophenol 0.07 (coupler), ethanolamine 3.00 (alkalizing agent), ammonium hydroxide (27.5%) 9.00 (alkalizing agent), oleic acid 12.50, erythorbic acid 0.20 (antioxidant), sodium sulfite 0.50 (antioxidant), ethoxydiglycol 5.00 (solvent), Hypnea musciformis ext., Gellidiala acerosa ext., Sargassum filipendula ext., and sorbitol 0.80 (conditioner), Oleth-20 1.00, Steareth-21 0.70, emulsifying wax 2.00, meadowfoam seed oil 0.75 (conditioner), Polyquaternium-10 0.20 (conditioner), Polyquaternium-28 0.50 (conditioner), hydrolyzed wheat protein 0.50 (conditioner), oleyl alc. 0.40 (conditioner, thickener), cetearyl alc. 4.00 (thickener), tetrasodium EDTA (38%) 0.80 (chelating agent), ammonium lauryl sulfate (28%) 2.00, mica, titanium dioxide (67:33) 0.30, sodium benzotriazolyl butylphenol sulfonate buteth-3, tri-Bu citrate 0.50 (UV absorber), fragrance 1.25%, and water 52.75%. A developer compn. contained (by wt.) water 69.72%, methylparaben 0.05%, EDTA 0.02%, mineral oil 0.60%, cetearyl alc./ceteareth-20 (80:20) 3.75%, cetearyl alc. 0.80%, ceteareth-20 0.40%, cyclomethicone/trimethylsiloxysilicate (50:50) 0.01%, trimethylsilylamodimethicone, C11-15 pareth-7 C12-16 pareth 9, trideceth-12, glycerin, water (20:6:4:2:3:65) 2.00%, disodium phosphate 0.03%, **phosphoric acid** 0.02%, **hydrogen peroxide** (35%) 22.50%, and Steareth-10 allyl ether/acrylates copolymer 0.10%. Groups of 95% gray hair swatches were colored with a coloring compn. prepd. by mixing 1 part of the dye compn. and 1.5 parts of the developer compn. The coloring process was interrupted at 2, 5, 10, 20, and 30 min. The coloring compn. was capable of depositing more color on the hair in a faster time than com. products.

AB An aq. compn. for oxidatively coloring hair in 12 min or less is prepd. by combining a dye mixt. and a developer mixt. The compn. comprises (i) 0.01-2.0% (combined wt.) of primary intermediates and couplers, (ii) a free alky. concn. of 0.20-0.75 meq/g, and (iii) a hydrogen peroxide concn. of 4-6% by wt. A hair coloring kit contg. sep. containers of an aq. dye mixt. and an aq. developing mixt. is also described. For example, a hair dye mixt. in the level 5, medium brown, shade was prepd. contg. (by wt.) p-phenylenediamine 0.63% (primary intermediate), N,N-bis(2-hydroxyethyl)-p-phenylenediamine sulfate 0.11 (primary intermediate), resorcinol 0.50 (coupler), 1-naphthol 0.04 (coupler), m-aminophenol 0.07 (coupler), ethanolamine 3.00 (alkalizing agent), ammonium hydroxide (27.5%) 9.00 (alkalizing agent), oleic acid 12.50, erythorbic acid 0.20 (antioxidant), sodium sulfite 0.50 (antioxidant), ethoxydiglycol 5.00 (solvent), Hypnea musciformis ext., Gellidiala acerosa ext., Sargassum filipendula ext., and sorbitol 0.80 (conditioner), Oleth-20 1.00, Steareth-21 0.70, emulsifying wax 2.00, meadowfoam seed oil 0.75 (conditioner), Polyquaternium-10 0.20 (conditioner), Polyquaternium-28 0.50 (conditioner), hydrolyzed wheat protein 0.50 (conditioner), oleyl alc. 0.40 (conditioner, thickener), cetearyl alc. 4.00 (thickener), tetrasodium EDTA (38%) 0.80 (chelating agent), ammonium lauryl sulfate (28%) 2.00, mica, titanium dioxide (67:33) 0.30, sodium benzotriazolyl butylphenol sulfonate buteth-3, tri-Bu citrate 0.50 (UV absorber), fragrance 1.25%, and water 52.75%. A developer compn. contained (by wt.) water 69.72%, methylparaben 0.05%, EDTA 0.02%, mineral oil 0.60%, cetearyl alc./ceteareth-20 (80:20) 3.75%, cetearyl alc. 0.80%, ceteareth-20 0.40%, cyclomethicone/trimethylsiloxysilicate (50:50) 0.01%,

trimethylsilylamodimethicone, C11-15 pareth-7 C12-16 pareth 9, trideceth-12, glycerin, water (20:6:4:2:3:65) 2.00%, disodium phosphate 0.03%, **phosphoric acid** 0.02%, **hydrogen peroxide** (35%) 22.50%, and Steareth-10 allyl ether/acrylates copolymer 0.10%. Groups of 95% gray hair swatches were colored with a coloring compn. prepd. by mixing 1 part of the dye compn. and 1.5 parts of the developer compn. The coloring process was interrupted at 2, 5, 10, 20, and 30 min. The coloring compn. was capable of depositing more color on the hair in a faster time than com. products.

IT 90-15-3, 1-Naphthol 106-50-3, p-Phenylenediamine, biological studies 108-46-3, Resorcinol, biological studies 141-43-5, Ethanolamine, biological studies 591-27-5, m-Aminophenol 1336-21-6, Ammonium hydroxide 7722-84-1, Hydrogen peroxide, biological studies 54381-16-7, N,N-Bis(2-hydroxyethyl)-p-phenylenediamine **sulfate**
 RL: COS (Cosmetic use); BIOL (Biological study); USES (Uses)
 (aq. hair coloring compn. contg. primary intermediates, couplers, alkalizing agents, and **hydrogen peroxide**)

L3 ANSWER 10 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2002:216240 CAPLUS

DN 136:249443

TI Thickened liquid bleaching detergent compositions with good bleaching detergency

IN Maki, Masataka; Ozaki, Kazuyoshi; Aoyagi, Muneo

PA Kao Corp., Japan

SO Jpn. Kokai Tokkyo Koho, 8 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 2002080895	A2	20020322	JP 2000-268465	20000905
PRAI	JP 2000-268465		20000905		

AB The compn. comprises (a) hydrogen peroxide, (b) 0.1-10% compd. having .gtoreq.1 C8-20 alkyl or alkenyl, .gtoreq.1 polyoxy(C2-4 alkylene) chain, and wt. av. mol. wt. 2500-50,000, and (c) 0.01-10% poly(C2-4 alkylene glycol), wherein the ratio of c to b is 0.001-10. Thus a bleaching detergent comprising H2O2 2, Emulgen 2200 (polyoxyethylene cetyl ether) 2, polyethylene glycol 1, C12H25O(C2H4O)6H 5, N-tetradecyl -N,N-dimethylammonium Me **sulfate** 0.5, Sodium 4-(dodecanoyloxy)benzenesulfonate 0.5, **Dequest** 2010 (acetodiphosphonic acid) 0.4 parts and water balanced showed viscosity (at 20.degree.) 25 mPa-s and good bleachability.

AB The compn. comprises (a) hydrogen peroxide, (b) 0.1-10% compd. having .gtoreq.1 C8-20 alkyl or alkenyl, .gtoreq.1 polyoxy(C2-4 alkylene) chain, and wt. av. mol. wt. 2500-50,000, and (c) 0.01-10% poly(C2-4 alkylene glycol), wherein the ratio of c to b is 0.001-10. Thus a bleaching detergent comprising H2O2 2, Emulgen 2200 (polyoxyethylene cetyl ether) 2, polyethylene glycol 1, C12H25O(C2H4O)6H 5, N-tetradecyl -N,N-dimethylammonium Me **sulfate** 0.5, Sodium 4-(dodecanoyloxy)benzenesulfonate 0.5, **Dequest** 2010 (acetodiphosphonic acid) 0.4 parts and water balanced showed viscosity (at 20.degree.) 25 mPa-s and good bleachability.

L3 ANSWER 11 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2002:885993 CAPLUS

DN 137:371758

TI Process for manufacturing granules of stable adducts of sodium sulphate, hydrogen peroxide and sodium chloride as whiteners in detergents

IN Forner, Benito Juan; Artigas, Puerto Ramon

PA FMC Foret S.A., Spain

SO Eur. Pat. Appl., 8 pp.

CODEN: EPXXDW

DT Patent
LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 1258454	A1	20021120	EP 2001-500122	20010517
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO, MK, CY, AL, TR				
PRAI	EP 2001-500122		20010517		
AB	The process comprises directly granulating an intimate mixt. of Na ₂ SO ₄ and NaCl having a grain size .ltoreq.100 .mu.m with H ₂ O ₂ ; or directly granulating Na ₂ SO ₄ with grain size .ltoreq.100 .mu.m by means of spraying with a soln. of NaCl in H ₂ O ₂ ; or directly granulating a blend formed by adding NaCl (solid or in soln.) to a paste contg. Na ₂ SO ₄ and H ₂ O ₂ . The granulated adduct 4Na ₂ SO ₄ .2H ₂ O ₂ .NaCl is intrinsically stable and, once dry, only contains .ltoreq.4.6 % active oxygen.				
RE.CNT	5 THERE ARE 5 CITED REFERENCES AVAILABLE FOR THIS RECORD ALL CITATIONS AVAILABLE IN THE RE FORMAT				
ST	sodium sulfate hydrogen peroxide adduct granule manuf.; detergent whitener sodium chloride granule stability				
IT	Detergents (powd.; process for manufg. granules of stable adducts of sodium sulfate, hydrogen peroxide and sodium chloride as whiteners in detergents)				
IT	Stabilizing agents Whitening agents (process for manufg. granules of stable adducts of sodium sulfate, hydrogen peroxide and sodium chloride as whiteners in detergents)				
IT	7722-84-1, Hydrogen peroxide , uses 13759-07-4, Thenardite 14762-51-7, Halite RL: PEP (Physical, engineering or chemical process); PYP (Physical process); TEM (Technical or engineered material use); PROC (Process); USES (Uses) (process for manufg. granules of stable adducts of sodium sulfate, hydrogen peroxide and sodium chloride as whiteners in detergents)				
IT	7487-88-9, Magnesium sulfate , uses 12058-66-1, Sodium stannate 15827-60-8, Dequest 2060 RL: TEM (Technical or engineered material use); USES (Uses) (stabilizer; process for manufg. granules of stable adducts of sodium sulfate, hydrogen peroxide and sodium chloride as whiteners in detergents)				

L3 ANSWER 12 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2002:503326 CAPLUS

DN 137:67887

TI Hair bleach composition and hair dye composition

IN Matsuo, Takashi; Miyabe, Hajime; Shibata, Yutaka; Ito, Yoshiaki; Monda, Keiji; Misu, Daisuke

PA Kao Corporation, Japan

SO Eur. Pat. Appl., 18 pp.

CODEN: EPXXDW

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 1219285	A2	20020703	EP 2001-130241	20011219
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO, MK, CY, AL, TR				
	JP 2002193773	A2	20020710	JP 2000-400808	20001228
	JP 2002193774	A2	20020710	JP 2000-400875	20001228
	JP 2002193770	A2	20020710	JP 2000-400876	20001228

	JP 2002193771	A2	20020710	JP 2000-400877	20001228
	JP 2002201118	A2	20020716	JP 2000-402454	20001228
	JP 2002226340	A2	20020814	JP 2001-27704	20010205
	JP 2002265338	A2	20020918	JP 2001-61696	20010306
	US 2002139957	A1	20021003	US 2001-25762	20011226
PRAI	JP 2000-400808	A	20001228		
	JP 2000-400875	A	20001228		
	JP 2000-400876	A	20001228		
	JP 2000-400877	A	20001228		
	JP 2000-402454	A	20001228		
	JP 2001-27704	A	20010205		
	JP 2001-61696	A	20010306		
AB	<p>A hair bleach or dye compn. is described which, during use, includes a mixt. of a first compn. contg. an alkali agent and a second compn. contg. an oxidizing agent. The compn. contains (A) an org. solvent having a partition coeff. log P in octanol-water at 25.degree. of 0.3-6 and a mol. wt. of 200 or less, in an amt. of 1-70 wt.%; (B) an org. solvent having a log P of less than 0.3, in an amt. of 0-8 wt.% and less than that of component (A); (C) an alkali agent, e.g., ammonia, in an amt. of 0.1-10 wt.%; (D) an oxidizing agent in an amt. of 0.1-12 wt.% as reduced to H2O2; (E) water in an amt. of 20-70 wt.%, and further may contain (F) a cationic surfactant in an amt. of 0.01-10 wt.%. A hair dye compn. further comprises (G) an oxidn.-type dye intermediate or (H) a direct dye. The compn. has a pH of 7.5-12 after mixing of the first compn. and the second compn. The bleach or dye compn. provides less offensive odor, exhibit excellent hair-bleaching power and hair-dyeing power, and exhibit excellent hair-conditioning effect during and after treatment. For example, an oxidn.-type hair bleach compn. was prepd. contg. (by wt.) first compn. - 2-benzyloxyethanol 10%, ethanol 3%, monoethanolamine 6%, oleyl alc. 2%, polyoxyethylene (20) octyl dodecyl ether 18%, polyoxyethylene (9) oleyl ether 6%, polyoxyethylene (3) tridecyl ether 15%, polyethylene glycol 400 8%, liq. paraffin 6%, perfume 0.4%, and water 25.6%; and second compn. - 35% aq. hydrogen peroxide 17%, polyoxyethylene (9) oleyl ether 25%, oleyl alc. 15%, 8-quinolinol sulfate 0.04%, 75% phosphoric acid 0.02%, and water 42.94%. The first compn. and the second compn. were mixed at a ratio of 1:1.5 (by wt.), to thereby prep. a hair bleach compn. When used, the compn. provided a weak offensive odor and excellent bleach power.</p>				
AB	<p>A hair bleach or dye compn. is described which, during use, includes a mixt. of a first compn. contg. an alkali agent and a second compn. contg. an oxidizing agent. The compn. contains (A) an org. solvent having a partition coeff. log P in octanol-water at 25.degree. of 0.3-6 and a mol. wt. of 200 or less, in an amt. of 1-70 wt.%; (B) an org. solvent having a log P of less than 0.3, in an amt. of 0-8 wt.% and less than that of component (A); (C) an alkali agent, e.g., ammonia, in an amt. of 0.1-10 wt.%; (D) an oxidizing agent in an amt. of 0.1-12 wt.% as reduced to H2O2; (E) water in an amt. of 20-70 wt.%, and further may contain (F) a cationic surfactant in an amt. of 0.01-10 wt.%. A hair dye compn. further comprises (G) an oxidn.-type dye intermediate or (H) a direct dye. The compn. has a pH of 7.5-12 after mixing of the first compn. and the second compn. The bleach or dye compn. provides less offensive odor, exhibit excellent hair-bleaching power and hair-dyeing power, and exhibit excellent hair-conditioning effect during and after treatment. For example, an oxidn.-type hair bleach compn. was prepd. contg. (by wt.) first compn. - 2-benzyloxyethanol 10%, ethanol 3%, monoethanolamine 6%, oleyl alc. 2%, polyoxyethylene (20) octyl dodecyl ether 18%, polyoxyethylene (9) oleyl ether 6%, polyoxyethylene (3) tridecyl ether 15%, polyethylene glycol 400 8%, liq. paraffin 6%, perfume 0.4%, and water 25.6%; and second compn. - 35% aq. hydrogen peroxide 17%, polyoxyethylene (9) oleyl ether 25%, oleyl alc. 15%, 8-quinolinol sulfate 0.04%, 75% phosphoric acid 0.02%, and water 42.94%. The first compn. and the second compn. were mixed at a ratio of 1:1.5 (by wt.), to thereby prep. a hair bleach compn. When used, the compn. provided a weak offensive odor and excellent bleach power.</p>				

L3 ANSWER 13 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2002:961452 CAPLUS
DN 138:16454
TI Hair bleaches containing ubiquinones
PA Goldwell GmbH, Germany
SO Ger. Gebrauchsmusterschrift, 13 pp.
CODEN: GGXXFR
DT Patent
LA German
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 20110784	U1	20021219	DE 2001-20110784	20010629
PRAI	DE 2001-20110784		20010629		
OS	MARPAT 138:16454				

AB The invention concerns hair bleaches that contain ubiquinones, peroxides and ammonium compds. Thus a hair bleach contained a bleaching powder, a hydrogen peroxide lotion and a cream base; they were mixed at a ratio of 1:2.5:2.5. The bleaching powder was composed of (wt./wt.%): silica 16.27; ammonium chloride 14.70; sodium carbonate 10.00; sodiummetasilicate 2.30; phthalimidoperhexanoic acid 41.50; sodium persulfate 15.20; Coenzyme Q6 0.03. The **hydrogen peroxide** lotion contained (wt./wt.%): **hydrogen peroxide** 6.00; cetylstearyl alc. 1.70; **phosphoric acid** 0.50; sodium lauryl **sulfate** 0.20; Coenzyme Q6 0.05; disodium hydrogen phosphate 0.10; salicylic acid 0.10; water to 100. The cream base included (wt./wt.%): cetylstearylalc. 11.0; Oleth-5 5.0; oleic acid 2.5; stearamide MEA 2.3; cocoamide MEA 2.3; sodium cetyl stearyl sulfate 1.2; 1,2-propyleneglycol 1.0; 1,2-propyleneglycol stearate 0.6; sodium lauryl sulfate 0.5; wheat protein hydrolyzate 0.9; organopolysiloxane 0.4; panthenol 0.6; perfume 0.4; complexing agent 0.2; water to 100.

AB The invention concerns hair bleaches that contain ubiquinones, peroxides and ammonium compds. Thus a hair bleach contained a bleaching powder, a hydrogen peroxide lotion and a cream base; they were mixed at a ratio of 1:2.5:2.5. The bleaching powder was composed of (wt./wt.%): silica 16.27; ammonium chloride 14.70; sodium carbonate 10.00; sodiummetasilicate 2.30; phthalimidoperhexanoic acid 41.50; sodium persulfate 15.20; Coenzyme Q6 0.03. The **hydrogen peroxide** lotion contained (wt./wt.%): **hydrogen peroxide** 6.00; cetylstearyl alc. 1.70; **phosphoric acid** 0.50; sodium lauryl **sulfate** 0.20; Coenzyme Q6 0.05; disodium hydrogen phosphate 0.10; salicylic acid 0.10; water to 100. The cream base included (wt./wt.%): cetylstearylalc. 11.0; Oleth-5 5.0; oleic acid 2.5; stearamide MEA 2.3; cocoamide MEA 2.3; sodium cetyl stearyl sulfate 1.2; 1,2-propyleneglycol 1.0; 1,2-propyleneglycol stearate 0.6; sodium lauryl sulfate 0.5; wheat protein hydrolyzate 0.9; organopolysiloxane 0.4; panthenol 0.6; perfume 0.4; complexing agent 0.2; water to 100.

L3 ANSWER 14 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2002:66696 CAPLUS
DN 136:120256
TI Storage-stable aqueous bleaching composition
IN Cruz, Mercedes Mendoza; Pons, Jaume Josa; Catalina, Adolf Arranz
PA Henkel KGaA, Germany
SO Ger. Offen., 8 pp.
CODEN: GWXXBX
DT Patent
LA German
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 10032589	A1	20020124	DE 2000-10032589	20000707
PRAI	DE 2000-10032589		20000707		

OS MARPAT 136:120256

AB The problems of unstable viscosity of aq. bleaching agents during storage and of increased oxidative load of textile laundered with peroxide-contg. detergents or pretreatment compns. were solved by use of aq. bleaching agent comprising H2O2 3-8, nonionic surfactant 0.2-2, anionic surfactant 0.5-5%, radical scavenger 0.01-0.2, H2O-miscible solvent or dispersant for the radical scavenger .ltoreq.1, 1st complexing agent for heavy metals with .gtoreq.1 phosphonic acid group 0.1-2, a 2nd complexing agent for heavy metals with .gtoreq.1 carboxylic acid group 0.05-2%, and a polysaccharide thickener in an amt. necessary to produce viscosity 200-5000 mPa.cntdot.s at 20.degree. and 20 rpm (Brookfield rotational viscometer), the balance being H2O with pH 2.5-7.5. A typical bleach-contg. liq. detergent having pH 5.5 contained H2O2 5, ethoxylated (2 EO) C12-14 fatty alc. sulfate triethanolamine salts 2, Arlipon FT (nonionic surfactant) 0.53, Keltron T 0.61, 1,6-di-tert-butyl-4-methylphenol 0.03, EtOH 0.2, 1-hydroxyethanediphosphonic acid 0.14, methylglycinediacetic acid 0.33, isothiazolone perfume/dye/preservative 0.15 parts and H2O balance.

AB The problems of unstable viscosity of aq. bleaching agents during storage and of increased oxidative load of textile laundered with peroxide-contg. detergents or pretreatment compns. were solved by use of aq. bleaching agent comprising H2O2 3-8, nonionic surfactant 0.2-2, anionic surfactant 0.5-5%, radical scavenger 0.01-0.2, H2O-miscible solvent or dispersant for the radical scavenger .ltoreq.1, 1st complexing agent for heavy metals with .gtoreq.1 phosphonic acid group 0.1-2, a 2nd complexing agent for heavy metals with .gtoreq.1 carboxylic acid group 0.05-2%, and a polysaccharide thickener in an amt. necessary to produce viscosity 200-5000 mPa.cntdot.s at 20.degree. and 20 rpm (Brookfield rotational viscometer), the balance being H2O with pH 2.5-7.5. A typical bleach-contg. liq. detergent having pH 5.5 contained H2O2 5, ethoxylated (2 EO) C12-14 fatty alc. sulfate triethanolamine salts 2, Arlipon FT (nonionic surfactant) 0.53, Keltron T 0.61, 1,6-di-tert-butyl-4-methylphenol 0.03, EtOH 0.2, 1-hydroxyethanediphosphonic acid 0.14, methylglycinediacetic acid 0.33, isothiazolone perfume/dye/preservative 0.15 parts and H2O balance.

L3 ANSWER 15 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2002-713562 [77] WPIDS

DNC C2002-202385

TI New colloidal copper compounds, useful as agricultural fungicides, are stable, adhere well to plants and are effective at low application rates.

DC C03

IN HUATO, J; OGURA, T

PA (HUAT-I) HUATO J; (OGUR-I) OGURA T; (DERM-N) DERMET SA DE CV

CYC 100

PI WO 2002075007 A2 20020926 (200277)* ES 22p
 RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
 NL OA PT SD SE SL SZ TR TZ UG ZM ZW
 W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CO CR CU CZ DE DK
 DM DZ EC EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR
 KZ LC LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ OM PH PL PT
 RO RU SD SE SG SI SK SL TJ TM TN TR TT TZ UA UG US UZ VN YU ZA ZM
 ZW

US 2002136781 A1 20020926 (200277)

ADT WO 2002075007 A2 WO 2002-MX22 20020320; US 2002136781 A1 US 2001-811610 20010320

PRAI US 2001-811610 20010320

AB WO 200275007 A UPAB: 20021129
 NOVELTY - Colloidal copper compounds (I) for use as agricultural fungicides are new.
 DETAILED DESCRIPTION - Colloidal copper compounds of formula (I) are new.
 $x+y = 0-2$, with $mx + ny$ equal to 2;
 $m+n$ = oxidation numbers of the anions A and B, respectively;

A = chloride, bromide, iodide, fluoride, sulfate, phosphate, RCOO anion, tartrate, citrate or amino acid residue; and

R = hydrogen, 1-20C linear or branched hydrocarbyl chain, or aryl; B is not defined.

An INDEPENDENT CLAIM is also included for preparation of (I) by purifying a solution of copper(II) ions then increasing the pH of the solution.

ACTIVITY - Fungicide.

When tomato plants were treated on day 48 with a suspension of 50 million spores of an unspecified Fusarium in 100 ml water, all plants were dead by day 69. When the plants were treated on day 45 with a colloidal copper citrate (50 mg copper/l), then with the spores on day 48, they showed a level of infection on the leaves, of only 5%.

MECHANISM OF ACTION - None given in the source material.

USE - (I) are used for control of fungal infections on plants.

ADVANTAGE - (I) produce stable colloidal particles (specifically because iron and aluminum impurities have been removed) that adhere well to leaves and require a lower rate of application than conventional copper fungicides.

Dwg.0/0

TECH

UPTX: 20021129

TECHNOLOGY FOCUS - INORGANIC CHEMISTRY - Preferred Process: The copper(II) ion solution is prepared from copper(II) sulfate pentahydrate and purification comprises:

(1) adding an oxidizing agent (e.g. hydrogen peroxide or chlorine) and phosphoric acid;

(2) adjusting to pH 3 (especially with aqueous sodium carbonate); and

(3) heating the solution and removing solids (iron and aluminum impurities)...

L3 ANSWER 16 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2002-732868 [79] WPIDS

DNN N2002-577844 DNC C2002-207444

TI Formation of surface conductivity on substrate by treating the substrate with monomer and oxidant mixture, and reacting the monomer and oxidant in contact with and on the substrate surface.

DC A85 A92 L03 U11 V04 X12 X25

IN GE, H

PA (HWEL-N) HW ELECTROCHEM TECHNOLOGY PTE LTD

CYC 1

PI WO 2002074841 A1 20020926 (200279)* EN 15p

W: SG

ADT WO 2002074841 A1 WO 2001-SG63 20010316

PRAI WO 2001-SG63 20010316

AB WO 200274841 A UPAB: 20021209

NOVELTY - A surface conductivity is formed on a substrate by treating the substrate with a mixture of a monomer and an oxidant and allowing the monomer and oxidant to react in contact with and on the surface of the substrate to provide an intrinsically conductive polymeric surface on the substrate.

USE - The process is used for forming of surface conductivity on substrate. The resulting substrate is used in the packaging of microchips, circuit boards and other electronic components. It is can also be used to form a Faraday cage, in conductive flooring for electronic component production plants, wafer storage and shipping, personal electrostatic protection devices, work stations in semiconductor or electronic factories, conductive paint, cable shielding material and in the formation of conductive textiles.

ADVANTAGE - The resulting substrate has a surface conductivity of 1-1 multiply 10power11 (preferably 1 multiply 10power2-1 multiply 10power8) Omega /square. The process provides an improved production of conductive surface coating on non-conductive substrates.

Dwg.0/0

TECH. . .

Preferred Oxidants: The oxidant is a compound which can change its valence. It is iron (III) chloride (FeCl_3), iron (III) **sulfate** ($\text{Fe}_2(\text{SO}_4)_3$), ammonium sulfur oxide-containing compound ($(\text{NH}_4)_2\text{S}_2\text{O}_8$), sodium sulfur oxide-containing compound ($\text{Na}_2\text{S}_2\text{O}_8$), potassium dichromate ($\text{K}_2\text{Cr}_2\text{O}_7$), **hydrogen peroxide** (H_2O_2), nitric acid (HNO_3), perchloric acid (HClO_4), quinone, potassium ferricyanide ($\text{K}_3(\text{Fe}(\text{CN})_6$), **phosphoric acid** (H_3PO_4), molybdenum (VI) oxide (MoO_3), tungsten oxide (WO_3), chromium (VI) oxide (CrO_3), ammonium cerium (III) sulfate ($(\text{NH}_4)\text{Ce}(\text{NO}_3)_6$), cerium sulfate ($\text{Ce}(\text{SO}_4)_2$), . . .

L3 ANSWER 17 OF 158 WPIDS (C) 2003 THOMSON DERWENT
 AN 2002-682664 [73] WPIDS
 DNN N2002-539008 DNC C2002-192528
 TI Injectable bone like implant used for repairing bone defect and injury comprises bone like compound and hydrophobic carrier or degradable component.
 DC A96 B04 D22 P34
 IN WIRONEN, J F
 PA (WIRO-I) WIRONEN J F; (REGE-N) REGENERATION TECHNOLOGIES INC
 CYC 96
 PI WO 2002058755 A2 20020801 (200273)* EN 15p
 RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
 NL OA PT SD SE SL SZ TR TZ UG ZM ZW
 W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CO CR CU CZ DE DK
 DM DZ EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ
 LC LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ PL PT RO RU SD
 SE SG SI SK SL TJ TM TR TT TZ UA UG UZ VN YU ZA ZW
 US 2002193883 A1 20021219 (200303)
 ADT WO 2002058755 A2 WO 2002-US3092 20020125; US 2002193883 A1 Provisional US 2001-263972P 20010125, US 2002-56217 20020125
 PRAI US 2001-263972P 20010125; US 2002-56217 20020125
 AB WO 200258755 A UPAB: 20021113
 NOVELTY - A bone-like implant comprises at least one bone-like compound and a hydrophobic carrier or at least one degradable component.
 DETAILED DESCRIPTION - An INDEPENDENT CLAIM is included for production of the implant which comprises mixing at least one bone-like compound in a hydrophobic carrier or a degradable component and concurrently or subsequently combining with an aqueous phase to form a combined mixture.
 ACTIVITY - Osteopathic.
 MECHANISM OF ACTION - None given in the source material.
 USE - Used for repairing a bone defect and injury.
 ADVANTAGE - The implant is capable of aqueous sintering or curing and increasing its porosity in situ.
 Dwg.0/0
 TECH. . .
 - INORGANIC CHEMISTRY - Preferred Components: The bone-like compound is tricalcium phosphate, dicalcium phosphate or monocalcium phosphate, potassium phosphate, calcium **sulfate**, hydroxyapatite and/or bioactive glass. The degradable component is a degradable gas-producing compound (preferably sodium bicarbonate and/or calcium bicarbonate, **hydrogen peroxide** and peroxidase) and an acid (preferably citric acid, formic acid, acetic **phosphoric acid** or hydrochloric acid).
 The aqueous component is water, saline and/or blood.
 Preferred Method: The method involves reconstituting at least one bone-like compound. . .

L3 ANSWER 18 OF 158 WPIDS (C) 2003 THOMSON DERWENT
 AN 2002-713305 [77] WPIDS
 DNN N2002-562777 DNC C2002-202161
 TI Cleaning wipe, for removing stains from inanimate surfaces, e.g. carpets, comprises cleaning composition, which is composed of surfactant system(s), preservative system(s) and carrier.

DC A97 D25 E19 F04 P73
IN DURDEN, C; MAURO, A J; MICCICHE, R P; TRIPATHI, U
PA (INPL) PLAYTEX PROD INC
CYC 97
PI WO 2002055650 A1 20020718 (200277)* EN 33p
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
NL OA PT SD SE SL SZ TR TZ UG ZM ZW
W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CO CR CU CZ DE DK
DM DZ EC EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR
KZ LC LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ PL PT RO RU
SD SE SG SI SK SL TJ TM TR TT TZ UA UG UZ VN YU ZA ZW
US 2002174500 A1 20021128 (200281)
ADT WO 2002055650 A1 WO 2002-US1124 20020111; US 2002174500 A1 Provisional US
2001-261399P 20010112, US 2002-43872 20020110
PRAI US 2002-43872 20020110; US 2001-261399P 20010112
AB WO 200255650 A UPAB: 20021129
NOVELTY - A cleaning wipe comprises a cleaning composition which is
composed of surfactant system(s), preservative system(s) and a carrier,
and which is adjusted to a pH of 7.5-10.5. It has a loading level ratio of
1:1-10:1, based on the total weight of the cleaning composition to a total
weight of the wipe.
DETAILED DESCRIPTION - An INDEPENDENT CLAIM is also included for a
method of producing the above cleaning wipe by placing the wipe in a
container; dispensing the composition into the container; and sealing the
container.
USE - For removing stains from inanimate surfaces, e.g. carpets,
fabrics, countertop, drawer, shelf, or metal fixture.
ADVANTAGE - The inventive wipe provides superior cleaning and ease of
use while maintaining chemical stability, resisting microbial growth and
maintaining cleaning solution distribution during storage.
Dwg.0/0
TECH. . .
skin softening and conditioning agent, alcohol, and soil resist. The pH
control agent is the salt of an acid, preferably 1-hydroxyethylidene-1,1-
diphosphonic acid, citric acid, sodium bicarbonate, sodium
citrate, sodium **sulfate**, sodium phosphate, and sodium carbonate.
The peroxygen compound is **hydrogen peroxide**, t-butyl
hydroperoxide, sodium percarbonate, and sodium perborate. Preferred skin
softening and conditioning agent is lanolin, lanolin derivative, polyol,
botanical derivative, . . .

L3 ANSWER 19 OF 158 WPIDS (C) 2003 THOMSON DERWENT
AN 2002-508636 [54] WPIDS
DNC C2002-144661
TI Peroxosilicate compound useful in the disinfection of water in e.g.
swimming pools, is obtained by reacting alkali or alkaline earth
metasilicate with an oxygen-liberating compound.

DC D15 D22 E37 M14
IN AURIOL, M; BOUKARI, M; AURIOL, S
PA (EOTE-N) EOTEC; (INSF) INST FRANCAIS DU PETROLE
CYC 93
PI WO 2002048030 A1 20020620 (200254)* FR 28p
RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE TR
W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CO CR CU CZ DE DK
DM DZ EC EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR
KZ LC LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ PL PT RO RU
SD SE SG SI SK SL TJ TM TR TT TZ UA UG US UZ VN YU ZA ZW
FR 2818264 A1 20020621 (200254)
AU 2002019295 A 20020624 (200267)
ADT WO 2002048030 A1 WO 2001-FR3974 20011213; FR 2818264 A1 FR 2000-16367
20001215; AU 2002019295 A AU 2002-19295 20011213
FDT AU 2002019295 A Based on WO 200248030
PRAI FR 2000-16367 20001215
AB WO 200248030 A UPAB: 20020823

NOVELTY - A peroxosilicate compound with anti-scale and anti-corrosive properties, partly soluble in water, is obtained by the reaction of an alkali or alkaline earth metasilicate, preferably pentahydrated sodium or potassium metasilicate, with an oxygen-liberating compound active at a neutral pH.

DETAILED DESCRIPTION - An INDEPENDENT CLAIM is included for a method of preparation of the compound.

USE - The compounds can be used for water treatment, particularly water in swimming pools. The combination of ingredients having the functions of: disinfection by progressive liberation of active O₂ in the water; prolonged stabilization enabling O₂ liberation over at least 3 weeks; autoregulation of pH due to a buffer effect over a wide pH range (6 - 8); antiscaling and lowering of hardness by complexation with Ca; or non-adherent scale formation eliminated by filtration; anti-corrosion; algicidal effects; and flocculation of materials in suspension or colloids, easily eliminated by filtration.

ADVANTAGE - The use of chlorine compounds, which are harmful to the environment and impart a corrosive character to the water treated, is avoided.

Dwg.0/0

TECH

UPTX: 20020823

TECHNOLOGY FOCUS - INORGANIC CHEMISTRY - Preferred Compounds: The reaction is carried out in the presence of **phosphoric acid**, sodium phosphate or acid sodium **sulfate**. The O₂-liberating compound is **H₂O₂**, the monopersulfates and acid monopersulfates of Na and K are sodium peroxodisulfate, peroxides of Li, Na, Ba, sodium peroxoborate and potassium peroxodisulfate; more preferably **H₂O₂** or potassium monopersulfate. Stabilized by reaction with a stabilizer, preferably sodium or potassium hexametaphosphate, before reaction with the metasilicate, to. . .

L3 ANSWER 20 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2002-712485 [77] WPIDS

DNN N2002-562016 DNC C2002-202023

TI Treatment of lead-contaminated waste materials or soil involves adding ferrous iron to the contaminated waste materials or soil and oxidizing the ferrous iron by drying or chemical oxidation.

DC D15 E37 J01 P35

IN STANFORTH, R R

PA (STAN-I) STANFORTH R R

CYC 1

PI US 2002115899 A1 20020822 (200277)* 9p

ADT US 2002115899 A1 Provisional US 1999-172574P 19991220, US 2000-736307 20001215

PRAI US 1999-172574P 19991220; US 2000-736307 20001215

AB US2002115899 A UPAB: 20021129

NOVELTY - Lead-contaminated waste material or soil is treated by adding ferrous iron to the contaminated waste material or soil. The ferrous iron is oxidized in the waste materials or soil by drying or chemical oxidation to form a ferric compound incorporating the lead.

USE - For treating lead-contaminated waste materials or soil (claimed).

ADVANTAGE - The inventive method reduces lead bioavailability to control leaching in the natural environment and during digestion after accidental ingestion. It effectively treats hazardous lead contaminated material. The lead level after treatment is less than 0.75 mg/L as measured by the TCLP lead leaching test.

Dwg.0/0

TECH.

to the total weight of the treated waste materials or soil.
Preferred Component: The soluble ferrous salt is ferrous nitrate, ferrous **sulfate** or ferrous chloride. The chemical oxidant is **hydrogen peroxide**. The source of orthophosphate is **phosphoric acid**, sodium phosphate, and/or triple

superphosphate. The alkaline material is calcium carbonate, calcium hydroxide, calcium oxide, magnesium hydroxide or magnesium oxide.

L3 ANSWER 21 OF 158 WPIDS (C) 2003 THOMSON DERWENT
AN 2002-706439 [76] WPIDS
CR 2002-146963 [19]; 2002-674176 [72]; 2002-674177 [72]
DNN N2002-556998 DNC C2002-200294
TI Memory device, e.g. dynamic random access memory, has titanium nitride layer which is planarized using planarizing medium containing fixed-abrasive planarizing pad and clean planarizing solution.
DC L03 P61 U11
IN CHOPRA, D; SABDE, G
PA (MICR-N) MICRON TECHNOLOGY INC
CYC 1
PI US 2002115384 A1 20020822 (200276)* 18p
ADT US 2002115384 A1 Div ex US 1999-339735 19990624, US 2002-115675 20020404
PRAI US 1999-339735 19990624; US 2002-115675 20020404
AB US2002115384 A UPAB: 20021125
NOVELTY - A memory device comprises an array of memory cells, row and column access circuits, and an address decoder circuit. At least one of the memory cells, access circuits and address decoder circuit has a titanium nitride layer which is planarized by abrading with a planarizing surface of a medium containing fixed-abrasive planarizing pad and clean planarizing solution.

DETAILED DESCRIPTION - INDEPENDENT CLAIMS are also included for the following:

(a) the production of a memory device by forming an array of memory cells, coupling row and column access circuits to the memory cells, respectively, and coupling an address decoder circuit to the access circuits;

(b) a memory module comprising a support, leads extending from the support, a command link coupled to at least one of the leads, data links coupled to the leads, and at least one memory device contained on the support and coupled to the command link;

(c) a memory system comprising a controller, command and data links respectively coupled to the controller, and the above memory device coupled to the command link and the data link; and

(d) the production of a memory module and memory system.

USE - As e.g. dynamic random access memory, static random access memory, or flash memory.

ADVANTAGE - The memory device exhibits physical characteristics different from the physical characteristics inherent in planarizing titanium nitride using non-abrasive planarizing pad and abrasive slurry planarizing solution.

DESCRIPTION OF DRAWING(S) - The figure is a block diagram of a computer system.

Dwg.9/9

TECH.

or ca. 2 wt.% oxidizer.

Preferred Components: The etchant comprises at least one etching agent, i.e. oxalic acid, ascorbic acid, or **phosphoric acid**.

The buffer comprises at least one buffering agent, i.e. ammonium acetate, ammonium oxalate, ammonium phosphate, or diammonium phosphate. The oxidizer comprises at least one oxidizing agent, i.e. ammonium persulfate, ammonium heptamolybdate, ceric ammonium nitrate, ceric ammonium **sulfate**, or **hydrogen peroxide**.

L3 ANSWER 22 OF 158 WPIDS (C) 2003 THOMSON DERWENT
AN 2002-674176 [72] WPIDS
CR 2002-146963 [19]; 2002-674177 [72]; 2002-706439 [76]; 2003-045815 [04]
DNN N2002-533091 DNC C2002-189849
TI Electronic system as memory system, contains circuit module having semiconductor die which contains integrated circuit having titanium nitride layer planarized with planarizing surface of preset planarizing

medium.

DC L03 P61 U11
IN CHOPRA, D; SABDE, G
PA (MICR-N) MICRON TECHNOLOGY INC
CYC 1
PI US 2002106975 A1 20020808 (200272)* 16p
ADT US 2002106975 A1 Div ex US 1999-339735 19990624, US 2002-116585 20020404
PRAI US 1999-339735 19990624; US 2002-116585 20020404
AB US2002106975 A UPAB: 20030117

NOVELTY - Electronic system contains circuit module having several leads coupled to processor and semiconductor die (SD) coupled to lead portion(s). SD contains substrate (12) supporting integrated circuit having titanium nitride layer (TNL). TNL is planarized by abrading TNL with planarizing surface (142) of planarizing medium having fixed-abrasive planarizing pad (140) and clean planarizing solution (44).

DETAILED DESCRIPTION - An INDEPENDENT CLAIM is included for a method of forming the electronic system. The method involves coupling a circuit module to a processor.

USE - As memory system, computer system, clock, television, cell phone, personal computer, automobile, industrial control system and aircraft.

ADVANTAGE - The circuit module provides unilateral or bilateral communication and control. The integrated circuit device exhibits reduced performance variability. The computer system which includes the memory system as a sub-component incorporates processor and memory system on a single integrated circuit. Hence the communication time between the processor and the memory circuit is reduced.

DESCRIPTION OF DRAWING(S) - The figure shows a schematic view of fixed-pad format planarizing machine.

Substrate 12
Planarizing solution 44
Planarizing pad 140
Planarizing surface 142

Dwg.2B/9

TECH. . .
agent selected from oxalic acid, ascorbic acid and phosphoric acid. The buffer comprises at least one buffering agent selected from **phosphoric acid**, ammonium acetate, ammonium oxalate, ammonium phosphate and diammonium phosphate. The planarizing solution has a pH of approximately 1-5, preferably less. . . 3. The oxidizer comprises at least one oxidizing agent selected from ammonium persulfate, ammonium heptamolybdate, ceric ammonium nitrate, ceric ammonium **sulfate** and **hydrogen peroxide**. The planarizing solution is an aqueous solution comprising approximately 1-5 wt.% of oxalic acid and approximately 2-10 wt.% of ammonium. . .

L3 ANSWER 23 OF 158 WPIDS (C) 2003 THOMSON DERWENT
AN 2002-360919 [39] WPIDS
DNN N2002-281998 DNC C2002-102140
TI Etching solution for chemical mechanical planarization or spin etch planarization of copper/tantalum/tantalum nitride surface, includes oxidizing reactant, co-reactant, and other additives.
DC A85 E19 L03 U11 X12
IN LEVERT, J A; MUKHERJEE, S P; TOWERY, D L; ZHANG, F; LEVERT, J; MUKHERJEE, S; TOWERY, D
PA (LEVE-I) LEVERT J A; (MUKH-I) MUKHERJEE S P; (TOWE-I) TOWERY D L; (ZHAN-I) ZHANG F; (HONE) HONEYWELL INT INC
CYC 92
PI US 2002020833 A1 20020221 (200239)* 10p
WO 2002059393 A1 20020801 (200260) EN
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
NL OA PT SD SE SL SZ TR TZ UG ZM ZW
W: AE AL AM AT AU AZ BA BB BG BR BY CA CH CN CR CU CZ DE DK DM EE ES
FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ LC LK LR LS

LT LU LV MA MD MG MK MN MW MX NO NZ PL PT RO RU SD SE SG SI SK SL
TJ TM TR TT TZ UA UG UZ VN YU ZA ZW

ADT US 2002020833 A1 Cont of US 1999-357264 19990719, US 2000-745266 20001220;
WO 2002059393 A1 WO 2001-US48867 20011218

PRAI US 1999-357264 19990719; US 2000-745266 20001220

AB US2002020833 A UPAB: 20020621

NOVELTY - An etching solution comprises an oxidizing reactant from ammonium persulfate, **hydrogen peroxide**, and nitric acid; a co-reactant from **phosphoric acid**, sulfuric acid, nitric acid, oxalic acid, acetic acid, and organic acids; and other additives.

DETAILED DESCRIPTION - An etching solution comprises an oxidizing reactant from ammonium persulfate, **hydrogen peroxide**, and nitric acid; a co-reactant from **phosphoric acid**, sulfuric acid, nitric acid, oxalic acid, acetic acid, and organic acids; and other additives from hydrochloric acid, aliphatic alcohols, butylated hydroxytoluene, Agidol-2, 2,6-di-tert-butyl-4((dimethylamino)methyl)phenol, 2,6-di-tert-4N,N-dimethylaminomethylphenol, borax, poly(oxyethylene)lauryl ether, malic acid, HOOC(X2)nCOOH, 3% tartaric acid, 1% ethylene glycol, 1,2,4-triazole, 1,2,3-triazole, tetrazole, nonionic surfactant, ethanol, trifluoroethanol, silicon hexafluoride, organic salt surfactant, polyvinyl alcohol, diphenylsulfamic acid, sodium oxalate, benzotriazole, sodium lignosulfonate, glycol, gelatin carboxymethylcellulose, amines, heavy metal salts, salts of copper and tantalum, potassium chloride, copper chloride, tin chloride, propylene glycol, 2-ethyl-hexylamine, copper carbonate, low molecular weight alcohols, glycols, phenols, aliphatic alcohols, polyvinylalcohols, anionic surfactants, cationic surfactants, fluorocarbon-based surfactants, non-ionic surfactants, polyvinyl alcohol solution stabilizers and species inhibiting spontaneous decomposition of oxidizing agents, and wetting agents. The non-ionic surfactants have the properties of adhering to certain materials to modify the chemical reactivity.

X = OH, amine, or H;

n = 1-4.

USE - For chemical mechanical planarization or spin etch planarization of copper/tantalum/tantalum nitride surface in the fabrication of integrated circuits.

ADVANTAGE - The invention achieves 1:1 removal rate selectivity for copper and tantalum/tantalum nitride, and provides adequate material removal rates without excessive downforce being necessary on low dielectric components.

DESCRIPTION OF DRAWING(S) - The figure shows a schematic view of chemical mechanical planarization.
Dwg.1/1

AB US2002020833 UPAB: 20020621

NOVELTY - An etching solution comprises an oxidizing reactant from ammonium persulfate, **hydrogen peroxide**, and nitric acid; a co-reactant from **phosphoric acid**, sulfuric acid, nitric acid, oxalic acid, acetic acid, and organic acids; and other additives.

DETAILED DESCRIPTION - An etching solution comprises an oxidizing reactant from ammonium persulfate, **hydrogen peroxide**, and nitric acid; a co-reactant from **phosphoric acid**, sulfuric acid, nitric acid, oxalic acid, acetic acid, and organic acids; and other additives from hydrochloric acid, aliphatic alcohols, butylated.

TECH.

ferric chloride. The species can be sodium chlorate, iron nitrate, ammonium persulfate, CuNH₄Cl₃, sodium persulfate, potassium metabisulfite, ammonium fluoride, copper **sulfate**, ammonium hydroxide, potassium hydroxide, **hydrogen peroxide**, copper nitrate, and sodium ethylenediamine tetraacetic acid (EDTA) salt of wetting agent. It can be hydrogen fluoride, nitric acid, **hydrogen peroxide**, sulfuric acid, and lactic acid. It can also be sodium hydroxide,

potassium hydroxide, ammonium hydroxide, and hydrogen peroxide. The solution. . . thorium, yttrium, tungsten, zinc, and zirconium. It also includes citric acid, erythorbic acid, and triethanolamine, in aqueous solution; sulfuric acid, **hydrogen peroxide**, molybdenum salt, phenolsulfonic acid in aqueous solution; and mineral acid, and molybdenum salt. The abrasive particles are coated with cesium oxide and are produced by a sol method. The oxidizing reactants include approximately 50-approximately 70 parts/vol. concentrated **phosphoric acid**, approximately 24-approximately 40 parts/vol. concentrated aqueous acetic acid, and approximately 3-approximately 10 parts/vol. concentrated aqueous nitric acid. The solution also includes approximately 1-approximately 15 parts/vol. concentrated. . .

L3 ANSWER 24 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2002-734212 [80] WPIDS

DNC C2002-208016

TI Purification of biogas containing hydrogen sulfide and ammonia, removes hydrogen sulfide at least partially by absorption into alkaline wash solution.

DC H06 J01

IN PIEPER, S

PA (PIEP-I) PIEPER S

CYC 1

PI DE 10119991 A1 20021024 (200280)* 7p

ADT DE 10119991 A1 DE 2001-10119991 20010423

PRAI DE 2001-10119991 20010423

AB DE 10119991 A UPAB: 20021212

NOVELTY - Hydrogen sulfide is removed at least partially by absorption using an alkaline wash solution.

DETAILED DESCRIPTION - An INDEPENDENT CLAIM is included for the corresponding plant. Preferred features: The biogas contains H₂S and/or NH₃ as impurities, and the ammonia is removed at least partially, by absorption in an acid wash solution. The two gases are removed in two separate absorption stages (K1, K2), the first removing H₂S or NH₃ and vice versa respectively. Two columns and appropriate pipework, absorbent storage vessels (7, 15), pumps (P1, P2), valves (20) and instruments (21) are included. Sodium carbonate addition is controlled from pH measurement and its consumption is a function of the hydrogen sulfide concentration in the biogas.

USE - To purify biogas.

ADVANTAGE - Directly or indirectly, the presence of hydrogen sulfide, carbon dioxide and ammonia have the potential to cause corrosion in packaged power plants, and environmentally harmful emissions. The contaminant gases are removed from biogas by the new process, in a simple, cost-effective manner, improving on earlier processes. Power generation plant life and efficiency are increased. The plant, comprising essentially scrubbers and wash fluid storage vessels, can be constructed compactly and economically. Standard containers can be used. Running costs are small, using freely available reagents.

DESCRIPTION OF DRAWING(S) - A schematic plant flow diagram is presented. (Drawing includes non-English language text)

storage vessels 7, 15

valves 20

instruments 21

separate absorption stages K1, K2

pumps P1, P2

Dwg.1/1

TECH. . .

and water. Sufficient sodium hydroxide is dosed to produce sufficient sodium carbonate. In order to convert sodium sulfide into sodium **sulfate**, **hydrogen peroxide** is added under the control of redox potential measurement. Sulfuric acid absorption of NH₃, produces ammonium **sulfate** ((NH₄)₂SO₄). Acid dosing is controlled from pH. An alkaline wash solution of potassium hydroxide (KOH) and an

acid wash solution of **phosphoric acid** (H3PO4) are employed.

L3 ANSWER 25 OF 158 WPIDS (C) 2003 THOMSON DERWENT
AN 2003-211928 [21] WPIDS
DNC C2003-054247
TI Determination and optionally regulation of accelerator concentration in zinc phosphonating solution for steel, zinc (alloy)-plated steel and aluminum (alloy) uses biochemical hydrogen peroxide sensor giving electrical signal.
DC D16 E19 E36 M14
IN KLING, H; SCHENZLE, B; SEEMANN, J
PA (HENK) HENKEL KGAA
CYC 46
PI DE 10118552 A1 20021017 (200321)* 5p
WO 2002083983 A2 20021024 (200321) DE
RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE TR
W: AU BG BR BY CA CN CZ HU ID IN JP KR MX NO NZ PL RO RU SG SI SK UA
US UZ VN YU ZA
ADT DE 10118552 A1 DE 2001-10118552 20010414; WO 2002083983 A2 WO 2002-EP3793 20020405
PRAI DE 2001-10118552 20010414
AB DE 10118552 A UPAB: 20030328
NOVELTY - Determination (M1) and optionally regulation of accelerator concentration in zinc phosphonating solution containing steel, steel plated with zinc (Zn) (alloy) and/or aluminum (alloy) with a solution containing Zn ions, phosphate ions and specified accelerator(s) (I) comprising determining the (I) concentration using a biochemical hydrogen peroxide sensor that gives an electrical signal with a magnitude, is new.
DETAILED DESCRIPTION - Phosphation of metal surfaces, comprising steel, steel plated with zinc (Zn) or Zn alloy, aluminum (Al) and/or Al alloy involves spraying or dipping for 3 seconds to 8 minutes with a solution containing 0.2-3 g/l Zn ions and 3-50 g/l phosphate ions, such that the weight ratio of phosphate ions to Zn ions is not less than 3.7, and accelerator(s) (I). (I) comprises 0.3-4 g/l chlorate ions, 0.01-0.2 g/l nitrite ions, 0.05-2 g/l m-nitrobenzene **sulfonate** ions, 0.05-2 g/l m-nitrobenzoate ions, 0.05-2 g/l p-nitrophenol, 0.005-0.15 g/l free or combined **hydrogen peroxide**, 0.1-10 g/l free or combined hydroxylamine, 0.1-10 g/l reducing sugar, 0.05-4 g/l organic N-oxide, preferably N-methylmorpholine, and/or 0.5-5 g/l organic nitro compound selected from nitroguanidine, nitroarginine, its (m)ethyl and propyl esters and nitrofurfurylidene diacetate. The novelty is that the (I) concentration is determined using a biochemical hydrogen peroxide sensor that gives an electrical signal with a magnitude from which the (I) concentration can be determined.
USE - (M1) is useful for measuring the concentration of accelerator(s) in the phosphonation solution, especially the concentration of free or combined hydrogen peroxide (H2O2), nitrite or hydroxylamine; and for automatic determination and regulation of the accelerator concentration (all claimed).
ADVANTAGE - (M1) provides a new way of determining and preferably also regulating the accelerator content in zinc phosphonation solutions.
Dwg.0/0
AB . . .
not less than 3.7, and accelerator(s) (I). (I) comprises 0.3-4 g/l chlorate ions, 0.01-0.2 g/l nitrite ions, 0.05-2 g/l m-nitrobenzene **sulfonate** ions, 0.05-2 g/l m-nitrobenzoate ions, 0.05-2 g/l p-nitrophenol, 0.005-0.15 g/l free or combined **hydrogen peroxide**, 0.1-10 g/l free or combined hydroxylamine, 0.1-10 g/l reducing sugar, 0.05-4 g/l organic N-oxide, preferably N-methylmorpholine, and/or 0.5-5 g/l organic. . .
TT TT: DETERMINE OPTION REGULATE ACCELERATE CONCENTRATE ZINC
PHOSPHONATED SOLUTION STEEL ZINC ALLOY PLATE STEEL ALUMINIUM
ALLOY BIOCHEMICAL HYDROGEN PEROXIDE SENSE ELECTRIC

SIGNAL.

L3 ANSWER 26 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2002:858210 CAPLUS
DN 138:270980
TI Environmentally benign oxidation method with hydrogen peroxide
AU Sato, Kazuhiko
CS Research Institute for Green Technology, National Institute of Advanced
Industrial Science and Technology, Higashi, Tsukuba, 305-8565, Japan
SO Yuki Gosei Kagaku Kyokaishi (2002), 60(10), 974-982
CODEN: YGKKAE; ISSN: 0037-9980
PB Yuki Gosei Kagaku Kyokai
DT Journal; General Review
LA Japanese
AB A review. The oxidn. procedure must be high-yielding and with high
selectivity without any byproducts through a simple, safe operation using
a clean, well-behaving, and cheap oxidant. Aq. hydrogen peroxide is an
ideal oxidant, because the atom efficiency is excellent and water is a
sole theor. side product. The authors here report the no-solvent oxidn.
using aq. hydrogen peroxide under entirely halide-free conditions. A
catalytic system consisting of sodium tungstate and methyltrioctylammonium
hydrogen **sulfate** effects oxidn. of secondary alcs. to ketones
and primary alcs. to carboxylic acids using 3-30% **hydrogen**
peroxide without any org. solvents. The turnover no. is two
orders of magnitude higher than any previously reported hydrogen peroxide
oxidn. Substituted benzyl alcs. are oxidized selectively to benzaldehydes
and/or benzoic acids, depending on the property of the substituents.
Addn. of (aminomethyl)phosphonic acid accelerates epoxidn. of olefins.
These oxidn. methods are high-yielding, clean; safe, operationally simple,
and cost-effective and therefore meet with the requirements of
contemporary org. synthesis. Cyclohexene is converted directly to anal.
pure, cryst. adipic acid in an excellent yield. Sulfides are oxidized to
sulfoxides or sulfones in high yield. Aldehydes are oxidized to
carboxylic acids without affecting olefinic or alc. functions.

AB A review. The oxidn. procedure must be high-yielding and with high
selectivity without any byproducts through a simple, safe operation using
a clean, well-behaving, and cheap oxidant. Aq. hydrogen peroxide is an
ideal oxidant, because the atom efficiency is excellent and water is a
sole theor. side product. The authors here report the no-solvent oxidn.
using aq. hydrogen peroxide under entirely halide-free conditions. A
catalytic system consisting of sodium tungstate and methyltrioctylammonium
hydrogen **sulfate** effects oxidn. of secondary alcs. to ketones
and primary alcs. to carboxylic acids using 3-30% **hydrogen**
peroxide without any org. solvents. The turnover no. is two
orders of magnitude higher than any previously reported hydrogen peroxide
oxidn. Substituted benzyl alcs. are oxidized selectively to benzaldehydes
and/or benzoic acids, depending on the property of the substituents.
Addn. of (aminomethyl)phosphonic acid accelerates epoxidn. of olefins.
These oxidn. methods are high-yielding, clean; safe, operationally simple,
and cost-effective and therefore meet with the requirements of
contemporary org. synthesis. Cyclohexene is converted directly to anal.
pure, cryst. adipic acid in an excellent yield. Sulfides are oxidized to
sulfoxides or sulfones in high yield. Aldehydes are oxidized to
carboxylic acids without affecting olefinic or alc. functions.

IT Alcohols, reactions
RL: RCT (Reactant); RACT (Reactant or reagent)
(primary, sodium tungstate and methyltrioctylammonium hydrogen
sulfate for oxidn. of primary alcs. to carboxylic acids;
environmentally benign no-solvent oxidn. method with **hydrogen**
peroxide under halide-free conditions)

IT Alcohols, preparation
RL: SPN (Synthetic preparation); PREP (Preparation)
(secondary, sodium tungstate and methyltrioctylammonium hydrogen
sulfate for oxidn. of secondary alcs. to ketones;

environmentally benign no-solvent oxidn. method with **hydrogen peroxide** under halide-free conditions)

IT Carboxylic acids, preparation
 RL: SPN (Synthetic preparation); PREP (Preparation)
 (sodium tungstate and methyltrioctylammonium hydrogen **sulfate** for oxidn. of primary alcs. to carboxylic acids; environmentally benign no-solvent oxidn. method with **hydrogen peroxide** under halide-free conditions)

IT Ketones, preparation
 RL: SPN (Synthetic preparation); PREP (Preparation)
 (sodium tungstate and methyltrioctylammonium hydrogen **sulfate** for oxidn. of secondary alcs. to ketones; environmentally benign no-solvent oxidn. method with **hydrogen peroxide** under halide-free conditions)

IT Oxidation catalysts
 (sodium tungstate and methyltrioctylammonium hydrogen **sulfate** for oxidn. of secondary and primary alcs. to ketones and carboxylic acids; environmentally benign no-solvent oxidn. method with **hydrogen peroxide** under halide-free conditions)

IT Epoxidation catalysts
 (sodium tungstate, methyltrioctylammonium hydrogen, and (aminomethyl) **phosphonic acid** catalysts for epoxidn. of olefins.; environmentally benign no-solvent oxidn. method with **hydrogen peroxide** under halide-free conditions)

IT Alkenes, reactions
 Epoxides
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (sodium tungstate, methyltrioctylammonium hydrogen, and (aminomethyl) **phosphonic acid** catalysts for epoxidn. of olefins.; environmentally benign no-solvent oxidn. method with **hydrogen peroxide** under halide-free conditions)

L3 ANSWER 27 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2001:923577 CAPLUS

DN 136:42518

TI Hair bleaches and dyes containing alkalies and oxidants

IN Matsuo, Takashi; Miyabe, Hajime; Shibata, Yutaka

PA Kao Corporation, Japan

SO PCT Int. Appl., 32 pp.

CODEN: PIXXD2

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2001095869	A1	20011220	WO 2001-JP4835	20010608
	W: US				
	RW: DE, FR, GB				
	JP 2001354530	A2	20011225	JP 2000-175133	20000612
	JP 2001354531	A2	20011225	JP 2000-175134	20000612
	EP 1291006	A1	20030312	EP 2001-938562	20010608
	R: DE, FR, GB				
PRAI	JP 2000-175133	A	20000612		
	JP 2000-175134	A	20000612		
	WO 2001-JP4835	W	20010608		

AB An oxidn.-type hair bleach or dye which is composed of a first lotion contg. an alkali agent and a second lotion contg. an oxidizing agent, contains the following components (A), (B), (C), and (D) in amts. described below based on the whole of the mixt. of the first lotion with the second one, and has a pH of 8 to 12: (A) 8-40 % a water-compatible org. solvent exhibiting an octanol-water partition coeff. (logP) of 0.3 or above at 25.degree.C and having a mol. wt. of .ltoreq. 200, (B) 0.1-10 % an alkali agent, (C) 0.1-12 % an oxidizing agent, hydrogen peroxide, and (D) 25-70 % water. This hair bleach or dye has a high bleaching power,

can dye the hair in a good bright color, and is lowered in the irritant stench and the irritation to the scalp. A hair bleach comprised (1) a first lotion contg. Na polyoxyethylene lauryl sulfate 15, coco fatty acid diethanolamide 40, benzyl alc. 25, ammonia water (28 %) 7, and water 13 % and (2) a second lotion contg. Na polyoxyethylene lauryl sulfate 20, coco fatty acid diethanolamide 2, H2O2 soln. (35 %) 17, phosphoric acid soln. (75 %) 0.3, and water 60.7 %.

RE.CNT 4 THERE ARE 4 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB An oxidn.-type hair bleach or dye which is composed of a first lotion contg. an alkali agent and a second lotion contg. an oxidizing agent, contains the following components (A), (B), (C), and (D) in amts. described below based on the whole of the mixt. of the first lotion with the second one, and has a pH of 8 to 12: (A) 8-40 % a water-compatible org. solvent exhibiting an octanol-water partition coeff. (logP) of 0.3 or above at 25.degree.C and having a mol. wt. of .ltoreq. 200, (B) 0.1-10 % an alkali agent, (C) 0.1-12 % an oxidizing agent, hydrogen peroxide, and (D) 25-70 % water. This hair bleach or dye has a high bleaching power, can dye the hair in a good bright color, and is lowered in the irritant stench and the irritation to the scalp. A hair bleach comprised (1) a first lotion contg. Na polyoxyethylene lauryl sulfate 15, coco fatty acid diethanolamide 40, benzyl alc. 25, ammonia water (28 %) 7, and water 13 % and (2) a second lotion contg. Na polyoxyethylene lauryl sulfate 20, coco fatty acid diethanolamide 2, H2O2 soln. (35 %) 17, phosphoric acid soln. (75 %) 0.3, and water 60.7 %.

L3 ANSWER 28 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2001:693212 CAPLUS

DN 135:244563

TI Preparation of an aqueous solution of hydrogen peroxide from hydrogen and oxygen

IN Devic, Michel

PA Atofina, Fr.

SO PCT Int. Appl., 29 pp.

CODEN: PIXXD2

DT Patent

LA French

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2001068519	A1	20010920	WO 2001-FR449	20010215
	W:				
	AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CR, CU, CZ, DE, DK, DM, DZ, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, TZ, UA, UG, US, UZ, VN, YU, ZA, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM				
	RW:				
	GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, TR, BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG				
	FR 2806399	A1	20010921	FR 2000-3438	20000317
	FR 2806399	B1	20020913		
	EP 1263680	A1	20021211	EP 2001-907809	20010215
	R:				
	AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO, MK, CY, AL, TR				
	US 2003086853	A1	20030508	US 2002-221339	20020910
PRAI	FR 2000-3438	A	20000317		
	WO 2001-FR449	W	20010215		

OS MARPAT 135:244563

AB An aq. soln. of hydrogen peroxide is prepd. by injection of hydrogen and oxygen into an aq. acidic soln. in the presence of a dispersed catalyst and a tenside. The tenside stable under acidic and oxidizing conditions has the following general formula: CnF2n+1-Q-G or CnF2n+1-G with Q being a spacer and G a hydrophilic group. The concn. of the tenside in the

reaction mixt. is 5-10 ppm to avoid foaming. An alk. metal bromide or hydrobromic acid (20-100 ppm) and bromine (2-10 ppm) are added to inhibit decompn. of the hydrogen peroxide. The catalyst consists of metals, such as palladium, platinum, ruthenium, rhodium, iridium, osmium, holmium or gold, preferably of palladium as the main component and platinum as the minor component. The bimetallic catalyst is supported on silica. The catalyst is prepd. by mixing the silica with a conc. soln. of the metal salts forming a paste, filtration and drying of the filtrate under conditions supporting slow crystn., redn. with hydrogen at 200-400 .degree.C, treatment of the solid with an acidic soln. (pH = 1-3) contg. bromide (20-100 mg/l) and bromine (2-20 mg/l) at 10-80 .degree.C, and subsequent filtration and drying at 100-140 .degree.C. The reaction for the prodn. of H2O2 is carried out in a stirred or tubular reactor at 30-60 .degree.C and 10-100 bars with a molar hydrogen/oxygen ratio < 0.0416. The aq. soln. is sepd. from the catalyst and additives by inverse osmosis using a polyamide membrane.

RE.CNT 7 THERE ARE 7 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

IT Fatty acids, uses
Sulfonic acids, uses
RL: MOA (Modifier or additive use); NUU (Other use, unclassified); USES (Uses)
(surfactant, fluorinated; prepn. of an aq. soln. of **hydrogen peroxide** from hydrogen and oxygen)

IT 7664-38-2, **Phosphoric acid**, uses
RL: NUU (Other use, unclassified); USES (Uses)
(prepn. of an aq. soln. of **hydrogen peroxide** from hydrogen and oxygen)

L3 ANSWER 29 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2001:565171 CAPLUS
DN 135:139050
TI Bleaching detergents
IN Kahre, Joerg; Elsner, Michael; Koester, Rita; Schmid, Karl Heinz
PA Cognis Deutschland G.m.b.H., Germany
SO PCT Int. Appl., 28 pp.
CODEN: PIXXD2
DT Patent
LA German
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2001055289	A1	20010802	WO 2001-EP614	20010119
	W: US				
	RW: AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, TR				
	DE 10003751	A1	20010802	DE 2000-10003751	20000128
PRAI	DE 2000-10003751	A	20000128		

OS MARPAT 135:139050
AB The title detergents contain hydroxy mixed ethers, alk(en)yl oligoglycosides, bleaching agents, and optionally, nonionic and anionic surfactants. A typical cleaning compn. contained 2-hydroxydecyl ethers of ethoxylated (22 EO) - propoxylated (1 PO) C8-10 fatty alcs. 5.0, **H2O2** (30% aq.) 25.0, octyl **sulfate** 3.0 and hydroxyethane-1,1-**diphosphonic** acid 0.3% in H2O.

RE.CNT 6 THERE ARE 6 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB The title detergents contain hydroxy mixed ethers, alk(en)yl oligoglycosides, bleaching agents, and optionally, nonionic and anionic surfactants. A typical cleaning compn. contained 2-hydroxydecyl ethers of ethoxylated (22 EO) - propoxylated (1 PO) C8-10 fatty alcs. 5.0, **H2O2** (30% aq.) 25.0, octyl **sulfate** 3.0 and hydroxyethane-1,1-**diphosphonic** acid 0.3% in H2O.

L3 ANSWER 30 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2001:592261 CAPLUS
DN 135:137299
TI Regioselective oxidation process and catalyst system for the preparation
of 2,5-dichlorophenol from 1,4-dichlorobenzene
IN Henrick, Clive A.; Scheuerman, Randall A.
PA Syngenta Participations, Switz.
SO U.S., 6 pp.
CODEN: USXXAM
DT Patent
LA English
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 6274776	B1	20010814	US 1995-570837	19951212
PRAI	US 1995-570837		19951212		
OS	CASREACT 135:137299				

AB 2,5-Dichlorophenol, an intermediate in the manuf. of the herbicide dicamba, is prepd. by the regioselective oxidn. of 1,4-dichlorobenzene using a peroxo-, hydroperoxo-, superoxo-, or alkylperoxo-metal species in the presence of an .alpha.-hydroxy-, di-, tribasic carboxylic acid, or **sulfonic acid**, and **hydrogen peroxide**.

RE.CNT 4 THERE ARE 4 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB 2,5-Dichlorophenol, an intermediate in the manuf. of the herbicide dicamba, is prepd. by the regioselective oxidn. of 1,4-dichlorobenzene using a peroxo-, hydroperoxo-, superoxo-, or alkylperoxo-metal species in the presence of an .alpha.-hydroxy-, di-, tribasic carboxylic acid, or **sulfonic acid**, and **hydrogen peroxide**.

IT 13598-36-2, **phosphonic acid**

RL: CAT (Catalyst use); USES (Uses)

(alkyl-; regioselective oxidn. catalysts with metal-oxo species in the presence of **hydrogen peroxide** for the prepn. of 2,5-dichlorophenol from 1,4-dichlorobenzene)

L3 ANSWER 31 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2001:111500 CAPLUS
DN 134:164871
TI Stable compositions for removing stains from fabrics and carpets
IN Micciche, Robert P.; Lynch, Ann Marie; Tripathi, Uma; Belansky, Carol
PA Playtex Products, Inc., USA
SO U.S., 5 pp.
CODEN: USXXAM
DT Patent
LA English
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 6187738	B1	20010213	US 1999-243086	19990202
PRAI	US 1998-73447P	P	19980202		

AB A compn. inhibiting the resoiling of fabrics and carpets includes a peroxygen compd., a surfactant, and a polymeric or copolymeric soil resist. Thus, a compn. contained water 94.77, **H2O2** 3.00, an acrylate copolymer 0.60, Na lauryl **sulfate** 0.60, Pareth 7 (alc. ethoxylate) 0.25, Na lauroyl sarcosinate 0.23, lauramine oxide 0.07, a fragrance 0.15, **Dequest** 2010 0.25, and Surcide P [hexahydro-1,3,5-tris(2-hydroxyethyl)-s-triazine] 0.08 %.

RE.CNT 51 THERE ARE 51 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB A compn. inhibiting the resoiling of fabrics and carpets includes a peroxygen compd., a surfactant, and a polymeric or copolymeric soil resist. Thus, a compn. contained water 94.77, **H2O2** 3.00, an acrylate copolymer 0.60, Na lauryl **sulfate** 0.60, Pareth 7 (alc. ethoxylate) 0.25, Na lauroyl sarcosinate 0.23, lauramine oxide 0.07, a

fragrance 0.15, **Dequest** 2010 0.25, and Surcide P
[hexahydro-1,3,5-tris(2-hydroxyethyl)-s-triazine] 0.08 %.

L3 ANSWER 32 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2001:847374 CAPLUS
DN 135:372556
TI Aldehyde stripping agents for aminoplast or phenolic resin adhesives
IN Ohama, Chiaki
PA Nankyo Efnica K. K., Japan
SO Jpn. Kokai Tokkyo Koho, 5 pp.
CODEN: JKXXAF
DT Patent
LA Japanese
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 2001323240	A2	20011122	JP 2000-141452	20000515
PRAI	JP 2000-141452		20000515		

AB The title agents, useful for mixing in the title adhesives (e.g., Oshika Resin 5905) to eliminate aldehyde odor, contain **H2O2**, ammonia generating materials (e.g., ammonium **sulfate**, guanidine phosphate), and optionally **H2O2** stabilizers (e.g., Na pyrophosphate, H3PO4), polyamines, **H2O2** decomp. accelerators (e.g., NH3), and/or urotropin.

AB The title agents, useful for mixing in the title adhesives (e.g., Oshika Resin 5905) to eliminate aldehyde odor, contain **H2O2**, ammonia generating materials (e.g., ammonium **sulfate**, guanidine phosphate), and optionally **H2O2** stabilizers (e.g., Na pyrophosphate, H3PO4), polyamines, **H2O2** decomp. accelerators (e.g., NH3), and/or urotropin.

IT 7664-38-2, **Phosphoric acid**, uses 7722-88-5, Sodium pyrophosphate
RL: MOA (Modifier or additive use); TEM (Technical or engineered material use); USES (Uses)
(**hydrogen peroxide** stabilizers; aldehyde stripping agents for aminoplast or phenolic resin adhesives)

L3 ANSWER 33 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2001:717176 CAPLUS
DN 135:277385
TI Polluted soil remediation method using hydrogen peroxide
IN Osumi, Shoujiro; Tsubota, Jun; Shinohara, Masafumi
PA Japan
SO Jpn. Kokai Tokkyo Koho, 8 pp.
CODEN: JKXXAF
DT Patent
LA Japanese
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 2001269657	A2	20011002	JP 2000-88464	20000328
PRAI	JP 2000-88464		20000328		

AB Polluted soil is remediated by decomp. org. matter with H2O2 in the presence of catalyst of a metal ion possible to be have 2 different valence values and the remediation involves a step of adding an assisting agent capable of providing a cation easier to be adsorbed in the soil by ion exchange or to cause reaction with H3PO4 more easily than the catalyst before treatment with H2O2. Alternatively, different agents may be added as the assisting agent for adsorption and reaction, resp. Hardly decomposable org. substances, e.g., polycyclic arom. hydrocarbons and org. substances with high cation exchange capacity are highly efficiently decompd. to remediate soil polluted with these substances.

IT 7720-78-7, ferrous **sulfate**
RL: CAT (Catalyst use); USES (Uses)

(catalyst; soil remediation by **hydrogen peroxide** in the presence of catalyst using assisting agent)

IT 50-32-8, Benzo[a]pyrene, processes 7664-38-2, **Phosphoric acid**, processes
RL: PEP (Physical, engineering or chemical process); POL (Pollutant); REM (Removal or disposal); OCCU (Occurrence); PROC (Process)
(pollutants; soil remediation by **hydrogen peroxide** in the presence of catalyst using assisting agent)

L3 ANSWER 34 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2001:632218 CAPLUS

DN 135:186872

TI Fuel cell type electrochemical reactors and operation of the reactors for hydrogen peroxide manufacture

IN Otsuka, Kiyoshi; Yamanaka, Ichiro

PA Asahi Chemical Industry Co., Ltd., Japan

SO Jpn. Kokai Tokkyo Koho, 8 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 2001236968	A2	20010831	JP 2000-46114	20000223
PRAI	JP 2000-46114		20000223		

AB The reactors have an electrolyte in a chamber between a pos. electrode chamber and a neg. electrode chamber, and an electron conductor forming an external short circuit between the electrodes. H2O2 is prepd. by using the reactors by supplying a H source and O to the neg. and pos. electrode chambers, resp., to obtain H2O2 in the electrolyte chamber.

IT 1310-73-2, Sodium hydroxide, uses 7447-41-8, Lithium chloride, uses 7488-54-2, Rubidium **sulfate** 7601-90-3, Perchloric acid, uses 7647-01-0, Hydrochloric acid, uses 7647-14-5, Sodium chloride, uses 7664-38-2, **Phosphoric acid**, uses 7664-93-9, Sulfuric acid, uses 7757-82-6, Sodium **sulfate**, uses 7778-80-5, Potassium **sulfate**, uses 10294-54-9, Cesium **sulfate** 10377-48-7, Lithium **sulfate**

RL: NUU (Other use, unclassified); USES (Uses)

(electrolytes in fuel cell type electrochem. reactors for manuf. of **hydrogen peroxide**)

L3 ANSWER 35 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2001:255075 CAPLUS

DN 134:283923

TI Manufacture of inorganic-coated galvanized steel sheets with good press formability

IN Osugi, Yukio; Nakagishi, Tetsuyuki; Kawanishi, Yoshihiro; Kuroda, Toru

PA Sumitomo Metal Industries, Ltd., Japan

SO Jpn. Kokai Tokkyo Koho, 9 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 2001098383	A2	20010410	JP 1999-273151	19990927
PRAI	JP 1999-273151		19990927		

AB Aq. solns. (pH 1-7) contg. phosphoric acids and H2O2 at **H2O2/phosphoric acid** wt. ratios of 0.1-1 are applied on galvanized steel sheets pretreated with aq. solns. contg. alkalis and/or acids and baked at 50-300.degree. without washing with water to form inorg. films contg. 10-500 mg/m2 P to give galvanized steel sheets showing good press formability, chem. conversion coating formability, weldability, and adhesion properties. Inorg. films having

P-O-P network structures are formed.

AB Aq. solns. (pH 1-7) contg. phosphoric acids and H2O2 at H2O2/phosphoric acid wt. ratios of 0.1-1 are applied on galvanized steel sheets pretreated with aq. solns. contg. alkalis and/or acids and baked at 50-300.degree. without washing with water to form inorg. films contg. 10-500 mg/m2 P to give galvanized steel sheets showing good press formability, chem. conversion coating formability, weldability, and adhesion properties. Inorg. films having P-O-P network structures are formed.

IT Coating process
(coating of galvanized steel sheets with **phosphoric acid-H2O2** mixts. for good press formability)

IT Galvanized steel
RL: PEP (Physical, engineering or chemical process); TEM (Technical or engineered material use); PROC (Process); USES (Uses)
(electrogalvanized; coating of galvanized steel sheets with **phosphoric acid-H2O2** mixts. for good press formability)

IT Galvanized steel
RL: PEP (Physical, engineering or chemical process); TEM (Technical or engineered material use); PROC (Process); USES (Uses)
(hot-dip; coating of galvanized steel sheets with **phosphoric acid-H2O2** mixts. for good press formability)

IT 2466-09-3, Pyrophosphoric acid 7664-38-2, **Phosphoric acid**, processes 10380-08-2, Tripolyphosphoric acid
RL: PEP (Physical, engineering or chemical process); PROC (Process)
(coating of galvanized steel sheets with **phosphoric acid-H2O2** mixts. for good press formability)

IT 12727-73-0, processes
RL: PEP (Physical, engineering or chemical process); TEM (Technical or engineered material use); PROC (Process); USES (Uses)
(coating of galvanized steel sheets with **phosphoric acid-H2O2** mixts. for good press formability)

IT 1310-73-2, Sodium hydroxide, processes 7664-93-9, Sulfuric acid, processes 7733-02-0, Zinc **sulfate** 10124-49-9, Iron **sulfate**
RL: PEP (Physical, engineering or chemical process); PROC (Process)
(for pretreatment; coating of galvanized steel sheets with **phosphoric acid-H2O2** mixts. for good press formability)

L3 ANSWER 36 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2001:21329 CAPLUS

DN 134:90884

TI Hair dyes comprising light isoparaffins in the first preparation and oxidants in the second preparation

IN Takemoto, Yuka; Yonetani, Yoshinobu; Okamoto, Yoshihiro

PA Yamahatsu Sangyo Kaisha, Ltd., Japan

SO Jpn. Kokai Tokkyo Koho, 14 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 2001002538	A2	20010109	JP 1999-178528	19990624
PRAI	JP 1999-178528		19990624		

AB This present invention relates to hair dye compns. comprising light isoparaffins and monoisopropanolamine in the first agent and oxidants in the second agent. These compns. do not have an irritating odor and do not damage the hair. A hair dye first prepn. contained light isoparaffins 3, monoisopropanolamine 5, Na polyoxyethylene lauryl sulfate 20, coco fatty acid diethanolamide 2, oleic acid 5, propylene glycol 3, p-phenylenediamine 1, p-aminophenol 0.5, TGAA 0.4, disodium EDTA 0.2,

perfumes 0.3, and distd. water q.s. to 100 %. A second prepn. contained sodium polyoxyethylene lauryl sulfate 2, cetanol 1.5 phosphoric acid 0.0002, H2O2 (35 % soln.) 17, and distd. water q.s. to 100 %. A mixt. of the first and second prepn. in an equal amt. provided excellent hair-dyeing ability and hair feels without offensive odors.

AB This present invention relates to hair dye compns. comprising light isoparaffins and monoisopropanolamine in the first agent and oxidants in the second agent. These compns. do not have an irritating odor and do not damage the hair. A hair dye first prepn. contained light isoparaffins 3, monoisopropanolamine 5, Na polyoxyethylene lauryl sulfate 20, coco fatty acid diethanolamide 2, oleic acid 5, propylene glycol 3, p-phenylenediamine 1, p-aminophenol 0.5, TGAA 0.4, disodium EDTA 0.2, perfumes 0.3, and distd. water q.s. to 100 %. A second prepn. contained sodium polyoxyethylene lauryl sulfate 2, cetanol 1.5 phosphoric acid 0.0002, H2O2 (35 % soln.) 17, and distd. water q.s. to 100 %. A mixt. of the first and second prepn. in an equal amt. provided excellent hair-dyeing ability and hair feels without offensive odors.

L3 ANSWER 37 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2001:885371 CAPLUS

DN 136:8591

TI Supported palladium-platinum catalysts for direct synthesis of hydrogen peroxide from hydrogen and oxygen

IN Paparatto, Giuseppe; D'Aloisio, Rino; De Alberti, Giordano; Buzzoni, Roberto

PA Eni S.p.A., Italy; Enichem S.p.A.

SO Eur. Pat. Appl., 13 pp.

CODEN: EPXXDW

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 1160196	A1	20011205	EP 2001-201971	20010525
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO				
	IT 2000MI1219	A1	20011203	IT 2000-MI1219	20000601
	JP 2002029711	A2	20020129	JP 2001-164088	20010531
	US 2002028174	A1	20020307	US 2001-870826	20010601
PRAI	IT 2000-MI1219	A	20000601		

AB Direct synthesis of hydrogen peroxide from hydrogen and oxygen is carried out in a solvent contg. a halogenated promoter and an acid promoter in the presence of a supported bimetallic catalyst selected from platinum-group metals, esp. platinum and palladium. Palladium and platinum are present at 0.01-5 wt.% and 0.01-1 wt.% amts., resp., corresponding to a Pt-Pd at. ratio of 0.1-50:50-99.9, with preferred catalyst compns. of 0.4-2 wt.% Pd, 0.05-0.5 wt.% Pt, and a 1-30:70-99 at. ratio. The catalyst supports are selected from low-ash activated carbon, activated sulfonated carbon, silica, alumina, silica-alumina, and zeolites. with a catalyst surface area >300 m2/g. Suitable promoters include: (1) compds. that generate bromide ion (e.g., HBr, NaBr, and KBr), (2) acids (e.g., sulfuric acid, phosphoric acid, nitric acid, sulfonic acids, etc.). Hydrogen peroxide synthesis is carried out at 2-50.degree. and 50-300 atm. pressure, at 1:3-15 H2-O2 molar ratio (and <4.5% molar H2 concn.) in an inert gas. Hydrogen peroxide is typically produced at a 6.2-6.6 wt.% dil. concn. in a solvent (e.g., MeOH), corresponding to a molar selectivity of 70-73%, and can be used directly after synthesis as the dil. soln. for an oxidn. process (esp. for olefins, arom. hydrocarbons, ammonia, and carbonyl compds., using titanium silicate catalyst).

RE.CNT 9 THERE ARE 9 CITED REFERENCES AVAILABLE FOR THIS RECORD

ALL CITATIONS AVAILABLE IN THE RE FORMAT

IT Alkali metal bromides

Halides

Sulfonic acids, uses

RL: CAT (Catalyst use); USES (Uses)

(promoter; supported palladium-platinum catalysts for direct synthesis of **hydrogen peroxide** from hydrogen and oxygen)

IT 7647-15-6, Sodium bromide, uses 7664-38-2, **Phosphoric acid**, uses 7664-93-9, Sulfuric acid, uses 7697-37-2, Nitric acid, uses 7758-02-3, Potassium bromide, uses 7789-38-0, Sodium bromate 10035-10-6, Hydrogen bromide, uses 12124-97-9, Ammonium bromide 24959-67-9, Bromide, uses
RL: CAT (Catalyst use); USES (Uses)
(promoter; supported palladium-platinum catalysts for direct synthesis of **hydrogen peroxide** from hydrogen and oxygen)

L3 ANSWER 38 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2001:885369 CAPLUS

DN 136:8590

TI Continuous prodn. of dilute solutions of hydrogen peroxide by direct synthesis from hydrogen and oxygen over supported palladium-platinum catalysts

IN Paparatto, Giuseppe; Rivetti, Franco; Andrigo, Pietro; De Alberti, Giordano

PA Eni S.p.A., Italy; Enichem S.p.A.

SO Eur. Pat. Appl., 12 pp.

CODEN: EPXXDW

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 1160195	A1	20011205	EP 2001-201961	20010525
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO				
	IT 2000MI1218	A1	20011203	IT 2000-MI1218	20000601
	SG 89399	A1	20020618	SG 2001-3242	20010530
	JP 2002029712	A2	20020129	JP 2001-164132	20010531
	US 2002025293	A1	20020228	US 2001-870494	20010601
PRAI	IT 2000-MI1218	A	20000601		

AB Continuous prodn. of alc. or aq. alc. solns. of hydrogen peroxide are prepd. by: (1) providing a reactor contg. a supported platinum-palladium catalyst dispersed in a reaction solvent, in the presence of an acid and a halogen-contg. promoter, (2) feeding a H₂-O₂-inert gas (e.g., N₂) into the reactor, in which the H₂ and O₂ concns. are <4.5 vol.% and <21 vol.% (with the balance being the inert gas), (3) reacting the gas feedstock to produce a dil. (1-10 wt.%) H₂O₂ product, and (4) removing the liq. product and a gaseous stream composed essentially of unreacted H₂ and O₂ as well as the inert gas. The palladium and platinum catalysts are present at 0.4-2:0.01-1 wt.% ratio of Pd-Pt, corresponding to a 1-20:100-200 Pd-Pt at. ratio, optionally in the presence of another Group VIIIB and Group IB elements. Suitable supports include a low-ash charcoal or carbon, SiO₂, Al₂O₃, SiO₂-Al₂O₃, zeolites, with a surface area >600 mL/g. Hydrogen peroxide synthesis is carried out at 0-40.degree. and 40-150 bars, in a reaction solvent (e.g., MeOH). Suitable promoters include: (1) compds. that generate bromide ion (e.g., HBr, NaBr, and KBr), (2) acids (e.g., sulfuric acid, phosphoric acid, nitric acid, sulfonic acids, etc.). The dil. hydrogen peroxide soln., typically produced at a 6.8-7.3 wt.% concn. in a solvent (e.g., MeOH), corresponding to a molar selectivity of 67-74%, can be used directly after synthesis for an oxidn. process (esp. for olefins, arom. hydrocarbons, ammonia, and carbonyl compds., using titanium silicate catalyst).

RE.CNT 5 THERE ARE 5 CITED REFERENCES AVAILABLE FOR THIS RECORD

ALL CITATIONS AVAILABLE IN THE RE FORMAT

IT Alkali metal bromides

Sulfonic acids, uses

RL: CAT (Catalyst use); USES (Uses)
(promoter; continuous prodn. of dil. solns. of **hydrogen peroxide** by direct synthesis from hydrogen and oxygen over supported palladium-platinum catalysts)

IT 7664-38-2, **Phosphoric acid**, uses 7664-93-9, Sulfuric acid, uses 7697-37-2, Nitric acid, uses 10035-10-6, Hydrogen bromide, uses 24959-67-9, Bromide, uses

RL: CAT (Catalyst use); USES (Uses)
(promoter; continuous prodn. of dil. solns. of **hydrogen peroxide** by direct synthesis from hydrogen and oxygen over supported palladium-platinum catalysts)

L3 ANSWER 39 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2001:729698 CAPLUS

DN 135:277720

TI Oxidative hair dye compositions

IN Lorenz, Heribert

PA Goldwell G.m.b.H., Germany

SO Eur. Pat. Appl., 6 pp.

CODEN: EPXXDW

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 1138316	A2	20011004	EP 2001-106800	20010319
	EP 1138316	A3	20020213		
	EP 1138316	B1	20030604		
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO				
	DE 10016497	A1	20011018	DE 2000-10016497	20000401
	US 2001037531	A1	20011108	US 2001-809931	20010316
PRAI	DE 2000-10016497	A	20000401		

AB Oxidative hair dye compns. contain a developer and a coupler substance and 1 or more metal compds. and have an alk. pH of 7.0-12.0. The compn. addnl. comprise H2O2 and 1-hydroxyethane-1,1-**diphosphonic acid** or its alkali metal or ammonium salt. Thus, an oxidative hair dye preformulation contained cetylstearyl alc. 10.00, coco fatty acid monoethanolamide 2.00, stearic acid monoethanolamide 2.00, stearic acid diethanolamide 1.00, p-toluenediamine sulfate 0.50, p-amino-o-cresol 0.40, p-aminophenol 0.10, monoethanolamine 6.20, NH4Cl 0.20, sodium lauryl sulfate 0.30, CuCl2 0.0004, sodium sulfite 0.25, EDTA 0.20, perfume 0.20, and water to 100%. This compn. was mixed in a 1:1 ratio with the compn. comprising H2O2 6.0, cetylstearyl alc. 1.80, sodium lauryl sulfate 0.20, 1-hydroxyethane-1,1-**diphosphonic acid** 0.30, salicylic acid 0.10, and water to 100.0%.

AB Oxidative hair dye compns. contain a developer and a coupler substance and 1 or more metal compds. and have an alk. pH of 7.0-12.0. The compn. addnl. comprise H2O2 and 1-hydroxyethane-1,1-**diphosphonic acid** or its alkali metal or ammonium salt. Thus, an oxidative hair dye preformulation contained cetylstearyl alc. 10.00, coco fatty acid monoethanolamide 2.00, stearic acid monoethanolamide 2.00, stearic acid diethanolamide 1.00, p-toluenediamine sulfate 0.50, p-amino-o-cresol 0.40, p-aminophenol 0.10, monoethanolamine 6.20, NH4Cl 0.20, sodium lauryl sulfate 0.30, CuCl2 0.0004, sodium sulfite 0.25, EDTA 0.20, perfume 0.20, and water to 100%. This compn. was mixed in a 1:1 ratio with the compn. comprising H2O2 6.0, cetylstearyl alc. 1.80, sodium lauryl sulfate 0.20, 1-hydroxyethane-1,1-**diphosphonic acid** 0.30, salicylic acid 0.10, and water to 100.0%.

L3 ANSWER 40 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2001:78011 CAPLUS

DN 134:133280

TI Manufacture of epoxides by epoxidation of olefins with hydrogen peroxide

IN Paparatto, Giuseppe; Forlin, Anna; Tegon, Paolo
 PA Enichem S.p.A., Italy
 SO Eur. Pat. Appl., 11 pp.
 CODEN: EPXXDW
 DT Patent
 LA English
 FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 1072600	A1	20010131	EP 2000-202658	20000724
	EP 1072600	B1	20030402		
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO				
	IT 99MI1658	A1	20010129	IT 1999-MI1658	19990727
	AT 236140	E	20030415	AT 2000-202658	20000724
	JP 2001072673	A2	20010321	JP 2000-225897	20000726
	US 6300506	B1	20011009	US 2000-626456	20000726
PRAI	IT 1999-MI1658	A	19990727		

OS MARPAT 134:133280

AB A continuous process for the prepn. of olefin oxides comprises direct epoxidn. of an olefin with hydrogen peroxide or compds. capable of producing hydrogen peroxide under the reaction conditions in a solvent in the presence of a catalytic system consisting of a zeolite contg. titanium and a buffer system with a pH controlled within the values of 5.5 to 8.0 and consisting of a nitrogen-contg. base and a salt thereof with an org. or inorg. acid. The process allows high conversions and selectivities of the olefin into the corresponding oxide with a catalytic activity stable over a period of time.

RE.CNT 6 THERE ARE 6 CITED REFERENCES AVAILABLE FOR THIS RECORD
 ALL CITATIONS AVAILABLE IN THE RE FORMAT

IT 50-21-5D, Lactic acid, amine salts 64-17-5, Ethanol, uses 64-18-6D, Formic acid, amine salts, uses 64-19-7D, Acetic acid, amine salts, uses 67-56-1, Methanol, uses 67-63-0, Isopropyl alcohol, uses 67-64-1, Acetone, uses 74-89-5, Methylamine, uses 75-04-7, Ethylamine, uses 75-50-3, Trimethylamine, uses 75-65-0, tert-Butyl alcohol, uses 78-93-3, Methyl ethyl ketone, uses 79-09-4D, Propionic acid, amine salts 79-14-1D, Glycolic acid, amine salts 98-86-2, Acetophenone, uses 102-71-6, Triethanolamine, uses 107-10-8, n-Propylamine, uses 107-92-6D, Butyric acid, amine salts 108-93-0, Cyclohexanol, uses 109-99-9, Tetrahydrofuran, uses 111-42-2, Diethanolamine, uses 124-40-3, Dimethylamine, uses 141-43-5, Ethanolamine, uses 142-96-1, Butyl ether 540-69-2, Ammonium formate 631-61-8, Ammonium acetate 7664-38-2D, **Phosphoric acid**, amine salts, uses 7664-41-7, Ammonia, uses 7664-93-9D, Sulfuric acid, amine salts, uses 7783-20-2, Ammonium **sulfate**, uses

RL: NUU (Other use, unclassified); USES (Uses)
 (manuf. of epoxides by epoxidn. of olefins with **hydrogen peroxide**)

L3 ANSWER 41 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2002-106046 [14] WPIDS

DNC C2002-032498

TI Aqueous liquid composition of matter used for conversion coating on metal surface, comprises zirconium compounds, inorganically bonded fluorine, organic polymer molecules, and peroxy compounds.

DC A82 G02 M13 M14

IN FICK, T H; MCKENZIE, C D; ROCHFORD, G L

PA (HENK) HENKEL CORP

CYC 95

PI WO 2001083850 A1 20011108 (200214)* EN 22p

RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
 NL OA PT SD SE SL SZ TR TZ UG ZW

W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CO CR CU CZ DE DK
 DM DZ EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ

LC LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ PL PT RO RU SD
SE SG SI SK SL TJ TM TR TT TZ UA UG US UZ VN YU ZA ZW

AU 2001057492 A 20011112 (200222)

ADT WO 2001083850 A1 WO 2001-US14135 20010502; AU 2001057492 A AU 2001-57492
20010502

FDT AU 2001057492 A Based on WO 200183850

PRAI US 2000-201100P 20000502

AB WO 200183850 A UPAB: 20020301

NOVELTY - An aqueous liquid composition of matter comprises component of transition metal compounds that contains zirconium and/or hafnium; inorganically bonded fluorine; organic polymer molecules; and peroxy compounds.

DETAILED DESCRIPTION - An aqueous liquid composition of matter comprises (A) a component of transition metal compounds that contains zirconium (Zr), hafnium, or both and also includes titanium (Ti), provided that Zr constitutes at least 30% of the moles of Zr and Ti; (B) a component of compound(s) that contains inorganically bonded fluorine and is not part of component (A); (C) a component of organic polymer molecules; and (D) a component of molecules, other than oxygen gas, that contain oxygen-atom-to-oxygen-atom chemical bonds.

INDEPENDENT CLAIMS are also included for (A) a process for forming a transparent conversion coating on a metal surface comprising contacting the metal surface with an aqueous liquid composition of matter; (B) a metal article having a surface; and (C) a method of making an aqueous liquid composition of matter for forming on a metal surface with which the aqueous liquid composition is contacted a transparent conversion coating comprising dissolving in water at least (A) component of transition metal compounds that contains zirconium (Zr), hafnium, or both and also includes titanium (Ti), provided that Zr constitutes at least 30% of the moles of Zr and Ti; (B) a component of compound(s) that contains inorganically bonded fluorine and is not part of component (A); (C) a component of organic polymer molecules; and (D) a component of molecules, other than oxygen gas, that contain oxygen-atom-to-oxygen-atom chemical bonds.

USE - Used for conversion coating on metal surface.

ADVANTAGE - The conversion coatings formed had improved heat stability compared to conventional conversion coatings.

Dwg.0/0

TECH.

soluble neutral or acid salt of HF, tetrafluoroboric acid or salt of tetrafluoroboric acid. The polar moiety can be carboxylate, **phosphonate, sulfate, amino, nitrogen or phosphino** phosphorus. Component (D) comprises **hydrogen peroxide**. Preferred Composition: The concentration of component (D) in g/l, measured as its stoichiometric equivalent as H₂O₂, has a ratio to the concentration of component (C) in g/l that is at least 1:1 and not more than 35:1.

L3 ANSWER 42 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-549950 [61] WPIDS

DNN N2001-408535 DNC C2001-163661

TI New formulation for use in chemiluminescent assays and detection procedures.

DC B02 B03 B05 B07 D16 S03

IN GIRI, B P

PA (GIRI-I) GIRI B P

CYC 95

PI WO 2001055446 A1 20010802 (200161)* EN 43p

RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
NL OA PT SD SE SL SZ TR TZ UG ZW

W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CR CU CZ DE DK DM
DZ EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ LC
LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ PL PT RO RU SD SE
SG SI SK SL TJ TM TR TT TZ UA UG UZ VN YU ZA ZW

AU 2001034612 A 20010807 (200174)

US 2001046688 A1 20011129 (200202)

EP 1254252 A1 20021106 (200281) EN

R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT
RO SE SI TR

ADT WO 2001055446 A1 WO 2001-US2779 20010126; AU 2001034612 A AU 2001-34612
20010126; US 2001046688 A1 Provisional US 2000-178626P 20000128, US
2001-770592 20010126; EP 1254252 A1 EP 2001-906739 20010126, WO
2001-US2779 20010126

FDT AU 2001034612 A Based on WO 200155446; EP 1254252 A1 Based on WO 200155446

PRAI US 2000-178626P 20000128; US 2001-770592 20010126

AB WO 200155446 A UPAB: 20011024

NOVELTY - A stabilized substrate formulation, (F), for use in
chemiluminescent assays and detection procedures comprising at least one
of a chemiluminescent organic compound, an oxidizing agent, a stabilizing
agent and one suitable buffer, where the formulation having between pH 5.5
- 12.0, is new.

DETAILED DESCRIPTION - INDEPENDENT CLAIMS are include for the
following:

- (1) conducting chemiluminescent detection; and
- (2) a kit for use in luminescent assays comprising:
 - (a) an organic chemiluminescent compound;
 - (b) an enhancer;
- (c) an oxidant;
 - (d) a stabilizing agent;
 - (e) a solubilizing agent; and
 - (f) a buffer. The substrate pH is between 5.5-12.0.

USE - The formulation is useful to create novel stabilized
chemiluminescent formulations.

ADVANTAGE - Chemiluminescence generates cold light by direct
transformation of chemicals into electronic energy.

Dwg.0/7

TECH.

cetyltrimethyl ammonium chloride, cetyldimethylethyl ammonium bromide,
Bacitracin, BSA, KLH, HSA, Trypsin inhibitor, polyethylene glycol,
polymeric phosphonium salts, polymeric ammonium salts, **DEQUEST**
2060s, Coenzyme A, inorganic pyrophosphate, cytidine nucleotides,
ethylenediaminetetraacetic acid (EDTA), diethylenetriaminepentaacetic
acid, ethylenebis(oxyethylenenitrolo) tetraacetic acid and other related
macromolecules, surfactants and polymers. The oxidizing agent is selected
from the group consisting of **hydrogen peroxide**, urea
hydrogen peroxide, sodium carbonate **hydrogen**
peroxide, a perborate salt and mixtures of these. F comprises
between:

(a) 0.1 microM - 200 mM, volume to volume, (v/v) of. . . an in situ
peroxide forming system, preferably the oxidising agent is a mixture of
D-glucose and glucose oxidase which produces **hydrogen**
peroxide in situ. The stabilizing agent is selected from
polyethylene branched nonylphenyl ether, lauryl **sulfate**,
domiphen bromide, CTAB, EDTA and mixtures of these. Preferred Method: (1)
comprises:

(a) providing a probe which reacts with a substrate. . .

L3 ANSWER 43 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-580850 [65] WPIDS

DNN N2001-432624 DNC C2001-172142

TI Composition for planarizing the surfaces of a substrate such as a
semiconductor wafer comprises a liquid carrier, a chemical accelerator,
and solids comprising fumed metal oxide and abrasive particles.

DC E16 E37 L03 P61 U11

IN DIRKSEN, J A; FANG, M; MUELLER, B L

PA (CABO) CABOT MICROELECTRONICS CORP

CYC 95

PI WO 2001036554 A1 20010525 (200165)* EN 26p

RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ

NL OA PT SD SE SL SZ TR TZ UG ZW
W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CR CU CZ DE DK DM
DZ EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ LC
LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ PL PT RO RU SD SE
SG SI SK SL TJ TM TR TT TZ UA UG UZ VN YU ZA ZW

AU 2001016600 A 20010530 (200165)

US 6293848 B1 20010925 (200165)

EP 1250390 A1 20021023 (200277) EN

R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT
RO SE SI TR

KR 2002056913 A 20020710 (200304)

US 6527817 B1 20030304 (200320)

JP 2003514949 W 20030422 (200336) 31p

CN 1399668 A 20030226 (200337)

ADT WO 2001036554 A1 WO 2000-US31653 20001115; AU 2001016600 A AU 2001-16600
20001115; US 6293848 B1 US 1999-440401 19991115; EP 1250390 A1 EP
2000-979195 20001115; WO 2000-US31653 20001115; KR 2002056913 A KR
2002-706172 20020514; US 6527817 B1 CIP of US 1999-440401 19991115, US
2000-625142 20000725; JP 2003514949 W WO 2000-US31653 20001115, JP
2001-539035 20001115; CN 1399668 A CN 2000-815722 20001115

FDT AU 2001016600 A Based on WO 200136554; EP 1250390 A1 Based on WO
200136554; JP 2003514949 W Based on WO 200136554

PRAI US 2000-625142 20000725; US 1999-440401 19991115

AB WO 200136554 A UPAB: 20011108

NOVELTY - Composition for planarizing or polishing a surface comprises a
liquid carrier, a chemical accelerator, and solids having a packing
density of at least 0.1 and comprising 5-90 wt. % fumed metal oxide and
10-95 wt. % abrasive particles. About 90% of the number of abrasive
particles have particle size of at most 100 nm.

DETAILED DESCRIPTION - An INDEPENDENT CLAIM is given for a method of
planarizing or polishing a surface, comprising contacting the surface with
the above composition, where the surface is a (preferably,
nickel-phosphorus) memory or rigid disk surface.

USE - Polishing of various surfaces, especially a semiconductor or
metal layer of a memory or rigid disk (claimed), for example, and in the
manufacture of integrated circuits and semiconductors.

ADVANTAGE - Good planarization efficiency, uniformity, removal rate
and low defectivity, i.e. reduced surface imperfections and damage to
underlying structures, during the polishing and planarization of
substrates.

Dwg. 0/0

TECH.

carrier is preferably water.

The chemical accelerator is a sulfate, a persulfate, or a nitrate, and is
preferably selected from ammonium sulfate, iron (III) nitrate,
and hydroxylamine nitrate.

The composition can also include an oxidizing agent (preferably,
hydrogen peroxide) and a complexing agent selected from
an amine-containing compound (preferably, glycine), a source of phosphate
ions, a source of phosphonate ions, a carboxylate, and their
mixtures.

Desirable abrasive particles are metal oxides such as alumina, silica,
titania, ceria, zirconia and.

L3 ANSWER 44 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-211364 [21] WPIDS

DNC C2001-062918

TI Production of papermaking pulp from nonwood fiber source material by
extracting source material with alkaline solution, reducing pH of source
material, treating source material with ozone, and bleaching source
material.

DC F09

IN BYRD, M V; HURTER, R W

PA (HURT-N) HURTERCONSULT INC; (UYN-C) UNIV NORTH CAROLINA STATE

CYC 95
PI WO 2001016423 A1 20010308 (200121)* EN 55p
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
NL OA PT SD SE SL SZ TZ UG ZW
W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CR CU CZ DE DK DM
DZ EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ LC
LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ PL PT RO RU SD SE
SG SI SK SL TJ TM TR TT TZ UA UG UZ VN YU ZA ZW
AU 2000066480 A 20010326 (200137)
US 6302997 B1 20011016 (200164)
EP 1242677 A1 20020925 (200271) EN
R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT
RO SE SI
CN 1371439 A 20020925 (200305)
ADT WO 2001016423 A1 WO 2000-US22921 20000821; AU 2000066480 A AU 2000-66480
20000821; US 6302997 B1 US 1999-385154 19990830; EP 1242677 A1 EP
2000-954145 20000821, WO 2000-US22921 20000821; CN 1371439 A CN
2000-812212 20000821
FDT AU 2000066480 A Based on WO 200116423; EP 1242677 A1 Based on WO 200116423
PRAI US 1999-385154 19990830
AB WO 200116423 A UPAB: 20010418
NOVELTY - A papermaking pulp is produced from a nonwood fiber source
material by (a) extracting the source material with an alkaline pulping
solution for at least atmospheric pressure; (b) reducing the pH of the
source material to an acidic pH with an acid solution; (c) treating the
acidic source material with ozone; and (d) treating the ozone-treated
source material with a bleaching solution.
USE - For producing a papermaking pulp from a nonwood fiber source
material.
ADVANTAGE - The method is cost effective, environmentally sound at
small scale and odor-free. It keeps the number of processing steps or
stages to a minimum, minimizes water usage, and uses chemicals which
permit recovery of all internal filtrate streams. It removes high content
of non-fibrous materials often found in nonwoods, e.g., pith, parenchymal
cells or fines. The papermaking pulp produced has a Kappa number of at
least 5, a brightness level of at least 70%, preferably at least 80%,
International Standards Organization, and a freeness level of at least
400, preferably at least 550, ml Canadian Standard Freeness.
Dwg.0/6
TECH. . .
comprises a mineral acid and/or an organic acid, and a chelating agent.
Preferably, the acid is sulfuric acid, nitric acid, **phosphoric
acid**, and/or acetic acid. The bleaching solution is an alkaline
peroxide bleaching solution or a chlorine-based bleaching solution,
preferably chlorine dioxide. The alkaline peroxide bleaching solution
comprises NaOH, **hydrogen peroxide**, magnesium
sulfate, sodium silicate, and a chelant.

L3 ANSWER 45 OF 158 WPIDS (C) 2003 THOMSON DERWENT
AN 2001-168474 [17] WPIDS
DNN N2001-121492 DNC C2001-050307
TI Etching solution for spin etch planarization of surfaces for fabricating
integrated circuits comprises an oxidizing reactant for forming a
passivation layer.
DC E19 L03 U11
IN LEVERT, J; TOWERY, D L; LEVERT, J A; TOWERY, D
PA (ALLC) ALLIED-SIGNAL INC; (LEVE-I) LEVERT J A; (TOWE-I) TOWERY D L;
(LEVE-I) LEVERT J; (TOWE-I) TOWERY D
CYC 85
PI WO 2001006555 A1 20010125 (200117)* EN 38p
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
NL OA PT SD SE SL SZ TZ UG ZW
W: AL AM AT AU AZ BA BB BG BR BY CA CH CN CU CZ DE DK EE ES FI GB GE
GH GM HU ID IL IN IS JP KE KG KP KR KZ LC LK LR LS LT LU LV MD MG

MK MN MW MX NO NZ PL PT RO RU SD SE SG SI SK SL TJ TM TR TT UA UG
UZ VN YU ZW

AU 2000060809 A 20010205 (200128)

US 2001054706 A1 20011227 (200206)

EP 1198827 A1 20020424 (200235) EN

R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT
RO SE SI

KR 2002016907 A 20020306 (200261)

JP 2003505859 W 20030212 (200321) 30p

CN 1382305 A 20021127 (200322)

US 2003073311 A1 20030417 (200329)

ADT WO 2001006555 A1 WO 2000-US18723 20000710; AU 2000060809 A AU 2000-60809
20000710; US 2001054706 A1 US 1999-356487 19990719; EP 1198827 A1 EP
2000-947151 20000710, WO 2000-US18723 20000710; KR 2002016907 A KR
2002-700824 20020119; JP 2003505859 W WO 2000-US18723 20000710, JP
2001-510911 20000710; CN 1382305 A CN 2000-812842 20000710; US 2003073311
A1 Div ex US 1999-356487 19990719, US 2002-222230 20020815

FDT AU 2000060809 A Based on WO 200106555; EP 1198827 A1 Based on WO
200106555; JP 2003505859 W Based on WO 200106555

PRAI US 1999-356487 19990719; US 2002-222230 20020815

AB WO 200106555 A UPAB: 20010328

NOVELTY - An etching solution (9) comprises an oxidizing reactant for
forming a passivation layer; a depassivating co-reactant terminating the
passivation property; and adjusting mechanism for diffusion of the
reactants such that the rates of reactions in depressed regions of the
surface are diffusion-limited and have slower rates than reactions at
elevated regions of the surface.

DETAILED DESCRIPTION - An INDEPENDENT CLAIM is also included for a
method of planarization of a substrate comprising spinning the substrate
having the face to be planarized upward; directing an etching solution on
the face of the substrate; and adjusting the diffusion of the oxidizing
reactant and the depassivating co-reactant.

USE - For spin etch planarization (SEP) of surfaces (8) for
fabricating integrated circuits.

ADVANTAGE - The invention provides planarization of a surface without
mechanical contact or mechanical abrasion.

DESCRIPTION OF DRAWING(S) - The figure shows a schematic view of a
SEP.

Surfaces 8

Etching solution 9

Dwg.2/3

TECH. UPTX: 20010328

TECHNOLOGY FOCUS - INORGANIC CHEMISTRY - Preferred Material: The surface
is copper or tantalum. The oxidizing reactant is **hydrogen
peroxide (H2O2)**, nitric acid (HNO3), and/or sodium
chlorate.

Preferred Reactant: The depassivating reactant is **phosphoric
acid**, sulfuric acid, ammonium fluoride, copper (II)
sulfate, HNO3, hydrogen fluoride, **H2O2**, sodium hydroxide
and/or potassium hydroxide.

Preferred Component: The adjusting mechanism comprises an additive from
hydrochloric acid, borax, zinc **sulfate**, silicon hexafluoride,
heavy metal salts, salts of copper and tantalum, copper carbonate, copper
(I) chloride, iron chloride, and/or potassium chloride.. . .

L3 ANSWER 46 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-138078 [14] WPIDS

DNN N2001-100561 DNC C2001-040625

TI Reagent used for clearing contaminants from soil or groundwater comprises
metal catalyst in aqueous solution and pH modifying agent.

DC C07 D15 H03 P43

IN ANDREWS, T; GREENBERG, R S

PA (ANDR-I) ANDREWS T; (GREE-I) GREENBERG R S

CYC 92

PI WO 2001002105 A1 20010111 (200114)* EN 30p
 RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
 NL OA PT SD SE SL SZ TZ UG ZW
 W: AE AL AM AT AU AZ BA BB BG BR BY CA CH CN CR CU CZ DE DK DM EE ES
 FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ LC LK LR LS
 LT LU LV MA MD MG MK MN MW MX NO NZ PL PT RO RU SD SE SG SI SK SL
 TJ TM TR TT TZ UA UG US UZ VN YU ZA ZW
 AU 2000057593 A 20010122 (200125)
 US 6319328 B1 20011120 (200174)
 US 2002002983 A1 20020110 (200208)
 EP 1212149 A1 20020612 (200239) EN
 R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT
 RO SE SI
 JP 2003503197 W 20030128 (200309) 38p
 ADT WO 2001002105 A1 WO 2000-US17231 20000622; AU 2000057593 A AU 2000-57593
 20000622; US 6319328 B1 US 1999-345922 19990701; US 2002002983 A1 Div ex
 US 1999-345922 19990701, US 2001-943130 20010830; EP 1212149 A1 EP
 2000-943068 20000622, WO 2000-US17231 20000622; JP 2003503197 W WO
 2000-US17231 20000622, JP 2001-507588 20000622
 FDT AU 2000057593 A Based on WO 200102105; US 2002002983 A1 Div ex US 6319328;
 EP 1212149 A1 Based on WO 200102105; JP 2003503197 W Based on WO 200102105
 PRAI US 1999-345922 19990701; US 2001-943130 20010830
 AB WO 200102105 A UPAB: 20010312

NOVELTY - Reagent (I) comprises:

(a) an aqueous solution comprising metal catalyst comprising at least one of iron(II) salts, Fe(III) salts, Fe(II) chelates and/or Fe(III) chelates; and

(b) an optional pH modifying agent in an amount to provide a pH for the reagent of 5-8.

DETAILED DESCRIPTION - An INDEPENDENT CLAIM is included for a method of treating contaminants in an in situ environment which comprises:

(a) preparing an aqueous solution comprising a metal catalyst comprising Fe(II) salts, Fe salts, Fe(II) chelates and/or Fe(III) chelates;

(b) optionally adding a pH-modifying;

(c) adding a source of an oxidizing agent to the in situ environment in an amount to treat the contaminants and

(d) adding the aqueous catalyst solution to the in situ environment in the presence of the source of the oxidizing agent in an amount to promote formation of the oxidizing agent in an amount to treat the contaminants.

USE - Used in the treatment of contaminants in an in situ environment. The reagent is used to remove contaminants such as pesticides, herbicides, liquid hydrocarbons, other organic compounds, lubricants, chlorinated solvents and metal cyanides, from soils, fractured bedrock and groundwater with moderate to high or moderate permeability. The contaminants to be treated are less dense or denser than water.

ADVANTAGE - Reactive species are efficiently generated, to give a cost efficient and effective method of oxidizing contaminants in soil or groundwater.

Dwg.0/0

TECH. - ENVIRONMENT - Preferred components: The pH-modifying agent comprises water or a base. The metal catalyst is Fe(II) or Fe(III) **sulfate**, Fe(II) or Fe(III) EDTA chelate. The pH-modifying agent is water or sodium hydroxide. The oxidizing agent is a peroxide comprising **hydrogen peroxide**, sodium peroxide or calcium peroxide. The oxidizing agent is stabilized with a stabilizer comprising acids and/or salts. The stabilizer comprises **phosphoric acid** and/or monopotassium phosphate. Preferred method: The method comprises alternately adding the source of the oxidizing agent and the aqueous catalyst solution, . . .

AN 2002-033900 [04] WPIDS

DNC C2002-009422

TI In-situ treatment and remediation of contaminated underground area comprises injecting reactive solutions comprising ferrous **sulfate**, **hydrogen peroxide** and sulfuric or **phosphoric acid** in the ground.

DC D15

IN BRYANT, J D; WILSON, J T

PA (BRYA-I) BRYANT J D; (WILS-I) WILSON J T

CYC 1

PI US 2001042722 A1 20011122 (200204)* 11p

ADT US 2001042722 A1 Provisional US 2000-205326P 20000516, US 2001-855442 20010515

PRAI US 2000-205326P 20000516; US 2001-855442 20010515

AB US2001042722 A UPAB: 20020117

NOVELTY - An underground area contaminated with hexavalent chromium and other heavy metals, semi-metals or cyanide is in-situ treated and remediated by injecting first and second reactive solutions comprising ferrous **sulfate** and sulfuric acid or **phosphoric acid**, and **hydrogen peroxide** and sulfuric or **phosphoric acid**, respectively, into injectors inserted in the ground.

DETAILED DESCRIPTION - In-situ treatment and remediation of an underground area contaminated with hexavalent chromium and other heavy metals, semi-metals or cyanide comprises:

(a) determining an amount and concentration of a first reactive solution comprising ferrous sulfate and an acid comprising sulfuric acid or phosphoric acid required to decomplex chromium and other heavy metals, semi-metals or cyanide from mineral surfaces; initiating reduction of hexavalent chromium to trivalent chromium, and heavy metals and semi-metals to their least soluble valence state; and precipitating trivalent chromium, heavy metals and semi-metals as insoluble compounds;

(b) determining an amount and concentration of a second reactive solution comprising **hydrogen peroxide** and an acid comprising sulfuric acid or **phosphoric acid** required to destroy organic ligands and enhance decomplexation;

(c) injecting the first reactive solution and subsequently the second reactive solution at a flow rate in excess of a sustainable yield and with a pulsating pressure into injector(s) that are inserted into the ground, the injectors are sealed and positioned so as to assure liquid flow and dispersion of the reactive solutions through the contaminated area; and

(d) allowing the first and second reactive solutions to flow through the contaminated area to react chemically with the contaminants in the area.

USE - For in-situ treatment and remediation of an underground area contaminated with hexavalent chromium and other heavy metals, semi-metals or cyanide.

ADVANTAGE - The inventive method is economical, thorough, uniform, and verifiable.

Dwg.0/4

TI In-situ treatment and remediation of contaminated underground area comprises injecting reactive solutions comprising ferrous **sulfate**, **hydrogen peroxide** and sulfuric or **phosphoric acid** in the ground.

AB

other heavy metals, semi-metals or cyanide is in-situ treated and remediated by injecting first and second reactive solutions comprising ferrous **sulfate** and sulfuric acid or **phosphoric acid**, and **hydrogen peroxide** and sulfuric or **phosphoric acid**, respectively, into injectors inserted in the ground.

DETAILED DESCRIPTION - In-situ treatment and remediation of an underground area contaminated. . . chromium, heavy metals and semi-metals as insoluble compounds;

(b) determining an amount and concentration of a second reactive solution comprising **hydrogen peroxide** and an acid comprising sulfuric acid or **phosphoric acid** required to destroy organic ligands and enhance decomplexation;

(c) injecting the first reactive solution and subsequently the second reactive solution. . .

TT: SITU TREAT CONTAMINATE UNDERGROUND AREA COMPRISE INJECTION REACT SOLUTION COMPRISE FERROUS **HYDROGEN PEROXIDE** **PHOSPHORIC ACID** GROUND.

L3 ANSWER 48 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-342033 [36] WPIDS

DNC C2001-105773

TI Dyeing of cellulosic containing textile material with fiber reactive dyestuffs by using specified water baths from scouring the textile material to an after-scouring's final rinse.

DC A87 E19 E24 F06

IN HERRERA, F A A

PA (ATAL-I) ATALA F A

CYC 1

PI US 6193764 B1 20010227 (200136)* 12p

ADT US 6193764 B1 US 1999-295294 19990420

PRAI US 1999-295294 19990420

AB US 6193764 B UPAB: 20010628

NOVELTY - Providing a process which provides outstanding lot-to-lot color reproduction and a reduced hydrolysis dyeing mechanism resulting in more fiber reactive dye being bonded to the substrate thus better color yield.

DETAILED DESCRIPTION - A cellulosic containing textile material is dyed with fiber reactive dyestuffs by using at most 6 water baths from scouring the textile material to an after-scouring's final rinse of the cellulosic containing textile material, and the elapsed time is less than 5 hours.

Dyeing of a cellulosic containing textile material with fiber reactive dyestuffs by:

(1) scouring the textile material in a scouring bath;

(2) dropping the scouring bath, and without rinsing, sequentially adding:

(a) acetic acid;

(b) a dye assist system;

(c) a dyestuff formula;

(d) 5 - 200 g of electrolyte/L of the dye bath; and

(e) a stable buffered alkaline material to the dye bath, after the required dyeing temperature of the dyestuffs has been reached, resulting in a pH of 8-12.5; subjecting the material to the dye bath; and

(3) cold rinsing, after-scouring with acetic acid and an after-scouring agent at a temperature of 80-105 deg. C for 5-20 minutes, optionally hot water rinsing; and

(4) finally cold water rinsing the material.

The scouring bath comprises a scour assistant that is stable to high alkali, and a stable buffered alkaline material that permits proper pH to the scouring bath. The scouring bath has a pH of 8-11. The dye assist system comprises sodium meta nitrobenzene sulfonate and sodium polyacrylate. The dye stuff formula contains fiber reactive dyestuff(s). The stable buffered alkaline material permits proper fixation of dyestuff. The process uses at most 6 water baths from scouring the textile material to an after-scouring's final rinse of the cellulosic containing textile material, and the elapsed time is less than 5 hours.

USE - For dyeing cellulosic containing textile materials, e.g. fibers, yarns, and fabrics, with fiber reactive dyestuffs.

ADVANTAGE - The invention is a water, energy, and time efficient reactive rapid dyeing process suitable for all types of fiber reactive dyestuffs. It increases the rate of production of fiber reactive dyed cellulose containing materials per dyeing machine by as much as 60%. It decreases the cost of production. It provides an outstanding lot-to-lot

color reproduction, and a reduced hydrolysis dyeing mechanism resulting in more fiber reactive dye being bonded to the substrate, providing a better color yield. It further provides positive effects on the environment due to reduced volume of effluents as a result of decreased water consumption of the process (reduced to as much as 50%).
Dwg.0/0

TECH.

CHEMISTRY - Preferred Component: The stable buffered alkali material is a blend of alkali metal hydroxides, alkali metal carbonates, and **phosphoric acid**.

The electrolyte comprises sodium **sulfate** or sodium chloride.

The scouring bath further comprises **hydrogen peroxide** that permits required degree of whiteness of the cellulosic containing textile material, and peroxide stabilizing agent.

L3 ANSWER 49 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-583739 [66] WPIDS

DNC C2001-173157

TI Composition for dyeing keratin-containing fibers, especially human hair, containing formyl-1-methylquinolinium tosylate, giving strong, fast dyeings in wide range of colors.

DC A60 D21 E13 E24 F06

IN HOEFFKES, H; MOELLER, H; OBERKOBUSCH, D

PA (HENK) HENKEL KGAA

CYC 33

PI DE 19962875 A1 20010628 (200166)* 12p

AU 2001030127 A 20010709 (200166)

WO 2001047483 A1 20010705 (200166) DE

RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE TR

W: AU BR CA CN CZ HU JP NO PL RU SK US VN

EP 1239817 A1 20020918 (200269) DE

R: AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU MC NL PT SE TR

ADT DE 19962875 A1 DE 1999-19962875 19991224; AU 2001030127 A AU 2001-30127

20001215; WO 2001047483 A1 WO 2000-EP12816 20001215; EP 1239817 A1 EP

2000-990772 20001215, WO 2000-EP12816 20001215

FDT AU 2001030127 A Based on WO 200147483; EP 1239817 A1 Based on WO 200147483

PRAI DE 1999-19962875 19991224

AB DE 19962875 A UPAB: 20011113

NOVELTY - A composition (A) for dyeing keratin-containing fibers contains a formyl-1-methylquinolinium tosylate of formula (I), as dyeing component. (I) include the new compound 4-formyl-1-methyl-quinolinium p-toluenesulfonate (Ia).

DETAILED DESCRIPTION - A composition (A) for dyeing keratin-containing fibers contains a formyl-1-methylquinolinium tosylate of formula (I) (where the CHO group is in the 2- or 4-position), optionally in the form of its hydrate, as dyeing component. (I) include the new compound 4-formyl-1-methyl-quinolinium p-toluenesulfonate (Ia).

INDEPENDENT CLAIMS are included for:

(1) a method of dyeing keratin containing fibers, involving applying a dyeing composition (A) containing (I) and other conventional components to the fibers, leaving (A) in contact with the fibers for a suitable time (generally ca. 30 minutes) and rinsing or washing with shampoo; and

(2) new compound 4-formyl-1-methyl-quinolinium p-toluenesulfonate of formula (Ia).

USE - (A) is especially used for dyeing human hair (claimed), but may also be used to dye wool, furs or feathers. In principle (A) may further be used for dyeing other natural fibers (e.g. cotton, jute, sisal, linen or silk), modified natural fibers (e.g. regenerated, nitro, alkyl, hydroxyalkyl or acetyl cellulose) and synthetic fibers (e.g. polyamide, polyacrylonitrile, polyurethane or polyester). (I) is used for dyeing keratin-containing fibers (claimed).

ADVANTAGE - (A) give bright dyeings with good fastness (e.g. light, wash and rub-fastness) in a wide range of colors, i.e. violet or blue if (I) is used alone or yellow via orange, red, violet and brown to

blue-black or black if (I) is used with other dyeing components. The dyeings have color depth, gray covering power and fastness at least comparable with those obtained using conventional oxidation hair dyes, even when (A) are used in the absence of oxidizing agents. (A) have little or no sensitizing potential and do not discolor the skin. (I) precipitate in crystalline form, and are thus readily obtained in the required purity and easily handled.

Dwg.0/0

TECH.

(iii) salts selected from ammonium, alkali metal, alkaline earth metal, aluminum, manganese, iron, cobalt, copper and zinc formates, carbonates, halides, **sulfates**, butyrates, valerates, caproates, acetates, lactates, glycolates, tartrates, citrates, gluconates, propionates, phosphates and **phosphonates**;

(iv) oxidizing agents, preferably **hydrogen peroxide**, at 0.01-6 wt. % of the ready-for-use solution; or

(v) anionic, zwitterionic or nonionic surfactants.

Preparation: (Ia) is prepared by reacting. . .

L3 ANSWER 50 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-466382 [51] WPIDS

DNC C2001-140829

TI Composition for dyeing keratin-containing fibers, especially human hair, contains carbazole aldehyde or ketone derivative.

DC A60 B07 D18 D21 E24 F06

IN HOEFFKES, H; MOELLER, H; OBERKOBUSCH, D

PA (HENK) HENKEL KGAA

CYC 21

PI DE 19951135 A1 20010426 (200151)* 11p

WO 2001030312 A1 20010503 (200151) DE

RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE

W: AU JP US

AU 2000077884 A 20010508 (200153)

ADT DE 19951135 A1 DE 1999-19951135 19991023; WO 2001030312 A1 WO 2000-EP10198 20001017; AU 2000077884 A AU 2000-77884 20001017

FDT AU 2000077884 A Based on WO 200130312

PRAI DE 1999-19951135 19991023

AB DE 19951135 A UPAB: 20010910

NOVELTY - Composition (A) contains at least one carbazole-aldehyde or carbazolyl-alkyl or -aryl ketone (I) as a dyeing component.

DETAILED DESCRIPTION - Composition (A) contains at least one carbazole derivative of formula (I) as a dyeing component.

R1 = H, 1-4C alkyl or aryl;

R2-R5 = H, halo, 1-4C alkyl, 1-4C alkoxy, 1-4C hydroxyalkoxy, OH, NO2, NH2 (optionally substituted by 1-4C alkyl) or 1-4C acyl, or two of R2-R5 = a fused aromatic ring;

R6 = 1-4C alkyl, 1-4C hydroxyalkyl, 1-4C carboxyalkyl, 1-4C aryl, aralkyl, aralkyl or heteroaryl.

INDEPENDENT CLAIMS are included for the following:

(1) the use of (I) as a dyeing component in oxidation hair dye compositions, and

(2) dyeing keratin containing fibers, involving applying a dyeing composition (A') containing (I), at least one compound (II) and other conventional components to the fibers, leaving (A') in contact with the fibers for a suitable time (generally 30 minutes) and rinsing or washing with shampoo. (II) Comprises compounds containing primary or secondary amino groups or hydroxy groups (i.e. primary or secondary aromatic amines, aromatic hydroxy compounds) and/or active CH compounds.

USE - (A) Is used for dyeing keratin containing fibers, especially for dyeing human hair (claimed), but may also be used to dye wool, furs or feathers. (A) May also be used for dyeing other natural fibers (e.g. cotton, jute, sisal, linen or silk), modified natural fibers (e.g. regenerated, nitro, alkyl, hydroxyalkyl or acetyl cellulose) and synthetic fibers (e.g. polyamide, polyacrylonitrile, polyurethane or polyester).

ADVANTAGE - (A) Give bright dyeings with good fastness (e.g. light, wash and rub-fastness) in a wide range of colors, i.e. yellow-brown via orange, brown-orange, medium brown, dark brown, olive green to blue-black or black. The dyeings have color depth, gray covering power and fastness at least comparable with those obtained using conventional oxidation hair dyes, even when (A) are used in the absence of oxidizing agents. (A) have little or no sensitizing potential.
Dwg.0/0

TECH.

(A);
(iii) salts selected from ammonium, alkali metal, alkaline earth metal, aluminum, manganese, iron, cobalt, copper and zinc formates, carbonates, halides, **sulfates**, butyrates, valerates, caproates, acetates, lactates, glycolates, tartrates, citrates, gluconates, propionates, phosphates and **phosphonates**;
(iv) oxidizing agents, preferably **hydrogen peroxide**, at 0.01-6 wt. % of the ready-for-use solution; and/or
(v) anionic, zwitterionic or nonionic surfactants.

L3 ANSWER 51 OF 158 WPIDS (C) 2003 THOMSON DERWENT
AN 2001-419145 [45] WPIDS
DNC C2001-126982
TI Composition for dyeing keratin-containing fibers, especially human hair, contains aromatic aldehyde or ketone and active CH compound.
DC B07 D18 D21 E24 F06
IN HOEFFKES, H; MOELLER, H; OBERKOBUSCH, D
PA (HENK) HENKEL KGAA
CYC 32
PI DE 19951134 A1 20010426 (200145)* 14p
WO 2001034106 A1 20010517 (200145) DE
RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE
W: AU BR CA CN CZ HU JP NO PL RU SK US VN
AU 2001010252 A 20010606 (200152)
EP 1235549 A1 20020904 (200266) DE
R: AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU MC NL PT SE
JP 2003513898 W 20030415 (200328) 42p
ADT DE 19951134 A1 DE 1999-19951134 19991023; WO 2001034106 A1 WO 2000-EP10125 20001014; AU 2001010252 A AU 2001-10252 20001014; EP 1235549 A1 EP 2000-971379 20001014, WO 2000-EP10125 20001014; JP 2003513898 W WO 2000-EP10125 20001014, JP 2001-536107 20001014
FDT AU 2001010252 A Based on WO 200134106; EP 1235549 A1 Based on WO 200134106; JP 2003513898 W Based on WO 200134106
PRAI DE 1999-19951134 19991023
AB DE 19951134 A UPAB: 20010813
NOVELTY - Composition (A) contains a combination of aromatic aldehydes or ketones (I) with N-heterocyclic or other active CH compounds (II) and/or (III) (and/or a product of reaction of (I) with (II) and/or (III)) as dyeing component.
DETAILED DESCRIPTION - Composition (A) contains a combination of aromatic aldehydes or ketones of formula (I) with active CH compounds of formula (II) and/or R10-CH2-C(Z)-R11 (III) (and/or a product of reaction of (I) with (II) and/or (III)) as dyeing component.
R1 = H, 1-4C alkyl or aryl;
R2-R4 = H, halo, 1-4C alkyl, OH, 1-4C alkoxy, 1-4C hydroxyalkoxy or NO2;
R5 = H, 1-4C alkyl, 2-4C hydroxyalkyl, 2-4C alkenyl or aryl; or
R5 + one of R2-R4 = a fused 5-7 membered heterocyclic ring; or
R5 + attached O = an olate group, the negative charge of which is balanced by an alkali metal or ammonium ion;
n = 0-2;
R10 = 1-10C alkyl, 2-4C alkenyl, 2-4C hydroxyalkyl, 2-4C carboxyalkyl, 2-4C sulfoalkyl or aralkyl;
R7, R8 = H, 1-4C alkyl, halo, OH, 1-4C alkoxy or NO2, or
R7 + R8 = a fused aromatic ring;

R9 = H, 1-4C alkyl or aryl;
 X = O, S, CH=CH or NR12;
 R12 = 1-4C alkyl, 2-4C carboxyalkyl, 2-4C sulfoalkyl, 2-4C sulfoxyalkyl, 2-4C hydroxyalkyl or aralkyl;
 Y = halide, 1-4C alkylsulfate, 1-4C alkanesulfonate, arenesulfonate, 1-4C perfluoroalkanesulfonate, tetrafluoroborate, perhalate, sulfate, hydrogen sulfate or carboxylate anion;
 R10 = 1-4C acyl, aroyl, 1-4C alkylsulfonyl, 1-4C alkylsulfinyl, mono- or di-(1-4C alkyl) amino, vinylcarbonyl, methanimino, nitrile, ester or carboxamide group (optionally substituted by 1-4C alkyl, 2-4C hydroxyalkyl or aryl);
 R11 = 1-4C acyl, 1-4C alkoxy, mono- or di-(1-4C alkyl)-amino or 1-4C acylamino, or
 R10 + R11 = a group completing a 5-7 membered heterocycle selected from thiazolidine-2,5-dione, thiazolidine-2-thion-5-one, perhydropyrimidine-2,4,6-trione, perhydropyrimidine-2-thione-4,6-dione, cyclopentane-1,3-dione, cyclohexane-1,3-dione, indane-1,3-dione, 2-pyrazolin-5-one, 1,2-dihydro-6-hydroxy-2-hydroxypyridine (sic) or their enol esters;

Z = O, S or C(CN)2.

INDEPENDENT CLAIMS are included for the following:

(i) the use of the (I)/((II) and/or (III)) combination and/or reaction product as a dyeing component in oxidation hair dye compositions; and

(ii) a method of dyeing keratin containing fibers, involving applying a dyeing composition (A) containing the above dyeing component and other conventional components to the fibers, leaving (A) in contact with the fibers for a suitable time (generally 30 minutes) and rinsing or washing with shampoo.

USE - (A) is used for dyeing keratin containing fibers, especially for dyeing human hair (claimed), but may also be used to dye wool, furs or feathers. (A) May also be used for dyeing other natural fibers (e.g. cotton, jute, sisal, linen or silk), modified natural fibers (e.g. regenerated, nitro, alkyl, hydroxyalkyl or acetyl cellulose) and synthetic fibers (e.g. polyamide, polyacrylonitrile, polyurethane or polyester).

ADVANTAGE - (A) Give bright dyeings with good fastness (e.g. light, wash and rub-fastness) in a wide range of colors. The dyeings have color depth, gray covering power and fastness at least comparable with those obtained using conventional oxidation hair dyes, even when (A) are used in the absence of oxidizing agents. (A) have little or no sensitizing potential.

Dwg.0/0

TECH.

of (A);

(iii) salts comprising ammonium, alkali metal, alkaline earth metal, aluminum, manganese, iron, cobalt, copper and zinc formates, carbonates, halides, **sulfates**, butyrates, valerates, caproates, acetates, lactates, glycolates, tartrates, citrates, gluconates, propionates, phosphates or **phosphonates**;

(iv) oxidizing agents, preferably **hydrogen peroxide**, at 0.01-6 wt. % of the ready-for-use solution, and/or

(v) anionic, zwitterionic or nonionic surfactants.

L3 ANSWER 52 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-357026 [38] WPIDS

DNC C2001-110895

TI Depolymerization of cellulose ether involves acid hydrolysis of a concentrated aqueous slurry at above the flocculation temperature, used to coat seeds, pharmaceuticals and for use in cosmetics and food.

DC A11 A96 A97 B07 D13 D21 G02

IN HAMMES, A

PA (CLRN) CLARIANT GMBH

CYC 23

PI DE 19941893 A1 20010308 (200138)*

7p

WO 2001018062 A1 20010315 (200138) DE
RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE
W: JP KR MX US

EP 1237931 A1 20020911 (200267) DE
R: AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU MC NL PT SE

KR 2002033784 A 20020507 (200271)

JP 2003508597 W 20030304 (200319) 23p

ADT DE 19941893 A1 DE 1999-19941893 19990903; WO 2001018062 A1 WO 2000-EP8538
20000901; EP 1237931 A1 EP 2000-960577 20000901, WO 2000-EP8538 20000901;
KR 2002033784 A KR 2002-702908 20020304; JP 2003508597 W WO 2000-EP8538
20000901, JP 2001-522284 20000901

FDT EP 1237931 A1 Based on WO 200118062; JP 2003508597 W Based on WO 200118062

PRAI DE 1999-19941893 19990903

AB DE 19941893 A UPAB: 20010711

NOVELTY - A method for the depolymerization of hot water-flocculatable cellulose ethers (I) by hydrolytic degradation with acid, in which degradation is carried out at a temperature above the flocculation point with (I) in the form of a concentrated aqueous slurry.

DETAILED DESCRIPTION - An INDEPENDENT CLAIM is also included for methylhydroxypropylcellulose (MHPC) with a Hoppler viscosity of not more than 50 mPa.s (measured on a 2.0% solution of the (dry) substance in water at 20 deg. C), which shows a degree of whiteness of more than 50% with a particle size distribution in where not more than 50% of the particles have a particle size of less than 125 micro m.

USE - For coating pharmaceuticals or seeds, for use in cosmetics or food and for use in suspension polymerisation (claimed).

ADVANTAGE - Enables the production of low-viscosity cellulose ethers with (as far as possible) a uniform degree of polymerisation, a low content of by-products, a very low salt content and a high degree of whiteness, forming solutions with a high transmission value.

Dwg.0/0

TECH.

to adjust the aqueous solution to pH 5.5-8.0.

TECHNOLOGY FOCUS - INORGANIC CHEMISTRY - Preferred Reagents: Hydrochloric, sulfuric, nitric and/or **phosphoric acid** are used as mineral acids, peroxy compounds, perborates, sodium chlorite, halogens and/or halogen oxides are used as oxidizing agents (especially **hydrogen peroxide**), and sodium carbonate, bicarbonate, **sulfate** and/or bisulfate are used as salts for washing the product.

L3 ANSWER 53 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-203652 [21] WPIDS

DNC C2001-060667

TI Dyeing composition giving deep, fast shades on keratin fibers, especially human hair, contains quaternized heterocyclic aldehyde or ketone and amino, hydroxy or active carbon-hydrogen compound.

DC B07 D21 E19 E23 E24

IN HOEFFKES, H; MOELLER, H; OBERKOBUSCH, D

PA (HENK) HENKEL KGAA

CYC 21

PI DE 19933187 A1 20010118 (200121)* 11p

WO 2001005359 A2 20010125 (200121) DE

RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE

W: AU JP US

AU 2000064324 A 20010205 (200128)

ADT DE 19933187 A1 DE 1999-19933187 19990715; WO 2001005359 A2 WO 2000-EP6399
20000706; AU 2000064324 A AU 2000-64324 20000706

FDT AU 2000064324 A Based on WO 200105359

PRAI DE 1999-19933187 19990715

AB DE 19933187 A UPAB: 20010418

NOVELTY - A composition (A) for dyeing keratin fibers contains at least one quaternized pyrazine, pyrimidine or pyridazine aldehyde or ketone

derivative (I) and at least one compound (II) selected from compounds having primary or secondary amino or hydroxy groups (i.e. primary or secondary aromatic amines, N-containing heterocycles or aromatic hydroxy compounds) and active CH compounds.

DETAILED DESCRIPTION - A composition (A) for dyeing keratin fibers contains at least one quaternized heterocyclic compound of formula (I) and at least one compound (II) selected from compounds having primary or secondary amino or hydroxy groups (i.e. primary or secondary aromatic amines, N-containing heterocycles or aromatic hydroxy compounds) and active CH compounds:

R1 = H, 1-4C alkyl, aryl or heteroaryl;

R2-R4 = H, halo, 1-4C alkyl, 1-4C alkoxy, 1-4C hydroxyalkoxy, OH, SO₃H or amino (optionally substituted by 1-4C alkyl) or 1-4C acyl; or two adjacent groups together form a fused benzene ring;

R5 = 1-4C alkyl, 1-4C hydroxyalkyl, 1-4C carboxyalkyl, 1-4C sulfoalkyl, aryl, aralkyl, heteroaryl or N-oxide residue;

two of Q1-Q3 = C and the other is optionally quaternized N;

X = direct bond or optionally substituted vinylene;

Y = halide, benzenesulfonate, p-toluenesulfonate, methanesulfonate, trifluoromethanesulfonate, perfluoroalkanoate, hydrogen sulfate, tetrachlorozincate or N-oxide residue.

An INDEPENDENT CLAIM is included for a method for dyeing keratin fibers, involving contacting the fibers for a suitable time (e.g. ca. 30 minutes) with a dyeing composition containing (I), (II) and conventional cosmetic components then rinsing or washing using shampoo, where (II) is selected from compounds having primary or secondary amino or hydroxy groups (i.e. primary or secondary aromatic amines, N-containing heterocycles or aromatic hydroxy compounds) and active CH compounds.

USE - (A) is especially used for dyeing human hair (claimed), but may also be used with wool, furs, feathers or fibers of materials other than keratin, e.g. cotton, jute, sisal, linen, silk, regenerated cellulose, nitro-, alkyl-, hydroxyalkyl or acetyl-cellulose, polyamide, polyacrylonitrile, polyurethane or polyester. The combinations of (I) and (II) generally give orange-brown, red-brown or violet-brown shades.

ADVANTAGE - (I) gives bright, deep dyeings in a wide range of colors and have little or no sensitizing potential. Good color depth, grey covering power and fastness properties are obtained even in the absence of oxidizing agents (such as hydrogen peroxide).

Dwg. 0/0

TECH.

alkaline earth metal (e.g. magnesium, calcium, strontium or barium), aluminum, manganese, iron, cobalt, copper, zinc or ammonium formate, carbonate, halide, sulfate, butyrate, valerate, caproate, acetate, lactate, glycolate, tartrate, citrate, gluconate, propionate, phosphate or phosphonate; 0.01-6 wt. % of an oxidizing agent, specifically hydrogen peroxide; and/or anionic, zwitterionic or nonionic surfactants.

L3 ANSWER 54 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-192436 [20] WPIDS

DNC C2001-057897

TI Dyeing composition giving deep, fast shades on keratin fibers, especially human hair, contains N-vinyl-isatin derivative as dyeing component.

DC B07 D21 E19 E23 E24

IN ROSE, D

PA (HENK) HENKEL KGAA

CYC 32

PI DE 19932567 A1 20010118 (200120)* 11p

WO 2001003660 A1 20010118 (200120) DE

RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE

W: AU BR CA CN CZ HU JP NO PL RU SK US VN

AU 2000058240 A 20010130 (200127)

EP 1194116 A1 20020410 (200232) DE

R: AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU MC NL PT SE

ADT DE 19932567 A1 DE 1999-19932567 19990713; WO 2001003660 A1 WO 2000-EP6218
20000704; AU 2000058240 A AU 2000-58240 20000704; EP 1194116 A1 EP
2000-943979 20000704, WO 2000-EP6218 20000704
FDT AU 2000058240 A Based on WO 200103660; EP 1194116 A1 Based on WO 200103660
PRAI DE 1999-19932567 19990713
AB DE 19932567 A UPAB: 20010410

NOVELTY - Composition (A) contains at least one N-vinyl-isatin derivative
(I) as dyeing component.

DETAILED DESCRIPTION - Composition (A) contains at least one isatin
derivative of formula (I) or its salt as dyeing component.

R1 = vinyl (optionally substituted by 1 or 2 1-4C alkyl);
R2-R5 = H, OH, halo, NO2, SO3H, COOH, 1-4C alkyl, 1-4C alkoxy or
NR6R7; or two adjacent groups together form 1-4C alkylenedioxy;
R5, R6 = H, 1-4C alkyl or 2-4C hydroxyalkyl.

INDEPENDENT CLAIMS are included for the following:

(i) the use of (I) or their salts as dyeing components in oxidation
hair dyes; and

(ii) a method for dyeing keratin fibers involving contacting the
fibers for a suitable time (e.g. 30 minutes) with a dyeing composition
containing (I) (or its salt), at least one compound (II) and conventional
cosmetic components then rinsing or washing using shampoo. (II) Is
selected from compounds having primary or secondary amino or hydroxy
groups (i.e. primary or secondary aromatic amines, N-containing
heterocycles, aminoacids, oligopeptides formed from 2-9 aminoacids or
aromatic hydroxy compounds) and active CH compounds.

USE - (A) Is especially used for dyeing human hair (claimed), but may
also be used with wool, furs, feathers or fibers of materials other than
keratin e.g. cotton, jute, sisal, linen, silk, regenerated cellulose,
nitro-, alkyl-, hydroxyalkyl or acetyl-cellulose, polyamide,
polyacrylonitrile, polyurethane or polyester.

(I) Give orange or brown shades when used alone, but in combination
with (II) may give shades from yellow via yellow-brown, orange,
brown-orange, medium brown, olive green, dark brown, violet, dark violet
and blue-black to black.

ADVANTAGE - (I) Gives bright, deep dyeings in a wide range of colors
and have little or no sensitizing potential. Good color depth, grey
covering power and fastness properties are obtained even in the absence of
oxidizing agents (such as hydrogen peroxide).

Dwg.0/0

TECH.

alkaline earth metal (e.g. magnesium, calcium, strontium or barium),
aluminum, manganese, iron, cobalt, copper, zinc or ammonium formate,
carbonate, halide, **sulfate**, butyrate, valerate, caproate,
acetate, lactate, glycolate, tartrate, citrate, gluconate, propionate,
phosphate or **phosphonate**; 0.01-6 wt. % of an oxidizing agent,
specifically **hydrogen peroxide**; and/or anionic,
zwitterionic or nonionic surfactants.

L3 ANSWER 55 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-212101 [22] WPIDS

DNC C2001-063394

TI Dyeing composition giving deep, fast shades on keratin fibers, especially
human hair, contains nitroso-benzene derivative as dyeing component.

DC B07 D21 E19 E24

IN HOEFFKES, H; MOELLER, H; OBERKOBUSCH, D

PA (HENK) HENKEL KGAA

CYC 21

PI DE 19932566 A1 20010118 (200122)* 13p

WO 2001003661 A1 20010118 (200122) DE

RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE

W: AU JP US

AU 2000065606 A 20010130 (200127)

ADT DE 19932566 A1 DE 1999-19932566 19990713; WO 2001003661 A1 WO 2000-EP6219
20000704; AU 2000065606 A AU 2000-65606 20000704

FDT AU 2000065606 A Based on WO 200103661
PRAI DE 1999-19932566 19990713
AB DE 19932566 A UPAB: 20010421

NOVELTY - A composition (A) for dyeing keratin fibers contains at least one nitroso-benzene derivative (I) as dyeing component.

DETAILED DESCRIPTION - A composition (A) for dyeing keratin fibers contains at least one nitroso-benzene derivative of formula (I) as dyeing component.

R1-R4 = H, halo, 1-4C alkyl, 1-4C alkoxy, 1-4C hydroxyalkoxy, OH, SO₃H or amino (optionally substituted by 1-4C alkyl, 1-4C hydroxyalkoxy or aryl; or two adjacent groups together form a fused aromatic or heteroaromatic ring.

INDEPENDENT CLAIMS are included for:

(1) the use of (I) as dyeing components in oxidation hair dyes; and
(2) a method for dyeing keratin fibers, involving contacting the fibers for a suitable time (e.g. ca. 30 minutes) with a dyeing composition containing (I), at least one compound (II) and conventional cosmetic components then rinsing or washing using shampoo, where (II) is selected from compounds having primary or secondary amino or hydroxy groups (i.e. primary or secondary aromatic amines, N-containing heterocycles or aromatic hydroxy compounds) and active CH compounds.

USE - (A) is especially used for dyeing human hair (claimed), but may also be used with wool, furs, feathers or fibers of materials other than keratin, e.g. cotton, jute, sisal, linen, silk, regenerated cellulose, nitro-, alkyl-, hydroxyalkyl or acetyl-cellulose, polyamide, polyacrylonitrile, polyurethane or polyester. (I) generally give yellow shades when used alone, but in combination with (II) may give shades from yellow via yellow-brown, orange, brown-orange, medium brown, olive green, dark brown, violet, dark violet and blue-black to black.

ADVANTAGE - (I) gives bright, deep dyeing in a wide range of colors and have little or no sensitizing potential. Good color depth, grey covering power and fastness properties are obtained even in the absence of oxidizing agents (such as hydrogen peroxide).

Dwg.0/0

TECH. . . .

alkaline earth metal (e.g. magnesium, calcium, strontium or barium), aluminum, manganese, iron, cobalt, copper, zinc or ammonium formate, carbonate, halide, **sulfate**, butyrate, valerate, caproate, acetate, lactate, glycolate, tartrate, citrate, gluconate, propionate, phosphate or **phosphonate**; 0.01-6 wt. % of an oxidizing agent, specifically **hydrogen peroxide**; and/or anionic, zwitterionic or nonionic surfactants.

L3 ANSWER 56 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-192435 [20] WPIDS

DNC C2001-057896

TI Dyeing composition giving deep, fast shades on keratin fibers, especially human hair, contains N-halo-imine derivative as dyeing component.

DC B07 D21 E19 E24

IN HOEFFKES, H; MOELLER, H; OBERKOBUSCH, D

PA (HENK) HENKEL KGAA

CYC 22

PI DE 19932565 A1 20010118 (200120)* 11p

WO 2001003651 A2 20010118 (200120) DE

RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE

W: AU JP US

AU 2000061543 A 20010130 (200127)

EP 1194117 A2 20020410 (200232) DE

R: AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU MC NL PT SE

ADT DE 19932565 A1 DE 1999-19932565 19990713; WO 2001003651 A2 WO 2000-EP6220

20000704; AU 2000061543 A AU 2000-61543 20000704; EP 1194117 A2 EP

2000-947919 20000704, WO 2000-EP6220 20000704

FDT AU 2000061543 A Based on WO 200103651; EP 1194117 A2 Based on WO 200103651

PRAI DE 1999-19932565 19990713

AB DE 19932565 A UPAB: 20010410

NOVELTY - A composition (A) for dyeing keratin fibers contains at least one N-halo-p-benzoquinoneimine derivative (I) as dyeing component.

DETAILED DESCRIPTION - A composition (A) for dyeing keratin fibers contains at least one N-halo-imine of formula (I) as dyeing component:

R1-R4 = H, halo, 1-4C alkyl or 1-4C alkoxy; or two adjacent groups together form a fused aromatic or heteroaromatic ring;

X = O or arylimine;

Y = F, Cl, Br or I.

INDEPENDENT CLAIMS are included for:

(1) the use of (I) as dyeing components in oxidation hair dyes; and
(2) a method for dyeing keratin fibers, involving contacting the fibers for a suitable time (e.g. ca. 30 minutes) with a dyeing composition containing (I), at least one compound (II) and conventional cosmetic components then rinsing or washing using shampoo, where (II) is selected from compounds having primary or secondary amino or hydroxy groups (i.e. primary or secondary aromatic amines, N-containing heterocycles or aromatic hydroxy compounds) and active CH compounds.

USE - (A) is especially used for dyeing human hair (claimed), but may also be used with wool, furs, feathers or fibers of materials other than keratin, e.g. cotton, jute, sisal, linen, silk, regenerated cellulose, nitro-, alkyl-, hydroxyalkyl or acetyl-cellulose, polyamide, polyacrylonitrile, polyurethane or polyester. (I) generally give orange or brown shades when used alone, but in combination with (II) may give shades from yellow via yellow-brown, orange, brown-orange, medium brown, olive green, dark brown, violet, dark violet and blue-black to black.

ADVANTAGE - (I) gives bright, deep dyeings in a wide range of colors and have little or no sensitizing potential. Good color depth, grey covering power and fastness properties are obtained even in the absence of oxidizing agents (such as hydrogen peroxide).

Dwg.0/0

TECH.

alkaline earth metal (e.g. magnesium, calcium, strontium or barium), aluminum, manganese, iron, cobalt, copper, zinc or ammonium formate, carbonate, halide, **sulfate**, butyrate, valerate, caproate, acetate, lactate, glycolate, tartrate, citrate, gluconate, propionate, phosphate or **phosphonate**; 0.01-6 wt. % of an oxidizing agent, specifically **hydrogen peroxide**; and/or anionic, zwitterionic or nonionic surfactants.

L3 ANSWER 57 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2001:222335 CAPLUS

DN 135:76499

TI Oxidation of sulfides to sulfoxides and sulfones with 30% hydrogen peroxide under organic solvent- and halogen-free conditions

AU Sato, K.; Hyodo, M.; Aoki, M.; Zheng, X.-Q.; Noyori, R.

CS Department of Chemistry and Research Center for Materials Science, Nagoya University, Chikusa, Nagoya, 464-8602, Japan

SO Tetrahedron (2001), 57(13), 2469-2476

CODEN: TETRAB; ISSN: 0040-4020

PB Elsevier Science Ltd.

DT Journal

LA English

OS CASREACT 135:76499

AB Arom. and aliph. sulfides are oxidized to sulfoxides or sulfones in high yield with 30% hydrogen peroxide under org. solvent- and halogen-free conditions. Dialkyl and alkyl aryl sulfides are cleanly oxidized to sulfoxides using aq. hydrogen peroxide without catalysts. The best catalyst for the sulfone synthesis consists of sodium tungstate, phenylphosphonic acid, and methyltrioctylammonium hydrogen sulfate. Co-existing primary or secondary alc. or olefinic moieties are unaffected under such conditions.

RE.CNT 79 THERE ARE 79 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

IT 791-28-6, Triphenylphosphine oxide 1066-51-9 1707-03-5 7664-38-2,
Phosphoric acid, uses
 RL: CAT (Catalyst use); MOA (Modifier or additive use); USES (Uses)
 (oxidn. of sulfides to sulfoxides and sulfones with 30%
hydrogen peroxide under org. solvent-free and
 halogen-free conditions)

IT 59158-14-4P, Methyltriethylammonium hydrogen **sulfate**
 RL: CAT (Catalyst use); SPN (Synthetic preparation); PREP (Preparation);
 USES (Uses)
 (oxidn. of sulfides to sulfoxides and sulfones with 30%
hydrogen peroxide under org. solvent-free and
 halogen-free conditions)

L3 ANSWER 58 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 2001:401614 CAPLUS
 DN 135:121480
 TI Antimicrobial treatments for minimally processed cantaloupe melon
 AU Sapers, G. M.; Miller, R. L.; Pilizota, V.; Mattrazzo, A. M.
 CS Agricultural Res. Service, Eastern Regional Res. Center, U.S. Dep.
 Agriculture, Wyndmoor, PA, 19038, USA
 SO Journal of Food Science (2001), 66(2), 345-349
 CODEN: JFDSA; ISSN: 0022-1147
 PB Institute of Food Technologists
 DT Journal
 LA English
 AB Efficacy of decontamination treatments in reducing endogenous microbial
 populations on cantaloupe and in extending fresh-cut shelf-life were
 investigated. Composite rind plug samples were washed with water or
 solns. of sodium hypochlorite, **H2O2**, com. detergent formulations
 contg. dodecylbenzene **sulfonic acid** and **phosphoric**
acid, or trisodium phosphate, and surviving microbial populations
 detd. Fresh-cut cubes were prepd. aseptically from whole melons given
 similar treatments, and their visual appearance and bacterial population
 detd. during storage at 4 .degree.C. Population redns. on washed rind
 plugs were < 1 log with water, 1 to 2 logs with washing and sanitizing
 agents applied individually, and 3 logs with some sequential treatments
 with **H2O2**. **H2O2** applied at 50 .degree.C was superior to other whole-melon
 treatments, yielding a fresh-cut shelf-life of > 2 wk.

RE.CNT 21 THERE ARE 21 CITED REFERENCES AVAILABLE FOR THIS RECORD
 ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB Efficacy of decontamination treatments in reducing endogenous microbial
 populations on cantaloupe and in extending fresh-cut shelf-life were
 investigated. Composite rind plug samples were washed with water or
 solns. of sodium hypochlorite, **H2O2**, com. detergent formulations
 contg. dodecylbenzene **sulfonic acid** and **phosphoric**
acid, or trisodium phosphate, and surviving microbial populations
 detd. Fresh-cut cubes were prepd. aseptically from whole melons given
 similar treatments, and their visual appearance and bacterial population
 detd. during storage at 4 .degree.C. Population redns. on washed rind
 plugs were < 1 log with water, 1 to 2 logs with washing and sanitizing
 agents applied individually, and 3 logs with some sequential treatments
 with **H2O2**. **H2O2** applied at 50 .degree.C was superior to other whole-melon
 treatments, yielding a fresh-cut shelf-life of > 2 wk.

L3 ANSWER 59 OF 158 CAPLUS COPYRIGHT 2003 ACS DUPLICATE 2
 AN 2000:420915 CAPLUS
 DN 133:48977
 TI Hydrogen peroxide disinfectant with increased activity
 IN Rochon, Michael J.
 PA Virox Technologies Inc., Can.
 SO PCT Int. Appl., 22 pp.
 CODEN: PIXXD2
 DT Patent
 LA English

FAN.CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2000035289	A1	20000622	WO 1999-CA1132	19991126
	W: AE, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CR, CU, CZ, DE, DK, DM, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, TZ, UA, UG, UZ, VN, YU, ZA, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM				
	RW: GH, GM, KE, LS, MW, SD, SL, SZ, TZ, UG, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG				
	US 6346279	B1	20020212	US 1999-356345	19990719
	BR 9915987	A	20010904	BR 1999-15987	19991126
	EP 1139762	A1	20011010	EP 1999-973389	19991126
	EP 1139762	B1	20021002		
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO				
	AU 741104	B2	20011122	AU 2000-13694	19991126
	CA 2344471	C	20020108	CA 1999-2344471	19991126
	JP 2002532398	T2	20021002	JP 2000-587617	19991126
	JP 3350526	B2	20021125		
	AT 225128	E	20021015	AT 1999-973389	19991126
	ES 2185428	T3	20030416	ES 1999-973389	19991126
	US 2002142051	A1	20021003	US 2002-67809	20020208
PRAI	US 1998-112047P	P	19981214		
	US 1999-356345	A	19990719		
	WO 1999-CA1132	W	19991126		
AB	An acidic aq. hydrogen peroxide soln. is provided, with improved disinfectant activity. Concd. solns. preferably contain .ltoreq.8 % and as-used concns. contain 0.5 % peroxide. The soln. also contains from 0.1 to 5.0 % of at least one acid compd., e.g. phosphoric and/or a phosphonate, with 1-5 phosphonic acid groups, and 0.02-5 % of at least one anionic surfactant. The surfactant is selected from C8 to C16 alkylarylsulfonic acids, sulfonated C12 to C22 carboxylic acids, C8 to C22 alkyl di-Ph oxide sulfonic acids, naphthalene sulfonic acids, C8 to C22 alkylsulfonic acids, and alkali metal and ammonium salts thereof, and alkali metal C8 to C18 alkyl sulfates, and mixts. thereof. Most preferably, the soln. has an emulsifier, e.g. a salt of an alkylated di-Ph oxide. The soln. may also contain corrosion inhibitors and/or lower alcs.				
RE.CNT	5 THERE ARE 5 CITED REFERENCES AVAILABLE FOR THIS RECORD				
	ALL CITATIONS AVAILABLE IN THE RE FORMAT				
IT	Carboxylic acids, biological studies				
	RL: BUU (Biological use, unclassified); THU (Therapeutic use); BIOL (Biological study); USES (Uses)				
	(C12-C22, sulfonated ; activity enhancers for hydrogen peroxide disinfectant)				
IT	27176-87-0, Dodecylbenzene sulfonic acid				
	RL: BUU (Biological use, unclassified); THU (Therapeutic use); BIOL (Biological study); USES (Uses)				
	(Biosoft; activity enhancer for hydrogen peroxide disinfectant)				
IT	6419-19-8, Briquest 301-50A 7664-38-2, Phosphoric acid , biological studies 9002-93-1, Triton X-405 61332-13-6, Dowfax				
	RL: BUU (Biological use, unclassified); THU (Therapeutic use); BIOL (Biological study); USES (Uses)				
	(activity enhancer for hydrogen peroxide disinfectant)				
IT	1330-69-4, Dodecylbenzene sulfonate 7722-84-1, Hydrogen peroxide , biological studies 25155-19-5, Naphthalenesulfonic acid				
	RL: BUU (Biological use, unclassified); THU (Therapeutic use); BIOL (Biological study); USES (Uses)				

(activity enhancers for hydrogen peroxide
disinfectant)

L3 ANSWER 60 OF 158 CAPLUS COPYRIGHT 2003 ACS DUPLICATE 3
AN 2000:278077 CAPLUS
DN 132:295400
TI Thickened liquid hydrogen peroxide bleach compositions
IN Mercado, Hedeliza Malonzo
PA Colgate-Palmolive Co., USA
SO PCT Int. Appl., 12 pp.
CODEN: PIXXD2
DT Patent
LA English
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2000023555	A1	20000427	WO 1999-US24910	19991021
	W:	AE, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CR, CU, CZ, DE, DK, DM, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, UZ, VN, YU, ZA, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM			
	RW:	GH, GM, KE, LS, MW, SD, SL, SZ, TZ, UG, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG			
	BR 9914702	A	20010710	BR 1999-14702	19991021
	EP 1123375	A1	20010816	EP 1999-970679	19991021
	R:	AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO			

PRAI US 1998-177688 A 19981022
WO 1999-US24910 W 19991021

AB A thickened liq. bleaching compn. is provided comprising **hydrogen peroxide**; a surfactant mixt. comprising alkyl dimethylamine oxide and soap formed by the neutralization of a C6-C18 fatty acid with an alkali metal hydroxide; an electrolyte comprising an alkyl metal **sulfate**; a **phosphonate** stabilizer for **hydrogen peroxide**; and alkali metal hydroxide in an amt. sufficient to effect in-situ neutralization of the C6-C18 fatty acid to form the soap and also provide a pH in the range of from about 8 to 9.5.

RE.CNT 11 THERE ARE 11 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB A thickened liq. bleaching compn. is provided comprising **hydrogen peroxide**; a surfactant mixt. comprising alkyl dimethylamine oxide and soap formed by the neutralization of a C6-C18 fatty acid with an alkali metal hydroxide; an electrolyte comprising an alkyl metal **sulfate**; a **phosphonate** stabilizer for **hydrogen peroxide**; and alkali metal hydroxide in an amt. sufficient to effect in-situ neutralization of the C6-C18 fatty acid to form the soap and also provide a pH in the range of from about 8 to 9.5.

IT 61792-09-4, Pentasodium salt of diethylenetriamine penta (methylene **phosphonic acid**)

RL: MOA (Modifier or additive use); USES (Uses)
(stabilizer; thickened liq. **hydrogen peroxide** bleach compns.)

IT 532-32-1, Sodium benzoate 7722-84-1, Hydrogen peroxide, uses
7757-82-6, Sodium **sulfate**, uses 7778-80-5, Potassium **sulfate**, uses

RL: TEM (Technical or engineered material use); USES (Uses)
(thickened liq. **hydrogen peroxide** bleach compns.)

L3 ANSWER 61 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2000:589878 CAPLUS
DN 133:160846

TI Ready-to-use aqueous hard surface cleaning and disinfecting compositions
containing hydrogen peroxide
IN Monticello, Michael Vincent; Mayerhauser, George Robert
PA Reckitt Benckiser Inc., USA
SO U.S., 9 pp., Cont.-in-part of U.S. 5,891,392.
CODEN: USXXAM

DT Patent
LA English

FAN.CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 6106774	A	20000822	US 1999-227464	19990108
	GB 2319179	A1	19980520	GB 1996-23473	19961112
	US 5891392	A	19990406	US 1997-928097	19970912
PRAI	GB 1996-23473	A	19961112		
	US 1997-928097	A2	19970912		

AB The title compns. include 0.1-20% by wt. C1-6 monohydric alc., 1.0-10% by wt. glycol ether, or butoxypropanol or propoxypropanol, 0.1-12% by wt. anionic, cationic, nonionic and amphoteric surfactant, 0.1-10% by wt. hydrogen peroxide, 0.1-7% by wt. acid and to 100% by wt. water. The compn. is at an acidic pH. The compn. may include minor amts. of further conventional additives.

RE.CNT 7 THERE ARE 7 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

IT 64-17-5, Ethanol, uses 77-92-9, Citric acid, uses 79-14-1, Glycolic acid, uses 110-94-1, Glutaric acid 112-34-5, Dowanol DB 151-21-3, Sodium lauryl **sulfate**, uses 1643-20-5, Ammonyx LO 5329-14-6, Sulfamic acid 6915-15-7, Malic acid 7664-38-2, **Phosphoric acid**, uses
RL: MOA (Modifier or additive use); USES (Uses)
(ready-to-use aq. hard surface cleaning and disinfecting compns. contg. **hydrogen peroxide** and)

L3 ANSWER 62 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2000:201116 CAPLUS
DN 132:238745

TI Stable compositions for removing stains from fabrics and carpets and inhibiting the resoiling of same

IN Micciche, Robert P.; Lynch, Ann Marie; Tripathi, Uma
PA Playtex Products, Inc., USA
SO U.S., 5 pp.
CODEN: USXXAM

DT Patent
LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 6043209	A	20000328	US 1998-3272	19980106
PRAI	US 1998-3272		19980106		

AB A compn. includes a water miscible org. solvent, a peroxygen compd., a surfactant, a polymeric or copolymeric soil resist, and a fluorinated hydrocarbon soil resist. Thus, a cleaning agent contained water 92.87, **H2O2** 3.00, an acrylate copolymer 0.60, Na lauryl **sulfate** 0.60, propylene glycol Me ether 1.00, dipropylene glycol Me ether 1.00, Na lauroyl sarcosinate 0.23, lauramine oxide 0.07, C11-15 Pareth 7 0.25, **Dequest** 2010 0.12, fragrance 0.15, Zoner 5180 0.03, and Surcide-D 0.08%.

RE.CNT 72 THERE ARE 72 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB A compn. includes a water miscible org. solvent, a peroxygen compd., a surfactant, a polymeric or copolymeric soil resist, and a fluorinated hydrocarbon soil resist. Thus, a cleaning agent contained water 92.87, **H2O2** 3.00, an acrylate copolymer 0.60, Na lauryl **sulfate** 0.60, propylene glycol Me ether 1.00, dipropylene glycol Me ether 1.00, Na

lauroyl sarcosinate 0.23, lauramine oxide 0.07, C11-15 Pareth 7 0.25, Dequest 2010 0.12, fragrance 0.15, Zoner 5180 0.03, and Surcide-D 0.08%.

L3 ANSWER 63 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2000:712940 CAPLUS
DN 133:286190
TI Compositions for dyeing and decoloring keratin fibers
IN Matsunaga, Kenichi
PA Kao Corp., Japan
SO Jpn. Kokai Tokkyo Koho, 7 pp. ~
CODEN: JKXXAF
DT Patent
LA Japanese
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 2000281543	A2	20001010	JP 2000-20591	20000128
	JP 3343232	B2	20021111		
PRAI	JP 1999-21116	A	19990129		

AB The invention relates to a compn. for dyeing or decoloring keratin fibers while providing improved keratin fiber-conditioning effect, wherein the compn. contains (1) an oxidizing agent (2) a nonionic surfactant contg. C8-36 branched acyl, branched alkyl, or branched alkenyl group, and having HLB of 2-12, wherein the compn. has a pH of 7-12. A hair dye 1st agent contg. monomethyl branched isostearyl glyceryl ether (HLB = 5.3) 1.5, pentaerythritol glyceryl isostearyl ether (HLB = 9) 4, toluene-2,5-diamine 0.5, 5-amino-o-cresol 0.6, 28 % ammonia 6, ethanol 15, propylene glycol 10, oleic acid 6, oleic acid diethanol amide 8, polyoxyethylene octyldodecyl ether 10, ammonium chloride q.s. to pH = 9.5, sodium sulfite 0.5, tetrasodium edetate 0.1, and water q.s. to 100 %, and a 2nd agent contg. cetanol 2, glycerin 1, sodium alkyl sulfate 1, H2O2 6, Me paraben 0.1, phosphoric acid q.s. to pH = 3.5, and water q.s. to 100 % were prepd.

AB The invention relates to a compn. for dyeing or decoloring keratin fibers while providing improved keratin fiber-conditioning effect, wherein the compn. contains (1) an oxidizing agent (2) a nonionic surfactant contg. C8-36 branched acyl, branched alkyl, or branched alkenyl group, and having HLB of 2-12, wherein the compn. has a pH of 7-12. A hair dye 1st agent contg. monomethyl branched isostearyl glyceryl ether (HLB = 5.3) 1.5, pentaerythritol glyceryl isostearyl ether (HLB = 9) 4, toluene-2,5-diamine 0.5, 5-amino-o-cresol 0.6, 28 % ammonia 6, ethanol 15, propylene glycol 10, oleic acid 6, oleic acid diethanol amide 8, polyoxyethylene octyldodecyl ether 10, ammonium chloride q.s. to pH = 9.5, sodium sulfite 0.5, tetrasodium edetate 0.1, and water q.s. to 100 %, and a 2nd agent contg. cetanol 2, glycerin 1, sodium alkyl sulfate 1, H2O2 6, Me paraben 0.1, phosphoric acid q.s. to pH = 3.5, and water q.s. to 100 % were prepd.

L3 ANSWER 64 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 2000:150487 CAPLUS
DN 132:180289
TI Preparation of oximes using hydrogen peroxide stabilizers
IN Saizawa, Chiharu; Nomura, Toshihiro; Nishimura, Yoshio; Matsuya, Naoko
PA Mitsubishi Gas Chemical Co., Ltd., Japan
SO Jpn. Kokai Tokkyo Koho, 4 pp.
CODEN: JKXXAF
DT Patent
LA Japanese
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 2000072738	A2	20000307	JP 1998-247281	19980901
PRAI	JP 1998-247281		19980901		

OS CASREACT 132:180289

AB Oximes are prepd. by treatment of ketones, quinones, or aldehydes with NH₃ and H₂O₂ contg. 1-1000 ppm (based on H₂O₂) alkali metal-free stabilizers in the presence of titanosilicate catalysts. MEK was autoclaved with aq. NH₃, titanosilicate, and aq. H₂O₂ cong. 30 ppm H₃PO₄ at 75.degree. in Me₃COH to give 80.9% MEK oxime.

ST oxime prepn titanosilicate catalyst; ammoximation aldehyde ketone quinone titanosilicate catalyst; MEK oxime prepn ammoximation; **phosphoric acid stabilizer hydrogen peroxide** ammoximation; hydrogen peroxide ammonia ammoximation catalyst titanosilicate

IT Chelating agents
(**sulfonic acid-type**; prepn. of oximes by ammoximation using **hydrogen peroxide stabilizers**)

IT 6419-19-8, **Dequest 2000**
RL: NUU (Other use, unclassified); USES (Uses)
(prepn. of oximes by ammoximation using **hydrogen peroxide stabilizers**)

IT 2466-09-3, Pyrophosphoric acid 7664-38-2, **Phosphoric acid**, uses 10124-31-9, Ammonium phosphate 22690-73-9, Ammonium pyrophosphate
RL: NUU (Other use, unclassified); USES (Uses)
(stabilizer; prepn. of oximes by ammoximation using **hydrogen peroxide stabilizers**)

L3 ANSWER 65 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2000:835429 CAPLUS

DN 133:351817

TI Cleaning formulation for textile laundering

IN Michel, Jacques; Chevalier, Olivier

PA Yplon S.A., Belg.

SO Brit. UK Pat. Appl., 20 pp.
CODEN: BAXXDU

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	GB 2346153	A1	20000802	GB 2000-2099	20000131
	GB 2346153	B2	20030423		
	EP 1038947	A2	20000927	EP 2000-300750	20000131
	EP 1038947	A3	20020102		
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO				
PRAI	GB 1999-1957	A	19990129		

AB A formulation having pH <6 comprises a surface active agent, H₂O₂, and a pH jump system contg. a borate and polyhydroxy compd. On diln. it forms an alk. soln. suitable for peroxide bleaching. Thus, a heavy duty laundry liq. was prepd. from H₂O 39.28, borax decahydrate 5.0, NaOH 1.2, sorbitol 15.0, Aromox BW220 (C₁₂ amine oxide surfactant) 8.33, Lutensol AO 7 12.5, Lutensol AO 3 5.0, Irganox L 134 (amine free radical scavenger) 0.05, Dequest 2060S 0.5, peroxclean 12.14, and Durazym 16.0L 1.00 parts.

IT **Sulfonates**
RL: TEM (Technical or engineered material use); USES (Uses)
(alkanesulfonates; cleaning formulation for textile laundering contg. surface active agent and **hydrogen peroxide** and pH jump system)

IT **Sulfonic acids**, uses
RL: TEM (Technical or engineered material use); USES (Uses)
(arenesulfonic, salts; cleaning formulation for textile laundering contg. surface active agent and **hydrogen peroxide** and pH jump system)

IT Enzymes, uses
Sulfonates

RL: MOA (Modifier or additive use); USES (Uses)
(cleaning formulation for textile laundering contg. surface active agent and **hydrogen peroxide** and pH jump system)

IT 50-70-4, Sorbitol, uses 1300-72-7, Sodium xylene **sulfonate** 1303-96-4, Borax decahydrate 11128-98-6, Ammonium borate 12045-88-4, Tincalconite (B₄Na₂O₇·5H₂O) 15827-60-8, **Dequest** 2060S 22042-96-2, **Dequest** 2066 26183-44-8 27176-87-0, Dodecyl benzene **sulfonic acid** 305807-97-0, Irganox L 134

RL: TEM (Technical or engineered material use); USES (Uses)
(cleaning formulation for textile laundering contg. surface active agent and **hydrogen peroxide** and pH jump system)

L3 ANSWER 66 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 2000:749450 CAPLUS

DN 133:283719

TI Purification of aqueous hydrogen peroxide solutions by passage through silica-bound complexing agents

IN Guillard, Roger; Denat, Franck; Pernelet, Olivier; Tripier, Raphael; Ledon, Henry; Devos, Catherine

PA L'Air Liquide Societe Anonyme pour L'Etude et l'Exploitation des ProcedesGeo, Fr.

SO Fr. Demande, 35 pp.

CODEN: FRXXBL

DT Patent

LA French

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	FR 2789380	A1	20000811	FR 1999-1499	19990209
	FR 2789380	B1	20010309		
PRAI	FR 1999-1499		19990209		

AB Silica gel-supported coordinative chelating nitrogen compds. for prepn. of ultrapure solns. of hydrogen peroxide are of general formula A1-R1-Si(X1X2)-O-(silica gel), in which SiX1X2 is a bridging or coupling group [X1 and X2 = H, C1-4-alkyl, halogen, or OR2 (OR2 = H or C1-4-alkyl or bound to silica gel)]; R1 is a divalent satd. or unsatd. C1-10-hydrocarbylene which can be substituted or interrupted by divalent structures, such as O, S, -O-C(:O)-, -N(R3)-C(:O)-, or -N(R3)-, in which R6 = H, C1-6-aliph, benzyl, or phenethyl (or substituted with Cl, OH, or OR4 (R4 = C1-4-alkyl, benzyl, or phenethyl)); and A1 is the chelating group (e.g., a polyazamacrocyclic or a nitrogen-substituted aliph. chain, contg. addnl. coordinative functional groups). The chelating or sequestering agents can be either synthesized by attaching the silyl coupling agent to the nitrogen macrocycle or complexing agent followed by grafting onto end-capped silica gel, or by treating the silica gel with the coupling agent followed by grafting the nitrogen macrocycle or complexing agent onto the treated silica gel. Treatment of aq. hydrogen peroxide through adsorbent beds results in H₂O₂ solns. contg. <10 ppb of heavy metal ions.

IT 7429-90-5, Aluminum, processes 7439-89-6D, Iron, cations, processes 7440-31-5D, Tin, cations, processes 14265-44-2, Phosphate, processes 14797-55-8, Nitrate, processes 14808-79-8, **Sulfate**, processes 22537-23-1, Aluminum(3+), processes

RL: REM (Removal or disposal); PROC (Process)
(removal of; purifn. of aq. **hydrogen peroxide** solns. by passage through silica-bound complexing agents)

IT 1184-84-5DP, Vinylsulfonic acid, reaction products with (3-chloropropyl)triethoxysilane, chlorotrimethylsilane, and silica gel, derivs. 1746-03-8DP, **Phosphonic acid**, vinyl-, reaction products with (3-chloropropyl)triethoxysilane, chlorotrimethylsilane, and silica gel, derivs. 299199-34-1DP, reaction products with silane coupling agents, silane end-capping agents, and silica gel

RL: RCT (Reactant); SPN (Synthetic preparation); PREP (Preparation); RACT

(Reactant or reagent)
(sequestering agents; purifn. of aq. **hydrogen peroxide** solns. by passage through silica-bound complexing agents)

L3 ANSWER 67 OF 158 WPIDS (C) 2003 THOMSON DERWENT
AN 2001-040882 [05] WPIDS
DNN N2001-030511 DNC C2001-011805
TI Paper product useful in treatment of illness or injury, waste disposal, food packaging, storage is impregnated with a chemical material.
DC D22 E19 E37 F09 P34
IN AAMODT, J A; COLVIN, J W; AAMODT, J; COLVIN, J
PA (AAMO-I) AAMODT J; (AAMO-I) AAMODT J A; (COLV-I) COLVIN J W; (COLV-I) COLVIN J
CYC 90
PI WO 2000066185 A1 20001109 (200105)* EN 36p
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW NL
OA PT SD SE SL SZ TZ UG ZW
W: AE AL AM AT AU AZ BA BB BG BR BY CA CH CN CR CU CZ DE DK DM EE ES
FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ LC LK LR LS
LT LU LV MA MD MG MK MN MW MX NO NZ PL PT RO RU SD SE SG SI SK SL
TJ TM TR TT TZ UA UG UZ VN YU ZA ZW
AU 2000044940 A 20001117 (200111)
US 6325969 B1 20011204 (200203)
US 2002110483 A1 20020815 (200256)
ADT WO 2000066185 A1 WO 2000-US11263 20000427; AU 2000044940 A AU 2000-44940
20000427; US 6325969 B1 US 1999-302937 19990430; US 2002110483 A1 Cont of
US 1999-302937 19990430, US 2001-6192 20011203
FDT AU 2000044940 A Based on WO 200066185
PRAI US 1999-302937 19990430; US 2001-6192 20011203
AB WO 200066185 A UPAB: 20010124
NOVELTY - A method for impregnating a porous paper product comprises exposing the porous product, which has pores and a surface to at least one chemical.
DETAILED DESCRIPTION - INDEPENDENT CLAIMS are included for the following:
(A) method for sterilizing an area comprising the steps of:
(a) impregnating a porous paper product with a chemical material,
(b) placing the product in the area and
(c) causing reaction of the impregnated chemical material to produce a biocidal compound;
(B) a composition producing chlorine dioxide comprising a porous paper product impregnated with an acid or metal salt and at least one sodium chlorite and sodium chlorate; and
(C) a composition producing a peracid comprising a porous paper product impregnated with hydrogen peroxide and an acid.
USE - In a variety of sources e.g. wood pulp, kenaf, flax or hemp (claimed). Provides sterile environment for activities such as food packaging or storage, treatment of illness or injury or waste disposal; in industries and household tasks.
ADVANTAGE - Inhibits the growth of microorganisms, cheap and easy to manufacture. The diffusion of volatile biocidal chemical out of pores in the paper creates no growth zone on and immediately surrounds the impregnated paper. Cuts down on noxious odors and danger of infection to waste-handlers. Conveys odor controlling, biocidal and chlorine dioxide destruction properties.
Dwg.0/1
TECH. . . .
with sodium chlorite and acetic acid. The porous paper product in the form of sheets or pellets is impregnated with **hydrogen peroxide** and aqueous d-limonene (5%) or an acid (preferably acetic acid (99%), propionic acid, citric acid (50%), benzoic acid, **phosphoric acid**, lactic acid, butyric acid, pentenoic acid, succinic acid, glutaric acid, sorbic acid and glycolic acid). The

paper is impregnated with. . . form. The step of exposing the surface of the product to at least one chemical comprises exposing the surface to **hydrogen peroxide** or metal salt (preferably ferric chloride or ferric **sulfate**) and an acid or d-limonene or at least one chemical (preferably potassium hydroxide, sodium sulfite and sodium bisulfite) with chlorine. . .

L3 ANSWER 68 OF 158 WPIDS (C) 2003 THOMSON DERWENT
AN 2001-015725 [02] WPIDS
DNC C2001-004171
TI An anti-tartar dental product comprising a combination of a water soluble calcium phosphate at a pH of less than 7 and separately stored combination of an alkaline material and an anti-carries fluoride ion source with a pH of more than 7.5.
DC B05 D21 E19 E37
IN BARROW, S R; LEE, G J; WILLIAMS, D R; ZIEMKIEWICZ, A G; BARROW, S; WILLIAMS, D; ZIEMKIEWICZ, A
PA (LEEG-I) LEE G J; (UNIL) UNILEVER HOME & PERSONAL CARE USA DIV CO; (CHEO) CHESEBROUGH PONDS USA CO DIV CONOPCO INC; (HIND-N) HINDUSTAN LEVER LTD; (UNIL) UNILEVER NV; (UNIL) UNILEVER PLC
CYC 91
PI WO 2000062749 A1 20001026 (200102)* EN 31p
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW NL
OA PT SD SE SL SZ TZ UG ZW
W: AE AL AM AT AU AZ BA BB BG BR BY CA CH CN CR CU CZ DE DK DM EE ES
FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ LC LK LR LS
LT LU LV MA MD MG MK MN MW MX NO NZ PL PT RO RU SD SE SG SI SK SL
TJ TM TR TT TZ UA UG UZ VN YU ZA ZW
AU 2000034326 A 20001102 (200107)
US 6207139 B1 20010327 (200119)
US 6248310 B1 20010619 (200137)
EP 1178773 A1 20020213 (200219) EN
R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT
RO SE SI
ADT WO 2000062749 A1 WO 2000-EP2758 20000328; AU 2000034326 A AU 2000-34326
20000328; US 6207139 B1 Provisional US 1999-129779P 19990416, US
1999-395064 19990913; US 6248310 B1 Provisional US 1999-129779P 19990416,
Div ex US 1999-395064 19990913, US 2000-538564 20000329; EP 1178773 A1 EP
2000-912656 20000328, WO 2000-EP2758 20000328
FDT AU 2000034326 A Based on WO 200062749; EP 1178773 A1 Based on WO 200062749
PRAI US 1999-129779P 19990416; US 1999-395064 19990913; US 2000-538564
20000329
AB WO 200062749 A UPAB: 20010124
NOVELTY - An anti-tartar dental product comprising a combination of a water soluble calcium phosphate at a pH of less than 7 and separately stored combination of an alkaline material and an anti-carries fluoride ion source with a pH of more than 7.5, is new.
DETAILED DESCRIPTION - An anti-tartar dental product comprising a container, an oral preparation of a formulation comprising 0.01 to 30% of a water soluble calcium phosphate and/or monolithic combination of water soluble calcium and phosphate salts at a pH of less than 7 and a composition including 0.01 to 30% of an alkaline material and an anti-carries fluoride ion source with a pH of greater than 7.5 stored separately from the first composition, and instructions for use, is new.
An INDEPENDENT CLAIM is also included for a method for controlling dental tartar comprising brushing the teeth with the product.
ACTIVITY - Antimicrobial.
The first component typically comprised glycerin (40%), pluronic F-127 (20%), monocalcium phosphate monohydrate (1.6%), 35% **hydrogen peroxide** (4.285%), **phosphoric acid** (0.4%), FD and C blue no.1 (0.01%) and water. The second component typically comprised 70% sorbitol (47%), hydrated silica (15%), sodium hydrogen carbonate (10%), sylox 15X (6%), polyethylene glycol 1450 (5%), ethanol (2.84%), sodium lauryl sulfate (2.98%), flavor (1.1%),

cellulose gum (0.8%), sodium saccharin (0.54%), menthol (0.5%), sodium fluoride (0.44%), titanium dioxide (0.30%) and water. Use of the products reduced calculus formation by up to 44%.

MECHANISM OF ACTION - None given.

USE - The composition is useful for controlling dental tartar.

Dwg.0/0

AB

the product.

ACTIVITY - Antimicrobial.

The first component typically comprised glycerin (40%), pluronic F-127 (20%), monocalcium phosphate monohydrate (1.6%), 35% hydrogen peroxide (4.285%), phosphoric acid (0.4%), FD and C blue no.1 (0.01%) and water. The second component typically comprised 70% sorbitol (47%), hydrated silica (15%),.

TECH.

pH of the second composition is preferably 7.2 to 11. The pH of the first composition results from inclusion of hydrogen peroxide, inorganic acids and/or 2-20C carboxylic acids. The monolithic combination of water soluble calcium salts is preferably calcium chloride, calcium sulfate or calcium acetate and the respective phosphate salts are preferably sodium phosphate, ammonium phosphate or sodium ammonium phosphate. The composition.

L3 ANSWER 69 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-069699 [08] WPIDS

DNN N2001-052662 DNC C2001-019276

TI Manufacture of conductive polymeric nanocomposite for batteries, involves performing oxidative polymerization of reaction mixture containing layered silicate, to form composite having specific conductivity.

DC A25 A85 E19 L03 U11 U12

IN LIAO, C; LIN, L K

PA (INTE-N) IND TECHNOLOGY RES INST; (KOGY-N) ZH KOGYO GIJUTSU KENKYUHI

CYC 3

PI US 6136909 A 20001024 (200108)* 5p

DE 19930947 A1 20001130 (200108)

JP 2001011310 A 20010116 (200119) 5p

DE 19930947 C2 20010531 (200130)

JP 3187025 B2 20010711 (200140) 5p

ADT US 6136909 A US 1999-334633 19990617; DE 19930947 A1 DE 1999-19930947 19990705; JP 2001011310 A JP 1999-219879 19990803; DE 19930947 C2 DE 1999-19930947 19990705; JP 3187025 B2 JP 1999-219879 19990803

FDT JP 3187025 B2 Previous Publ. JP 2001011310

PRAI TW 1999-108290 19990520

AB US 6136909 A UPAB: 20010207

NOVELTY - A reaction mixture comprising water, an aniline monomer, a protonic acid, an oxidizing agent and a layered silicate is obtained. The mixture is subjected to oxidative polymerization, to form a nanocomposite having layered silicate dispersed in a polymeric matrix of polyaniline. The conductivity of the nanocomposite is 0.1 S/cm.

DETAILED DESCRIPTION - A reaction mixture comprising water, an aniline monomer, a protonic acid, an oxidizing agent and a layered silicate is obtained. The silicate is subjected to an acid treatment or is intercalated with polyethylene glycol. The reaction mixture is subjected to oxidative polymerization, to form a nanocomposite having layered silicate dispersed in a polymeric matrix of polyaniline. The conductivity of the nanocomposite is 0.1 S/cm.

USE - The nanocomposite is used for batteries, displays, optics, sensors, light emitting diodes and in the aeronautical industry.

ADVANTAGE - A highly conductive polymeric nanocomposite containing layered silicate is obtained. Stiffness, strength and heat resistance of the nanocomposite are enhanced. Moisture absorption, flammability and permeability of the nanocomposite are reduced.

Dwg.0/2

TECH.

of silicon fluoride, sodium fluoride and/or lithium fluoride. The fluoromica is treated with sulfuric acid, hydrochloric acid, nitric acid or **phosphoric acid**.

Preferred Oxidizing Agent: The oxidizing agent is ammonium persulfate, ferric chloride, ferrous chloride or **hydrogen peroxide**

Preferred Protonic Acid: The protonic acid is hydrochloric acid, sulfuric acid, **phosphoric acid**, organic **sulfonic acid** or organic **phosphorous-containing acids**.

TECHNOLOGY FOCUS - POLYMERS - Preferred Glycol: The molecular weight of polyethylene glycol is 100-50000.

Preferred Monomer: The aniline monomer is aniline, . . .

L3 ANSWER 70 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-183908 [19] WPIDS

DNC C2001-055186

TI Dyeing composition for keratin fibers, especially human hair, containing xanthene derivative dye, e.g. pyrogallol red, giving strong, fast shades even in the absence of oxidizing agents.

DC A60 B07 D21 E19 E23 F06

IN MEINIGKE, B; MOELLER, H

PA (HENK) HENKEL KGAA

CYC 31

PI DE 19926377 A1 20001214 (200119)* 13p

WO 2000076466 A1 20001221 (200119) DE

RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE

W: AU BR CA CN CZ HU JP NO PL RU SK US VN

AU 2000058106 A 20010102 (200121)

ADT DE 19926377 A1 DE 1999-19926377 19990610; WO 2000076466 A1 WO 2000-EP5044 20000602; AU 2000058106 A AU 2000-58106 20000602

FDT AU 2000058106 A Based on WO 200076466

PRAI DE 1999-19926377 19990610

AB DE 19926377 A UPAB: 20010405

NOVELTY - A composition (A) for dyeing keratin fibers contains a 9-phenyl-xanthene derivative (I) or its salt as dyeing component.

DETAILED DESCRIPTION - A composition (A) for dyeing keratin fibers contains at least one xanthene derivative of formula (I) or its salt as dyeing component.

R1 = H, 1-4C carboxyl (sic) or sulfo;

R2-R5 = H, 1-4C alkyl, 1-4C alkoxy or OH;

(i) X1 = OH (optionally in alkali metal, alkaline earth metal or ammonium salt form); and

X2 = O; or

(ii) X1 = amino (optionally mono- or disubstituted by methyl or forming a fused 5-7 membered ring in combination with R2 and/or R3); and

X2 = -N+A1A2.U-;

A1, A2 = H or Me; or form a fused 5-7 membered ring in combination with R4 and/or R5;

U- = 1 equivalent of a halide, sulfate, sulfonate or carboxylate anion.

INDEPENDENT CLAIMS are included for:

(i) the use of (I) or their salts as dyeing components in hair-dyeing compositions; and

(ii) a method for dyeing keratin fibers, involving applying a composition containing (a) (I) (or its salt), at least one compound (II) having primary or secondary amino, OH or active CH groups and conventional cosmetic components to the fibers, leaving the composition in contact with the fibers for a suitable time (usually ca. 30 minutes) and rinsing or shampooing. (II) is selected from primary or secondary aliphatic amines, N-containing heterocycles, aminoacids, oligopeptides formed from 2-9 aminoacids, aromatic hydroxy compounds and active CH compounds.

USE - For dyeing keratin fibers, especially human hair (claimed), but

also e.g. wool, skins or feathers. (I) may also be applicable to the dyeing of other natural fibers (e.g. cotton, jute, sisal, linen or silk), modified natural fibers (e.g. regenerated cellulose or nitro-, alkyl-, hydroxyalkyl- or acetyl cellulose) or synthetic fibers (e.g. polyamide, polyacrylonitrile, polyurethane or polyester). (I) generally give red-orange, red-brown or violet shades when used alone, but in combination with (II) may give orange, red, brown-orange, medium brown, violet, dark violet, blue-black or black colorations.

ADVANTAGE - (I) give a wide range of bright, deep colorations, even in the absence of oxidizing agents. The dyeings have good greyness covering and fastness properties (at least comparable with those obtained using conventional oxidation dyes), e.g. good light-, wash- and rub-fastness.

Dwg.0/0

TECH.

%); salts selected from alkali metal, alkaline earth metal, aluminum, manganese, iron, cobalt, copper, zinc or ammonium formates, carbonates, halides, **sulfates**, butyrates, valerates, caproates, acetates, lactates, glycolates, tartrates, citrates, gluconates, propionates, phosphates and **phosphonates**; oxidizing agents, specifically **hydrogen peroxide**, at 0.01-6 wt. % based on the ready-for-use solution; and/or anionic, zwitterionic or nonionic surfactants.

The formulation may contain a. . .

L3 ANSWER 71 OF 158 CAPLUS COPYRIGHT 2003 ACS DUPLICATE 4

AN 1999:670046 CAPLUS

DN 131:276767

TI Compositions containing peroxides for lightening and highlighting hair

IN Newell, Gerald; Fowler, Margie

PA Helene Curtis, Inc., USA

SO U.S., 6 pp.

CODEN: USXXAM

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 5968486	A	19991019	US 1997-854829	19970512
PRAI	US 1997-854829		19970512		

AB There is described a shampoo compn. for lightening and highlighting hair which comprises (1) a peroxygen compd; and (2) an anionic sulfonate; said compn. having a pH less than 5. A method for lightening and highlighting hair comprises: (1) applying water to the hair; (2) applying to the hair a lightening and highlighting effective amt. of the above; (3) lathering; and (4) rinsing the hair with water. A hair lightening prepn. (pH 3) contained Polyquaternium-10 0.525, Polymer JR 30M 0.3, Na C14-16 olefin **sulfonate** 35, lauryl alc. 0.5, lauramine oxide 3.7, **phosphoric acid** 1, DMDM hydantoin 0.1, NaCl 0.8, fragrances 0.6, H2O2 (35 %) 4.6, and water q.s. to 100 %.

RE.CNT 6 THERE ARE 6 CITED REFERENCES AVAILABLE FOR THIS RECORD

ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB There is described a shampoo compn. for lightening and highlighting hair which comprises (1) a peroxygen compd; and (2) an anionic sulfonate; said compn. having a pH less than 5. A method for lightening and highlighting hair comprises: (1) applying water to the hair; (2) applying to the hair a lightening and highlighting effective amt. of the above; (3) lathering; and (4) rinsing the hair with water. A hair lightening prepn. (pH 3) contained Polyquaternium-10 0.525, Polymer JR 30M 0.3, Na C14-16 olefin **sulfonate** 35, lauryl alc. 0.5, lauramine oxide 3.7, **phosphoric acid** 1, DMDM hydantoin 0.1, NaCl 0.8, fragrances 0.6, H2O2 (35 %) 4.6, and water q.s. to 100 %.

IT **Sulfonates**

RL: BUU (Biological use, unclassified); BIOL (Biological study); USES

(Uses)

(1-alkene; hair-lightening prepns. contg. **hydrogen peroxide** and olefin **sulfonates**)

IT **Sulfonic acids**, biological studies

RL: BUU (Biological use, unclassified); BIOL (Biological study); USES
(Uses)

(C14-16-1-alkenesulfonic, sodium salts; hair-lightening prepns. contg. **hydrogen peroxide** and olefin **sulfonates**)

IT Hair preparations

(bleaches; hair-lightening prepns. contg. **hydrogen peroxide** and olefin **sulfonates**)

IT 1643-20-5, Lauramine oxide

RL: BUU (Biological use, unclassified); BIOL (Biological study); USES
(Uses)

(foam enhancer; hair-lightening prepns. contg. **hydrogen peroxide** and olefin **sulfonates**)

IT 7722-84-1, **Hydrogen peroxide**, biological studies

RL: BUU (Biological use, unclassified); BIOL (Biological study); USES
(Uses)

(hair-lightening prepns. contg. **hydrogen peroxide** and olefin **sulfonates**)

IT 7664-38-2, **Phosphoric acid**, biological studies

7664-93-9, Sulfuric acid, biological studies

RL: BUU (Biological use, unclassified); BIOL (Biological study); USES
(Uses)

(pH adjuster; hair-lightening prepns. contg. **hydrogen peroxide** and olefin **sulfonates**)

IT 112-53-8, Lauryl alcohol

RL: BUU (Biological use, unclassified); BIOL (Biological study); USES
(Uses)

(viscosity enhancer; hair-lightening prepns. contg. **hydrogen peroxide** and olefin **sulfonates**)

L3 ANSWER 72 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1999:811317 CAPLUS

DN 132:50891

TI Polyester resin compositions

IN Okamoto, Chihiro; Taniguchi, Shunro

PA Kuraray Co., Ltd., Japan

SO PCT Int. Appl., 59 pp.

CODEN: PIXXD2

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9965987	A1	19991223	WO 1998-JP2627	19980615
	W: AU, CA, CN, KR, NO, SG, US				
	RW: AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
	CA 2334665	AA	19991223	CA 1998-2334665	19980615
	AU 9876752	A1	20000105	AU 1998-76752	19980615
	AU 750299	B2	20020718		
	EP 1111006	A1	20010627	EP 1998-924624	19980615
	R: CH, DE, FR, GB, IT, LI, NL, SE				
	US 6429243	B1	20020806	US 2000-719174	20001215
PRAI	WO 1998-JP2627	A	19980615		

AB A polyester resin compn. (A) comprising a polyethylene terephthalate copolymer comprising 5 to 40 % by mole of naphthalenedicarboxylic acid units (copolymer PET), a cobalt compd. and/or a manganese compd., an olefin polymer, and a predetd. amt. of a compatibilizing agent; and another polyester resin compn. (B) contg. predetd. amts. of a cobalt compd. and a manganese compd. in the copolymer PET. The compns. are excellent in resistance to hydrogen peroxide, flavor barrier properties,

gas barrier properties, heat sealability, and mech. properties such as elongation and strength, so that they can be utilized in various applications including packaging materials for containers for beverages contg. a perfume component, such as fruit juice beverages.

RE.CNT 13 THERE ARE 13 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

IT 127-09-3, Sodium acetate 2180-18-9, Manganese acetate 5931-89-5,
Cobalt acetate 7664-38-2, **Phosphoric acid**, uses
7785-87-7, Manganese **sulfate** 10124-43-3, Cobalt
sulfate

RL: MOA (Modifier or additive use); USES (Uses)
(for prepg. polyester resin compns. excellent in resistance to
hydrogen peroxide, flavor barrier properties, gas
barrier properties, heat sealability, and mech. properties)

L3 ANSWER 73 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1999:614071 CAPLUS

DN 131:244857

TI Method for gently bleaching fabrics

IN Josa, Jaume; De Jorge, Elisabet; Arranz, Adolf

PA Henkel Kommanditgesellschaft Auf Aktien, Germany

SO PCT Int. Appl., 15 pp.

CODEN: PIXXD2

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9947632	A1	19990923	WO 1999-EP1411	19990304
	W: CZ, HU, JP, PL, SK, US				
	RW: AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
	DE 19810885	A1	19991007	DE 1998-19810885	19980313
	EP 1062313	A1	20001227	EP 1999-907595	19990304
	R: AT, BE, DE, ES, FR, GB, IT, NL				
	JP 2002506922	T2	20020305	JP 2000-536815	19990304
PRAI	DE 1998-19810885	A	19980313		
	WO 1999-EP1411	W	19990304		

AB The invention relates to a method for gently bleaching textile fabrics, esp. coloreds and delicates, by bringing them into contact with aq. prepn. contg. hydrogen peroxide. The method is characterized in that the agents are free of heavy metal and/or chloride ions. The absence of these ions prevents pitting of the metal walls of the washing machine. Sulfates are used instead of chlorides and red dyes are used instead of blue dyes to eliminate heavy metal ions in the bleaching agents. Optionally, the aq. bleaching agents contain N-, P-, or I-contg. inhibitors.

RE.CNT 7 THERE ARE 7 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

IT Amine oxides

RL: NUU (Other use, unclassified); USES (Uses)

(**phosphonic acid** derivs.; **hydrogen peroxide** bleaching of colored delicate fabrics in absence of heavy metal and(or) chloride ions)

IT 13598-36-2D, **Phosphonic acid**, amine oxide derivs.

RL: NUU (Other use, unclassified); USES (Uses)

(Sequion; **hydrogen peroxide** bleaching of colored delicate fabrics in absence of heavy metal and(or) chloride ions)

IT 139-13-9 7487-88-9, Magnesium **sulfate**, uses 7647-14-5,
Sodium chloride, uses 7790-28-5, Sodium periodate 13598-36-2,
Phosphonic acid

RL: NUU (Other use, unclassified); USES (Uses)

(**hydrogen peroxide** bleaching of colored delicate fabrics in absence of heavy metal and(or) chloride ions)

L3 ANSWER 74 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1999:326022 CAPLUS

DN 130:353959

TI O-substituted N,N-diacylhydroxylamine bleach activators and compositions
bleaching soiled fabrics and dishware

IN Miracle, Gregory Scot; Dykstra, Robert Richard

PA The Procter & Gamble Company, USA

SO PCT Int. Appl., 45 pp.

CODEN: PIXXD2

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9924537	A2	19990520	WO 1998-US23767	19981109
	WO 9924537	A3	19990729		
	W: BR, CA, CN, JP, MX, US				
	RW: AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
	CA 2309592	AA	19990520	CA 1998-2309592	19981109
	EP 1032631	A2	20000906	EP 1998-958488	19981109
	EP 1032631	B1	20021023		
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, PT, IE, FI				
	BR 9812782	A	20001003	BR 1998-12782	19981109
	JP 2001522866	T2	20011120	JP 2000-520533	19981109
	AT 226622	E	20021115	AT 1998-958488	19981109
	US 6291413	B1	20010918	US 2000-554203	20000510
	US 2001046953	A1	20011129	US 2001-861133	20010518
	US 6423676	B2	20020723		
	US 6514925	B1	20030204	US 2002-154005	20020523
PRAI	US 1997-64973P	P	19971110		
	WO 1998-US23767	W	19981109		
	US 2000-554203	A1	20000510		
	US 2001-861133	A3	20010518		

OS MARPAT 130:353959

AB The title activators R1CON(OR2)CO[CO]eXfR3 (X = O, NR16 and S; e = 0 or 1; f = 0 or 1; R16 = H and linear or branched, satd. or unsatd. C1-4-alkyl; and R1 = Ph or linear or branched chain, satd. or unsatd. C7-13-alkyl; R2 = branched or unbranched, satd. or unsatd. C1-10-alkyl; and R3 = linear or branched chain, satd. or unsatd. C1-12-alkyl) with hydrophilic/hydrophobic groups are prepd. for bleach compns. based on H2O2. An example tile cleaner contained bleach activator 5.0, H2O2 10, LAS 5.0, ethoxylated alkyl sulfate salt 1.5, amine oxide 1.0, Dequest 2060 0.5, citric acid 6.0%, HCl, and the balance water.

AB The title activators R1CON(OR2)CO[CO]eXfR3 (X = O, NR16 and S; e = 0 or 1; f = 0 or 1; R16 = H and linear or branched, satd. or unsatd. C1-4-alkyl; and R1 = Ph or linear or branched chain, satd. or unsatd. C7-13-alkyl; R2 = branched or unbranched, satd. or unsatd. C1-10-alkyl; and R3 = linear or branched chain, satd. or unsatd. C1-12-alkyl) with hydrophilic/hydrophobic groups are prepd. for bleach compns. based on H2O2. An example tile cleaner contained bleach activator 5.0, H2O2 10, LAS 5.0, ethoxylated alkyl sulfate salt 1.5, amine oxide 1.0, Dequest 2060 0.5, citric acid 6.0%, HCl, and the balance water.

L3 ANSWER 75 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1999:384070 CAPLUS

DN 131:33273

TI Carpet cleaning powders containing active oxygen sources

IN Jonke, Hermann; Ditze, Alexander; Poethkow, Daniela; Schaffrath, Norbert; Ifland, Werner

PA Henkel K.-G.a.A., Germany

SO Ger. Offen., 8 pp.

CODEN: GWXXBX

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 19753700	A1	19990610	DE 1997-19753700	19971204
	WO 9929817	A1	19990617	WO 1998-EP7616	19981125

W: JP, US

RW: AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE

EP 1036148	A1	20000920	EP 1998-965183	19981125
------------	----	----------	----------------	----------

R: AT, BE, DE, ES, FR, GB, IT, NL

JP 2001526302	T2	20011218	JP 2000-524394	19981125
---------------	----	----------	----------------	----------

PRAI DE 1997-19753700 A 19971204

WO 1998-EP7616 W 19981125

AB The title powders, which clean entire surfaces or stained spots without damaging the carpet and have good storage stability, contain powd. cellulose, active O sources, and no wood flour. An aq. slurry of powd. cellulose 43, H2O2 0.9, fatty alkyl sulfate 0.85, EtOH 3, and absorbent flocs 0.55% (plus phosphonic acid and perfume) removed red wine, coffee, and blueberry juice stains completely from polyamide and wool carpets when applied at 100 and 200 g/m2, resp.; vs. no removal in the absence of H2O2.

AB The title powders, which clean entire surfaces or stained spots without damaging the carpet and have good storage stability, contain powd. cellulose, active O sources, and no wood flour. An aq. slurry of powd. cellulose 43, H2O2 0.9, fatty alkyl sulfate 0.85, EtOH 3, and absorbent flocs 0.55% (plus phosphonic acid and perfume) removed red wine, coffee, and blueberry juice stains completely from polyamide and wool carpets when applied at 100 and 200 g/m2, resp.; vs. no removal in the absence of H2O2.

L3 ANSWER 76 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 1999-570607 [48] WPIDS

DNN N1999-420347 DNC C1999-166472

TI Compositions for removing polymeric coatings from hard surfaces, e.g. paints, varnishes, enamels, sealants or lacquers on aircraft, cars and buildings.

DC A35 A82 E19 G02 P43

IN FREDRICKSEN, J R; JASPER, J M; SMITH, P E; ZINNECKER, W R

PA (ELDO-N) ELDORADO CHEM CO INC

CYC 1

PI US 5954890 A 19990921 (199948)* 8p

ADT US 5954890 A Div ex US 1995-549495 19951027, US 1997-938222 19970926

FDT US 5954890 A Div ex US 5830836

PRAI US 1995-549495 19951027; US 1997-938222 19970926

AB US 5954890 A UPAB: 19991122

NOVELTY - A method of making a polymeric coating on a surface softened-to-release at ambient temperature comprises applying to the surface, a composition comprising hydrogen peroxide or an organic peroxide, water, a corrosion inhibitor and a thickener, such that the composition has a pH of 6.5-11.0 and minimal corrosiveness during use.

USE - For removing polymeric coatings, e.g. a paint, varnish, enamel, sealant or lacquer, on aircraft and cars at ambient temperature (all claimed) and also buildings. Polymeric coatings may be removed from any hard, nonporous surface, e.g. metal, certain types of stone and acrylic siding.

ADVANTAGE - The compositions use peroxide in a way that minimizes corrosion of metals and minimizes hazardous reactions. The shelf life of the product is maximized (e.g. a peroxide content of 16% after storage for 3 months), and significantly longer than that of prior art compositions, through the use of 2 storage-stable stock solutions, which are mixed inline immediately before use. The solution applied has greater activity for oxidation of coatings.

Dwg.0/1

TECH.

The peroxide solution may further contains a surfactant, a stabilizer or a chelator, or a supplementary activator. The peroxide is **hydrogen peroxide** or methyl ethyl ketone peroxide. The oxidation catalyst is sodium molybdate or manganese **sulfate**. Two preferred acid peroxide stock solutions, (A) and (B), comprise water (0.40 or 7.50 wt.% respectively), sodium ethylenediamine tetraacetate (0.50 wt.% for both), 75% **phosphoric acid** (0.10 or 1.00 wt.% respectively), gamma-butyrolactone (16.0 and 0 wt.% respectively), hexylene glycol (30.0 and 0 wt.% respectively), N-methyl-2-pyrrolidone (0. . .

L3 ANSWER 77 OF 158 WPIDS (C) 2003 THOMSON DERWENT

AN 1999-231366 [20] WPIDS

DNC C1999-068150

TI Peroxygen bleach composition containing salt-free zwitterionic betaine surfactant to reduce fabric damage and color loss.

DC D25 E19

IN ANTONIOLI, S; DEL DUCA, V; GIUNTI, S

PA (PROC) PROCTER & GAMBLE CO

CYC 83

PI EP 908510 A1 19990414 (199920)* EN 14p

R: AL AT BE CH DE DK ES FI FR GB GR IE IT LI LT LU LV MC NL PT RO SE SI

WO 9918179 A1 19990415 (199922) EN

RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW NL OA PT SD SE SZ UG ZW

W: AL AM AT AU AZ BA BB BG BR BY CA CH CN CU CZ DE DK EE ES FI GB GE GH GM HR HU ID IL IS JP KE KG KP KR KZ LC LK LR LS LT LU LV MD MG MK MN MW MX NO NZ PL PT RO RU SD SE SG SI SK SL TJ TM TR TT UA UG US UZ VN YU ZW

AU 9910726 A 19990427 (199936)

EP 1021505 A1 20000726 (200037) EN

R: AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU NL PT SE

BR 9812902 A 20000808 (200044)

US 6316400 B1 20011113 (200173)

JP 2001519458 W 20011023 (200202)

37p

ADT EP 908510 A1 EP 1997-870151 19971008; WO 9918179 A1 WO 1998-US21240

19981007; AU 9910726 A AU 1999-10726 19981007; EP 1021505 A1 EP

1998-953321 19981007, WO 1998-US21240 19981007; BR 9812902 A BR 1998-12902

19981007, WO 1998-US21240 19981007; US 6316400 B1 WO 1998-US21240

19981007, US 2000-509915 20000619; JP 2001519458 W WO 1998-US21240

19981007, JP 2000-514979 19981007

FDT AU 9910726 A Based on WO 9918179; EP 1021505 A1 Based on WO 9918179; BR

9812902 A Based on WO 9918179; US 6316400 B1 Based on WO 9918179; JP

2001519458 W Based on WO 9918179

PRAI EP 1997-870151 19971008

AB EP 908510 A UPAB: 20011203

NOVELTY - Addition of salt-free zwitterionic betaine surfactant to peroxygen bleach to reduce tensile strength loss and color damage to fabric

DETAILED DESCRIPTION - A liquid bleaching composition comprises a peroxygen bleach (A) and a salt-free zwitterionic betaine surfactant (B).

USE - The composition is used for bleaching fabrics (claimed).

ADVANTAGE - Addition of (B) reduces loss of tensile strength and color damage to fabrics by the peroxide bleach (claimed). The composition also has excellent stain removing activity on various stains (including greasy) and is physically and chemically stable on prolonged storage.

Dwg.0/0

TECH.

a ring;

R4 = a group which, with N+forms a hydrophilic group (especially 1-10C alkylene, hydroxyalkylene or polyalkoxy);

X = carboxylate or **sulfonate**.

provided that there is no more than 1 OH group in any (C(Rb)2).

(A) is H₂O₂, or a water soluble source of it, e.g. percarbonates, persulfates, persulfates, peroxyacids, hydroperoxides and/or aromatic and aliphatic diacyl peroxides (preferably t-Bu hydroperoxide, cumyl hydroperoxide, 2,4,4-trimethylpentyl-2-hydroperoxide, di-isopropylbenzenemmonohydroperoxide, t-amylhydroperoxide, 2,5-dimethylhexane-2,5-dihydroperoxide, dilauroyl peroxide, didecanoyl peroxide, dimyristoyl peroxide, benzoyl peroxide and/or especially H₂O₂). The composition comprises 0.01-20 (especially 2-10) wt. %

(A).

The composition may also comprise up to 30 (especially 0.5-20) wt. % an ethoxylated nonionic surfactant of formula (II) and up to 5 (especially 0.01-1.5) wt. % chelating agent (preferably a **phosphonate**, amino carboxylate, other carboxylate or polyfunctionally substituted aromatic chelating agent and/or ethylenediamine N,N'-disuccinic acid (especially amino aminotri(methylene phosphonic acid), diethylene-triamino-pentaacetic. . .

L3 ANSWER 78 OF 158 JAPIO COPYRIGHT 2003 JPO

AN 1999-315398 JAPIO

TI FORMATION OF TITANIUM ANODICALLY OXIDIZED FILM FOR PHOTOCATALYST

IN ITO SEISHIRO; TADA HIROAKI; KURAMOTO ATSUSHI

PA DAIWA HOUSE IND CO LTD

PI JP 11315398 A 19991116 Heisei

AI JP 1999-42857 (JP11042857 Heisei) 19990222

PRAI JP 1998-38436 19980220

SO PATENT ABSTRACTS OF JAPAN (CD-ROM), Unexamined Applications, Vol. 1999

AB PROBLEM TO BE SOLVED: To allow a film to exhibit photocatalytic operation even by ultraviolet rays and visible rays and to obtain antibacterial, deodorant and contamination preventive effects by immersing titanium into an electrolytic bath composed of sulfuric acid, **phosphoric acid** and **hydrogen peroxide**, increasing the voltage to the prescribed one by a prescribed direct constant current and holding it for a prescribed time to form a titanium anodically oxidized film essentially consisting of an anatase form.

SOLUTION: This is a method for forming a titanium anodically oxidized film for a photocatalyst exhibiting photocatalytic operation even by visible rays, in which titanium is immersed into an electrolytic bath composed of sulfuric acid, **phosphoric acid**, **hydrogen peroxide** and cobalt **sulfate**, the voltage is increased to the prescribed one by a prescribed direct constant current, and it is held for a prescribed time to form a TiO₂-CoO series titanium anodically oxidized film. Moreover, titanium is immersed into an electrolytic bath composed of sulfuric acid, **phosphoric acid**, **hydrogen peroxide** and zinc **sulfate**, the voltage is increased to the prescribed one, and it is held for a prescribed time to form a TiO₂-ZnO series titanium anodically oxidized film. The titanium anodically oxidized film to be formed is used for the interior and exterior materials of buildings, apparatus for cooking, tableware or the like.

COPYRIGHT: (C)1999,JPO

AB . . . and to obtain antibacterial, deodorant and contamination preventive effects by immersing titanium into an electrolytic bath composed of sulfuric acid, **phosphoric acid** and **hydrogen peroxide**, increasing the voltage to the prescribed one by a prescribed direct constant current and holding it for a prescribed time. . . exhibiting photocatalytic operation even by visible rays, in which titanium is immersed into an electrolytic bath composed of sulfuric acid, **phosphoric acid**, **hydrogen peroxide** and cobalt **sulfate**, the voltage is increased to the prescribed one by a prescribed direct constant current, and it is held for a . . . form a TiO₂-CoO series titanium anodically oxidized film. Moreover, titanium is immersed into an electrolytic bath composed of sulfuric acid, **phosphoric acid**, **hydrogen peroxide** and zinc

sulfate, the voltage is increased to the prescribed one, and it is held for a prescribed time to form a $\text{TiO}_2\text{-ZnO}$.

L3 ANSWER 79 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1999:550538 CAPLUS

DN 131:322773

TI A Heterogeneous Tungsten Catalyst for Epoxidation of Terpenes and Tungsten-Catalyzed Synthesis of Acid-Sensitive Terpene Epoxides

AU Villa de P., Aida L.; Sels, Bert F.; De Vos, Dirk E.; Jacobs, Pierre A.

CS Center for Surface Chemistry and Catalysis, Katholieke Universiteit Leuven, Heverlee, B-3001, Belg.

SO Journal of Organic Chemistry (1999), 64(19), 7267-7270

CODEN: JOCEAH; ISSN: 0022-3263

PB American Chemical Society

DT Journal

LA English

OS CASREACT 131:322773

AB A heterogeneous, reusable tungsten catalyst for the selective epoxidn. of a series of terpene olefins with hydrogen peroxide is reported.

RE.CNT 35 THERE ARE 35 CITED REFERENCES AVAILABLE FOR THIS RECORD

ALL CITATIONS AVAILABLE IN THE RE FORMAT

IT 1066-51-9, (Aminomethyl)phosphonic acid 11120-01-7, Tungstic acid sodium salt dihydrate 59158-14-4, Methyltrioctylammonium hydrogen sulfate 92768-58-6D, Venturello anion, macroreticular Amberlite IRA-900 complex

RL: CAT (Catalyst use); USES (Uses)

(heterogeneous tungsten catalyst for epoxidn. of terpenes with hydrogen peroxide)

L3 ANSWER 80 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1999:643380 CAPLUS

DN 132:63758

TI A Practical Method for Alcohol Oxidation with Aqueous Hydrogen Peroxide under Organic Solvent- and Halide-Free Conditions

AU Sato, Kazuhiko; Aoki, Masao; Takagi, Junko; Zimmermann, Klaus; Noyori, Ryoji

CS Dep. Chem. and Res. Cent. for Mater. Sci., Nagoya University, Chikusa-ku, Nagoya, 464-8602, Japan

SO Bulletin of the Chemical Society of Japan (1999), 72(10), 2287-2306

CODEN: BCSJA8; ISSN: 0009-2673

PB Chemical Society of Japan

DT Journal

LA English

AB A catalytic system consisting of sodium tungstate and methyltrioctylammonium hydrogen sulfate effects oxidn. of simple secondary alcs. to ketones using 3-30% H_2O_2 without any org. solvents. The oxidn. can be conducted under entirely halide-free, mildly acidic conditions. A combination of tungstic acid and an appropriate quaternary ammonium salt also effects the alc. dehydrogenation. The org./aq. biphasic reaction allows easy product/catalyst sepn. The turnover no., defined as mols of product per mol of catalyst, approaches 77700 (2-octanol) or 179000 (1-phenylethanol), two orders of magnitude higher than any previously reported. Ester, alkyl and t-butyldimethylsilyl ether, epoxy, carbonyl, N-alkyl carboxamide, and nitrile groups are tolerated under the reaction conditions. Secondary alcs. are preferentially oxidized over terminal olefins. Primary alkanols are oxidized directly to carboxylic acids in a moderate to high yield. Benzylic alcs. are selectively oxidized to benzaldehydes or benzoic acids under suitable conditions. This method is high-yielding, clean, safe, operationally simple, and cost-effective, and therefore suitable for practical org. synthesis. The mechanistic origin of the catalytic efficiency is discussed.

RE.CNT 89 THERE ARE 89 CITED REFERENCES AVAILABLE FOR THIS RECORD

ALL CITATIONS AVAILABLE IN THE RE FORMAT

IT 123-03-5, N-Hexadecylpyridinium chloride 1112-67-0, Tetrabutylammonium

chloride 2760-18-1, N,N,N-Tris(decyl)-1-decanaminium hydrogen
sulfate 7783-03-1, Tungstic acid 13472-45-2, Sodium tungstate
 17756-56-8, Tetrahexylammonium hydroxide 32503-27-8, Tetrabutylammonium
 hydrogen **sulfate** 32503-34-7, Tetrahexylammonium hydrogen
sulfate 57340-65-5, Tetradecylammonium hydroxide
 RL: CAT (Catalyst use); USES (Uses)

(oxidn. of alcs. in aq. **hydrogen peroxide** under
 solvent-free and halide-free conditions)

IT 59158-14-4P, Methyltriocetylammmonium hydrogen **sulfate**
 RL: CAT (Catalyst use); SPN (Synthetic preparation); PREP (Preparation);
 USES (Uses)

(oxidn. of alcs. in aq. **hydrogen peroxide** under
 solvent-free and halide-free conditions)

IT 67-63-0, 2-Propanol, reactions 75-65-0, tert-Butyl alcohol, reactions
 76-83-5, Triphenylmethyl chloride 88-65-3, 2-Bromobenzoic acid
 89-95-2, 2-Methylbenzyl alcohol 94-96-2, 2-Ethyl-1,3-hexanediol
 96-22-0, 3-Pentanone 96-41-3, Cyclopentanol 98-52-2,
 4-tert-Butylcyclohexanol 98-85-1, .alpha.-Methylbenzenemethanol
 100-51-6, Benzyl alcohol, reactions 104-76-7, 2-Ethyl-1-hexanol
 105-13-5, 4-Methoxybenzyl alcohol 108-93-0, Cyclohexanol, reactions
 111-26-2, Hexylamine 111-66-0, 1-Octene 111-87-5, 1-Octanol, reactions
 123-96-6, 2-Octanol 143-08-8, 1-Nonanol 507-70-0 589-18-4,
 4-Methylbenzyl alcohol 619-73-8, 4-Nitrobenzyl alcohol 868-46-2,
 3,4-Diethyl-3-hexene 873-75-6, 4-Bromobenzyl alcohol 873-76-7,
 4-Chlorobenzyl alcohol 920-39-8, Isopropylmagnesium bromide 1066-51-9,
 (Aminomethyl)**phosphonic acid** 1592-38-7,
 2-Naphthalenemethanol 1653-30-1, 2-Undecanol 2046-21-1, 6-Oxoheptanoic
 acid methyl ester 3128-07-2, 6-Oxoheptanoic acid 3174-74-1,
 5,6-Dihydro-2H-pyran 4048-42-4, 1-Dodecen-3-ol 5337-72-4,
 2,6-Dimethylcyclohexanol 6032-29-7, 2-Pentanol 7154-66-7,
 2-Bromobenzoyl chloride 13491-79-7, 2-tert-Butylcyclohexanol
 14609-79-1, 2,2,4,4-Tetramethyl-3-pentanol 14850-22-7, (Z)-3-Octene
 14919-01-8, (E)-3-Octene 18162-48-6, tert-Butyldimethylsilyl chloride
 18516-37-5, 2-Methyl-1-undecene 21951-49-5, 11-Dodecen-2-ol
 23381-92-2, 2-Methyl-2-decene 95465-45-5 199916-09-1,
 5-Cyclohexadecen-1-ol 253309-47-6, 11-Hydroxy-9-nonadecanone
 RL: RCT (Reactant); RACT (Reactant or reagent)

(oxidn. of alcs. in aq. **hydrogen peroxide** under
 solvent-free and halide-free conditions)

L3 ANSWER 81 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1998:568901 CAPLUS

DN 129:177238

TI Rinse aid solutions and process for stainless steel

IN Peterson, Joseph C.

PA Crown Technology, Inc., USA

SO PCT Int. Appl., 12 pp.

CODEN: PIXXD2

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9836044	A1	19980820	WO 1998-US1721	19980128
	W: CA, JP, MX				
	RW: AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
	US 5821212	A	19981013	US 1997-947994	19971009
PRAI	US 1997-40075P	P	19970213		
	US 1997-40186P	P	19970312		
	US 1997-947994	A	19971009		
AB	The solns., for acid pickled stainless steel, comprise H2O2 2-20, phosphoric acid 3-4 and Na 2-ethylhexyl sulfate 3-4% and balanced water.				

RE.CNT 5 THERE ARE 5 CITED REFERENCES AVAILABLE FOR THIS RECORD

ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB The solns., for acid pickled stainless steel, comprise **H2O2**
2-20, **phosphoric acid** 3-4 and Na 2-ethylhexyl
sulfate 3-4% and balanced water.

ST stainless steel rinse aid soln; **hydrogen peroxide**
phosphoric acid rinsing aid; ethylhexyl sulfate aq soln
rinsing steel

L3 ANSWER 82 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 1998:293593 CAPLUS
DN 129:5897
TI Carpet cleaning compositions and method for cleaning carpets
IN Silvaggi, Lorena; Raso, Floriana; Scialla, Stefano
PA Procter & Gamble Co., USA; Silvaggi, Lorena; Raso, Floriana; Scialla,
Stefano
SO PCT Int. Appl., 38 pp.
CODEN: PIXXD2
DT Patent
LA English
FAN.CNT 1

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
WO 9818894	A1	19980507	WO 1997-US19792	19971030
W: AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GE, GH, HU, ID, IL, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, US, UZ, VN, YU, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM RW: GH, KE, LS, MW, SD, SZ, UG, ZW, BF, BJ, CF, CG, CI, CM, GA, GN, ML, MR, NE, SN, TD, TG				
EP 839900	A1	19980506	EP 1996-870136	19961031
R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO				
AU 9852425	A1	19980522	AU 1998-52425	19971030
JP 2000507643	T2	20000620	JP 1998-520785	19971030
US 6177395	B1	20010123	US 1999-297432	19990430
PRAI EP 1996-870136	A	19961031		
WO 1997-US19792	W	19971030		

OS MARPAT 129:5897

AB The title compns. having improved overall carpet stain removal
performance, e.g., on particulate stains, greasy, oily, bleachable and/or
enzymic stains, comprise a peroxygen bleach, a C1-4 linear or branched,
(un)satd. alc. and a 2nd solvent selected from a hydrophilic solvent
comprising .gtoreq.1 ether groups and having a soly. in H2O >10 mL per 100
mL at 25 .degree.C, a polyol hydrophilic solvent having a soly. in H2O >10
mL per 100 mL at 25 .degree.C, a hydrophobic solvent having a soly. in H2O
<10 mL per 100 mL at 25 .degree.C, and their mixts., at a wt. ratio of the
alc. to the 2nd solvent of 1:(10-1.1). A typical compn. contained
H2O2 5.0, Na alkyl **sulfate** 3.0, ethoxylated
tetraethylenepentamine 1.0, salicylic acid 0.1, **phosphonate**
chelant 0.1, Me2CHOH 1.0, EtOCH2CH2OH 2.0, Pr gallate 0.01 and H2O
balance.

RE.CNT 2 THERE ARE 2 CITED REFERENCES AVAILABLE FOR THIS RECORD

ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB The title compns. having improved overall carpet stain removal
performance, e.g., on particulate stains, greasy, oily, bleachable and/or
enzymic stains, comprise a peroxygen bleach, a C1-4 linear or branched,
(un)satd. alc. and a 2nd solvent selected from a hydrophilic solvent
comprising .gtoreq.1 ether groups and having a soly. in H2O >10 mL per 100
mL at 25 .degree.C, a polyol hydrophilic solvent having a soly. in H2O >10
mL per 100 mL at 25 .degree.C, a hydrophobic solvent having a soly. in H2O
<10 mL per 100 mL at 25 .degree.C, and their mixts., at a wt. ratio of the
alc. to the 2nd solvent of 1:(10-1.1). A typical compn. contained
H2O2 5.0, Na alkyl **sulfate** 3.0, ethoxylated

tetraethylenepentamine 1.0, salicylic acid 0.1, **phosphonate**
chelant 0.1, Me2CHOH 1.0, EtOCH2CH2OH 2.0, Pr gallate 0.01 and H2O
balance.

L3 ANSWER 83 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 1998:208484 CAPLUS
DN 128:232412
TI Method for producing chlorine dioxide using methanol, chloride, and
hydrogen peroxide as reducing agents
IN Nonni, Arthur J.; Graff, Robert J.; Liu, Raymond; Voss, Jerry N.; Hammond,
Timothy R.
PA International Paper Co., USA; Nonni, Arthur J.; Graff, Robert J.; Liu,
Raymond; Voss, Jerry N.; Hammond, Timothy R.
SO PCT Int. Appl., 21 pp.
CODEN: PIXXD2
DT Patent
LA English
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE			
PI	WO 9813296	A1	19980402	WO 1997-US18156	19970929			
	W:	AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GE, GH, HU, ID, IL, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, US, UZ, VN, YU, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM						
	RW:	GH, KE, LS, MW, SD, SZ, UG, ZW, AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, BF, BJ, CF, CG, CI, CM, GA, GN, ML, MR, NE, SN, TD, TG						
	AU 9746732	A1	19980417	AU 1997-46732	19970929			
PRAI	US 1996-720087		19960927					
	WO 1997-US18156		19970929					
AB	ClO2 is produced by reacting an alkali metal chlorate (e.g., NaClO3) with CH3OH, metal chloride (e.g., NaCl), and H2O2 (at <50%) as reducing agents in an aq. acidic medium (e.g., H2SO4, HCl, H3PO4, HNO3, HClO3) at 20-140.degree.C, esp 50-75.degree.C. A catalyst may be used, selected from AgNO3, MnSO4, V2O5, Ru oxide, Rh oxide and PdO. The reaction can be carried out at atm. pressure or sub-atm. pressure. The combination of these reducing agents causes an unexpected enhancement in ClO2 generation.							
RE.CNT	5	THERE ARE 5 CITED REFERENCES AVAILABLE FOR THIS RECORD						
	ALL CITATIONS AVAILABLE IN THE RE FORMAT							
IT	1314-08-5, Palladium oxide	1314-62-1, Vanadium pentoxide, uses						
	7761-88-8, Silver nitrate, uses	7785-87-7, Manganese sulfate						
	11113-84-1, Ruthenium oxide	12680-36-3, Rhodium oxide						
RL:	CAT (Catalyst use); USES (Uses)							
	(chlorine dioxide prodn. using methanol, chloride, and hydrogen peroxide as reducing agents)							
IT	67-56-1, Methanol, reactions	7647-01-0, Hydrochloric acid, reactions						
	7647-01-0D, Hydrogen chloride, salts	7647-14-5, Sodium chloride, reactions						
	7664-38-2, Phosphoric acid , reactions							
	7664-93-9, Sulfuric acid, reactions	7697-37-2, Nitric acid, reactions						
	7722-84-1, Hydrogen peroxide, reactions	7775-09-9, Sodium chlorate						
	7790-93-4, Chloric acid	7790-93-4D, Chloric acid, alkali metal salts						
RL:	RCT (Reactant); RACT (Reactant or reagent)							
	(chlorine dioxide prodn. using methanol, chloride, and hydrogen peroxide as reducing agents)							

L3 ANSWER 84 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 1998:478952 CAPLUS
DN 129:86049
TI Germicidal compositions for the treatment of animal infectious diseases of the hoof
IN Meisters, George J.; Gradle, Charles D.
PA Babson Bros. Co., USA

SO U.S., 7 pp.
CODEN: USXXAM
DT Patent
LA English
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 5780064	A	19980714	US 1997-938013	19970912
	WO 9913888	A1	19990325	WO 1998-US18751	19980909
	W: AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GE, GH, GM, HU, ID, IL, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, UZ, VN, YU, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM				
	RW: GH, GM, KE, LS, MW, SD, SZ, UG, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG				
	AU 9892270	A1	19990405	AU 1998-92270	19980909
	AU 739193	B2	20011004		
	EP 1027062	A1	20000816	EP 1998-944820	19980909
	R: DE, FR, GB, NL				
	BR 9812437	A	20011120	BR 1998-12437	19980909
	NZ 503340	A	20020828	NZ 1998-503340	19980909

PRAI US 1997-938013 A 19970912
WO 1998-US18751 W 19980909

AB An aq. germicidal compn. for the treatment or prevention of infectious diseases of the hoof in animals, comprising a copper salt, a quaternary ammonium compd., and a peroxide is disclosed. An aq. compn. contained water 69.9, hydroxyacetic acid 0.5, Barlox 12 1.0, BTC 2125M 1.0, 75% **phosphoric acid** 0.6, **copper sulfate** pentahydrate 20.0, and 35% **hydrogen peroxide** 10.0%.
Efficacy of the compn. in treatment of dairy cows infected with papillomatous digital dermatitis was shown.

RE.CNT 36 THERE ARE 36 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

AB An aq. germicidal compn. for the treatment or prevention of infectious diseases of the hoof in animals, comprising a copper salt, a quaternary ammonium compd., and a peroxide is disclosed. An aq. compn. contained water 69.9, hydroxyacetic acid 0.5, Barlox 12 1.0, BTC 2125M 1.0, 75% **phosphoric acid** 0.6, **copper sulfate** pentahydrate 20.0, and 35% **hydrogen peroxide** 10.0%.
Efficacy of the compn. in treatment of dairy cows infected with papillomatous digital dermatitis was shown.

ST germicide pharmaceutical animal infectious disease hoof; BTC 2125M
hydrogen peroxide pharmaceutical germicide; **copper sulfate**
pharmaceutical germicide **hydrogen peroxide**; dermatitis
germicide pharmaceutical animal hoof

L3 ANSWER 85 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1998:397851 CAPLUS

DN 129:71675

TI Reduction of S-(2-Diisopropylamino)ethyl methylphosphonothioate ion with hydrogen peroxide

IN Yang, Yu-chu; Wagner, George W.

PA United States Dept. of the Army, USA

SO U.S., 5 pp.
CODEN: USXXAM

DT Patent
LA English
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 5763737	A	19980609	US 1997-884492	19970627
PRAI	US 1997-884492		19970627		

AB A method for the redn. of toxicity of S-(2-diisopropylamino)ethyl methylphosphonothioate ions in a medium or mixt. by reaction with hydrogen peroxide and a strong inorg. acid is described.

RE.CNT 2 THERE ARE 2 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

IT **Sulfonic acids, processes**
RL: PEP (Physical, engineering or chemical process); PROC (Process)
(redn. of S-(2-Diisopropylamino)ethyl methylphosphonothioate ion with
hydrogen peroxide)

IT 64-19-7, Acetic acid, processes 993-13-5, Methylphosphonic acid
7647-01-0, Hydrochloric acid, processes 7664-38-2, **Phosphoric acid**, processes 7697-37-2, Nitric acid, processes 7722-84-1,
Hydrogen peroxide (H2O2), processes 65332-44-7 128869-82-9
RL: PEP (Physical, engineering or chemical process); PROC (Process)
(redn. of S-(2-Diisopropylamino)ethyl methylphosphonothioate ion with
hydrogen peroxide)

L3 ANSWER 86 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1998:786116 CAPLUS

DN 130:28042

TI Mixture solution containing hydrogen peroxide and liquid metal-cleaning composition using it

IN Owa, Kyosuke

PA Horiuchi K. K., Japan

SO Jpn. Kokai Tokkyo Koho, 4 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 10324508	A2	19981208	JP 1997-152875	19970526
PRAI	JP 1997-152875		19970526		

AB The mixt. soln. contains a H2O2 soln, a weak acid, and a dihydrogen Na dihydrate of an weak acid. The cleaning compn. contains the above mixt. soln, a cleaning agent, a chelating agent, and a surfactant. Reactivity of H2O2 is mild in the soln., so that the mixt. soln. and the cleaning soln. are easy to handle with high safety.

IT 60-00-4, EDTA, uses 77-92-9, Citric acid, uses 126-92-1D, Sodium 2-ethylhexyl **sulfate**, alkyl esters 5329-14-6, Sulfamic acid 7664-38-2, **Phosphoric acid**, uses 9004-99-3, Polyoxyethylene monostearate 9072-19-9, Fucoidin 25322-68-3D, alkyl ethers

RL: TEM (Technical or engineered material use); USES (Uses)

(metal-cleaning soln. compn. contg. **H2O2** mixt. soln.,
cleaning agent, chelating agent, and surfactant for safe handling)

L3 ANSWER 87 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1998:334868 CAPLUS

DN 129:42639

TI Bleach compositions free from sharp odors and lasting effects for preventing mildew redeposition

IN Saiga, Koji; Yamada, Hiroyuki; Tsukuda, Kazunori

PA Kao Corp., Japan

SO Jpn. Kokai Tokkyo Koho, 9 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 10140190	A2	19980526	JP 1996-302744	19961114
PRAI	JP 1996-302744		19961114		

OS MARPAT 129:42639

AB The title compns. comprise hydrogen peroxide or peroxides generating hydrogen peroxide in aq. solns., bleach activators R1C6H4(R4)pO(AO)nR2CO2R3SO3M and cationic surfactants chosen from R5R6N+Me2 X- and R7Me2N+CH2Ph X-, wherein R1 = H, C1-10 alkyl, etc.; R2, R3 = C1-8 alkylene, etc.; R4 = C1-5 alkylene, etc.; p = 0,1; A = C2-4 alkylene, etc.; n = 1-100; M = alkali metal, etc.; R5 = C6-18 alkyl, etc.; R6 = Me, C6-18 alkyl, etc.; R7 = C6-18 alkyl, etc.; X- = anion. A compn. comprised **hydrogen peroxide 6**, PhO(C2H4O)2CH2CO2CH2CH2SO3Na 22, **Dequest 2015DN 0.5**, Na alkyl **sulfate 1**, and water to 100%.

AB The title compns. comprise hydrogen peroxide or peroxides generating hydrogen peroxide in aq. solns., bleach activators R1C6H4(R4)pO(AO)nR2CO2R3SO3M and cationic surfactants chosen from R5R6N+Me2 X- and R7Me2N+CH2Ph X-, wherein R1 = H, C1-10 alkyl, etc.; R2, R3 = C1-8 alkylene, etc.; R4 = C1-5 alkylene, etc.; p = 0,1; A = C2-4 alkylene, etc.; n = 1-100; M = alkali metal, etc.; R5 = C6-18 alkyl, etc.; R6 = Me, C6-18 alkyl, etc.; R7 = C6-18 alkyl, etc.; X- = anion. A compn. comprised **hydrogen peroxide 6**, PhO(C2H4O)2CH2CO2CH2CH2SO3Na 22, **Dequest 2015DN 0.5**, Na alkyl **sulfate 1**, and water to 100%.

L3 ANSWER 88 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1998:314875 CAPLUS
 DN 129:29425
 TI Bleaching agent compositions for mildewcides
 IN Saika, Koshi; Yamada, Hiroyuki; Tsukuda, Kazunori
 PA Kao Corp., Japan
 SO Jpn. Kokai Tokkyo Koho, 8 pp.
 CODEN: JKXXAF
 DT Patent
 LA Japanese
 FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 10130694	A2	19980519	JP 1996-286415	19961029
PRAI	JP 1996-286415		19961029		

AB The bleaching agent compns. contain (a) H2O2 or peroxides which generate H2O2 in water, (b) bleaching agent activators R1C6H4R4pO(AO)nR2CO2R3SO3M (I; R1 = H, C1-10 alkyl, alkenyl, acyl; R2 = C1-8 alkylene, phenylene which may be substituted with C1-5 alkyl; R3 = C1-8 alkylene, R4 = C1-5 alkylene; p = 0, 1; A = C2-4 alkylene; n = 1-100; M = alkali metal, alk. earth metal, ammonium, alkylammonium, alkanolammonium), and (c) R1C6H4R4pO(AO)nR2CO2M1 (II; R1-R4, A, p, n = same as above; M1 = H, alkali metal, alk. earth metal, ammonium, alkylammonium, alkanolammonium). Thus, a storage-stable 2-pot-type bleaching agent compn. was composed of agent-A comprising **H2O2 6**, I (R1 = H, p = 0, AO = CH2CH2O, n = 2, R2 = CH2, R3 = CH2CH2, M = Na) 22, **Dequest 2015DN 0.5**, Na C12-14 alkyl **sulfonate 1**, and II (R1 = H, p = 0, AO = CH2CH2O, n = 2, R2 = CH2, M1 = H) 0.1% and balance water and agent-B contg. 20% K2CO3 and balance water.

AB The bleaching agent compns. contain (a) H2O2 or peroxides which generate H2O2 in water, (b) bleaching agent activators R1C6H4R4pO(AO)nR2CO2R3SO3M (I; R1 = H, C1-10 alkyl, alkenyl, acyl; R2 = C1-8 alkylene, phenylene which may be substituted with C1-5 alkyl; R3 = C1-8 alkylene, R4 = C1-5 alkylene; p = 0, 1; A = C2-4 alkylene; n = 1-100; M = alkali metal, alk. earth metal, ammonium, alkylammonium, alkanolammonium), and (c) R1C6H4R4pO(AO)nR2CO2M1 (II; R1-R4, A, p, n = same as above; M1 = H, alkali metal, alk. earth metal, ammonium, alkylammonium, alkanolammonium). Thus, a storage-stable 2-pot-type bleaching agent compn. was composed of agent-A comprising **H2O2 6**, I (R1 = H, p = 0, AO = CH2CH2O, n = 2, R2 = CH2, R3 = CH2CH2, M = Na) 22, **Dequest 2015DN 0.5**, Na C12-14 alkyl **sulfonate 1**, and II (R1 = H, p = 0, AO = CH2CH2O, n = 2, R2 = CH2, M1 = H) 0.1% and balance water and agent-B contg. 20% K2CO3 and balance water.

L3 ANSWER 89 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1998:314370 CAPLUS
 DN 129:42635
 TI Thickened liquid bleaching agents
 IN Ogura, Nobuyuki; Ozaki, Kazuyoshi; Aoyagi, Muneo
 PA Kao Corp., Japan
 SO Jpn. Kokai Tokkyo Koho, 12 pp.
 CODEN: JKXXAF
 DT Patent
 LA Japanese
 FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 10130693	A2	19980519	JP 1997-166944	19970624
	TW 416986	B	20010101	TW 1997-86111382	19970808
	CN 1176302	A	19980318	CN 1997-118258	19970905
PRAI	JP 1996-236218	A	19960906		

OS MARPAT 129:42635

AB Bleaching agents contain H2O2, thickening agents, and surfactants and have viscosity 10-1000 mPa.s at 30.degree., and the thickening agents have polyoxyalkylene chains, alkyl or alkenyl groups, mol. wt. >600, and polyoxyalkylene chain-alkyl or alkenyl group wt. ratio >2. Thus, a bleaching agent contained H2O2 5, Emanon 3170 2, C12 alkylbenzenesulfonic acid 0.5, polyethylene glycol lauryl ether 2.0, Dequest 2010 0.1%, and H2O.

AB Bleaching agents contain H2O2, thickening agents, and surfactants and have viscosity 10-1000 mPa.s at 30.degree., and the thickening agents have polyoxyalkylene chains, alkyl or alkenyl groups, mol. wt. >600, and polyoxyalkylene chain-alkyl or alkenyl group wt. ratio >2. Thus, a bleaching agent contained H2O2 5, Emanon 3170 2, C12 alkylbenzenesulfonic acid 0.5, polyethylene glycol lauryl ether 2.0, Dequest 2010 0.1%, and H2O.

IT 2809-21-4, Dequest 2010

RL: MOA (Modifier or additive use); USES (Uses)
 (bleaching agents contg. **hydrogen peroxide** and thickening agents and surfactants)

IT 7722-84-1, Hydrogen peroxide, uses 9002-92-0, Polyethylene glycol lauryl ether 19309-23-0, N,N-Dimethyl-N-ethyl-N-tetradecylammonium ethyl sulfate 27176-87-0, Laurylbenzenesulfonic acid 51889-11-3, Sodium p-Sulfophenyl hexanoate 89740-13-6, p-Sulfophenyl laurate sodium salt

RL: TEM (Technical or engineered material use); USES (Uses)
 (bleaching agents contg. **hydrogen peroxide** and thickening agents and surfactants)

L3 ANSWER 90 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1998:287106 CAPLUS
 DN 129:29413
 TI Storage stable liquid bleaching compositions
 IN Ogura, Nobuyuki; Ozaki, Kazuyoshi; Aoyagi, Muneo
 PA Kao Corp., Japan
 SO Jpn. Kokai Tokkyo Koho, 9 pp.
 CODEN: JKXXAF
 DT Patent
 LA Japanese
 FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 10121100	A2	19980512	JP 1996-276283	19961018
PRAI	JP 1996-276283		19961018		

OS MARPAT 129:29413

AB Bleaching compns. contain H2O2, heavy metal chelating agents, and nonionic surfactants, which contain <5% polyoxyalkylenes YO(AO)nZ, where A = linear

or branched C2-3 alkylenes, Z, Y = H or C1-3 alkyls, and n = 3-50. Thus, a bleaching agent contained **H2O2** 5, polyethylene glycol lauryl ether 5, p-sulfophenyl laurate Na salt 0.5, **Dequest** 2010 0.2, N,N-dimethyl-N-ethyl-N-tetradecylammonium Et **sulfate** 0.5%, and H2O.

AB Bleaching compns. contain H2O2, heavy metal chelating agents, and nonionic surfactants, which contain <5% polyoxyalkylenes YO(AO)nZ, where A = linear or branched C2-3 alkylenes, Z, Y = H or C1-3 alkyls, and n = 3-50. Thus, a bleaching agent contained **H2O2** 5, polyethylene glycol lauryl ether 5, p-sulfophenyl laurate Na salt 0.5, **Dequest** 2010 0.2, N,N-dimethyl-N-ethyl-N-tetradecylammonium Et **sulfate** 0.5%, and H2O.

IT 2809-21-4, **Dequest** 2010

RL: MOA (Modifier or additive use); USES (Uses)
(**Dequest** 2010; storage stable liq. bleaching compns. contg. **hydrogen peroxide** and chelating agents and nonionic surfactants)

IT 7722-84-1, Hydrogen peroxide, uses 9002-92-0, Polyethylene glycol lauryl ether 19309-23-0, N,N-Dimethyl-N-ethyl-N-tetradecylammonium ethyl **sulfate** 27176-87-0, Laurylbenzenesulfonic acid 70424-62-3 89740-13-6, p-Sulfophenyl laurate sodium salt

RL: TEM (Technical or engineered material use); USES (Uses)
(storage stable liq. bleaching compns. contg. **hydrogen peroxide** and chelating agents and nonionic surfactants)

L3 ANSWER 91 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1998:157561 CAPLUS

DN 128:232437

TI Manufacture of high-purity hydrogen peroxide

IN Nankawa, Yoshitsugu; Kokubu, Atsushi; Murakami, Seishi

PA Mitsubishi Gas Chemical Co., Inc., Japan

SO Jpn. Kokai Tokkyo Koho, 6 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
------------	------	------	-----------------	------

PI JP 10067505	A2	19980310	JP 1996-225296	19960827
----------------	----	----------	----------------	----------

PRAI JP 1996-225296		19960827		
---------------------	--	----------	--	--

AB The title process comprises mixing impurity-contg. aq. **H2O2** with acids having pKa .ltoreq.5, and treating it with F salt-type anion exchangers, optionally **sulfonic acid**-contg. strongly acidic cation exchangers and bicarbonate-type anion exchangers properly. The process is useful for removal of Si from H2O2.

AB The title process comprises mixing impurity-contg. aq. **H2O2** with acids having pKa .ltoreq.5, and treating it with F salt-type anion exchangers, optionally **sulfonic acid**-contg. strongly acidic cation exchangers and bicarbonate-type anion exchangers properly. The process is useful for removal of Si from H2O2.

IT 7647-01-0, Hydrochloric acid, processes 7664-38-2, **Phosphoric acid**, processes 7664-93-9, Sulfuric acid, processes 7697-37-2, Nitric acid, processes

RL: PEP (Physical, engineering or chemical process); PROC (Process)
(treatment with acid and anion and cation exchanger in removal of Si from **H2O2**)

L3 ANSWER 92 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1998:204436 CAPLUS

DN 128:259114

TI Hydrogen peroxide complexes of inorganic salts and their synthesis as pastes for disinfection

IN Chen, Xiaolan; Jacobs, Paul T.; Lin, Szu-min

PA Johnson & Johnson Medical, Inc., USA

SO Eur. Pat. Appl., 63 pp.

CODEN: EPXXDW

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 831056	A2	19980325	EP 1997-307263	19970918
	EP 831056	A3	19990107		
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, FI				
	US 5820841	A	19981013	US 1996-716055	19960919
	SG 90704	A1	20020820	SG 1997-3365	19970912
	NO 9704306	A	19980320	NO 1997-4306	19970918
	JP 10182113	A2	19980707	JP 1997-270354	19970918
	ZA 9708423	A	19990318	ZA 1997-8423	19970918
	RU 2185321	C2	20020720	RU 1997-115762	19970918
	CA 2216061	AA	19980319	CA 1997-2216061	19970919
	CN 1196886	A	19981028	CN 1997-121487	19970919
	CN 1104201	B	20030402		
	BR 9704759	A	19990105	BR 1997-4759	19970919
PRAI	US 1996-716055	A	19960919		

AB H2O2 complexes with inorg. salts are made by mixing the desired salt with sufficient water to form a soft hydrate paste, mixing the paste with a 12-80% aq. H2O2 soln., and then drying the H2O2-contg. paste by oven or vacuum drying. The dried paste is heated for use as a solid disinfectant in vapor sterilization, e.g., of medical instruments or devices. The salt may be a phosphate salt or condensed phosphate salt of Na, Ca or Mg, such as Na4P2O7, which is complexed with H2O2. In examples, 30-90% aq. H2O2 was sprayed dropwise onto 98% Na4P2O7, the mixt. was incubated at 10-45.degree.C for 1-16 h, then dried by vacuum drying at 25-60.degree.C or oven drying at 60.degree.C, forming Na2P2O7.nH2O2 where n=1-4.

IT 7722-84-1DP, Hydrogen peroxide (H2O2), complexes, preparation
9002-89-5DP, Polyvinylalcohol, compd. with hydrogen peroxide
9003-01-4DP, compd. with hydrogen peroxide 9003-05-8DP, Polyacrylamide, compd. with hydrogen peroxide 9003-09-2DP, Polyvinylmethylether, compd. with hydrogen peroxide 9003-20-7DP, Polyvinylacetate, compd. with hydrogen peroxide 9003-39-8DP, Polyvinylpyrrolidone, compd. with hydrogen peroxide 9004-35-7DP, Cellulose acetate, compd. with hydrogen peroxide 9005-22-5DP, Cellulose sulfate, sodium salt, compd. with hydrogen peroxide 9005-38-3DP, Sodium alginate, compd. with hydrogen peroxide 13769-85-2P, Diphosphoric acid, tetrasodium salt, compd. with hydrogen peroxide (H2O2) (1:3) 14479-85-7P, Urea, compd. with hydrogen peroxide 15039-07-3P, Diphosphoric acid, tetrasodium salt, compd. with hydrogen peroxide (H2O2) (1:2) 15362-88-6P, Diphosphoric acid, tetrasodium salt, compd. with hydrogen peroxide (H2O2) 15523-00-9P, Phosphoric acid, disodium salt, compd. with hydrogen peroxide (H2O2) (1:2) 15609-73-1P 15630-89-4P 16884-59-6P, Carbonic acid, disodium salt, compd. with hydrogen peroxide (H2O2) 17032-34-7P 18441-01-5P, L-Histidine, compd. with hydrogen peroxide (H2O2) 22464-71-7P, Phosphoric acid, trisodium salt, compd. with hydrogen peroxide (H2O2) 22653-27-6P, Triphosphoric acid, pentasodium salt, compd. with hydrogen peroxide (H2O2) 25038-54-4DP, Nylon 6, compd. with hydrogen peroxide 25215-15-0P, Phosphoric acid, disodium salt, compd. with hydrogen peroxide (H2O2) (1:1) 25232-41-1DP, Poly(4-vinylpyridine), compd. with hydrogen peroxide 25609-36-3P 30347-99-0P 32131-17-2DP, Nylon 6.6, compd. with hydrogen peroxide 33960-94-0P 33960-95-1P, Diphosphoric acid, tetrapotassium salt, compd. with hydrogen peroxide (H2O2) 36563-87-8P, Phosphoric acid, tripotassium salt, compd. with hydrogen peroxide (H2O2) (1:2) 39041-45-7P, Phosphoric acid, tripotassium salt, compd. with hydrogen

peroxide (H2O2) (1:3) 51770-74-2P, Diphosphoric acid, tetrapotassium salt, compd. with hydrogen peroxide (H2O2) (1:6) 122040-69-1P, Phosphoric acid, trisodium salt, compd. with hydrogen peroxide (H2O2) (1:5) 189760-18-7P, Phosphoric acid, monopotassium salt, compd. with hydrogen peroxide (H2O2) (1:1) 189760-19-8P 189760-20-1P 189760-55-2P, Phosphoric acid, tripotassium salt, compd. with hydrogen peroxide (H2O2) 189760-61-0P 197566-45-3P 197566-46-4P 205367-43-7P 205367-47-1P 205367-51-7P 205367-54-0P 205367-58-4P 205367-74-4P 205367-75-5P 205367-78-8P 205367-80-2P 205367-87-9P 205367-91-5P 205367-95-9P 205367-99-3P 205368-03-2P 205368-05-4P 205368-07-6P 205368-11-2P 205368-12-3P 205368-14-5P 205368-15-6P 205368-17-8P 205368-20-3P 205368-23-6P
 RL: NUU (Other use, unclassified); SPN (Synthetic preparation); PREP (Preparation); USES (Uses)
 (hydrogen peroxide complexes of inorg. salts as pastes for disinfection)

L3 ANSWER 93 OF 158 JAPIO COPYRIGHT 2003 JPO
 AN 1998-218610 JAPIO
 TI PURIFICATION OF HYDROGEN PEROXIDE AND PLANT THEREFOR
 IN DHALLUIN JEAN MARIE; WAWRZYNIAK JEAN JACQUES; LEDON HENRY
 PA CHEMOXAL SA
 PI JP 10218610 A 19980818 Heisei
 AI JP 1997-280503 (JP09280503 Heisei) 19971014
 PRAI FR 1996-12569 19961015
 SO PATENT ABSTRACTS OF JAPAN (CD-ROM), Unexamined Applications, Vol. 1998
 AB PROBLEM TO BE SOLVED: To provide a method for purifying hydrogen peroxide, enabling to remove cations and organic acids from an aqueous hydrogen peroxide solution by adding a microligand having one or more functional groups selected from carboxyl group, sulfonyl group, phosphoryl group and a nitrogen-containing functional group such as an aromatic or non-aromatic amine functional group or an N-oxidized amine functional group to an aqueous hydrogen peroxide solution and subsequently subjecting the mixture to an ultrafiltration treatment.
 SOLUTION: This method for purifying hydrogen peroxide comprises adding one or more kinds of microligands selected from 4-vinylpyridine homopolymer, styrene/4-vinylpyridine 2/8 and 1/9 copolymer, an acryl/sulfate/sulfonate copolymer (MW=50,000), an acryl/sulfate copolymer (MW=50,000) and a polyvinyl phosphoric acid (MW=30,000) and subsequently forcibly passing the mixture through a ultrafiltration membrane.
 COPYRIGHT: (C)1998,JPO
 AB . . . an aqueous hydrogen peroxide solution and subsequently subjecting the mixture to an ultrafiltration treatment.
 SOLUTION: This method for purifying hydrogen peroxide comprises adding one or more kinds of microligands selected from 4-vinylpyridine homopolymer, styrene/4-vinylpyridine 2/8 and 1/9 copolymer, an acryl/sulfate/sulfonate copolymer (MW=50,000), an acryl/sulfate copolymer (MW=50,000) and a polyvinyl phosphoric acid (MW=30,000) and subsequently forcibly passing the mixture through a ultrafiltration membrane.
 COPYRIGHT: (C)1998,JPO
 L3 ANSWER 94 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1997:805823 CAPLUS
 DN 128:49881
 TI Detergent compositions for laundry and dish washing with delayed alkalinity system release
 IN Hall, Robin Gibson
 PA Procter & Gamble Company, USA; Hall, Robin Gibson
 SO PCT Int. Appl., 76 pp.
 CODEN: PIXXD2

DT Patent
LA English
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9745524	A1	19971204	WO 1997-US8233	19970515
	W: BR, CA, MX, US				
	RW: AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
	GB 2313603	A1	19971203	GB 1996-11325	19960531
	EP 912703	A1	19990506	EP 1997-925579	19970515
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, PT, IE, FI				
	BR 9709404	A	19990810	BR 1997-9404	19970515
PRAI	GB 1996-11325		19960531		
	WO 1997-US8233		19970515		

OS MARPAT 128:49881

AB The title compns. comprise a cationic ester surfactant and an alky. system and a means for delaying the release of the alky. system relatively to the release of the cationic ester surfactant, in such a way that in the T50 test (defined), the time to achieve a concn. that is 50% of the ultimate concn. of the cationic ester surfactant is at least 120 s less than the time to achieve a concn. that is 50% of the ultimate concn. of the alky. system. A laundry detergent comprised Na dodecylbenzenesulfonate 8, C12-15 alc. ethoxylate 3.4, RCO2CH2CH2CH2N+Me3 (R = C11-13) 6, zeolite 18.1, amorphous Na carbonate 13, amorphous Na silicate 1.4, Na sulfate 26.1, NaBO2.3H2O.H2O2 5, Na percarbonate coated with Na silicate for slow release 4, tetraacetythylenediamine 1.5, Dequest 2060 0.25, 1,1-hydroxyethanediphosphonic acid 0.3, protease 0.26, amylase 0.1, maleic acid-acrylic acid copolymer 0.3, CMC 0.2, photoactivated bleach 15 (ppm), disodium 4,4'-bis(2-sulfostryryl)biphenyl 0.09, perfume 0.3, and silicone antifoam 0.5 part.

AB The title compns. comprise a cationic ester surfactant and an alky. system and a means for delaying the release of the alky. system relatively to the release of the cationic ester surfactant, in such a way that in the T50 test (defined), the time to achieve a concn. that is 50% of the ultimate concn. of the cationic ester surfactant is at least 120 s less than the time to achieve a concn. that is 50% of the ultimate concn. of the alky. system. A laundry detergent comprised Na dodecylbenzenesulfonate 8, C12-15 alc. ethoxylate 3.4, RCO2CH2CH2CH2N+Me3 (R = C11-13) 6, zeolite 18.1, amorphous Na carbonate 13, amorphous Na silicate 1.4, Na sulfate 26.1, NaBO2.3H2O.H2O2 5, Na percarbonate coated with Na silicate for slow release 4, tetraacetythylenediamine 1.5, Dequest 2060 0.25, 1,1-hydroxyethanediphosphonic acid 0.3, protease 0.26, amylase 0.1, maleic acid-acrylic acid copolymer 0.3, CMC 0.2, photoactivated bleach 15 (ppm), disodium 4,4'-bis(2-sulfostryryl)biphenyl 0.09, perfume 0.3, and silicone antifoam 0.5 part.

L3 ANSWER 95 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1997:757095 CAPLUS

DN 128:36426

TI Detergent composition

IN Asano, Kaori; Scheibel, Jeffrey John; Baba, Hajime; Baeck, Andre Cesar; Minamikawa, Kaori; Hartshorn, Richard Timothy; Speed, Lynda Anne; Katsuda, Rinko; Kvietok, Frank Andrej; Mao, Mark Hsiang-kuen; Moss, Michael Alan John; Murata, Susumu; Ohtani, Royohei; Panandiker, Rajan Keshav; Pramod, Kakumanu; Sarnaik, Khizar Mohamed Khan; Thoen, Christian Arthur Jacques Kamiel; Kott, Kevin Lee; Willman, Kenneth William; Cron, Laura; Cripe, Thomas Anthony; Curry, John Downing; Foley, Peter Robert; Dodd, Ian Martin; Manohar, Sanjeev Krishnadas; Okamoto, Mitsuyo

PA The Procter & Gamble Company, USA; Asano, Kaori; Scheibel, Jeffrey John; Baba, Hajime; Baeck, Andre Cesar; Minamikawa, Kaori; Hartshorn, Richard Timothy; Speed, Lynda Anne; Katsuda, Rinko; Kvietok, Frank Andrej

SO PCT Int. Appl., 95 pp.

CODEN: PIXXD2

DT Patent

LA English

FAN.CNT 45

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9743388	A2	19971120	WO 1997-US8444	19970516
	W: BR, CA, CN, CZ, HU, JP, MX, TR, US				
	RW: AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
	ZA 9704226	A	19981228	ZA 1997-4226	19970515
	CA 2255005	AA	19971120	CA 1997-2255005	19970516
	CA 2255006	AA	19971120	CA 1997-2255006	19970516
	CA 2255003	AA	19971127	CA 1997-2255003	19970516
	EP 912686	A2	19990506	EP 1997-924766	19970516
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, PT, IE, FI				
	CN 1225122	A	19990804	CN 1997-196381	19970516
	CN 1225116	A	19990804	CN 1997-196385	19970516
	CN 1225115	A	19990804	CN 1997-196402	19970516
	CN 1225121	A	19990804	CN 1997-196403	19970516
	CN 1225677	A	19990811	CN 1997-196468	19970516
	CN 1225678	A	19990811	CN 1997-196475	19970516
	CN 1225673	A	19990811	CN 1997-196481	19970516
	CN 1225675	A	19990811	CN 1997-196494	19970516
	BR 9710678	A	19990817	BR 1997-10678	19970516
	JP 11510554	T2	19990914	JP 1997-541172	19970516
	CN 1230213	A	19990929	CN 1997-196466	19970516
	CN 1232489	A	19991020	CN 1997-196467	19970516
PRAI	US 1996-17883P	P	19960517		
	WO 1997-US8444	W	19970516		

OS MARPAT 128:36426

AB The title compns. comprise a bis-alkoxylated quaternary ammonium (AQA) cationic surfactant R₁R₂N⁺(ApR₄)(A'qR₄) X⁻ (R₁ = linear, branched, or substituted C₈-18 alkyl, alkenyl, aryl, alkaryl, ether or glycidyl ether moiety; R₂ = C₁-3 alkyl; R₃, R₄ = H, Me, Et; X⁻ = anion; A, A' = C₁-4 alkoxy; p, q = 1-30) , a non-AQA surfactant and enzymes. A laundry detergent was prepd. from blown powder comprising anhyd. sodium tripolyphosphate 24, Na C₁₄-15 linear alkyl sulfate 8, 1:4 maleic acid-acrylic acid copolymer 2, Na dodecylbenzensulfonate 6, Na tallow alkyl sulfate 1.5, ethoxylated cocoalkylmethylbis(hydroxyethyl)ammonium chloride 1.5, silicate 7, CMC 1, brightener 0.2, soap 1, and **Dequest** 2060 0.4 part; spray-on comprising C₁₄-15 alc. ethoxylate 2.5, C₁₂-15 alc. ethoxylate 2.5, silicone antifoam 0.3, and perfume 0.3 part; dry additives comprising Na carbonate 6, NaBO₂.3H₂O.H₂O₂ 18, NaBO₂.H₂O₂ 4, tetraacetythylenediamine 3, photoactivated bleach (**sulfonated** zinc phthalocyanine encapsulated in bleach dextrin sol. polymer) 0.02, protease 1, lipase 0.4, amylase 0.25; dry mixed Na sulfate 3.0; and moisture and others to 100 parts.

AB The title compns. comprise a bis-alkoxylated quaternary ammonium (AQA) cationic surfactant R₁R₂N⁺(ApR₄)(A'qR₄) X⁻ (R₁ = linear, branched, or substituted C₈-18 alkyl, alkenyl, aryl, alkaryl, ether or glycidyl ether moiety; R₂ = C₁-3 alkyl; R₃, R₄ = H, Me, Et; X⁻ = anion; A, A' = C₁-4 alkoxy; p, q = 1-30) , a non-AQA surfactant and enzymes. A laundry detergent was prepd. from blown powder comprising anhyd. sodium tripolyphosphate 24, Na C₁₄-15 linear alkyl sulfate 8, 1:4 maleic acid-acrylic acid copolymer 2, Na dodecylbenzensulfonate 6, Na tallow alkyl sulfate 1.5, ethoxylated cocoalkylmethylbis(hydroxyethyl)ammonium chloride 1.5, silicate 7, CMC 1, brightener 0.2, soap 1, and **Dequest** 2060 0.4 part; spray-on comprising C₁₄-15 alc. ethoxylate 2.5, C₁₂-15 alc. ethoxylate 2.5, silicone antifoam 0.3, and perfume 0.3 part; dry additives comprising Na carbonate 6, NaBO₂.3H₂O.H₂O₂ 18, NaBO₂.H₂O₂ 4, tetraacetythylenediamine 3, photoactivated bleach (**sulfonated** zinc phthalocyanine encapsulated in bleach dextrin sol. polymer) 0.02, protease 1, lipase 0.4, amylase 0.25; dry mixed Na sulfate 3.0; and moisture and others to 100 parts.

AN 1997:757094 CAPLUS

DN 128:36425

TI Detergent compositions providing excellent cleaning and whiteness performances on enzyme-sensitive and greasy soils and cleaning and building perfumes on surfaces using the same

IN Asano, Kaori; Scheibel, Jeffrey John; Baba, Hajime; Baeck, Andre Cesar; Minamikawa, Kaori; Cripe, Thomas Anthony; Curry, John Downing; Cron, Laura; Dodd, Ian Martin; Hartshorn, Richard Timothy; Speed, Lynda Anne; Katsuda, Rinko; Kvietok, Frank Andrej; Mao, Mark Hsiang-kuen; Moss, Michael Alan John; Murata, Susumu; Ohtani, Royohei; Panandiker, Rajan Keshav; Pramod, Kakumanu; Sarnaik, Khizar Mohamed Khan; Thoen, Christian Arthur Jacques Kamiel; Kott, Kevin Lee; Willman, Kenneth William; Okamoto, Mitsuyo; Foley, Peter Robert; Manohar, Sanjeev Krishnadas

PA Procter and Gamble Company, USA; Asano, Kaori; Scheibel, Jeffrey John; Baba, Hajime; Baeck, Andre Cesar; Minamikawa, Kaori; Cripe, Thomas Anthony; Curry, John Downing; Cron, Laura; Dodd, Ian Martin

SO PCT Int. Appl., 83 pp.

CODEN: PIXXD2

DT Patent

LA English

FAN.CNT 45

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9743387	A2	19971120	WO 1997-US8436	19970516
	WO 9743387	A3	19980219		
	W: BR, CA, CN, CZ, HU, JP, MX, TR, US				
	RW: AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
	ZA 9704225	A	19981228	ZA 1997-4225	19970515
	CA 2255008	AA	19971120	CA 1997-2255008	19970516
	CA 2255009	AA	19971120	CA 1997-2255009	19970516
	CA 2255012	AA	19971120	CA 1997-2255012	19970516
	EP 906400	A2	19990407	EP 1997-924761	19970516
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, PT, IE, FI				
	CN 1225679	A	19990811	CN 1997-196479	19970516
	CN 1225681	A	19990811	CN 1997-196480	19970516
	CN 1225670	A	19990811	CN 1997-196495	19970516
	CN 1225680	A	19990811	CN 1997-196499	19970516
	CN 1225671	A	19990811	CN 1997-196501	19970516
	CN 1225676	A	19990811	CN 1997-196502	19970516
	BR 9710668	A	19990817	BR 1997-10668	19970516
	JP 11511786	T2	19991012	JP 1997-541165	19970516
	US 6136769	A	20001024	US 1999-295421	19990421
PRAI	US 1996-17886P	P	19960517		
	WO 1997-US8436	W	19970516		

OS MARPAT 128:36425

AB The title compns. comprise an alkoxylated quaternary ammonium (AQA) cationic surfactant R1R2R3N+ApR4 X- (R1 = linear, branched, or substituted C8-18 alkyl, alkenyl, aryl, alkaryl, ether or glycidyl ether moiety; R2 = C1-3 alkyl; R3, R4 = H, Me, Et; X- = anion; A = C1-4 alkoxy; p = 2-30), a non-AQA surfactant and enzymes. A laundry detergent was prep'd. from blown powder comprising anhyd. sodium tripolyphosphate 24, Na C14-15 linear alkyl sulfate 8, 1:4 maleic acid-acrylic acid copolymer 2, Na dodecylbenzensulfonate 6, Na tallow alkyl sulfate 1.5, ethoxylated cocoalkyldimethylhydroxyethylammonium chloride 1.5, silicate 7, CMC 1, brightener 0.2, soap 1, and Dequest 2060 0.4 part; spray-on comprising C14-15 alc. ethoxylate 2.5, C12-15 alc. ethoxylate 2.5, silicone antifoam 0.3, and perfume 0.3 part; dry additives comprising Na carbonate 6, NaBO2.3H2O.H2O2 18, NaBO2.H2O2 4, tetraacetythylenediamine 3, photoactivated bleach (sulfonated zinc phthalocyanine encapsulated in bleach dextrin sol. polymer) 0.02, protease 1, lipase 0.4, amylase 0.25; dry mixed Na sulfate 3.0; and moisture and others to 100 parts.

AB The title compns. comprise an alkoxylated quaternary ammonium (AQA) cationic surfactant R1R2R3N+ApR4 X- (R1 = linear, branched, or substituted

C8-18 alkyl, alkenyl, aryl, alkaryl, ether or glycidyl ether moiety; R2 = C1-3 alkyl; R3, R4 = H, Me, Et; X- = anion; A = C1-4 alkoxy; p = 2-30), a non-AQA surfactant and enzymes. A laundry detergent was prepd. from blown powder comprising anhyd. sodium tripolyphosphate 24, Na C14-15 linear alkyl sulfate 8, 1:4 maleic acid-acrylic acid copolymer 2, Na dodecylbenzensulfonate 6, Na tallow alkyl sulfate 1.5, ethoxylated cocoalkyldimethylhydroxyethylammonium chloride 1.5, silicate 7, CMC 1, brightener 0.2, soap 1, and Dequest 2060 0.4 part; spray-on comprising C14-15 alc. ethoxylate 2.5, C12-15 alc. ethoxylate 2.5, silicone antifoam 0.3, and perfume 0.3 part; dry additives comprising Na carbonate 6, NaBO2.3H2O.H2O2 18, NaBO2.H2O2 4, tetraacetylenehydrazine 3, photoactivated bleach (sulfonated zinc phthalocyanine encapsulated in bleach dextrin sol. polymer) 0.02, protease 1, lipase 0.4, amylase 0.25; dry mixed Na sulfate 3.0; and moisture and others to 100 parts.

L3 ANSWER 97 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1997:366408 CAPLUS

DN 126:327995

TI Vapor sterilization using inorganic hydrogen peroxide complexes

IN Jacobs, Paul T.; Lin, Szu-min; Chen, Xiaolan

PA Johnson & Johnson Medical, Inc., USA

SO PCT Int. Appl., 96 pp.

CODEN: PIXXD2

DT Patent

LA English

FAN.CNT 6

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9715333	A1	19970501	WO 1996-US16570	19961023
	W: AU, BR, CA, JP, KR, MX, NO, SG				
	RW: AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
	US 5667753	A	19970916	US 1995-549425	19951027
	AU 9674462	A1	19970515	AU 1996-74462	19961023
	AU 704089	B2	19990415		
	EP 862466	A1	19980909	EP 1996-936565	19961023
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, FI				
	BR 9611152	A	19991228	BR 1996-11152	19961023
	JP 2002515770	T2	20020528	JP 1997-516655	19961023
PRAI	US 1995-549425	A	19951027		
	US 1994-234738	B2	19940428		
	US 1995-369786	A2	19950106		
	US 1996-716094	A	19960919		
	WO 1996-US16570	W	19961023		
AB	An app. and process for H2O2 vapor sterilization of medical instruments and similar devices make use of H2O2 vapor released from an inorg. H2O2 complex. The peroxide vapor can be released at room temp. and atm. pressure; however, the pressure used can be <50 torr and the temp. >86.degree. to facilitate the release of H2O2 vapor. Preferred H2O2 complexes for use in the invention include Na4P2O7.3H2O2 and KH2PO4.H2O2. The heating rate can be >5.degree./min. Optionally, a plasma can be used in conjunction with the vapor.				
IT	51-45-6D, Histamine, compd. with hydrogen peroxide (H2O2) 56-40-6D, Glycine, compd. with hydrogen peroxide (H2O2), biological studies 71-00-1D, L-Histidine, compd. with hydrogen peroxide (H2O2), biological studies 79-05-0D, Propionamide, compd. with hydrogen peroxide (H2O2) 96-31-1D, 1,3-Dimethylurea, compd. with hydrogen peroxide (H2O2) 108-19-0D, Biuret, compd. with hydrogen peroxide (H2O2) 124-43-6 144-55-8D, Sodium bicarbonate, compd. with hydrogen peroxide (H2O2) 583-52-8D, Potassium oxalate, compd. with hydrogen peroxide (H2O2) 584-08-7D, Potassium carbonate, compd. with hydrogen peroxide (H2O2) 584-09-8D, Rubidium carbonate, compd. with hydrogen peroxide (H2O2) 1305-62-0D, Calcium hydroxide (Ca(OH)2), compd. with hydrogen peroxide				

(H2O2), biological studies 7320-34-5D, Potassium pyrophosphate, compd. with hydrogen peroxide (H2O2) 7601-54-9D, compd. with hydrogen peroxide (H2O2) 7722-84-1, Hydrogen peroxide, biological studies 7722-88-5D, Tetrasodium pyrophosphate, compd. with hydrogen peroxide (H2O2) 7778-53-2D, Potassium orthophosphate, compd. with hydrogen peroxide (H2O2) 9002-89-5D, Poly(vinyl alcohol), compd. with hydrogen peroxide (H2O2) 9003-01-4D, Poly(acrylic acid), compd. with hydrogen peroxide (H2O2) 9003-05-8D, Polyacrylamide, compd. with hydrogen peroxide (H2O2) 9003-09-2D, Poly(vinyl methyl ether), compd. with hydrogen peroxide (H2O2) 9003-20-7D, Poly(vinyl acetate), compd. with hydrogen peroxide (H2O2) 9003-39-8D, PVP, compd. with hydrogen peroxide (H2O2) 9004-35-7D, Cellulose acetate, compd. with hydrogen peroxide (H2O2) 9005-22-5D, Cellulose sulfate sodium salt, compd. with hydrogen peroxide (H2O2) 9005-38-3D, Sodium alginate, compd. with hydrogen peroxide (H2O2) 15630-89-4, Sodium carbonate peroxide 25038-54-4D, Nylon 6, compd. with hydrogen peroxide (H2O2) 25038-87-3D, Poly(vinyl methyl ketone), compd. with hydrogen peroxide (H2O2) 25232-41-1D, Poly(4-vinylpyridine), compd. with hydrogen peroxide (H2O2) 32131-17-2D, Nylon 6,6, compd. with hydrogen peroxide (H2O2)
 RL: BAC (Biological activity or effector, except adverse); BSU (Biological study, unclassified); BIOL (Biological study)

(vapor sterilization using inorg. hydrogen peroxide complexes)

IT 7758-29-4DP, Triphosphoric acid pentasodium salt, compd. with hydrogen peroxide (H2O2) 13769-85-2P 15362-88-6P 22464-71-7P 22653-27-6P 33960-95-1P 52908-87-9P 122040-69-1P, Phosphoric acid, trisodium salt, compd. with hydrogen peroxide (H2O2) (1:5) 167874-81-9P 189760-17-6P 189760-18-7P 189760-19-8P 189760-20-1P 189760-21-2P 189760-22-3P 189760-24-5P 189760-26-7P 189760-28-9P 189760-29-0P 189760-31-4P 189760-33-6P 189760-35-8P 189760-37-0P 189760-39-2P 189760-43-8P 189760-45-0P 189760-48-3P 189760-49-4P 189760-51-8P 189760-53-0P 189760-55-2P 189760-56-3P 189760-57-4P 189760-58-5P 189760-59-6P 189760-60-9P 189760-61-0P

RL: BAC (Biological activity or effector, except adverse); BSU (Biological study, unclassified); SPN (Synthetic preparation); BIOL (Biological study); PREP (Preparation)

(vapor sterilization using inorg. hydrogen peroxide complexes)

L3 ANSWER 98 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1997:739355 CAPLUS

DN 128:41543

TI Photographic bleach solution

IN Fyson, John Richard

PA Eastman Kodak Co., USA

SO U.S., 8 pp., Cont. of U.S. Ser. No. 379,034, abandoned.

CODEN: USXXAM

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 5683858	A	19971104	US 1996-665991	19960619
PRAI	US 1993-94183		19930802		
	US 1995-379034		19950127		

OS MARPAT 128:41543

AB A photog. bleach soln. contains hydrogen peroxide or a compd. capable of releasing hydrogen peroxide and two or more sequestering agents capable of complexing with a transition metal and selected from the group consisting of polyalkylcarboxylic, polyalkylphosphonic, and polyalkylsulfonic acids. The bleach soln. has a pH in the alk. range. A method of using such a photog. bleach soln. is also disclosed.

IT 67-43-6, Diethylenetriaminepentaacetic acid 149-45-1, Tiron 2809-21-4,

1-Hydroxyethylidene-1,1-diphosphonic acid 30938-57-9
RL: TEM (Technical or engineered material use); USES (Uses)
(photog. bleach solns. contg. **hydrogen peroxide**
and)

IT 7758-98-7, Copper(II) **sulfate**, uses 7785-87-7, Manganese
sulfate 10028-22-5, Ferric **sulfate**
RL: TEM (Technical or engineered material use); USES (Uses)
(photog. bleach solns. contg. polyalkylcarboxylic, polyalkylphosphonic,
and polyalkylsulfonic acids and **hydrogen peroxide**
and)

L3 ANSWER 99 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1997:425940 CAPLUS

DN 127:142741

TI Processing silver halide photographic element with non-rehalogenating
peroxide bleaching composition

IN Haye, Shirleyanne Elizabeth; Bertucci, Sidney Joseph; Schmittou, Eric
Richard

PA Eastman Kodak Co., USA

SO U.S., 15 pp., Cont.-in-part of U.S. Ser. No. 230,365.

CODEN: USXXAM

DT Patent

LA English

FAN.CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 5641615	A	19970624	US 1995-546057	19951020
	JP 07295170	A2	19951110	JP 1995-93915	19950419
PRAI	US 1994-230365		19940420		

OS MARPAT 127:142741

AB An imagewise exposed and developed black-and-white or color silver halide
photog. element is effectively bleached using a non-rehalogenating
peroxide bleaching compn. This bleaching compn. comprises
hydrogen peroxide or a precursor that can release
hydrogen peroxide, at least one compd. of the formula
 $R(O)nSO_3M$ (R = an org. group having 1-10 C atoms; n = 0 or 1; M = H, an
alkali or alk. earth metal atom, or an ammonium ion), and a
phosphonic acid or a salt thereof and a pH of 2-6.

AB An imagewise exposed and developed black-and-white or color silver halide
photog. element is effectively bleached using a non-rehalogenating
peroxide bleaching compn. This bleaching compn. comprises
hydrogen peroxide or a precursor that can release
hydrogen peroxide, at least one compd. of the formula
 $R(O)nSO_3M$ (R = an org. group having 1-10 C atoms; n = 0 or 1; M = H, an
alkali or alk. earth metal atom, or an ammonium ion), and a
phosphonic acid or a salt thereof and a pH of 2-6.

IT Photographic processing
(nonrehalogenating bleaching compns. contg. **hydrogen**
peroxide, **sulfonic acids**, and **phosphonic**
acids for)

IT 7722-84-1, **Hydrogen peroxide**, uses

RL: TEM (Technical or engineered material use); USES (Uses)
(processing silver halide photog. elements with non-rehalogenating
bleaching compns. contg. **sulfonic acids**, **phosphonic**
acids and)

L3 ANSWER 100 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1997:231426 CAPLUS

DN 126:294769

TI Stabilized, silicate-free, hydrogen peroxide composition for use in
bleaching paper pulp, cotton and cotton blends

IN Moore, Samuel B.; Leuck, James F.; Turner, Edwin T.

PA Burlington Chemical Co., Inc., USA

SO U.S., 5 pp., Cont.-in-part of U.S. 5,464,563.

CODEN: USXXAM

DT Patent
LA English
FAN.CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 5616280	A	19970401	US 1995-553886	19951106
	US 5464563	A	19951107	US 1993-112582	19930825
PRAI	US 1993-112582		19930825		

AB The title bleaching compns. with freedom from conventionally used silicate stabilizer are obtained from NaOH, optical brighteners and an enhanced H₂O₂ (35-50%) contg. 0.05-1% Mg salt, 0.01-0.1% aminoalkylphosphonic acid and 0.01-0.1% dipicolinic acid as stabilizer, with the balance H₂O. Textile goods and pulp bleached with the compns. are soft, absorbent, silicate-free with a Hunter Scale whiteness of .gtorsim.85. Because a silicate-free stabilizer is used, low levels of extractable solids are obtained.

IT 142-72-3, Magnesium acetate 499-83-2, Dipicolinic acid 1310-58-3, Potassium hydroxide, uses 1310-73-2, Sodium hydroxide, uses 7487-88-9, Magnesium **sulfate**, uses 7722-84-1, Hydrogen peroxide, uses 13598-36-2D, **Phosphonic acid**, aminoalkyl derivs.

RL: TEM (Technical or engineered material use); USES (Uses)
(stabilized, silicate-free, **hydrogen peroxide**
compn. for use in bleaching paper pulp, cotton and cotton blends)

L3 ANSWER 101 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1997:537569 CAPLUS

DN 127:140186

TI Oxidative hair dyes containing glycerides

IN Okawa, Juichi; Hatanaka, Katsuto

PA Yamahatsu Sangyo Kaisha, Ltd., Japan

SO Jpn. Kokai Tokkyo Koho, 19 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 09169623	A2	19970630	JP 1996-100983	19960423
PRAI	JP 1995-272485		19951020		

AB Hair dyes comprise oxidative dyes, oxidants, and glycerol triesters. The prepn. provides effective hair dyeing at mild conditions, i.e. low pH, low concn. at a short time. A first prepn. contg. polyoxyethylene nonylphenyl ether 20, geraniol-modified alc. 10, and ammonia soln. (29 %) 5, p-phenylenediamine 0.3, m-phenylenediamine.cntdot.HCl 0.4, p-aminophenol 0.8, ammonium thioglycolate 0.4, Na edetate 0.2, triacetyl glycerin 2, and distd. water to 100 parts, was mixed with a second prepn. contg. Na polyoxyethylene lauryl **sulfate** 2, cetanol 1.5, **phosphoric acid** 0.0002, H₂O₂ soln. (35 %) 17, and distd. water to 100 parts. The final product (1:1 mixt.) showed improved hair dyeing effects.

AB Hair dyes comprise oxidative dyes, oxidants, and glycerol triesters. The prepn. provides effective hair dyeing at mild conditions, i.e. low pH, low concn. at a short time. A first prepn. contg. polyoxyethylene nonylphenyl ether 20, geraniol-modified alc. 10, and ammonia soln. (29 %) 5, p-phenylenediamine 0.3, m-phenylenediamine.cntdot.HCl 0.4, p-aminophenol 0.8, ammonium thioglycolate 0.4, Na edetate 0.2, triacetyl glycerin 2, and distd. water to 100 parts, was mixed with a second prepn. contg. Na polyoxyethylene lauryl **sulfate** 2, cetanol 1.5, **phosphoric acid** 0.0002, H₂O₂ soln. (35 %) 17, and distd. water to 100 parts. The final product (1:1 mixt.) showed improved hair dyeing effects.

L3 ANSWER 102 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1997:397122 CAPLUS
DN 127:36237
TI Liquid bleach compositions with good storability at low temperature
IN Miyamae, Yoshitaka; Shindo, Hiroyuki
PA Lion Corp., Japan
SO Jpn. Kokai Tokkyo Koho, 8 pp.
CODEN: JKXXAF
DT Patent
LA Japanese
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 09104900	A2	19970422	JP 1995-287834	19951009
PRAI	JP 1995-287834		19951009		

OS MARPAT 127:36237

AB The title compns. (pH 2.5-4.5) contain (A) surfactant(s) chosen from R1O(CH2CH2O)nH, R1C6H4O(CH2CH2O)nH, and R1N+Me2CH2CH2CH2SO3-, (B) surfactant(s) chosen from R2Me2N.fwdarw.O, R2Me2N+CH2CH2CH2CO2-, R2O(Glu)m glucoside, and R2CO2(Glu-OMe), linear alkylbenzenesulfonates, .alpha.-olefinsulfoantes, and polyethylene glycol alkyl or alkenyl ether **sulfate salts**; (C) 0.5-10% **hydrogen peroxide**, (D) 0.2-5% R3CO2C6H4Z peroxide precursor, and (E) 0.001-5% org. peroxy compds. R4COOOH (R1, R2 = C8-18 alkyl, alkenyl; n = 2-30; m = 1-10; R3 = C7-12 alkyl; Z = CO2H, CO2M, SO3H, SO3M; M = alkali metal; R4 = C2-12 alkyl) at B/A ratio 1/0 to 10/1 and A + B content 5-30%. A compn. (pH 3) comprised C13H27O(CH2CH2O)9H 20, dodecyldimethylamine oxide 2, **H2O2** 5, C7H15CO2C6H4SO3Na 1, peroctanoic acid 2, ethane-1-hydroxy-1,1-**diphosphonic** acid 0.1, and water to 100%.

AB The title compns. (pH 2.5-4.5) contain (A) surfactant(s) chosen from R1O(CH2CH2O)nH, R1C6H4O(CH2CH2O)nH, and R1N+Me2CH2CH2CH2SO3-, (B) surfactant(s) chosen from R2Me2N.fwdarw.O, R2Me2N+CH2CH2CH2CO2-, R2O(Glu)m glucoside, and R2CO2(Glu-OMe), linear alkylbenzenesulfonates, .alpha.-olefinsulfoantes, and polyethylene glycol alkyl or alkenyl ether **sulfate salts**; (C) 0.5-10% **hydrogen peroxide**, (D) 0.2-5% R3CO2C6H4Z peroxide precursor, and (E) 0.001-5% org. peroxy compds. R4COOOH (R1, R2 = C8-18 alkyl, alkenyl; n = 2-30; m = 1-10; R3 = C7-12 alkyl; Z = CO2H, CO2M, SO3H, SO3M; M = alkali metal; R4 = C2-12 alkyl) at B/A ratio 1/0 to 10/1 and A + B content 5-30%. A compn. (pH 3) comprised C13H27O(CH2CH2O)9H 20, dodecyldimethylamine oxide 2, **H2O2** 5, C7H15CO2C6H4SO3Na 1, peroctanoic acid 2, ethane-1-hydroxy-1,1-**diphosphonic** acid 0.1, and water to 100%.

L3 ANSWER 103 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1997:397123 CAPLUS
DN 127:36239
TI Liquid bleach compositions with good storability and fragrance retention
IN Miyamae, Yoshitaka; Fukano, Kazuaki
PA Lion Corp., Japan
SO Jpn. Kokai Tokkyo Koho, 11 pp.
CODEN: JKXXAF
DT Patent
LA Japanese
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 09104898	A2	19970422	JP 1995-287833	19951009
PRAI	JP 1995-287833		19951009		

OS MARPAT 127:36239

AB The title compns. (pH 2.5-4.5) contain (A) surfactant(s) chosen from R1O(CH2CH2O)nH, R1C6H4O(CH2CH2O)nH, and R1N+Me2CH2CH2CH2SO3-, (B) surfactant(s) chosen from R2Me2N.fwdarw.O, R2Me2N+CH2CH2CH2CO2-, R2O(Glu)m glucoside, and R2CO2(Glu-OMe), linear alkylbenzenesulfonates, .alpha.-olefinsulfoantes, and polyethylene glycol alkyl or alkenyl ether **sulfate salts**; (C) 0.5-10% **hydrogen peroxide**,

(D) 0.2-5% R3CO2C6H4Z peroxide precursor, and 0.001-5% phenolic compds. (R1, R2 = C8-18 alkyl, alkenyl; n = 2-30; m = 1-10; R3 = C7-12 alkyl; Z = CO2H, CO2M, SO3H, SO3M; M = alkali metal) at B/A ratio 1/0 to 10/1 and A + B content 5-30%. A compn. (pH 3) comprised C13H27O(CH2CH2O)9H 20, H2O2 5, C7H15CO2C6H4SO3Na 2, PhOH 2, ethane-1-hydroxy-1,1-diphosphonic acid 0.1, perfume mixt. 0.1, and water to 100%.

AB The title compns. (pH 2.5-4.5) contain (A) surfactant(s) chosen from R1O(CH2CH2O)nH, R1C6H4O(CH2CH2O)nH, and R1N+Me2CH2CH2CH2SO3-, (B) surfactant(s) chosen from R2Me2N.fwdarw.O, R2Me2N+CH2CH2CH2CO2-, R2O(Glu)m glucoside, and R2CO2(Glu-OMe), linear alkylbenzenesulfonates, .alpha.-olefinsulfoantes, and polyethylene glycol alkyl or alkenyl ether sulfate salts; (C) 0.5-10% **hydrogen peroxide**, (D) 0.2-5% R3CO2C6H4Z peroxide precursor, and 0.001-5% phenolic compds. (R1, R2 = C8-18 alkyl, alkenyl; n = 2-30; m = 1-10; R3 = C7-12 alkyl; Z = CO2H, CO2M, SO3H, SO3M; M = alkali metal) at B/A ratio 1/0 to 10/1 and A + B content 5-30%. A compn. (pH 3) comprised C13H27O(CH2CH2O)9H 20, H2O2 5, C7H15CO2C6H4SO3Na 2, PhOH 2, ethane-1-hydroxy-1,1-diphosphonic acid 0.1, perfume mixt. 0.1, and water to 100%.

L3 ANSWER 104 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1997:346911 CAPLUS

DN 127:19977

TI Thermally stable liquid hydrogen peroxide-based bleaching compositions

IN Miyamae, Yoshitaka; Fukano, Kazuaki

PA Lion Corp., Japan

SO Jpn. Kokai Tokkyo Koho, 6 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 09078094	A2	19970325	JP 1995-259348	19950912
PRAI	JP 1995-259348		19950912		

AB The compns. (pH 2-5), suitable for removal of fruit or vegetable stains from fabrics, contain (A) H2O2 2.0-10, (B) double bond-contg. anionic surfactants 0.5-10, (C) non-crosslinkable carboxylic acid-based polymers 1.0-10, and (D) org. P compds. 0-3.0%. Thus, an aq. soln. (pH 4.0) contg. H2O2 5, poly(Na acrylate) 3, a 6:4 mixt. of Na C14 .alpha.-olefinsulfonate and Na hydroxymyristylsulfonate 3, and Ferriox CY 115 (1-hydroxyethylidene-1,1-diphosphonic acid) 0.5% was applied on a fabric stained with grape juice to show good bleaching effect.

AB The compns. (pH 2-5), suitable for removal of fruit or vegetable stains from fabrics, contain (A) H2O2 2.0-10, (B) double bond-contg. anionic surfactants 0.5-10, (C) non-crosslinkable carboxylic acid-based polymers 1.0-10, and (D) org. P compds. 0-3.0%. Thus, an aq. soln. (pH 4.0) contg. H2O2 5, poly(Na acrylate) 3, a 6:4 mixt. of Na C14 .alpha.-olefinsulfonate and Na hydroxymyristylsulfonate 3, and Ferriox CY 115 (1-hydroxyethylidene-1,1-diphosphonic acid) 0.5% was applied on a fabric stained with grape juice to show good bleaching effect.

IT Sulfonic acids, uses

RL: TEM (Technical or engineered material use); USES (Uses)
(alkene, sodium salts; thermally stable liq. **hydrogen peroxide**-based bleaching compns. for fabrics)

IT 2809-21-4, 1-Hydroxyethylidene-1,1-diphosphonic acid

RL: TEM (Technical or engineered material use); USES (Uses)
(Ferriox CY 115; thermally stable liq. **hydrogen peroxide**-based bleaching compns. for fabrics)

L3 ANSWER 105 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1997:481079 CAPLUS

DN 127:92433

TI Assay utilizing hydrogen peroxide adduct

IN Kuzuya, Keiko; Yamauchi, Tadakazu

PA Mochida Pharmaceutical Co., Ltd., Japan
 SO Eur. Pat. Appl., 53 pp.
 CODEN: EPXXDW
 DT Patent
 LA English
 FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 781850	A2	19970702	EP 1996-120736	19961223
	EP 781850	A3	19980812		
	R: AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LI, LU, MC, NL, PT, SE				
	CA 2193918	AA	19970629	CA 1996-2193918	19961224
	JP 09234097	A2	19970909	JP 1996-349790	19961227
	US 5858645	A	19990112	US 1996-773181	19961227
PRAI	JP 1995-343822		19951228		

AB Improvement in assays utilizing at least hydrogen peroxide for one anal. reagent is provided. The assay of the present invention employs a stable hydrogen peroxide adduct in dry state that has no adverse effects on the assay and that has a high hydrogen peroxide-retaining ability. In the assay, an aq. soln. is added to an adduct in dry state of (1) at least one member selected from the group consisting of a carboxylic acid and a salt thereof, **phosphoric acid** and a salt thereof, and a **sulfonic acid** and a salt thereof and (2) **hydrogen peroxide** to generate peroxide, and the thus generated peroxide is used for the anal. reagent.

AB Improvement in assays utilizing at least hydrogen peroxide for one anal. reagent is provided. The assay of the present invention employs a stable hydrogen peroxide adduct in dry state that has no adverse effects on the assay and that has a high hydrogen peroxide-retaining ability. In the assay, an aq. soln. is added to an adduct in dry state of (1) at least one member selected from the group consisting of a carboxylic acid and a salt thereof, **phosphoric acid** and a salt thereof, and a **sulfonic acid** and a salt thereof and (2) **hydrogen peroxide** to generate peroxide, and the thus generated peroxide is used for the anal. reagent.

ST hydrogen peroxide adduct dry reagent analysis; carboxylate adduct hydrogen peroxide dry reagent; phosphate adduct hydrogen peroxide dry reagent; **sulfonate adduct hydrogen peroxide dry reagent**; immunoassay hydrogen peroxide dry reagent

IT **Sulfonic acids, uses**
 RL: ARG (Analytical reagent use); PEP (Physical, engineering or chemical process); ANST (Analytical study); PROC (Process); USES (Uses)
 (amino alkane, **hydrogen peroxide** adducts;
hydrogen peroxide adducts in stable dry anal. reagent)

IT **Sulfonic acids, uses**
 RL: ARG (Analytical reagent use); PEP (Physical, engineering or chemical process); ANST (Analytical study); PROC (Process); USES (Uses)
 (amino alkane, salts, **hydrogen peroxide** adducts;
hydrogen peroxide adducts in stable dry anal. reagent)

IT Carboxylic acids, uses
 Phosphates, uses
Sulfonates
Sulfonic acids, uses
 RL: ARG (Analytical reagent use); ANST (Analytical study); USES (Uses)
 (**hydrogen peroxide** adducts; **hydrogen peroxide** adducts in stable dry anal. reagent)

IT **Sulfonic acids, uses**
 RL: ARG (Analytical reagent use); PEP (Physical, engineering or chemical process); ANST (Analytical study); PROC (Process); USES (Uses)
 (hydroxy alkane, **hydrogen peroxide** adducts;
hydrogen peroxide adducts in stable dry anal.

reagent)

IT **Sulfonic acids, uses**
 RL: ARG (Analytical reagent use); PEP (Physical, engineering or chemical process); ANST (Analytical study); PROC (Process); USES (Uses)
 (hydroxy alkane, salts, **hydrogen peroxide** adducts;
hydrogen peroxide adducts in stable dry anal.
 reagent)

IT **Sulfonic acids, uses**
 RL: ARG (Analytical reagent use); PEP (Physical, engineering or chemical process); ANST (Analytical study); PROC (Process); USES (Uses)
 (hydroxy benzene, **hydrogen peroxide** adducts;
hydrogen peroxide adducts in stable dry anal.
 reagent)

IT **Sulfonic acids, uses**
 RL: ARG (Analytical reagent use); PEP (Physical, engineering or chemical process); ANST (Analytical study); PROC (Process); USES (Uses)
 (hydroxy benzene, salts, **hydrogen peroxide** adducts;
hydrogen peroxide adducts in stable dry anal.
 reagent)

IT 7664-38-2D, **Phosphoric acid, hydrogen peroxide** adducts, uses 7722-84-1D, Hydrogen peroxide, adducts 81398-06-3
 RL: ARG (Analytical reagent use); ANST (Analytical study); USES (Uses)
 (**hydrogen peroxide** adducts in stable dry anal.
 reagent)

IT 50-21-5D, hydrogen peroxide adducts, uses 50-81-7D, Ascorbic acid, hydrogen peroxide adducts, uses 50-99-7D, D-Glucose, hydrogen peroxide adducts, uses 56-41-7D, L-Alanine, hydrogen peroxide adducts, uses 56-45-1D, L-Serine, hydrogen peroxide adducts, uses 56-86-0D, L-Glutamic acid, hydrogen peroxide adducts, uses 57-13-6D, Urea, hydrogen peroxide adducts, uses 57-50-1D, hydrogen peroxide adducts 59-23-4D, Galactose, hydrogen peroxide adducts 60-27-5D, Creatinine, hydrogen peroxide adducts 60-35-5D, Acetamide, hydrogen peroxide adducts, uses 63-42-3D, Lactose, hydrogen peroxide adducts 64-19-7D, Acetic acid, hydrogen peroxide adducts, uses 69-65-8D, Mannitol, hydrogen peroxide adducts 69-79-4D, Maltose, hydrogen peroxide adducts 72-19-5D, L-Threonine, hydrogen peroxide adducts, uses 74-89-5D, Methylamine, hydrogen peroxide adducts 77-92-9D, hydrogen peroxide adducts, uses 79-09-4D, Propionic acid, hydrogen peroxide adducts, uses 79-14-1D, Glycolic acid, hydrogen peroxide adducts, uses 87-69-4D, L-(+)-Tartaric acid, hydrogen peroxide adducts, uses 88-99-3D, 1,2-Benzenedicarboxylic acid, hydrogen peroxide adducts, uses 97-67-6D, (S)-(-)-Malic acid, hydrogen peroxide adducts 98-98-6D, Picolinic acid, hydrogen peroxide adducts 99-20-7D, Trehalose, hydrogen peroxide adducts 100-21-0D, 1,4-Benzenedicarboxylic acid, hydrogen peroxide adducts, uses 107-35-7D, Taurine, hydrogen peroxide adducts 107-36-8D, Isethionic acid, hydrogen peroxide adducts 110-15-6D, Butanedioic acid, hydrogen peroxide adducts, uses 110-16-7, 2-Butenedioic acid (Z)-, uses 110-17-8D, 2-Butenedioic acid (E)-, hydrogen peroxide adducts, uses 110-99-6D, Diglycolic acid, hydrogen peroxide adducts 111-16-0D, Pimelic acid, hydrogen peroxide adducts 121-34-6D, Vanillic acid, hydrogen peroxide adducts 121-91-5D, 1,3-Benzenedicarboxylic acid, hydrogen peroxide adducts, uses 124-04-9D, Adipic acid, hydrogen peroxide adducts, uses 127-07-1D, Hydroxyurea, hydrogen peroxide adducts 127-17-3D, Pyruvic acid, hydrogen peroxide adducts, uses 133-37-9D, DL-Tartaric acid, hydrogen peroxide adducts 140-95-4D, Dimethylolurea, hydrogen peroxide adducts 141-43-5D, hydrogen peroxide adducts 141-82-2D, Malonic acid, hydrogen peroxide adducts, uses 142-45-0D, Acetylenedicarboxylic acid, hydrogen peroxide adducts 144-62-7D, Ethanedioic acid, hydrogen peroxide adducts, uses 156-38-7D, 4-Hydroxyphenylacetic acid, hydrogen peroxide adducts 328-42-7D, Oxaloacetic acid, hydrogen peroxide adducts 328-50-7D, 2-Ketoglutaric acid, hydrogen peroxide adducts 372-75-8D, L-Citrulline, hydrogen peroxide adducts 462-60-2D, Hydantoic acid, hydrogen peroxide adducts 473-81-4D, DL-Glyceric acid, hydrogen peroxide adducts 498-21-5D,

Methylsuccinic acid, hydrogen peroxide adducts 498-23-7D, Citraconic acid, hydrogen peroxide adducts 498-40-8D, L-Cysteic acid, hydrogen peroxide adducts 503-49-1D, 3-Hydroxy-3-methylglutaric acid, hydrogen peroxide adducts 503-66-2D, 3-Hydroxypropionic acid, hydrogen peroxide adducts 505-48-6D, Octanedioic acid, hydrogen peroxide adducts 516-05-2D, Methylmalonic acid, hydrogen peroxide adducts 526-95-4D, Gluconic acid, hydrogen peroxide adducts 526-99-8D, Mucic acid, hydrogen peroxide adducts 542-05-2D, 3-Ketoglutaric acid, hydrogen peroxide adducts 543-24-8D, N-Acetylglycine, hydrogen peroxide adducts 597-43-3D, 2,2-Dimethylsuccinic acid, hydrogen peroxide adducts 610-09-3D, cis-1,2-Cyclohexanedicarboxylic acid, hydrogen peroxide adducts 610-35-5D, 4-Hydroxyphthalic acid, hydrogen peroxide adducts 618-83-7D, hydrogen peroxide adducts 626-51-7D, 3-Methylglutaric acid, hydrogen peroxide adducts 636-46-4D, 4-Hydroxyisophthalic acid, hydrogen peroxide adducts 1076-97-7D, 1,4-Cyclohexanedicarboxylic acid, hydrogen peroxide adducts 1483-07-4D, Albizziine, hydrogen peroxide adducts 2305-32-0D, trans-1,2-Cyclohexanedicarboxylic acid, hydrogen peroxide adducts 3233-59-8D, hydrogen peroxide adducts 4023-65-8D, trans-Aconitic acid, hydrogen peroxide adducts 4534-68-3D, meso-Butane-1,2,3,4-tetracarboxylic acid, hydrogen peroxide adducts 4767-03-7D, hydrogen peroxide adducts 6236-09-5D, (S)-(+)-Citramalic acid, hydrogen peroxide adducts 6556-12-3D, Glucuronic acid, hydrogen peroxide adducts 6915-15-7D, DL-Malic acid, hydrogen peroxide adducts 6940-58-5D, 1,3,5-Pentanetricarboxylic acid, hydrogen peroxide adducts 7440-09-7D, Potassium, salts, uses 7440-23-5D, Sodium, salts, uses 7512-17-6D, hydrogen peroxide adducts 7632-05-5D, Sodium phosphate, hydrogen peroxide adducts 7647-14-5, Sodium chloride (NaCl), uses 9003-01-4D, Polyacrylic acid, hydrogen peroxide adducts 9004-32-4D, Sodium carboxymethylcellulose, hydrogen peroxide adducts 9004-54-0D, Dextran, hydrogen peroxide adducts, uses 9005-38-3D, Sodium alginate, hydrogen peroxide adducts 9082-07-9D, Sodium chondroitin sulfate, hydrogen peroxide adducts 16068-46-5D, Potassium phosphate, hydrogen peroxide adducts 23351-51-1D, Glucoheptonic acid, hydrogen peroxide adducts 25322-68-3D, hydrogen peroxide adducts 25525-21-7D, Glucaric acid, hydrogen peroxide adducts 26239-55-4D, N-(2-Acetamido)iminodiacetic acid, hydrogen peroxide adducts

RL: ARG (Analytical reagent use); PEP (Physical, engineering or chemical process); ANST (Analytical study); PROC (Process); USES (Uses)

(hydrogen peroxide adducts in stable dry anal. reagent)

L3 ANSWER 106 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1997:145194 CAPLUS

DN 126:145623

TI Laundry pre-treatment agent with improved fabric/color safety

IN Del, Duca Valerio; Ricci, Carlo; Trani, Marina

PA The Procter & Gamble Company, USA

SO Eur. Pat. Appl., 13 pp.

CODEN: EPXXDW

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 752469	A1	19970108	EP 1995-870082	19950705
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IE, IT, LI, LU, NL, PT, SE				
	CA 2226105	AA	19970123	CA 1996-2226105	19960614
	WO 9702331	A1	19970123	WO 1996-US10347	19960614
	W: AL, AM, AT, AU, AZ, BB, BG, BR, BY, CA, CH, CN, CZ, DE, DK, EE, ES, FI, GB, GE, HU, IS, JP, KE, KG, KP, KR, KZ, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK				
	RW: KE, LS, MW, SD, SZ, UG, BF, BJ, CF, CG, CI, CM, GA, GN, ML, MR, NE, SN, TD, TG				

AU 9661778	A1	19970205	AU 1996-61778	19960614
AU 706041	B2	19990610		
CN 1195366	A	19981007	CN 1996-196749	19960614
BR 9609555	A	19990302	BR 1996-9555	19960614
JP 11513050	T2	19991109	JP 1996-505154	19960614
RU 2145630	C1	20000220	RU 1998-102028	19960614
NZ 310889	A	20000228	NZ 1996-310889	19960614
FI 9800009	A	19980204	FI 1998-9	19980102
US 6001794	A	19991214	US 1998-981950	19980326
PRAI EP 1995-870082	A	19950705		
WO 1996-US10347	W	19960614		
AB	<p>The use of a liq. compn. comprising a peroxygen bleach and .gtoreq.1 radical scavenger, for pretreating soiled fabrics before the fabrics are washed, shows reduced loss of tensile strength in the fabrics and/or reduced color damage of the fabrics. A compn. contg. alkyl sulfate, Dobanol 91-10, Dobanol 23-3, Dequest, H2O2, Isofol 12, BHT, and water applied as pre-treatment agent to cotton ribbons showed tensile strength loss 18%, vs. 42% for a compn. not contg. BHT. Addn. of hydroquinone to the above compn. further reduced tensile strength loss to 6%.</p>			
AB	<p>The use of a liq. compn. comprising a peroxygen bleach and .gtoreq.1 radical scavenger, for pretreating soiled fabrics before the fabrics are washed, shows reduced loss of tensile strength in the fabrics and/or reduced color damage of the fabrics. A compn. contg. alkyl sulfate, Dobanol 91-10, Dobanol 23-3, Dequest, H2O2, Isofol 12, BHT, and water applied as pre-treatment agent to cotton ribbons showed tensile strength loss 18%, vs. 42% for a compn. not contg. BHT. Addn. of hydroquinone to the above compn. further reduced tensile strength loss to 6%.</p>			
L3	ANSWER 107 OF 158 CAPLUS COPYRIGHT 2003 ACS			
AN	1997:274146 CAPLUS			
DN	126:343432			
TI	A halide-free method for olefin epoxidation with 30% hydrogen peroxide			
AU	Sato, Kazuhiko; Aoki, Masao; Ogawa, Masami; Hashimoto, Tadashi; Panyella, David; Noyori, Ryoji			
CS	Dep. Chem. and Molecular Chirality Res. Unit, Nagoya Univ., Nagoya, 464-01, Japan			
SO	Bulletin of the Chemical Society of Japan (1997), 70(4), 905-915 CODEN: BCSJA8; ISSN: 0009-2673			
PB	Nippon Kagakkai			
DT	Journal			
LA	English			
OS	CASREACT 126:343432			
AB	<p>A catalytic system consisting of sodium tungstate dihydrate, (aminomethyl) phosphonic acid, and methyltrioctylammonium hydrogen sulfate, effects the epoxidn. of olefins using 30% hydrogen peroxide with a substrate-to-catalyst molar ratio of 50-500. The reaction proceeds in high yield without solvents, or, alternatively, with added toluene under entirely halide-free conditions. Lipophilic ammonium hydrogen sulfate, which replaces the conventional chloride, and an (.alpha.-aminoalkyl)phosphonic acid are crucial for the high reactivity. This method is operationally simple, environmentally benign, and much more economical than the oxidn. with m-chloroperbenzoic acid, allowing for a large-scale prepn. of epoxides. Various substrates including terminal olefins, 1,1- and 1,2-disubstituted olefins, cyclic olefins, and tri- and tetrasubstituted olefins as well as allylic alcs., esters, .alpha.,.beta.-unsatd. ketones, and ethers can be epoxidized in high yield. The scope and limitations of this new reaction system are discussed.</p>			
AB	<p>A catalytic system consisting of sodium tungstate dihydrate, (aminomethyl) phosphonic acid, and methyltrioctylammonium hydrogen sulfate, effects the epoxidn. of olefins using 30% hydrogen peroxide with a substrate-to-catalyst molar</p>			

ratio of 50-500. The reaction proceeds in high yield without solvents, or, alternatively, with added toluene under entirely halide-free conditions. Lipophilic ammonium hydrogen sulfate, which replaces the conventional chloride, and an (.alpha.-aminoalkyl)phosphonic acid are crucial for the high reactivity. This method is operationally simple, environmentally benign, and much more economical than the oxidn. with m-chloroperbenzoic acid, allowing for a large-scale prepn. of epoxides. Various substrates including terminal olefins, 1,1- and 1,2-disubstituted olefins, cyclic olefins, and tri- and tetrasubstituted olefins as well as allylic alcs., esters, .alpha.,.beta.-unsatd. ketones, and ethers can be epoxidized in high yield. The scope and limitations of this new reaction system are discussed.

IT Epoxidation catalysts

(halide-free epoxidn. of olefins with **hydrogen peroxide** catalyzed by sodium tungstate dihydrate, (aminomethyl) **phosphonic acid**, and methyltrioctylammonium hydrogen sulfate)

IT 1066-51-9, (Aminomethyl)**phosphonic acid** 13472-45-2, Sodium tungstate 59158-14-4, Methyltrioctylammonium hydrogen sulfate

RL: CAT (Catalyst use); USES (Uses)

(halide-free epoxidn. of olefins with **hydrogen peroxide**)

L3 ANSWER 108 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1996:676090 CAPLUS

DN 125:308644

TI Aerosols containing hydrogen peroxide

IN Kawase, Jiro

PA Kao Corp, Japan

SO Jpn. Kokai Tokkyo Koho, 18 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 08231345	A2	19960910	JP 1995-64984	19950227
PRAI	JP 1995-64984		19950227		

AB An aerosol is constructed such that the H2O2 compn. is stabilized for a long period and used for spraying odorous substances from hair bleach, hair color, waving lotion, cleanser, etc. A bag made of ethylene-vinyl alc. copolymer, polypropylene, polyacrylonitrile, polyethylene terephthalate, etc., contains H2O2 and it is placed in a metallic cylindrical container, and a compressed gas is filled in the space between the bag and the cylindrical container.

IT 134-31-6, Oxyquinoline **sulfate** 7664-38-2, **Phosphoric acid**, uses

RL: NUU (Other use, unclassified); USES (Uses)
(in aerosols contg. **hydrogen peroxide**)

L3 ANSWER 109 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1996:672616 CAPLUS

DN 125:308658

TI Double-layered wall bag-type aerosol containers for hydrogen peroxide-containing hair sprays

IN Kawase, Jiro

PA Kao Corp, Japan

SO Jpn. Kokai Tokkyo Koho, 17 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
--	------------	------	------	-----------------	------

PI JP 08230959 A2 19960910 JP 1995-64983 19950227
 PRAI JP 1995-64983 19950227

AB Hydrogen peroxide-contg. hair spray compns. are filled into double-layered wall bag-type aerosol containers to improve the storage stability. The containers comprise an outer can, an inner bag (as wall bag), and a valve for aerosol release. The hair aerosol compns. contained a propellant (i.e. liquefied petroleum gas), stabilizer selected from **phosphoric acid** and oxyquinolin **sulfate**, and pH 1.0-4.0 **hydrogen peroxide** soln. in addn. to other ingredients.

AB Hydrogen peroxide-contg. hair spray compns. are filled into double-layered wall bag-type aerosol containers to improve the storage stability. The containers comprise an outer can, an inner bag (as wall bag), and a valve for aerosol release. The hair aerosol compns. contained a propellant (i.e. liquefied petroleum gas), stabilizer selected from **phosphoric acid** and oxyquinolin **sulfate**, and pH 1.0-4.0 **hydrogen peroxide** soln. in addn. to other ingredients.

IT 134-31-6, Oxyquinoline **sulfate** 7664-38-2, **Phosphoric acid**, biological studies 7722-84-1, Hydrogen peroxide, biological studies

RL: BUU (Biological use, unclassified); BIOL (Biological study); USES (Uses)

(**hydrogen peroxide**-contg. hair spray compns. in double-layered wall bag-type aerosol containers to improve the storage stability)

L3 ANSWER 110 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1996:649311 CAPLUS

DN 125:279302

TI Oxygen bleach compositions with no irritating odor

IN Saiga, Koji; Yamada, Hiroyuki

PA Kao Corp, Japan

SO Jpn. Kokai Tokkyo Koho, 10 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 08209194	A2	19960813	JP 1995-17043	19950203
JP 1995-17043		19950203		

PI JP 08209194 A2 19960813 JP 1995-17043 19950203

PRAI JP 1995-17043 19950203

OS MARPAT 125:279302

AB Title compns., useful as household bleaches for removing molds, contain

(A) H₂O₂ or peroxides capable of forming H₂O₂ in H₂O 0.1-98, (B) bleaching activators 0.002-70, and (C) H₂O-sol. azides 0.00001-10%. Thus, an aq. soln. contg. H₂O₂ 6, PhO(CH₂CH₂O)2CH₂CO2CH₂CH₂SO₃Na 15, **Dequest** 2015DN 0.1% was mixed with an aq. soln. contg. 0.0003% NaN₃ and 20% K₂CO₃ to obtain a bleach showing good bleaching effects on an ABS resin plate stained by Cladosporium herbarum.

AB Title compns., useful as household bleaches for removing molds, contain (A) H₂O₂ or peroxides capable of forming H₂O₂ in H₂O 0.1-98, (B) bleaching activators 0.002-70, and (C) H₂O-sol. azides 0.00001-10%. Thus, an aq. soln. contg. H₂O₂ 6, PhO(CH₂CH₂O)2CH₂CO2CH₂CH₂SO₃Na 15, **Dequest** 2015DN 0.1% was mixed with an aq. soln. contg. 0.0003% NaN₃ and 20% K₂CO₃ to obtain a bleach showing good bleaching effects on an ABS resin plate stained by Cladosporium herbarum.

ST **hydrogen peroxide** azide **sulfonate** bleach odorless

IT Bleaching agents

Fungicides and Fungistats

(odorless bleaches contg. **hydrogen peroxide**, azides, and **sulfonate** derivs. for removing molds)

IT 135261-32-4 162136-44-9
RL: TEM (Technical or engineered material use); USES (Uses)
(bleaching activators; odorless bleaches contg. **hydrogen peroxide**, azides, and **sulfonate** derivs. for removing molds)

IT 7722-84-1, **Hydrogen peroxide**, uses 26628-22-8,
Sodium azide
RL: TEM (Technical or engineered material use); USES (Uses)
(odorless bleaches contg. **hydrogen peroxide**, azides, and **sulfonate** derivs. for removing molds)

L3 ANSWER 111 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1996:271293 CAPLUS

DN 124:320187

TI Bleaching detergent compositions containing hydrogen peroxide generators and peroxy acid precursors

IN Ogura, Nobuyuki; Aoyanagi, Muneo; Yamaguchi, Yukyoshi; Tsucha, Shigemi

PA Kao Corp, Japan

SO Jpn. Kokai Tokkyo Koho, 11 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 08041493	A2	19960213	JP 1994-177068	19940728
PRAI	JP 1994-177068		19940728		
OS	MARPAT 124:320187				

AB Title compns., which show high detergency for oily soils without bleaching dyed colors, contain 1-95% inorg. peroxides generating H₂O₂ in H₂O, 0.1-30% RCOL [R = ((substituted and/or connecting group-contg.) C₅-22 alkyl-, alkenyl-, C₁-18 alkyl-, alkenyl-substituted) aryl; L = leaving group] as precursors of peroxy acid bleaching agents, 0.1-100 ppm .gtoreq.1 metal cation(s) chosen from Mn, Cu, and Fe, and 0.1-50% anionic, nonionic, cationic, and/or amphoteric surfactants. The peroxy acid precursors show HLB .ltoreq.23 measured by the method of J. T. Davies [Proc. Intern. Cong. Surface Active Substances, 2nd (London), vol. 1, p426 (1957)] and generate .gtoreq.1 mM RCO₃COR (R = same as above) from aq. solns. contg. 10 mM the precursors and 50 mM H₂O₂ at 20.degree. 10 min after. Thus, Na₂C₂O₆ 80, 4-C₈H₁₇CO₂C₆H₄SO₃Na 5, C₁₂H₂₅OSO₃Na 5, Na ethane-1-hydroxy-1,1-diphosphonate 5 parts, and 30 ppm (as metal ion) Fe₃(PO₄)₂ were mixed to give a bleaching compn.

IT 60-00-4, EDTA, uses 151-21-3, Sodium dodecyl **sulfate**, uses 1313-13-9, Manganese oxide (MnO₂), uses 2211-98-5, Sodium 4-dodecylbenzenesulfonate 3645-16-7, Ferrous octanoate 7414-83-7, Sodium ethane-1-hydroxy-1,1-diphosphonate 7439-89-6, Iron, uses 7439-96-5, Manganese, uses 7440-50-8, Copper, uses 7773-01-5, Manganese chloride 9002-92-0, Polyethylene glycol monododecyl ether 13197-76-7 13444-72-9, Manganese **sulfate** [Mn₂(SO₄)₃] 14940-41-1, Iron phosphate [Fe₃(PO₄)₂] 17599-81-4, Copper **sulfate** (Cu₂SO₄) 60472-42-6, Sokalan CP 5
RL: TEM (Technical or engineered material use); USES (Uses)
(bleaching detergent compns. contg. **hydrogen peroxide** generators and peroxy acid precursors)

L3 ANSWER 112 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1996:241771 CAPLUS

DN 124:320205

TI Bleaching detergent compositions containing hydrogen peroxide-releasing agents and peroxy acid precursors

IN Ogura, Nobuyuki; Yamaguchi, Yukyoshi; Aoyanagi, Muneo; Tsucha, Shigemi; Hanada, Yoshuki; Saito, Shinya

PA Kao Corp, Japan

SO Jpn. Kokai Tokkyo Koho, 14 pp.

CODEN: JKXXAF

DT Patent
LA Japanese
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 08027484	A2	19960130	JP 1994-168144	19940720
	JP 3330230	B2	20020930		
PRAI	JP 1994-168144		19940720		

OS MARPAT 124:320205

AB Title compns., effective to lipophilic stains from fabrics, etc., contain (A) 1-95% inorg. peroxides releasing H₂O₂ in aq. soln., (B) 0.1-30% peroxy acid precursors RCOL [R = (substituted or linkage group-contg.) linear or branched C₇-22 alkyl, alkenyl, C₁-18 alkyl, (alkyl-substituted) aryl; L = leaving group; RCO₂Na corresponding to RCOL-derived org. peracids shows HLB (Davies method) .ltoreq.23; producing .gtoreq.1 mM RCO₃COR from aq. soln. of 10 mM RCOL and 50 mM H₂O₂ after 10 min at 20.degree.], (C) 0.05-10% (based on B) QSO₃M and 0.01-10% (based on B) QCO₂M [Q = (substituted or linkage-contg.) C₁-12 alkyl, alkenyl, (substituted or linkage-contg.) aryl; M = H, (in)org. cation], and (D) 0.1-50% anionic, nonionic, cationic, or amphoteric surfactants, in which, B and C are blended into the total compns. as their fine mixed powder prepd. by dissolving or dispersing in solvents of sp. dielec. const. .gtoreq.15 at 20.degree. or 25.degree.. Thus, a compn. of Na percarbonate 80, 50:0.5:0.5 powd. mixt. of p-C₁₁H₂₃CO₂C₆H₄SO₃Na (HLB 20.9, producing 3.2 mM diacyloyl peroxide), Na p-phenolsulfonate, and lauric acid 5, C₁₂H₂₅OSO₃Na 5, Na ethane-1-hydroxy-1,1-diphosphonate 5, and Na₂CO₃ 5% was added in a com. laundering detergent aq. soln. to show bleaching effect on stain of curry and prevention of discoloration on dyed color.

IT Bleaching agents
Detergents
Discoloration prevention
Surfactants

(detergent bleaching agents contg. **hydrogen peroxide**
-releasing compds., peroxy acid precursors, **sulfonic acids**,
carboxylic acids, and surfactants)

IT Peroxides, uses

RL: TEM (Technical or engineered material use); USES (Uses)
(detergent bleaching agents contg. **hydrogen peroxide**
-releasing compds., peroxy acid precursors, **sulfonic acids**,
carboxylic acids, and surfactants)

IT 50-78-2, Acetylsalicylic acid 56-40-6, Glycine, uses 56-41-7, Alanine, uses 65-85-0, Benzoic acid, uses 69-72-7, Salicylic acid, uses 75-75-2, Methanesulfonic acid 77-92-9, Citric acid, uses 79-14-1, Glycolic acid, uses 87-69-4, Tartaric acid, uses 88-99-3, Phthalic acid, uses 96-77-5, Phenoldisulfonic acid 98-67-9, p-Phenolsulfonic acid 99-96-7, p-Hydroxybenzoic acid, uses 100-21-0, Terephthalic acid, uses 104-15-4, p-Toluenesulfonic acid, uses 107-35-7, Taurine 107-36-8, Isethionic acid 110-15-6, Succinic acid, uses 110-16-7, Maleic acid, uses 110-17-8, Fumaric acid, uses 121-57-3, Sulfanilic acid 124-04-9, Adipic acid, uses 143-07-7, Lauric acid, uses 334-48-5, Capric acid 609-46-1, o-Phenolsulfonic acid 657-84-1, Sodium p-toluenesulfonate 825-90-1, Sodium p-phenolsulfonate 5138-18-1, Sulfosuccinic acid 6915-15-7, Malic acid

RL: MOA (Modifier or additive use); USES (Uses)
(detergent bleaching agents contg. **hydrogen peroxide**
-releasing compds., peroxy acid precursors, **sulfonic acids**,
carboxylic acids, and surfactants)

IT 3313-92-6, Sodium percarbonate

RL: TEM (Technical or engineered material use); USES (Uses)
(detergent bleaching agents contg. **hydrogen peroxide**
-releasing compds., peroxy acid precursors, **sulfonic acids**,
carboxylic acids, and surfactants)

IT 60-00-4, EDTA, uses 7414-83-7, Sodium ethane-1-hydroxy-1,1-

diphosphonate

RL: MOA (Modifier or additive use); USES (Uses)
(metal-masking agents; for detergent bleaching agents contg.
hydrogen peroxide releasers, peroxy acid precursors,
sulfonic acids, carboxylic acids, and surfactants)

IT 89740-11-4 89740-13-6 172284-81-0 175596-80-2 176310-20-6

RL: TEM (Technical or engineered material use); USES (Uses)
(peroxy acid precursors; detergent bleaching agents contg.
hydrogen peroxide-releasing compds., peroxy acid
precursors, **sulfonic** acids, carboxylic acids, and
surfactants)

IT 67-56-1, Methanol, uses 67-64-1, Acetone, uses 68-12-2, DMF, uses
75-05-8, Acetonitrile, uses 127-19-5, Dimethylacetamide 7732-18-5,
Water, uses 39799-78-5

RL: NUU (Other use, unclassified); USES (Uses)
(solvents; for detergent bleaching agents contg. **hydrogen**
peroxide releasers, peroxy acid precursors, **sulfonic**
acids, carboxylic acids, and surfactants)

IT 151-21-3, Sodium dodecyl **sulfate**, uses 9002-92-0,
Poly(oxyethylene) dodecyl ether 13197-76-7

RL: MOA (Modifier or additive use); USES (Uses)
(surfactants; detergent bleaching agents contg. **hydrogen**
peroxide-releasing compds., peroxy acid precursors,
sulfonic acids, carboxylic acids, and surfactants)

L3 ANSWER 113 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1996:591239 CAPLUS

DN 125:251042

TI The production of heavy duty fabric washing detergents containing hydrogen
peroxide

AU Croud, V. B.; Tompsett, S. J.

CS Warwick International Ltd, Clwyd, CH8 9HE, UK

SO World Surfactants Congress, 4th, Barcelona, June 3-7, 1996 (1996), Volume
2, 98-108 Publisher: Asociacion Espanola de Productores de Sustancias para
Aplicaciones Tensioactivas, Barcelona, Spain.

CODEN: 63KCAH

DT Conference

LA English

AB Incorporation of hydrogen peroxide into alk. liq. detergents is difficult
due to the inherent instability of the peroxide under alk. conditions and
the tendency of the peroxide to react with other formulation ingredients.
This latter interaction tends to cause a decrease in the soln. pH. The
peroxide can be stabilized under alk. conditions by use of an appropriate
choice of sequestrants and a pretreatment process. Sequestrants evaluated
include: diethylenetriamine pentamethylenephosphonate (DTPMP),
dipyridylamine (DPA), ethylene diamine tetramethylene phosphonate (EDTMP),
amino-tris(methylenephosphonate) (ATMP), hydroxyethylidene diphosphonic
acid (HEDP), diamino-cyclohexane tetramethylene phosphonate (DACHH), and
dimethylglyoxime (DMG). The efficacy of the sequestrants was evaluated
against Cu, Co, Fe, and Mn. It is possible to minimize the pH drift by
the choice of the surfactants and other active materials, and also by the
incorporation of suitable buffering materials. Combining these
techniques, enabled development of alk. liq. detergents contg. hydrogen
peroxide. Test formulations contained paraffin **sulfonate**, LAS,
alc. ethoxylate (7EO), **hydrogen peroxide**, NaOH, DTPMP,
DPA, DMG, and deionized water. A formulation for color-safe detergents
contained LABS [linear alkylbenzene **sulfonate**], 7EO, NaOH,
hydrogen peroxide, borax, sodium citrate, sodium
sesquicarbonate, DTPMP, and DPA.

AB Incorporation of hydrogen peroxide into alk. liq. detergents is difficult
due to the inherent instability of the peroxide under alk. conditions and
the tendency of the peroxide to react with other formulation ingredients.
This latter interaction tends to cause a decrease in the soln. pH. The
peroxide can be stabilized under alk. conditions by use of an appropriate

choice of sequestrants and a pretreatment process. Sequestrants evaluated include: diethylenetriamine pentamethylenephosphonate (DTPMP), dipyridylamine (DPA), ethylene diamine tetramethylene phosphonate (EDTMP), amino-tris(methylenephosphonate) (ATMP), hydroxyethylidene diphosphonic acid (HEDP), diamino-cyclohexane tetramethylene phosphonate (DACHH), and dimethylglyoxime (DMG). The efficacy of the sequestrants was evaluated against Cu, Co, Fe, and Mn. It is possible to minimize the pH drift by the choice of the surfactants and other active materials, and also by the incorporation of suitable buffering materials. Combining these techniques, enabled development of alk. liq. detergents contg. hydrogen peroxide. Test formulations contained paraffin **sulfonate**, LAS, alc. ethoxylate (7EO), **hydrogen peroxide**, NaOH, DTPMP, DPA, DMG, and deionized water. A formulation for color-safe detergents contained LABS [linear alkylbenzene **sulfonate**], 7EO, NaOH, **hydrogen peroxide**, borax, sodium citrate, sodium sesquicarbonate, DTPMP, and DPA.

ST hydrogen peroxide stabilization alk laundry detergent; sequestant stabilization hydrogen peroxide buffering detergent; amine **phosphonate** sequestant **hydrogen peroxide** detergent

L3 ANSWER 114 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1995:643678 CAPLUS
 DN 123:260506
 TI Liquid bleaching agent compositions
 IN Ogura, Nobuyuki; Aoyanagi, Muneo; Shimizu, Toshio
 PA Kao Corp, Japan
 SO Jpn. Kokai Tokkyo Koho, 14 pp.
 CODEN: JKXXAF
 DT Patent
 LA Japanese
 FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 07082591	A2	19950328	JP 1993-228791	19930914
	JP 3330200	B2	20020930		
PRAI	JP 1993-228791		19930914		

OS MARPAT 123:260506

AB Bleaching compns. with storage stability contain (a) nonionic surfactants, (b) **H2O2**, (c) certain kind of bleaching activating agents which are carboxylic acid esters having a carboxylate-, **sulfonate**-, or ammonium-contg. alc. part, and (d) quaternary ammonium compds. or anionic compds. contg. carboxylate, sulfonate, or **phosphonate** group. Thus, a bleaching agent contained 20% monolauryl ether of polyethylene glycol (av.-no. of oxyethylene 10), 1% C11H23CO2C6H4SO3Na, 2% C12H25Me3N+ Cl-, 0.1% ethane-1-hydroxy-1,1-diphosphonate, 5% **H2O2**, and balance H2O.

AB Bleaching compns. with storage stability contain (a) nonionic surfactants, (b) **H2O2**, (c) certain kind of bleaching activating agents which are carboxylic acid esters having a carboxylate-, **sulfonate**-, or ammonium-contg. alc. part, and (d) quaternary ammonium compds. or anionic compds. contg. carboxylate, sulfonate, or **phosphonate** group. Thus, a bleaching agent contained 20% monolauryl ether of polyethylene glycol (av.-no. of oxyethylene 10), 1% C11H23CO2C6H4SO3Na, 2% C12H25Me3N+ Cl-, 0.1% ethane-1-hydroxy-1,1-diphosphonate, 5% **H2O2**, and balance H2O.

L3 ANSWER 115 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1995:520581 CAPLUS
 DN 122:248024
 TI Hydrogen peroxide preparations for oxidative hair dyeing and bleaching
 IN Hoeffkes, Horst; Neuhaus, Winifried; Nelles, Karin
 PA Henkel K.-G.a.A., Germany
 SO Ger. Offen., 8 pp.

CODEN: GWXXBX

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 4333370	A1	19950406	DE 1993-4333370	19930930
	WO 9508978	A1	19950406	WO 1994-EP3118	19940919
	W: PL, US				
	RW: AT, BE, CH, DE, DK, ES, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
	EP 721321	A1	19960717	EP 1994-927624	19940919
	EP 721321	B1	19981111		
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, NL, PT, SE				
	AT 173151	E	19981115	AT 1994-927624	19940919
	ES 2124914	T3	19990216	ES 1994-927624	19940919
	PL 176726	B1	19990730	PL 1994-313770	19940919
	US 5651793	A	19970729	US 1996-619679	19960328
PRAI	DE 1993-4333370		19930930		
	WO 1994-EP3118		19940919		

AB Aq. H2O2 prepns. for oxidative hair dyeing and bleaching contain an anionic surfactant foaming agent, a carboxylated polymer thickening agent, and 0.001-0.05 wt.% foaming regulator comprising polysiloxane with a small amt. of finely divided, hydrophobic silicic acid. Thus, a compn. contg. 50% aq. H2O2 12.0, 28% aq. ethoxylated lauryl/myristyl ether Na sulfate 2.0, 25% aq. Et acrylate/methacrylic acid copolymer 15.0, dimethylpolysiloxane contg. 10-20 wt.% hydrophobic silicic acid 0.0067, 1-hydroxyethane-1,1-diphosphonic acid 1.0, ethoxylated hydrogenated castor oil 0.05, and water to 100 wt.% was readily miscible with an oxidative hair dye cream with little foaming on either mixing or rinsing. This compn. was stored at 45.degree. for 6 mo without loss of foam-regulating properties.

AB Aq. H2O2 prepns. for oxidative hair dyeing and bleaching contain an anionic surfactant foaming agent, a carboxylated polymer thickening agent, and 0.001-0.05 wt.% foaming regulator comprising polysiloxane with a small amt. of finely divided, hydrophobic silicic acid. Thus, a compn. contg. 50% aq. H2O2 12.0, 28% aq. ethoxylated lauryl/myristyl ether Na sulfate 2.0, 25% aq. Et acrylate/methacrylic acid copolymer 15.0, dimethylpolysiloxane contg. 10-20 wt.% hydrophobic silicic acid 0.0067, 1-hydroxyethane-1,1-diphosphonic acid 1.0, ethoxylated hydrogenated castor oil 0.05, and water to 100 wt.% was readily miscible with an oxidative hair dye cream with little foaming on either mixing or rinsing. This compn. was stored at 45.degree. for 6 mo without loss of foam-regulating properties.

L3 ANSWER 116 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1995:362384 CAPLUS

DN 122:141515

TI Treatment of photographic effluents by ultraviolet radiation and hydrogen peroxide for removal of organic compounds.

IN Gehin, Guy Michel; Caillaud, Olivier Andre Georges

PA Kodak-Pathe, Fr.; Eastman Kodak Co.

SO Eur. Pat. Appl., 12 pp.

CODEN: EPXXDW

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 622339	A1	19941102	EP 1994-420125	19940419
	EP 622339	B1	19980701		
	R: CH, DE, FR, GB, IT, LI				
	FR 2704543	A1	19941104	FR 1993-5412	19930430
	FR 2704543	B1	19950713		
	US 5439599	A	19950808	US 1994-231216	19940422

JP 07047374 A2 19950221 JP 1994-89777 19940427
PRAI FR 1993-5412 19930430

AB The present invention relates to a process for eliminating the org. compds. from aq. effluents having a COD >5 g/L. The process consists of treating the effluents with a sufficient quantity of light energy originating from a polychromatic UV source in the presence of H2O2. The quality of oxygen provided by the hydrogen peroxide expressed in grams per L is between 0.5 and 3 times the COD of the effluent to be treated and the pH is equal to or <10.5. This process applies to photog. effluents such as those from color development baths for silver halide photog. products.

IT 64-19-7, Acetic acid, processes 92-43-3, Phenidone 99-11-6, Citrazinic acid 100-51-6, Benzenemethanol, processes 106-50-3, p-Phenylene diamine, processes 106-50-3D, p-Phenylene diamine, derivs. 111-30-8, Glutaraldehyde 111-46-6, processes 123-31-9, Hydroquinone, processes 123-31-9D, Hydroquinone, derivs. 1310-58-3, Potassium hydroxide, processes 2235-43-0 7647-15-6, Sodium bromide, processes 7664-38-2, **Phosphoric acid**, processes 7681-11-0, Potassium iodide, processes 7757-83-7, Sodium sulfite 10039-54-0, Hydroxylamine **sulfate** 25646-71-3 27516-53-6, Pentenoic acid

RL: REM (Removal or disposal); PROC (Process)
(treatment of photog. effluents by UV radiation and **hydrogen peroxide** for removal of org. compds.)

L3 ANSWER 117 OF 158 CAPLUS COPYRIGHT 2003 ACS DUPLICATE 5
AN 1993:168345 CAPLUS
DN 118:168345

TI Agricultural and horticultural oxygen-supplying compositions containing hydrogen peroxide, inorganic salt fertilizers, and inorganic acids.

IN Takahashi, Naokazu
PA Nissan Chemical Industries, Ltd., Japan
SO Jpn. Kokai Tokkyo Koho, 3 pp.
CODEN: JKXXAF

DT Patent
LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 05004902	A2	19930114	JP 1991-158303	19910628
	JP 06096483	B4	19941130		
PRAI	JP 1991-158303		19910628		

AB Stable O-supplying compns. for agriculture and horticulture contain H2O2, inorg. salt fertilizers, and inorg. acids. A compn. (H2O2 content 4.88%) contg. 35% aq. H2O2 140, (NH4)2SO4 25, 85% H3PO4 2, and H2O 833 g was kept in white polyethylene bottles at room temp. for 4 mo to show 4.73% H2O2, vs. 4.19%, for the control formulated without H3PO4.

IT Phosphates, uses
Sulfates, uses

RL: USES (Uses)
(oxygen-supplying compns. contg. **hydrogen peroxide** and inorg. acids and, stable, for soils)

IT 7722-76-1, Ammonium primary phosphate 7783-20-2, Ammonium **sulfate**, uses

RL: USES (Uses)
(oxygen-supplying compns. contg. **hydrogen peroxide** and inorg. acids and, stable, for soils)

IT 7664-38-2, **Phosphoric acid**, uses 7664-93-9, Sulfuric acid, uses

RL: USES (Uses)
(oxygen-supplying compns. contg. **hydrogen peroxide** and inorg. salt fertilizers and, stable, for soils)

L3 ANSWER 118 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 1995:615086 CAPLUS
DN 123:11774

TI Finishing of coarse cellulose fabrics
 IN Kovalchuk, Lyudmila S.; Aftaeva, Nataliya S.; Doronina, Nataliya S.;
 Akulova, Lyudmila K.
 PA Tsentralnyj Nauchno-Issledovatel'skiy Institut Khlopchatobumazhnoj
 Promyshlennosti, USSR
 SO U.S.S.R.
 From: Izobreteniya 1993, (21), 66-7.
 CODEN: URXXAF
 DT Patent
 LA Russian
 FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	SU 1819926	A1	19930607	SU 1990-4885403	19901025
PRAI	SU 1990-4885403		19901025		

AB The consumption of water, amt. of wastewater, and amt. of silicate ppts. are decreased in upgrading gray cellulose textile material by alk. cooking with steaming, scouring in an aq. H₂SO₄ soln., alk.-peroxide brightening in the presence of a wetting agent and sodium silicate with specific d. 1.44 with steaming and counter-current intermittent washing between stages, and final washing with hot and cold water. The scouring soln. comprises H₂SO₄ 2-3; chloramine B or chloramine KhB 1.0-1.5; Sulfoxide 31 wetting agent (mixt. of hydroxyethylated amides of C₁₀-C₁₄ synthetic fatty acids with the degree of hydroxyethylation 5-6 and C₁₁-12 alkyl benzene **sulfonates**) 0.3-0.5; hydroxyethylidene **diphosphonic acid** (I) complexing agent 0.1-0.2; and H₂O₂ 0.3-0.5 g/L and is used at 80-958.degree.. Alk.-peroxide brightening is carried out in the presence of Mg chloride or sulfate stabilizer, I, and nonionic wetting agent syntanol BV. The impregnating brightening soln. comprises H₂O₂ 3-5; Na silicate with 1.44 specific d. 5-6; Mg chloride or **sulfate** 0.5-0.7; I 0.2-0.4; Syntanol BV 0.5-1.0; NaOH up to total alky. 1.2-1.4 g/L; and by intermittent washing after each cooking and souring with hot water.

AB The consumption of water, amt. of wastewater, and amt. of silicate ppts. are decreased in upgrading gray cellulose textile material by alk. cooking with steaming, scouring in an aq. H₂SO₄ soln., alk.-peroxide brightening in the presence of a wetting agent and sodium silicate with specific d. 1.44 with steaming and counter-current intermittent washing between stages, and final washing with hot and cold water. The scouring soln. comprises H₂SO₄ 2-3; chloramine B or chloramine KhB 1.0-1.5; Sulfoxide 31 wetting agent (mixt. of hydroxyethylated amides of C₁₀-C₁₄ synthetic fatty acids with the degree of hydroxyethylation 5-6 and C₁₁-12 alkyl benzene **sulfonates**) 0.3-0.5; hydroxyethylidene **diphosphonic acid** (I) complexing agent 0.1-0.2; and H₂O₂ 0.3-0.5 g/L and is used at 80-958.degree.. Alk.-peroxide brightening is carried out in the presence of Mg chloride or sulfate stabilizer, I, and nonionic wetting agent syntanol BV. The impregnating brightening soln. comprises H₂O₂ 3-5; Na silicate with 1.44 specific d. 5-6; Mg chloride or **sulfate** 0.5-0.7; I 0.2-0.4; Syntanol BV 0.5-1.0; NaOH up to total alky. 1.2-1.4 g/L; and by intermittent washing after each cooking and souring with hot water.

L3 ANSWER 119 OF 158 JAPIO COPYRIGHT 2003 JPO

AN 1993-148658 JAPIO

TI ELECTROLESS TIN PLATING METHOD

IN TAKATSU AKIO; NIHEI TOMOMICHI

PA SUMITOMO METAL MINING CO LTD

PI JP 05148658 A 19930615 Heisei

AI JP 1991-334454 (JP03334454 Heisei) 19911122

PRAI JP 1991-334454 19911122

SO PATENT ABSTRACTS OF JAPAN (CD-ROM), Unexamined Applications, Vol. 1993

AB PURPOSE: To enhance reliability in mounting a multi-pin TAB tape, etc., by treating the surface of a material to be plated with a liq. mixture of the acid, peroxide and org. high molecular compd. and then electroless-plating

the material.

CONSTITUTION: Copper as a material to be plated is electroless-plated with tin. In this case, the surface of the material is treated with a liq. mixture of the acid, peroxide and org. high molecular compd., and then the material is electroless-plated. The plated material is further annealed. A soln. contg. at least one kind among sulfuric acid, hydrochloric acid, nitric acid and **phosphoric acid** is used as the acid.

Ammonium persulfate or **hydrogen peroxide** is used as the peroxide. A compd. obtained by linking a straight-chain hydrocarbon or cyclic-carbon compd. to the **sulfate** group or phosphate group or the salt of the compd. are used as the org. compd. Consequently, any trouble is hardly caused in the electric-component circuit.

COPYRIGHT: (C)1993,JPO&Japio

AB . . . plated material is further annealed. A soln. contg. at least one kind among sulfuric acid, hydrochloric acid, nitric acid and **phosphoric acid** is used as the acid. Ammonium persulfate or **hydrogen peroxide** is used as the peroxide. A compd. obtained by linking a straight-chain hydrocarbon or cyclic-carbon compd. to the **sulfate** group or phosphate group or the salt of the compd. are used as the org. compd. Consequently, any trouble is. . .

L3 ANSWER 120 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1992:594626 CAPLUS

DN 117:194626

TI Continuous process for the manufacture of sodium percarbonate from sodium carbonate and hydrogen peroxide

IN Sasaki, Hiromi; Ueda, Hiroshi; Muneno, Yasushi; Kusunoki, Hirokazu; Yamamoto, Sunao

PA Central Glass Co., Ltd., Japan

SO Eur. Pat. Appl., 10 pp.

CODEN: EPXXDW

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 496430	A1	19920729	EP 1992-101194	19920124
	R: DE, FR, GB, SE				
	US 5294427	A	19940315	US 1993-15052	19930208
PRAI	JP 1991-7412		19910125		
	JP 1991-238368		19910918		
	US 1992-826531		19920127		

AB A Na₂CO₃ and a H₂O₂ soln. are continuously fed to a 1st vessel to form a seed crystal-contg. Na percarbonate slurry, the slurry is continuously fed into a 2nd vessel along with .gtoreq.1 of the Na₂CO₃ and H₂O₂ to form a 2nd slurry contg. final Na percarbonate crystals grown from the seed crystals, the 2nd slurry is continuously discharged, the final Na percarbonate crystals are sepd. from the mother liquor, and the Na percarbonate crystals are granulated to a particle size .ltoreq.100 .mu.m.

IT 60-00-4, EDTA, uses 1312-76-1, Potassium silicate 1344-09-8, Sodium silicate 7487-88-9, Magnesium **sulfate**, uses 7647-14-5, Sodium chloride, uses 7664-38-2, **Phosphoric acid**, uses 7757-82-6, Sodium **sulfate**, uses

RL: USES (Uses)

(solns. contg., in sodium percarbonate manuf. from sodium carbonate and **hydrogen peroxide**, continuous process for)

L3 ANSWER 121 OF 158 JAPIO COPYRIGHT 2003 JPO

AN 1992-114734 JAPIO

TI DEODORANT

IN YOSHIDA NORIKAZU; KIMURA YOSHIKI; SEO SUMIMASA; IGUCHI TSUTOMU

PA NIPPON KAYAKU CO LTD

PI JP 04114734 A 19920415 Heisei

AI JP 1990-233416 (JP02233416 Heisei) 19900905

PRAI JP 1990-233416 19900905
 SO PATENT ABSTRACTS OF JAPAN (CD-ROM), Unexamined Applications, Vol. 1992
 AB PURPOSE: To decompose a malodorous component, rapidly and to use a deodorant for a long period of time by adding manganese or a manganese compound and peroxide.
 CONSTITUTION: A deodorant containing at least one component selected from manganese and a manganese compound, peroxide and **phosphoric acid, phosphorous acid**, hypophosphorous acid or a salt thereof being an arbitrary component is obtained.
 As the manganese compound, oxide, carbonate, **sulfate**, oxalate and permanganate are pref. As peroxide, **hydrogen peroxide**, sodium peroxide, potassium peroxide, calcium peroxide, peracetic acid and benzoyl peroxide are pref. This deodorant rapidly decomposes a malodorous component and can be used for a long period of time.

COPYRIGHT: (C)1992,JPO&Japio

AB . . . and peroxide.
 CONSTITUTION: A deodorant containing at least one component selected from manganese and a manganese compound, peroxide and **phosphoric acid, phosphorous acid**, hypophosphorous acid or a salt thereof being an arbitrary component is obtained.
 As the manganese compound, oxide, carbonate, **sulfate**, oxalate and permanganate are pref. As peroxide, **hydrogen peroxide**, sodium peroxide, potassium peroxide, calcium peroxide, peracetic acid and benzoyl peroxide are pref. This deodorant rapidly decomposes a malodorous component. . .

L3 ANSWER 122 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1991:689786 CAPLUS
 DN 115:289786
 TI Electrochemical synthesis of hydrogen peroxide
 IN Berzins, Talvildis; Gosser, Lawrence Wayne
 PA du Pont de Nemours, E. I., and Co., USA
 SO PCT Int. Appl., 20 pp.
 CODEN: PIXXD2

DT Patent
 LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9114802	A1	19911003	WO 1991-US645	19910205
	W: AU, BR, CA, FI, JP, KR, NO				
	RW: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LU, NL, SE				
	US 5112702	A	19920512	US 1990-636868	19901231
	AU 9172417	A1	19911021	AU 1991-72417	19910205
	CN 1055963	A	19911106	CN 1991-102474	19910319
PRAI	US 1990-495674		19900319		
	US 1990-636868		19901231		
	WO 1991-US645		19910205		

AB The synthesis of H2O2 occurs by the electrochem. redn. of O in the presence of acid and halide, preferably bromide, conducted in an electrolytic cell or in the cathode compartment of a fuel cell, where H2O2 is produced in com. useful concns.

IT **Sulfonic acids**, uses and miscellaneous
 RL: USES (Uses)

(in electrochem. prodn. of **hydrogen peroxide**)

IT 7601-90-3, Perchloric acid, uses and miscellaneous 7647-01-0, Hydrochloric acid, uses and miscellaneous 7664-38-2, **Phosphoric acid**, uses and miscellaneous 7664-93-9, Sulfuric acid, uses and miscellaneous 7697-37-2, Nitric acid, uses and miscellaneous 16887-00-6, Chloride, uses and miscellaneous 24959-67-9, Bromide, uses and miscellaneous
 RL: USES (Uses)

(in electrochem. prodn. of **hydrogen peroxide**)

L3 ANSWER 123 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1991:650354 CAPLUS

DN 115:250354

TI Citric and tartaric acid-containing solutions as hydrogen peroxide stabilizer

IN Kelly, Felix; McKay, Colin; Steed, Brian Harold

PA Aquaclear International Ltd., UK

SO PCT Int. Appl., 34 pp.

CODEN: PIXXD2

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9108981	A2	19910627	WO 1990-GB1968	19901217
	WO 9108981	A3	19910905		
	W: AT, AU, BB, BG, BR, CA, CH, DE, DK, ES, FI, GB, HU, JP, KP, KR, LK, LU, MC, MG, MW, NL, NO, RO, SD, SE, SU, US				
	RW: AT, BE, BF, BJ, CF, CG, CH, CM, DE, DK, ES, FR, GA, GB, GR, IT, LU, ML, MR, NL, SE, SN, TD, TG				
	AU 9168997	A1	19910718	AU 1991-68997	19901217
	ZA 9101882	A	19911224	ZA 1991-1882	19910314
PRAI	ZA 1989-9597		19891215		
	WO 1990-GB1968		19901217		
AB	A H2O2-stabilizing soln. comprises 1-7% citric acid, 7-12% tartaric acid and a chelating and/or buffering agent. Ag+ are optionally present. The stabilized H2O2 is useful as a disinfectant for swimming pools and breweries. A stabilizer conc. comprises tartaric acid 80, citric acid 21, and 85% H3PO4 50 g in 1 L aq. soln.				
IT	77-92-9, Citric acid, uses and miscellaneous 87-69-4, uses and miscellaneous 7664-38-2, Phosphoric acid, uses and miscellaneous 7761-88-8, Silver nitrate, uses and miscellaneous 7783-90-6, Silver chloride, uses and miscellaneous 7783-99-5, Silver nitrite 10294-26-5, Silver sulfate 14701-21-4, Silver ion, uses and miscellaneous				
	RL: USES (Uses)				
	(stabilizers soln. contg., for hydrogen peroxide)				

L3 ANSWER 124 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1992:91132 CAPLUS

DN 116:91132

TI Low-foam oxidative hair dyes

IN Hoeffkes, Horst; Nelles, Karin; Neuhaus, Winifried

PA Henkel K.-G.a.A., Germany

SO Ger. Offen., 5 pp.

CODEN: GWXXBX

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 4018259	A1	19911212	DE 1990-4018259	19900607
	WO 9118584	A1	19911212	WO 1991-EP999	19910529
	W: FI, HU, JP, KR, PL, US				
	RW: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LU, NL, SE				
	EP 531391	A1	19930317	EP 1991-910298	19910529
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, NL, SE				
	JP 05507476	T2	19931028	JP 1991-509536	19910529
PRAI	DE 1990-4018259		19900607		
	WO 1991-EP999		19910529		
OS	MARPAT 116:91132				
AB	H2O2 compns. for oxidative hair dyeing comprise an anionic surfactant, a thickening CO2H-contg. polymer, and a nonionic surfactant RO(C2H4O)xR1 [R				

= C10-22 alkyl or alkenyl, C6-15 alkylphenyl; R1 = (C3H6O)yH, CnH2n+1 (C3H6O)yCnH2n+1; x = 1-10; y = 3-5; n = 1-4]. The compns. are free of the abundant foaming shown by conventional anionic surfactant-contg. prepns. A compn. contained 50% H2O2 12, 28% aq. Na lauryl/myristyl polyglycol ether sulfate 2, 25% aq. Latekoll D 16, coco fatty alkyl polyoxyethyl Bu ether 0.5, 1-hydroxyethane 1,1-diphosphonic acid 0.3, and water to 100% by wt. The H2O2 compn. was used in conjunction with known dye creams.

AB H2O2 compns. for oxidative hair dyeing comprise an anionic surfactant, a thickening CO2H-contg. polymer, and a nonionic surfactant RO(C2H4O)xR1 [R = C10-22 alkyl or alkenyl, C6-15 alkylphenyl; R1 = (C3H6O)yH, CnH2n+1 (C3H6O)yCnH2n+1; x = 1-10; y = 3-5; n = 1-4]. The compns. are free of the abundant foaming shown by conventional anionic surfactant-contg. prepns. A compn. contained 50% H2O2 12, 28% aq. Na lauryl/myristyl polyglycol ether sulfate 2, 25% aq. Latekoll D 16, coco fatty alkyl polyoxyethyl Bu ether 0.5, 1-hydroxyethane 1,1-diphosphonic acid 0.3, and water to 100% by wt. The H2O2 compn. was used in conjunction with known dye creams.

L3 ANSWER 125 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1991:188463 CAPLUS

DN 114:188463

TI Process for the regeneration of sulfuric acid from the purification of yellow phosphorus

IN Pieper, Werner; Biermann, Walter

PA Hoechst A.-G., Germany

SO Ger. Offen., 3 pp.

CODEN: GWXXBX

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 3926112	A1	19910214	DE 1989-3926112	19890808
	US 5122359	A	19920616	US 1990-557928	19900725
PRAI	DE 1989-3926112		19890808		

AB The process comprises mixing the contaminated H2SO4 with H2O2, and reacting the mixt. in the presence of a catalyst at 110.degree.-210.degree. in concd. H2SO4. This process is esp. suitable for regenerating the H2SO4 from the electrothermal process for the manuf. of yellow P. The resulting H2SO4 may afterwards are decolorized with H2O2 at 100-160.degree..

ST sulfuric acid purifn yellow phosphorus; hydrogen peroxide sulfuric acid purifn; ferrous sulfate catalyst hydrogen peroxide; manganese sulfate catalyst hydrogen peroxide; nickel sulfate catalyst hydrogen peroxide; catalyst hydrogen peroxide sulfuric acid

IT Color centers
(removal of, in purified sulfuric acid from yellow phosphorus manuf., with hydrogen peroxide)

IT 7720-78-7, Ferrous sulfate 7785-87-7, Manganese sulfate (MnSO4) 7786-81-4, Nickel sulfate

RL: CAT (Catalyst use); USES (Uses)

(catalyst, in purifn. of contaminated sulfuric acid from yellow phosphorus manuf., with hydrogen peroxide)

L3 ANSWER 126 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1992:70523 CAPLUS

DN 116:70523

TI Effect of wave form of power source on electrolytic coloring of titanium

AU Onaka, Takashi; Ito, Seisiro; Ishida, Shinichi; Shiokawa, Jiro

CS Tech. Dev. Dep., Nippon Alum. Mfg. Co., Ltd., Osaka, 532, Japan

SO Shikizai Kyokaishi (1991), 64(11), 690-5

CODEN: SKYOAQ; ISSN: 0010-180X

DT Journal

LA Japanese

AB Effect of the wave form of the power source on anodic oxidn. (electrolytic coloring) of titanium using a bath contg. H_3PO_4 , H_2SO_4 , and H_2O_2 with or without $Al_2(SO_4)_3$ or $CoSO_4$ was studied. As the wave forms of the power source, a complete flat wave and incomplete rectified waves, such as single-phase half-wave, two-phase half-wave, three-phase half-wave and six-phase half-wave, were used. As a result, colored thick films obtained from using the incomplete rectified wave forms in a short time were more uniform and clearer than these obtained from using the complete flat wave form. This reason is as follows: the electrolysis using the incomplete rectified wave forms is equal to that with high c.d. over $10 A/dm^2$, because the voltage and the current of the incomplete rectified wave forms have high peak value. Also the spark discharges of the incomplete rectified wave forms occur smoothly because their currents are intermittent. The thickness, the color and the surface of the film obtained from using the incomplete rectified wave forms are decided by the peak voltage.

IT Oxidation, electrochemical

(of titanium in bath contg. sulfuric and **phosphoric acid** and **hydrogen peroxide** with and without aluminum **sulfate** and cobalt **sulfate**, voltage and current wave in relation to)

IT Anodization

(of titanium with coloring in bath contg. sulfuric acid and **phosphoric acid** and **hydrogen peroxide** with and without aluminum **sulfate** and cobalt **sulfate**)

IT 7664-93-9, Sulfuric acid, uses

RL: USES (Uses)

(anodization of titanium with coloring in bath contg. **phosphoric acid** and **hydrogen peroxide** and, with and without aluminum **sulfate** and cobalt **sulfate**)

IT 7664-38-2, **Phosphoric acid**, uses

RL: USES (Uses)

(anodization of titanium with coloring in bath contg. sulfuric acid and **hydrogen peroxide** and, with and without aluminum **sulfate** and cobalt **sulfate**)

IT 10043-01-3, Aluminum **sulfate** 10124-43-3, Cobalt **sulfate** ($CoSO_4$)

RL: USES (Uses)

(anodization of titanium with coloring in bath contg. sulfuric acid and **phosphoric acid** and **hydrogen peroxide** with or without)

IT 7722-84-1, **Hydrogen peroxide** (H_2O_2), uses

RL: USES (Uses)

(anodization of titanium with coloring in bath contg. sulfuric acid and **phosphoric acid** and, with and without aluminum **sulfate** and cobalt **sulfate**)

L3 ANSWER 127 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1991:188023 CAPLUS

DN 114:188023

TI Thickened acidic liquid composition with amino group-containing fluorescent whitener for use as a bleaching agent vehicle

IN Peterson, David

PA Clorox Co., USA

SO Eur. Pat. Appl., 16 pp.

CODEN: EPXXDW

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 403062	A2	19901219	EP 1990-304660	19900430
	EP 403062	A3	19910109		
	EP 403062	B1	19950118		
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE				
	US 5106523	A	19920421	US 1989-367375	19890616
	ES 2069006	T3	19950501	ES 1990-304660	19900430
	CA 2016430	AA	19901216	CA 1990-2016430	19900511
	CA 2016430	C	19941108		
	JP 03026796	A2	19910205	JP 1990-137331	19900529
PRAI	US 1989-367375		19890616		
AB	The title compn., having good storagability and useful as a laundering aid, is an aq. soln. having pH <7 and contg. a thickening system comprising an amide, an alkyl sulfate and/or alkylarene sulfonate surfactant, and a normally insol., but solubilized, fluorescent whitener contg. amino groups. A soln. having pH 3.8 and contg. poly(vinylpyrrolidone) (I; K 30) 0.10, Phorwite RKH 0.32, ammonium lauryl sulfate 1.50, H2O2 3.50, Dequest 2000 0.12%, and the balance being water, had viscosity 325 cP after 2 wk of storage at 40.degree., vs. space sepn. without I.				
AB	The title compn., having good storagability and useful as a laundering aid, is an aq. soln. having pH <7 and contg. a thickening system comprising an amide, an alkyl sulfate and/or alkylarene sulfonate surfactant, and a normally insol., but solubilized, fluorescent whitener contg. amino groups. A soln. having pH 3.8 and contg. poly(vinylpyrrolidone) (I; K 30) 0.10, Phorwite RKH 0.32, ammonium lauryl sulfate 1.50, H2O2 3.50, Dequest 2000 0.12%, and the balance being water, had viscosity 325 cP after 2 wk of storage at 40.degree., vs. space sepn. without I.				

L3 ANSWER 128 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1990:593983 CAPLUS
 DN 113:193983
 TI Aqueous alkaline beaching compositions containing hydrogen peroxide, thickeners, and stabilizers
 IN Clarke, David Ellis
 PA Unilever PLC, UK; Unilever N. V.
 SO Eur. Pat. Appl., 9 pp.
 CODEN: EPXXDW
 DT Patent
 LA English
 FAN. CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 376704	A1	19900704	EP 1989-313623	19891227
	EP 376704	B1	19950215		
	R: CH, DE, ES, FR, GB, IT, LI, NL, SE				
	ZA 8909842	A	19910828	ZA 1989-9842	19891221
	ZA 8909843	A	19910828	ZA 1989-9843	19891221
	CA 2006530	AA	19900628	CA 1989-2006530	19891222
	CA 2006530	C	19961015		
	CA 2006531	AA	19900628	CA 1989-2006531	19891222
	CA 2006531	C	19960903		
	AU 8947271	A1	19900705	AU 1989-47271	19891222
	AU 623961	B2	19920528		
	AU 8947272	A1	19900705	AU 1989-47272	19891222
	AU 624209	B2	19920604		
	NO 8905260	A	19900629	NO 1989-5260	19891227
	NO 173885	B	19931108		
	NO 173885	C	19940216		
	NO 8905261	A	19900629	NO 1989-5261	19891227
	NO 172354	B	19930329		
	NO 172354	C	19930707		

IN 170708	A	19920509	IN 1989-BO355	19891227
IN 171127	A	19920725	IN 1989-BO354	19891227
ES 2067558	T3	19950401	ES 1989-313623	19891227
ES 2067559	T3	19950401	ES 1989-313625	19891227
JP 02227498	A2	19900910	JP 1989-345128	19891228
JP 2562064	B2	19961211		
JP 02227499	A2	19900910	JP 1989-345129	19891228
JP 07035520	B4	19950419		
BR 8906843	A	19900918	BR 1989-6843	19891228
BR 8906844	A	19900925	BR 1989-6844	19891228
PRAI GB 1988-30296		19881228		
OS	MARPAT 113:193983			
AB	The title compns. are thickened by an electrolyte-surfactant combination and are stabilized against decompn. of H2O2 by a phosphonate or colloidal hydrous stannic oxide. The surfactant is an alkyl ether sulfate or a combination of a nonionic or amphoteric surfactant and a C8-20 alkane sulfonate or C8-20 alkyl sulfate. A compn. contg. H2O2 5, tallow alkyl dimethylamine oxide 1.0, Na alkane sulfonate 0.5, perfume 1.0, borax 1.6, and Na2SnO3.3H2O 0.1%, the balance being NaOH (to give pH 9.0) and water, retained 96% of the H2O2 during 100 days in a plastic bottle at 37.degree..			
AB	The title compns. are thickened by an electrolyte-surfactant combination and are stabilized against decompn. of H2O2 by a phosphonate or colloidal hydrous stannic oxide. The surfactant is an alkyl ether sulfate or a combination of a nonionic or amphoteric surfactant and a C8-20 alkane sulfonate or C8-20 alkyl sulfate. A compn. contg. H2O2 5, tallow alkyl dimethylamine oxide 1.0, Na alkane sulfonate 0.5, perfume 1.0, borax 1.6, and Na2SnO3.3H2O 0.1%, the balance being NaOH (to give pH 9.0) and water, retained 96% of the H2O2 during 100 days in a plastic bottle at 37.degree..			
IT	7488-55-3, Stannous sulfate 7646-78-8, Stannic chloride, uses and miscellaneous 7772-99-8, Stannous chloride, uses and miscellaneous 12058-66-1, Sodium stannate 13598-36-2D, Phosphonic acid , amine derivs. 18282-10-5, Stannic oxide RL: USES (Uses) (stabilizers, for thickened hydrogen peroxide bleach solns.)			
IT	151-21-3, uses and miscellaneous 9004-82-4 54116-08-4, Polyethylene glycol monotridecyl ether sulfate sodium salt RL: USES (Uses) (thickening by electrolyte and, of hydrogen peroxide bleach solns.)			
L3	ANSWER 129 OF 158 CAPLUS COPYRIGHT 2003 ACS			
AN	1990:182424 CAPLUS			
DN	112:182424			
TI	Stabilizing agents for hydrogen peroxide			
IN	McMillen, Mark W.			
PA	Henkel Corp., USA			
SO	Eur. Pat. Appl., 8 pp. CODEN: EPXXDW			
DT	Patent			
LA	English			
FAN.CNT	1			
	PATENT NO.	KIND	DATE	APPLICATION NO. DATE
	-----	----	-----	-----
PI	EP 351772	A2	19900124	EP 1989-113144 19890718
	EP 351772	A3	19900704	
	R: AT, BE, CH, DE, ES, FR, GB, GR, IT, LI, NL, SE			
	AU 8938229	A1	19900125	AU 1989-38229 19890719
	BR 8903577	A	19900313	BR 1989-3577 19890719
	JP 02074506	A2	19900314	JP 1989-187240 19890719
PRAI	US 1988-221063		19880719	
AB	A stabilized H2O2 compn. comprises aq. H2O2 and a stabilizing agent			

selected from the group consisting of nonionic, anionic, cationic, and amphoteric surfactants, phenylacetic acid and its salts, salicylic acid and its salts, diethylenetriaminepentaacetic acid and its salts, sulfosalicylic acid and its salts, polymeric Ph sulfonates and their salts, and methyl 3,4,5-trihydroxybenzoate. The surfactant/H₂O₂ ratios are 1/10,000 to 1/1.

- ST hydrogen peroxide stabilization surfactant; **sulfonate** stabilization **hydrogen peroxide**; phenylacetic acid stabilization hydrogen peroxide; salicylic acid stabilization hydrogen peroxide; methyl trihydroxybenzoate stabilization hydrogen peroxide; sulfosalicylic acid stabilization hydrogen peroxide; diethylenetriaminepentaacetic acid stabilization hydrogen peroxide
- IT **Sulfonates**
RL: USES (Uses)
(alkylarene, stabilizers, for **hydrogen peroxide**)
- IT 67-43-6, Diethylene triamine pentaacetic acid 69-72-7, Salicylic acid, uses and miscellaneous 97-05-2, Sulfosalicylic acid 97-78-9, Hamposyl L 99-24-1, Methyl-3,4,5-trihydroxy benzoate 104-15-4, P-Toluene **sulfonic** acid, uses and miscellaneous 126-92-1, Sodium 2-ethylhexyl **sulfate** 151-21-3, Sodium dodecyl **sulfate**, uses and miscellaneous 156-38-7 1191-50-0 1300-72-7, Sodium xylene **sulfonate** 5324-84-5, Sodium 1-octane **sulfonate** 9002-93-1, Triton X-102 9010-41-7, Triton X 200 9016-45-9, Triton N-100 27176-87-0, Dodecylbenzene **sulfonic** acid 28519-02-0, Sodium dodecyl diphenyl oxide disulfonate 30260-72-1 39464-70-5 52019-36-0, Monafax 1214 80803-49-2, Monateric 1000 98227-97-5, Mirataine CBS 106392-12-5, Pluronic R 127 109210-46-0, Hamplex DPS 126775-93-7, Avanel BSA 74 126775-98-2, Chemfac PC 099 126776-59-8, Mirawet ASC 126776-60-1, Mona AT 1200 126776-61-2, Monafax 057 126776-62-3, Monafax 060
RL: USES (Uses)
(stabilizer, for **hydrogen peroxide**)
- IT 107-43-7D, alkyl derivs. 7664-38-2D, **Phosphoric acid**, alkyl and aryl esters
RL: USES (Uses)
(stabilizers, for **hydrogen peroxide**)
- L3 ANSWER 130 OF 158 JAPIO COPYRIGHT 2003 JPO
AN 1990-145542 JAPIO
TI DIBASIC ACID ESTER
IN MIYOSHI NORIYUKI; UEHATA MASASHI
PA OKAMURA SEIYU KK
PI JP 02145542 A 19900605 Heisei
AI JP 1988-301417 (JP63301417 Showa) 19881128
PRAI JP 1988-301417 19881128
SO PATENT ABSTRACTS OF JAPAN (CD-ROM), Unexamined Applications, Vol. 1990
AB NEW MATERIAL: A dibasic acid ester expressed by formula I (R<SP>1</SP> and R<SP>2</SP> each are lower alkyl).
EXAMPLE: Useful as an electrolyte for driving electrolytic capacitors, lubricating oil, cutting oil, plasticizer, etc., a raw material for producing glycidyl esters, polyesters, polyamides, etc., and further an epoxy curing agent, epoxy reaction diluents, etc.
PREPARATION: For example, cyclohexanone is reacted with **hydrogen peroxide** in the presence of an acid catalyst (preferably sulfuric or **phosphoric acid**) in a lower alcohol expressed by formula II (e.g. absolute methanol or isopropanol) to provide a styrene derivative expressed by formula III, which is then reacted in the presence of a metal salt, preferably ferrous **sulfate** to afford the compound expressed by formula I. Furthermore, the ratio of the cyclohexanone to the **hydrogen peroxide** used is 80-130 pts.wt., especially about 100-110 pts.wt. latter based on 100 pts.wt. former.
COPYRIGHT: (C)1990,JPO&Japio
- AB . . . polyamides, etc., and further an epoxy curing agent, epoxy

reaction diluents, etc.

PREPARATION: For example, cyclohexanone is reacted with **hydrogen peroxide** in the presence of an acid catalyst (preferably sulfuric or **phosphoric acid**) in a lower alcohol expressed by formula II (e.g. absolute methanol or isopropanol) to provide a styrene derivative expressed by formula III, which is then reacted in the presence of a metal salt, preferably ferrous **sulfate** to afford the compound expressed by formula I. Furthermore, the ratio of the cyclohexanone to the **hydrogen peroxide** used is 80-130 pts.wt., especially about 100-110 pts.wt. latter based on 100 pts.wt. former.

COPYRIGHT: (C)1990,JPO&Japio

L3 ANSWER 131 OF 158 JAPIO COPYRIGHT 2003 JPO

AN 1990-066175 JAPIO

TI. ELECTROLESS NICKEL PLATING SOLUTION

IN KOJIMA KAYOKO

PA SHIMADZU CORP

PI JP 02066175 A 19900306 Heisei

AI JP 1988-217657 (JP63217657 Showa) 19880830

PRAI JP 1988-217657 19880830

SO PATENT ABSTRACTS OF JAPAN (CD-ROM), Unexamined Applications, Vol. 1990

AB PURPOSE: To prolong the service life of an electroless Ni plating soln. by adding hydrogen peroxide to an aq. soln. contg. Ni ions, a hypophosphite and an org. acid to inhibit the deposition of black Ni powder.

CONSTITUTION: An aq. soln. contg. 35% **hydrogen peroxide** is added to an aq. soln. contg. nickel chloride or nickel **sulfate**, a hypophosphite such as sodium hypophosphite and an org. acid or org. acid salt and the resulting soln. is adjusted to about pH 4-6 to prepare an electroless Ni plating soln. When Ni plating is carried out with the plating soln., **phosphorous acid** does not accumulate over a certain concn. because it is oxidized to orthophosphoric acid by the **hydrogen peroxide** soln. Accordingly, the deposition of black Ni powder is inhibited and the service life of the plating soln. is prolonged.

COPYRIGHT: (C)1990,JPO&Japio

AB . . . hypophosphite and an org. acid to inhibit the deposition of black Ni powder.

CONSTITUTION: An aq. soln. contg. 35% **hydrogen peroxide** is added to an aq. soln. contg. nickel chloride or nickel **sulfate**, a hypophosphite such as sodium hypophosphite and an org. acid or org. acid salt and the resulting soln. is adjusted. . . about pH 4-6 to prepare an electroless Ni plating soln. When Ni plating is carried out with the plating soln., **phosphorous acid** does not accumulate over a certain concn. because it is oxidized to orthophosphoric acid by the **hydrogen peroxide** soln. Accordingly, the deposition of black Ni powder is inhibited and the service life of the plating soln. is prolonged.

L3 ANSWER 132 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1989:479852 CAPLUS

DN 111:79852

TI Flame-resistant nylon/cotton fabric durable to multiple launderings

IN Hansen, John H.

PA Burlington Industries, Inc., USA

SO PCT Int. Appl., 22 pp.

CODEN: PIXXD2

DT Patent

LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 8900217	A1	19890112	WO 1988-US1649	19880524

W: AU, BR, DK, FI, JP, KR, NO

RW: AT, BE, CH, DE, FR, GB, IT, LU, NL, SE

US 4812144 A 19890314 US 1987-70559 19870707

AU 8821239 A1 19890130 AU 1988-21239 19880524

CA 1273158 A1 19900828 CA 1988-568261 19880531

PRAI US 1987-70559 19870707

WO 1988-US1649 19880524

OS MARPAT 111:79852

AB The title fabric is prepd. by treating a blend contg. cotton and .gtoreq.40% nylon fibers with compns. contg. prepolymers of urea and tetrakis(hydroxymethyl)phosphonium salts, drying the fabric to H2O content .apprx.20%, treating the fabric with NH3 to form a flame-retardant polymer network within the cotton, treating the fabric with oxidizing agents, then treating the fabric with cyclic phosphonate esters I (x = 0-1), tetrakis(hydroxymethyl)phosphonium salts, and urea, and finally treating the fabric with oxidizing agents to give a fabric with P content .gtoreq.2.47%. Thus, a 52:48 nylon-cotton blend was printed, padded with an aq. compn. contg. 34.0% tetrakis(hydroxymethyl)phosphonium chloride-urea copolymer (Retardol AC) to pickup 60%, dried to H2O content 14-15%, treated with NH3(g) for 3 min at 3.12 m3/h, oxidized with aq. 15% H2O2 for 5 min, washed, centrifuged, and dried. The fabric was then padded with an aq. compn. contg. urea 13.0, tetrakis(hydroxymethyl)phosphonium sulfate 50.0, and Antiblaze 19-T (cyclic phosphonate ester) to pickup 69%, dried, cured 70 s at 182.degree., oxidized with aq. 15% H2O2, washed, and dried to give a fabric with P content 4.76 (initially), 3.98 (after 25 launderings), and 3.71% (after 50 laundering) and char length 8.31 (initially), 10.92 (after 25 laundering), and 11.63 cm (after 50 launderings).

AB The title fabric is prepd. by treating a blend contg. cotton and .gtoreq.40% nylon fibers with compns. contg. prepolymers of urea and tetrakis(hydroxymethyl)phosphonium salts, drying the fabric to H2O content .apprx.20%, treating the fabric with NH3 to form a flame-retardant polymer network within the cotton, treating the fabric with oxidizing agents, then treating the fabric with cyclic phosphonate esters I (x = 0-1), tetrakis(hydroxymethyl)phosphonium salts, and urea, and finally treating the fabric with oxidizing agents to give a fabric with P content .gtoreq.2.47%. Thus, a 52:48 nylon-cotton blend was printed, padded with an aq. compn. contg. 34.0% tetrakis(hydroxymethyl)phosphonium chloride-urea copolymer (Retardol AC) to pickup 60%, dried to H2O content 14-15%, treated with NH3(g) for 3 min at 3.12 m3/h, oxidized with aq. 15% H2O2 for 5 min, washed, centrifuged, and dried. The fabric was then padded with an aq. compn. contg. urea 13.0, tetrakis(hydroxymethyl)phosphonium sulfate 50.0, and Antiblaze 19-T (cyclic phosphonate ester) to pickup 69%, dried, cured 70 s at 182.degree., oxidized with aq. 15% H2O2, washed, and dried to give a fabric with P content 4.76 (initially), 3.98 (after 25 launderings), and 3.71% (after 50 laundering) and char length 8.31 (initially), 10.92 (after 25 laundering), and 11.63 cm (after 50 launderings).

L3 ANSWER 133 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1990:98357 CAPLUS

DN 112:98357

TI Simultaneous preparation of aldehydes and epoxides by oxidation of olefinic compounds

IN Inagaki, Hiroyuki; Nakamura, Teiji; Nishikawa, Eiichiro; Imai, Chihiro; Misonoo, Makoto

PA Tonen Sekiyu Kagaku K. K., Japan

SO Jpn. Kokai Tokkyo Koho, 11 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 01190647	A2	19890731	JP 1988-12621	19880125
	JP 08019027	B4	19960228		
PRAI	JP 1988-12621		19880125		
OS	MARPAT 112:98357				
AB	Aldehydes and epoxides are prepd. by oxidn. of compds. contg. olefinic double bonds with H2O2 in the presence of heteropoly acids and/or their salts and MgSO4 (I) to prevent formation of diols. A Bu3PO4 soln. of H2O2 was added dropwise to a mixt. of cyclopentene, H3PMo6W6O40, H3PO4, and I at 45.degree. over 20 min and the reaction mixt. was further stirred at 45.degree. for 6 h to give 18.3% CH2(CH2CHO)2 and 16.7% cyclopentene oxide, vs. 21.7 and 1.3%, resp., for a control without addn. of I.				
IT	7664-38-2, Phosphoric acid , uses and miscellaneous RL: USES (Uses) (catalyst system contg. heteropoly acids and, for oxidn. of olefinic compds. with hydrogen peroxide)				
IT	7487-88-9, Magnesium sulfate , uses and miscellaneous RL: USES (Uses) (dehydrating agent, for oxidn. of olefinic compds. with hydrogen peroxide to aldehydes and epoxides)				
IT	100-42-5, Styrene, reactions 110-83-8, Cyclohexene, reactions 142-29-0, Cyclopentene RL: RCT (Reactant); RACT (Reactant or reagent) (oxidn. of, by hydrogen peroxide , aldehyde and epoxide from, magnesium sulfate dehydrating agent for)				

L3 ANSWER 134 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 1990:38081 CAPLUS
DN 112:38081
TI Improved bleaching of cellulose fibers by alkaline hydrogen peroxide
IN Balland, Jean
PA Manufacture de Produits Chimiques PROTEX S. A., Fr.
SO Fr. Demande, 12 pp.
CODEN: FRXXBL
DT Patent
LA French
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	FR 2626020	A1	19890721	FR 1988-589	19880120
	FR 2626020	B1	19900504		
PRAI	FR 1988-589		19880120		
AB	In the title process, alkali metal, alk. earth, or NH4 persulfates and reducing agents are added to the bleaching bath. An ecru cotton fabric was bleached in an aq. soln. of 15 mL/L 35% H2O2, 10 mL/L 30% NaOH, Mg EDTA complex 10, softening agent 2, Na2S2O8 5, and 1:2:2 N2H4 Na (1-hydroxyethylidene)diphosphonate (I) 5 g/L at 103.degree. for 20 min to gave a fabric with whiteness 86%; vs. 81 without Na2S2O8, 80 without I, and 78 with neither.				
AB	In the title process, alkali metal, alk. earth, or NH4 persulfates and reducing agents are added to the bleaching bath. An ecru cotton fabric was bleached in an aq. soln. of 15 mL/L 35% H2O2, 10 mL/L 30% NaOH, Mg EDTA complex 10, softening agent 2, Na2S2O8 5, and 1:2:2 N2H4 Na (1-hydroxyethylidene)diphosphonate (I) 5 g/L at 103.degree. for 20 min to gave a fabric with whiteness 86%; vs. 81 without Na2S2O8, 80 without I, and 78 with neither.				
IT	Bleaching agents (hydrogen peroxide -peroxy sulfates -reducing agents, for cellulosic fibers)				
IT	7722-84-1 RL: USES (Uses) (bleaching agents, hydrogen peroxide -peroxy sulfates -reducing agents, for cellulosic fibers)				

IT 302-01-2, Hydrazine, uses and miscellaneous 7772-98-7, Sodium thiosulfate 7775-27-1, Disodium peroxydisulfate 7803-49-8, Hydroxylamine, uses and miscellaneous 29712-28-5, (Aminomethylene) diphosphonic acid 62607-44-7D, Sulfenic acid, hydroxyalkane derivs., salts 124766-02-5
 RL: USES (Uses)
 (hydrogen peroxide bleaching of cellulosic fibers in presence of)

L3 ANSWER 135 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1988:572094 CAPLUS
 DN 109:172094
 TI Flame-resistant cotton blend fabrics
 IN Hauser, Peter J.; Triplett, Benny L.; Sujarit, Chumpon; Hansen, John H.; Johnson, James R.
 PA Burlington Industries, Inc., USA
 SO PCT Int. Appl., 24 pp.
 CODEN: PIXXD2
 DT Patent
 LA English
 FAN.CNT 2.

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
	-----	---	-----	-----	-----
PI	WO 8802283	A1	19880407	WO 1987-US2432	19870925
	W: AU, DK, JP, NO				
	RW: AT, BE, CH, DE, FR, GB, IT, LU, NL, SE				
	US 4750911	A	19880614	US 1986-911720	19860926
	US 4732789	A	19880322	US 1986-923965	19861028
	AU 8780716	A1	19880421	AU 1987-80716	19870925
	AU 603391	B2	19901115		
	JP 02500454	T2	19900215	JP 1987-506118	19870925
	AT 104000	E	19940415	AT 1987-906715	19870925
	NO 8802310	A	19880725	NO 1988-2310	19880525
	NO 173790	B	19931025		
	NO 173790	C	19940202		
	DK 8802892	A	19880526	DK 1988-2892	19880526
PRAI	US 1986-911720		19860926		
	US 1986-923965		19861028		
	EP 1987-906715		19870925		
	WO 1987-US2432		19870925		

AB A 2-step process for imparting washfast flame resistance to polyester/cotton or nylon/cotton blends comprises steps of (A) applying to the textile a flame retarding amt. of a precondensate of urea and a tetrakis(hydroxymethyl)phosphonium (I) salt flame retardant which fixes to the cotton, exposing to a source of NH₃ to form an ammoniated prepolymer, then oxidizing to form a polymer network, and (B) applying a flame retarding amt. of a cyclic phosphonate ester or hexabromocyclododecane flame retardant, then curing to fix the flame retardant in the synthetic fibers. A 50/50 nylon/cotton textile was padded with a urea-I sulfate precondensate to 61% wet pickup, the textile equilibrated in a plastic bag for 1-2 h, exposed to NH₃, oxidized with H₂O₂ at pH 9-9.5, rinsed, dried, padded with a 15% soln. of a cyclic phosphonate ester, and heated to give a textile with char length (FTM 5903) 5 cm as-prepd. and 11.4 cm after 25 wash cycles.

AB A 2-step process for imparting washfast flame resistance to polyester/cotton or nylon/cotton blends comprises steps of (A) applying to the textile a flame retarding amt. of a precondensate of urea and a tetrakis(hydroxymethyl)phosphonium (I) salt flame retardant which fixes to the cotton, exposing to a source of NH₃ to form an ammoniated prepolymer, then oxidizing to form a polymer network, and (B) applying a flame retarding amt. of a cyclic phosphonate ester or hexabromocyclododecane flame retardant, then curing to fix the flame retardant in the synthetic fibers. A 50/50 nylon/cotton textile was padded with a urea-I sulfate precondensate to 61% wet pickup, the textile equilibrated

in a plastic bag for 1-2 h, exposed to NH₃, oxidized with H₂O₂ at pH 9-9.5, rinsed, dried, padded with a 15% soln. of a cyclic phosphonate ester, and heated to give a textile with char length (FTM 5903) 5 cm as-prepd. and 11.4 cm after 25 wash cycles.

L3 ANSWER 136 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1989:25349 CAPLUS
 DN 110:25349
 TI Flame-resistant polyester/cotton fabric and process for its production
 IN Johnson, James R.; Sujarit, Chumpon
 PA Burlington Industries, Inc., USA
 SO U.S., 8 pp. Cont.-in-part of U.S. Ser. No. 870,892:
 CODEN: USXXAM
 DT Patent
 LA English
 FAN.CNT 3

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 4748705	A	19880607	US 1987-23372	19870309
	IL 82367	A1	19900712	IL 1987-82367	19870428
	AU 8772465	A1	19871210	AU 1987-72465	19870504
	AU 597358	B2	19900531		
	EP 248553	A2	19871209	EP 1987-304250	19870513
	EP 248553	A3	19901114		
	EP 248553	B1	19930127		
	R: AT, BE, CH, DE, ES, FR, GB, GR, IT, LI, LU, NL, SE				
	CA 1273757	A1	19900911	CA 1987-537057	19870513
	AT 85094	E	19930215	AT 1987-304250	19870513
	ES 2038177	T3	19930716	ES 1987-304250	19870513
	FI 8702131	A	19871206	FI 1987-2131	19870514
	FI 92606	B	19940831		
	FI 92606	C	19941212		
	JP 62299574	A2	19871226	JP 1987-120341	19870519
	NO 8702198	A	19871207	NO 1987-2198	19870526
	NO 172655	B	19930510		
	NO 172655	C	19930818		
	CN 87104025	A	19871216	CN 1987-104025	19870604
	CN 1022193	B	19930922		
	BR 8702854	A	19880301	BR 1987-2854	19870604
	US 4902300	A	19900220	US 1988-195858	19880519
	US 4842609	A	19890627	US 1988-216185	19880707
	US 5238464	A	19930824	US 1992-855499	19920323
PRAI	US 1986-870892		19860605		
	US 1987-23372		19870309		
	EP 1987-304250		19870513		
	US 1987-52937		19870522		
	US 1988-195858		19880519		
	US 1989-446071		19891205		

OS MARPAT 110:25349

AB A process for imparting flame resistance to a polyester/cotton fabric contg. .gtoreq.40 wt.% polyester fiber comprises applying to the blend fiber a flame retarding amt. of I (x = 0, 1) which fixes onto the polyester fiber, applying a flame retarding amt. of a prepolymer condensate of urea and tetrakis (hydroxymethyl) phosphonium (II) salt flame retardant which fixes to cotton, exposing to NH₃ to form an ammoniated prepolymer to form a flame-retardant polymer network in the cotton fiber, applying a flame retarding amt. of II sulfate and urea to fabric, heating, and oxidizing with H₂O₂ to further improve flame resistance. A 40/60 polyester cotton heavy wt. twill was padded with a cyclic phosphonate, cured, treated with II-urea precondensate, treated with II sulfate urea, heated to reduce moisture content, exposed to NH₃, and oxidized with H₂O₂. The resulting fabric retained .apprx.80% of initial P after 50 washes and when evaluated by FR Federal Test Method 5903 gave a value of .apprx. 4.

AB A process for imparting flame resistance to a polyester/cotton fabric contg. .gtoreq.40 wt.% polyester fiber comprises applying to the blend fiber a flame retarding amt. of I (x = 0, 1) which fixes onto the polyester fiber, applying a flame retarding amt. of a prepolymer condensate of urea and tetrakis (hydroxymethyl) phosphonium (II) salt flame retardant which fixes to cotton, exposing to NH3 to form an ammoniated prepolymer to form a flame-retardant polymer network in the cotton fiber, applying a flame retarding amt. of II sulfate and urea to fabric, heating, and oxidizing with H2O2 to further improve flame resistance. A 40/60 polyester cotton heavy wt. twill was padded with a cyclic phosphonate, cured, treated with II-urea precondensate, treated with II sulfate urea, heated to reduce moisture content, exposed to NH3, and oxidized with H2O2. The resulting fabric retained .apprx.80% of initial P after 50 washes and when evaluated by FR Federal Test Method 5903 gave a value of .apprx. 4.

L3 ANSWER 137 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1988:226913 CAPLUS

DN 108:226913

TI Disinfectants containing hydrogen peroxide and silver salts

IN Goemoeri, Janos

PA Sanosil A.-G., Switz.

SO Ger. Offen., 4 pp.

CODEN: GWXXBX

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
	-----	----	-----	-----	-----
PI	DE 3620609	A1	19871105	DE 1986-3620609	19860620
	DE 3620609	C2	19930729		
	CH 673225	A	19900228	CH 1986-1629	19860422
	DE 3645266	C2	19950504	DE 1986-3645266	19860620
	US 4915955	A	19900410	US 1986-904055	19860905
	HU 45383	A2	19880728	HU 1986-4122	19860926
	HU 200944	B	19900928		
	AU 8767854	A1	19871029	AU 1987-67854	19870121
	AU 590379	B2	19891102		
	GB 2189394	A1	19871028	GB 1987-2506	19870204
	GB 2189394	B2	19910109		
	CA 1288334	A1	19910903	CA 1987-533260	19870330
	NL 8700749	A	19871116	NL 1987-749	19870331
	NL 192918	B	19980105		
	NL 192918	C	19980507		
	IN 168298	A	19910309	IN 1987-DE275	19870331
	SE 8701500	A	19871023	SE 1987-1500	19870410
	SE 468968	B	19930426		
	SE 468968	C	19930819		
	JP 62255401	A2	19871107	JP 1987-95501	19870420
	JP 08002769	B4	19960117		
	FR 2597347	A1	19871023	FR 1987-5589	19870421
	FR 2597347	B1	19940211		
	BE 1000083	A7	19880202	BE 1987-427	19870421
	BR 8701885	A	19880202	BR 1987-1885	19870421
	ES 2004931	A6	19890216	ES 1987-1150	19870421
	CZ 282207	B6	19970611	CZ 1987-2804	19870421
	SK 279184	B6	19980708	SK 1987-2804	19870421
	CZ 284378	B6	19981111	CZ 1996-2925	19870421
	DK 8702045	A	19871023	DK 1987-2045	19870422
	DK 167948	B1	19940110		
	AT 389802	B	19900212	AT 1987-1764	19870713
	AT 8701764	A	19890715		
	IN 173798	A	19940716	IN 1989-DE888	19891005
	JP 08225418	A2	19960903	JP 1995-173787	19950710

PRAI CH 1986-1629 A 19860422
 DE 1986-3620609 A3 19860620
 IN 1987-DE275 A1 19870331
 CS 1987-2804 A 19870421
 AB A process for the prepn. of a H2O2-miscible conc. comprises mixing an acidic aq. soln. wherein pH is adjusted to .ltoreq.1.6 with an inorg. acid at 50-60.degree. with a 100 .+- 5 g/L Ag salt or Ag complex. The acidic soln. is prepd. at 25-30.degree. and contains acid in an amt. at least equimolar with Ag, and the soln. contains an org. stabilizer and optionally gelatin, and it is finally homogenized. The conc. has unlimited storage stability and is mixed with H2O2 (I) to give a disinfectant. Sufficient 75% H3PO4 was added to 1 L H2O to give a soln. with pH .ltoreq.1.6 at 20.degree., the soln. was heated to 55.degree. and mixed with 140 g AgNO3, the soln. was cooled to 25-30.degree., and H3PO4 was added to a total amt. of 100 g. The mixt. was cooled to 20-25.degree., mixed with 50 g tartaric acid, 20 g gelatin, and homogenized. The conc. (5 L) was mixed with 50% by vol. I in a 1:99 ratio; the disinfectant was useful for the long-term disinfection of whirl-pools and circulation systems of water purifiers.
 IT 7440-22-4D, Silver, salts 7761-88-8, Silver nitrate, biological studies 7783-90-6, Silver chloride, biological studies 10294-26-5, Silver sulfate 66942-69-6
 RL: BIOL (Biological study)
 (disinfectants contg. hydrogen peroxide and)
 IT 65-85-0D, Benzoic acid, alkali salts 77-92-9, Citric acid, biological studies 87-69-4, Tartaric acid, biological studies 7664-38-2, Phosphoric acid, biological studies 7664-93-9, Sulfuric acid, biological studies 7697-37-2, Nitric acid, biological studies
 RL: BIOL (Biological study)
 (disinfectants contg. hydrogen peroxide and silver salts and)
 L3 ANSWER 138 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1987:428199 CAPLUS
 DN 107:28199
 TI Oxidative hair dyes
 IN Hollenberg, Detlef; Neuhaus, Winifried; Schrader, Dieter
 PA Henkel K.-G.a.A., Fed. Rep. Ger.
 SO Ger. Offen., 5 pp.
 CODEN: GWXXBX
 DT Patent
 LA German
 FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 3534471	A1	19870402	DE 1985-3534471	19850927
	EP 216334	A2	19870401	EP 1986-112978	19860919
	EP 216334	A3	19870527		
	EP 216334	B1	19901031		
	R: AT, BE, CH, DE, FR, GB, IT, LI, NL, SE				
	AT 57830	E	19901115	AT 1986-112978	19860919
	JP 62072608	A2	19870403	JP 1986-229463	19860927
PRAI	DE 1985-3534471		19850927		
	EP 1986-112978		19860919		
AB	An oxidative hair dye is made of an oxidative dye cream and an oxidizing prepn. The 2 are kept sep. in containers sepd. by a wall. Prior to use, the wall is destroyed mech. and the 2 components are mixed. Both components are oil-in-water emulsions. The oxidative cream contains 0.1-5% dye precursors and the oxidizing prepn. 1.5-15% H2O2. There is a 1-3:1 oxidative cream/oxidizing prepn. ratio. Thus, an oxidative dye cream contained C12-10 fatty alc. 10, C16-18 polyethoxylated fatty alc. polyglycol ether 0.75, 28% C12-14 diethoxylated fatty alc. polyglycol ether sulfate Na salt 20, 30% cocoacylamidopropyldimethylammonium				

glycinate 12.5, cationic cellulose deriv. 1, EDTA 0.2, Na2SO3 0.5, ascorbic acid 0.5, (NH4)2SO4 1.5, perfume 0.2, p-toluylenediamine 0.13, p-aminophenol 0.05, resorcinol 0.1, 2,4-diaminophenetol 0.002, .alpha.-naphthol 0.02 and water to 100% by wt. The pH was adjusted to 10 (NH3). The oxidizing prepn. contained cetyl alc. 1.5, C12-14 diethoxylated fatty alc. polyglycol ether 3, 25% C12-14 polyethoxylated fatty alc. polyglycol ether **sulfate** Na salt 8, NH4NO3 0.012, Na4P2O7 0.03, dipicolinic acid 0.1, 1-hydroxyethane-1,1-**diphosphonic** acid 1.5, 50% H2O2 12, and water to 100% by wt. The pH was adjusted to 4 (NH3).

AB An oxidative hair dye is made of an oxidative dye cream and an oxidizing prepn. The 2 are kept sep. in containers sepd. by a wall. Prior to use, the wall is destroyed mech. and the 2 components are mixed. Both components are oil-in-water emulsions. The oxidative cream contains 0.1-5% dye precursors and the oxidizing prepn. 1.5-15% H2O2. There is a 1-3:1 oxidative cream/oxidizing prepn. ratio. Thus, an oxidative dye cream contained C12-10 fatty alc. 10, C16-18 polyethoxylated fatty alc. polyglycol ether 0.75, 28% C12-14 diethoxylated fatty alc. polyglycol ether **sulfate** Na salt 20; 30% cocoacylamidopropyldimethylammonium glycinate 12.5, cationic cellulose deriv. 1, EDTA 0.2, Na2SO3 0.5, ascorbic acid 0.5, (NH4)2SO4 1.5, perfume 0.2, p-toluylenediamine 0.13, p-aminophenol 0.05, resorcinol 0.1, 2,4-diaminophenetol 0.002, .alpha.-naphthol 0.02 and water to 100% by wt. The pH was adjusted to 10 (NH3). The oxidizing prepn. contained cetyl alc. 1.5, C12-14 diethoxylated fatty alc. polyglycol ether 3, 25% C12-14 polyethoxylated fatty alc. polyglycol ether **sulfate** Na salt 8, NH4NO3 0.012, Na4P2O7 0.03, dipicolinic acid 0.1, 1-hydroxyethane-1,1-**diphosphonic** acid 1.5, 50% H2O2 12, and water to 100% by wt. The pH was adjusted to 4 (NH3).

L3 ANSWER 139 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1986:562188 CAPLUS

DN 105:162188

TI Forming color images

IN Iwano, Haruhiko

PA Fuji Photo Film Co., Ltd., Japan

SO Jpn. Kokai Tokkyo Koho, 15 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 61088259	A2	19860506	JP 1984-209562	19841005
PRAI	JP 1984-209562		19841005		

AB In a method for color image formation involving intensified development by use of a developer-intensifying soln. contg. H2O2 (or a H2O2-releasing compd.) and a color developer and substantially not contg. Br and I anions, Ag halide color photog. materials which have been imagewise-exposed are treated, prior to the intensifying process, with a pretreatment soln. of a compn. contg. a color developer. The improved method for color development provides good photog. image properties with short processing time. Thus, a Ag(Br,Cl) emulsion (av. grain size, 0.28 .mu.m) was mixed with a dispersion of a magenta dye coupler, 1-(2,4,6-trichlorophenyl)-3-(2-chloro-5-tetradecanamido)anilino-2-pyrazolin-5-one, and was coated on a support to give a photog. film. After being wedge-exposed, the film was pretreated with a soln. comprising K2SO3, K3PO4, K2HPO4, 1-hydroxyethane-1,1-**diphosphonic** acid, 5-nitrobenzotriazole, and 4-amino-3-methyl-N-ethyl-N-.beta.-(methanesulfonamido)ethylaniline **sulfate**, developed with an intensifier soln. contg. H2O2 in addn. to the above components, and fixed to give a magenta dye image, which showed high sensitivity and high color d. as compared with a control film which was developed by a process omitting the above pretreatment.

AB In a method for color image formation involving intensified development by use of a developer-intensifying soln. contg. H2O2 (or a H2O2-releasing compd.) and a color developer and substantially not contg. Br and I anions, Ag halide color photog. materials which have been imagewise-exposed are treated, prior to the intensifying process, with a pretreatment soln. of a compn. contg. a color developer. The improved method for color development provides good photog. image properties with short processing time. Thus, a Ag(Br,Cl) emulsion (av. grain size, 0.28 .mu.m) was mixed with a dispersion of a magenta dye coupler, 1-(2,4,6-trichlorophenyl)-3-(2-chloro-5-tetradecanamido)anilino-2-pyrazolin-5-one, and was coated on a support to give a photog. film. After being wedge-exposed, the film was pretreated with a soln. comprising K2SO3, K3PO4, K2HPO4, 1-hydroxyethane-1,1-diphosphonic acid, 5-nitrobenzotriazole, and 4-amino-3-methyl-N-ethyl-N-.beta.-(methanesulfonamido)ethylaniline sulfate, developed with an intensifier soln. contg. H2O2 in addn. to the above components, and fixed to give a magenta dye image, which showed high sensitivity and high color d. as compared with a control film which was developed by a process omitting the above pretreatment.

L3 ANSWER 140 OF 158 CAPLUS COPYRIGHT 2003 ACS
 AN 1984:42999 CAPLUS
 DN 100:42999
 TI Color photographic imaging method
 IN Hirai, Hiroyuki; Nakamura, Koichi; Iwano, Haruhiko
 PA Fuji Photo Film Co., Ltd. , Japan
 SO Ger. Offen., 71 pp.
 CODEN: GWXXBX
 DT Patent
 LA German
 FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 3302741	A1	19830804	DE 1983-3302741	19830127
	DE 3302741	C2	19920702		
	JP 58127926	A2	19830730	JP 1982-11176	19820127
	JP 03008535	B4	19910206		
	JP 58137837	A2	19830816	JP 1982-20046	19820210
	JP 63020330	B4	19880427		
	JP 58140741	A2	19830820	JP 1982-23140	19820216
	JP 63020331	B4	19880427		
	GB 2117914	A1	19831019	GB 1983-2027	19830125
	GB 2117914	B2	19850710		
	US 4469780	A	19840904	US 1983-461430	19830127
PRAI	JP 1982-11176		19820127		
	JP 1982-20046		19820210		
	JP 1982-23140		19820216		

AB The formation of color stain in photog. materials contg. extremely small amts. of Ag and which are processed by using a combined development and intensification bath can be reduced by posttreatment of the material in a dil. buffer soln. having a pH of 2 to 7. This process allows the elimination of the usual washing step. Thus, on a TiO2-coated polyethylene support there was coated a blue-sensitive gelatin-Ag(Br,Cl) (80 mol% Br-) emulsion layer (8 mg Ag/m2) contg. a yellow coupler, a gelatin layer, a green-sensitive gelatin-Ag (Br,Cl) (30 mol% Br-) emulsion layer (5 mg Ag/m2) contg. a magenta coupler, a gelatin layer contg. 2-(2-hydroxy-3-sec-butyl)-5-tert-butylphenyl)benzotriazole and dioctylhydroquinone, a red-sensitive gelatin-Ag(Br,Cl) (30 mol% Br-) emulsion layer (5 mg Ag/m2) contg. a cyan coupler, and a gelatin top layer. The resultant film was exposed and developed in a bath contg. K2SO3, K3PO4, K2HPO4, 1-hydroxyethane-1,1-diphosphonic acid, benzotriazole, 4-amino-3-methyl-N-ethyl-N-.beta.-(methanesulfonamido)ethylaniline 3/2 sulfate, 30% H2O2, and H2O at pH 11.0 for 1 min at 35.degree.. The developed film was then

washed in a buffer soln. (pH 3.2) contg. 0.1 N HCl 350 mL, Na citrate dihydrate 4.4 g, and H2O 1 L to give images with a blue min d. of 0.13, green min. d. of 0.13, and red min d. of 0.12 vs. 0.40, 0.30, and 0.23, resp., for a film just washed with H2O after development.

AB The formation of color stain in photog. materials contg. extremely small amts. of Ag and which are processed by using a combined development and intensification bath can be reduced by posttreatment of the material in a dil. buffer soln. having a pH of 2 to 7. This process allows the elimination of the usual washing step. Thus, on a TiO₂-coated polyethylene support there was coated a blue-sensitive gelatin-Ag(Br,Cl) (80 mol% Br-) emulsion layer (8 mg Ag/m²) contg. a yellow coupler, a gelatin layer, a green-sensitive gelatin-Ag (Br,Cl) (30 mol% Br-) emulsion layer (5 mg Ag/m²) contg. a magenta coupler, a gelatin layer contg. 2-(2-hydroxy-3-sec-butyl)-5-tert-butylphenylbenzotriazole and dioctylhydroquinone, a red-sensitive gelatin-Ag(Br,Cl) (30 mol% Br-) emulsion layer (5 mg Ag/m²) contg. a cyan coupler, and a gelatin top layer. The resultant film was exposed and developed in a bath contg. K₂SO₃, K₃PO₄, K₂HPO₄, 1-hydroxyethane-1,1-diphosphonic acid, benzotriazole, 4-amino-3-methyl-N-ethyl-N-.beta.-(methanesulfonamido)ethylaniline 3/2 sulfate, 30% H₂O₂, and H₂O at pH 11.0 for 1 min at 35.degree.. The developed film was then washed in a buffer soln. (pH 3.2) contg. 0.1 N HCl 350 mL, Na citrate dihydrate 4.4 g, and H₂O 1 L to give images with a blue min d. of 0.13, green min. d. of 0.13, and red min d. of 0.12 vs. 0.40, 0.30, and 0.23, resp., for a film just washed with H₂O after development.

L3 ANSWER 141 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1983:506726 CAPLUS

DN 99:106726

TI Bleaching of keratin-containing natural fibers

IN Bereck, Attila

PA Rheinisch-Westfaelischen Technischen Hochschule Aachen, Deutsches Wollforschungsinstitut, Fed. Rep. Ger.

SO Ger. Offen., 19 pp.

CODEN: GWXXBX

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 3149978	A1	19830630	DE 1981-3149978	19811217
	DE 3149978	C2	19850905		
PRAI	DE 1981-3149978		19811217		

AB Keratin-contg. natural fibers are bleached by a process which comprises mordanting with a Fe²⁺ salt soln. in the presence of a P compd. reducing agent and optionally washing .gtoreq.5 min at .gtoreq.70.degree. (or longer at lower temps. to .gtoreq.5 h at 10.degree.), and bleaching with a peroxide soln. Thus, 4 g each of black karakul wool and white merino wool were treated with a compn. contg. FeSO₄ 10, 50% H₃PO₂ 2, and a wetting agent 0.1 g/L at a 1:40 bath ratio at 80.degree. for 1 h (pH 3-3.5), the bath drained, washed with 2 g/L H₃PO₂ soln. at 80.degree. for 10 min, drained, and washed with 80.degree. H₂O for 10 min. The wool was treated at 55.degree. for 2 h at pH 8.5 in a bath contg. 10 mL/L H₂O₂ and 10 g/L tetra-Na pyrophosphate decahydrate to give whitened wool. The brightness of the karakul wool was 33 (Y coordinate of CIE color system) compared with 6.5 for the untreated material. This process is esp. useful for removal of pigmented single fiber in tech. white wool and gives wool with reduced alkali soly.

IT 7720-78-7

RL: USES (Uses)

(mordants, treatment of wool with phosphorus acid and, for improved bleaching properties. with hydrogen peroxide)

IT 6303-21-5

RL: USES (Uses)

(wool pretreatment by ferrous **sulfate** and, for improved bleachability with **hydrogen peroxide**)

L3 ANSWER 142 OF 158 JAPIO COPYRIGHT 2003 JPO

AN 1982-062251 JAPIO

TI PREPARATION OF CYSTEAMINE AND/OR CYSTAMINE

IN YAMAGUCHI HACHIRO

PA SOGO YATSUKOU KK

PI JP 57062251 A 19820415 Showa

AI JP 1980-135841 (JP55135841 Showa) 19801001

PRAI JP 1980-135841 19801001

SO PATENT ABSTRACTS OF JAPAN (CD-ROM), Unexamined Applications, Vol. 1982

AB PURPOSE: To prepare the titled compound useful as a perventive for radiation diseases or its intermediate, in high yield, without producing by-products, by reacting (2-aminoethyl) hydrogen **sulfate** with sodium thiosulfate, and hydrolyzing the product with a strong acid or **hydrogen peroxide**.

CONSTITUTION: 1mol of (2-aminoethyl) hydrogen sulfate which is a novel raw material having excellent economic value, reactivity, and safety, is made to react with 1∼5mol of sodium thiosulfate at 4∼7pH and about 100°C for several hours. The reaction product is hydrolyzed with an aqueous solution of a strong acid such as hydrochloric acid, sulfuric acid, **phosphoric acid**, etc. or **hydrogen**

peroxide solution in nitrogen gas stream to obtain the objective compound while suppressing the production of di(2-aminoethyl)sulfide.

COPYRIGHT: (C)1982,JPO&Japio

AB . . . useful as a perventive for radiation diseases or its intermediate, in high yield, without producing by-products, by reacting (2-aminoethyl) hydrogen **sulfate** with sodium thiosulfate, and hydrolyzing the product with a strong acid or **hydrogen peroxide**.

CONSTITUTION: 1mol of (2-aminoethyl) hydrogen sulfate which is a novel raw material having excellent economic value, reactivity, and safety, . . . hours. The reaction product is hydrolyzed with an aqueous solution of a strong acid such as hydrochloric acid, sulfuric acid, **phosphoric acid**, etc. or **hydrogen peroxide** solution in nitrogen gas stream to obtain the objective compound while suppressing the production of di(2-aminoethyl)sulfide.

COPYRIGHT: (C)1982,JPO&Japio

L3 ANSWER 143 OF 158 JAPIO COPYRIGHT 2003 JPO

AN 1980-098750 JAPIO

TI PHOTOGRAPHIC BLEACHING COMPOSITION

IN IDOTA YOSHIO; YAMADA MINORU

PA FUJI PHOTO FILM CO LTD

PI JP 55098750 A 19800728 Showa

AI JP 1979-6645 (JP54006645 Showa) 19790123

PRAI JP 1979-6645 19790123

SO PATENT ABSTRACTS OF JAPAN (CD-ROM), Unexamined Applications, Vol. 1980

AB PURPOSE: To always maintain stable bleaching capacity in treatment of a silver halide color photographic material by adding optionally substituted aromatic **sulfonic acid** or **sulfonate** to **hydrogen peroxide** and metal complex salt of organic acid.

CONSTITUTION: An adequate amt. of aromatic sulfonic acid or **sulfonate** optionally substituted by at least one of carboxyl group, nitro group, etc. is added to a bleaching bath containing 0.01∼10mol/l of **hydrogen peroxide** and a **hydrogen peroxide**-releasing substance such as perboric acid, perborate, percarbonic acid or percarbonate and metal complex salt of organic acid prepared by chelating amino-carboxylic acid, organic **phosphonic acid** or the like with metal ion such as iron or cobalt ion, e.g. 0.0001∼2mol/l of ferric ion complex salt. By

making use of the resulting bleaching composition, environmental pollution is controlled, and bleaching speed is increased. Staining and occurrence of inferior recoloring phenomenon are prevented.

COPYRIGHT: (C)1980,JPO&Japio

AB To always maintain stable bleaching capacity in treatment of a silver halide color photographic material by adding optionally substituted aromatic sulfonic acid or sulfonate to hydrogen peroxide and metal complex salt of organic acid.

CONSTITUTION: An adequate amt. of aromatic sulfonic acid or sulfonate optionally substituted by at least one of carboxyl group, nitro group, etc. is added to a bleaching bath containing 0.01~10mol/l of hydrogen peroxide and a hydrogen peroxide-releasing substance such as perboric acid, perborate, percarbonic acid or percarbonate and metal complex salt of organic acid prepared by chelating amino-carboxylic acid, organic phosphonic acid or the like with metal ion such as iron or cobalt ion, e.g. 0.0001~2mol/l of ferric ion complex salt. By .

L3 ANSWER 144 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1978:607306 CAPLUS

DN 89:207306

TI Photoresist removal solutions

IN Harita, Yoshiyuki; Hanaoka, Hideyuki; Harada, Kunihiro

PA Japan Synthetic Rubber Co., Ltd., Japan

SO Jpn. Kokai Tokkyo Koho, 5 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 53025102	A2	19780308	JP 1976-99492	19760820
	JP 59050054	B4	19841206		
PRAI	JP 1976-99492		19760820		

AB Photoresist removal solns. are composed of H₂O₂, an org. solvent, and .gtoreq.1 compd. selected from sulfonic acids, sulfuric acid esters, and phosphoric acid esters. The solns. are esp. useful for removal of photohardened polybutadiene-type resists. Thus, a photoresist compn. consisting of cyclized cis-1,4-polybutadiene (degree of cyclization 65%, .eta. = 0.45), 2,6-bis(4'-azidobenzal)cyclohexanone (a photocrosslinking agent; 3 wt.% of the polymer), 2,6-di-tert-butyl-p-cresol (a stabilizer; 1 wt.% of the polymer), and phenyl-.alpha.-naphthylamine (a stabilizer; 1 wt.% of the polymer) was coated on a Si wafer, imagewise exposed, developed, and hardened by heating at 200.degree. for 30 min to give a resist pattern. The wafer was then immersed in a mixt. of o-dichlorobenzene 50, dodecylbenzenesulfonic acid 30, and a 38% H₂O₂ soln. 20 g and heated at 60.degree. for 5 min to completely remove the resist pattern.

AB Photoresist removal solns. are composed of H₂O₂, an org. solvent, and .gtoreq.1 compd. selected from sulfonic acids, sulfuric acid esters, and phosphoric acid esters. The solns. are esp. useful for removal of photohardened polybutadiene-type resists. Thus, a photoresist compn. consisting of cyclized cis-1,4-polybutadiene (degree of cyclization 65%, .eta. = 0.45), 2,6-bis(4'-azidobenzal)cyclohexanone (a photocrosslinking agent; 3 wt.% of the polymer), 2,6-di-tert-butyl-p-cresol (a stabilizer; 1 wt.% of the polymer), and phenyl-.alpha.-naphthylamine (a stabilizer; 1 wt.% of the polymer) was coated on a Si wafer, imagewise exposed, developed, and hardened by heating at 200.degree. for 30 min to give a resist pattern. The wafer was then immersed in a mixt. of o-dichlorobenzene 50, dodecylbenzenesulfonic acid 30, and a 38% H₂O₂ soln. 20 g and heated at 60.degree. for 5 min to completely remove the resist pattern.

IT **Sulfonic acids, uses and miscellaneous**
 RL: USES (Uses)
 (polybutadiene photoresist removal solns. contg. **hydrogen peroxide**, org. solvent and)

IT **Rubber, butadiene, uses and miscellaneous**
 RL: USES (Uses)
 (of cis-1,4-configuration, photoresists, removal of hardened, solns. contg. **hydrogen peroxide**, org. solvent and sulfuric or **phosphoric acid esters** for)

IT **Resists**
 (photo-, polybutadiene, removal of hardened, solns. contg. **hydrogen peroxide**, org. solvent and sulfuric or **phosphoric acid esters** for)

L3 ANSWER 145 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1978:55113 CAPLUS

DN 88:55113

TI Storage-stable concentrates of functional compositions

IN Grosse, Boewing Walter; Mrozek, Hinrich; Schluessler, Hans Joachim; Tinnefeld, Bernd; Voegele, Peter

PA Henkel und Cie. G.m.b.H., Fed. Rep. Ger.

SO Ger. Offen., 13 pp.

CODEN: GWXXBX

DT Patent

LA German

FAN.CNT 4

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 2616049	A1	19771027	DE 1976-2616049	19760412
	DE 2616049	C2	19870625		
	SE 7608459	A	19770217	SE 1976-8459	19760726
	SE 438424	B	19850422		
	SE 438424	C	19850801		
	NO 7602605	A	19770217	NO 1976-2605	19760726
	NO 146697	B	19820816		
	NO 146697	C	19821124		
	DK 7603363	A	19770217	DK 1976-3363	19760726
	DK 156026	B	19890619		
	DK 156026	C	19891106		
	FI 7602128	A	19770217	FI 1976-2128	19760726
	FI 63222	B	19830131		
	FI 63222	C	19830510		
	NL 7608265	A	19770218	NL 1976-8265	19760726
	NL 7608266	A	19770218	NL 1976-8266	19760726
	US 4051059	A	19770927	US 1976-711206	19760803
	US 4051058	A	19770927	US 1976-711205	19760803
	CA 1050876	A1	19790320	CA 1976-258701	19760809
	CA 1050877	A1	19790320	CA 1976-258702	19760809
	BR 7605307	A	19770809	BR 1976-5307	19760813
	AT 7606032	A	19780915	AT 1976-6032	19760813
	AT 349653	B	19790425		
	AT 7606029	A	19781015	AT 1976-6029	19760813
	AT 350194	B	19790510		
	GB 1561680	A	19800227	GB 1976-33746	19760813
	GB 1563713	A	19800326	GB 1976-33745	19760813
	CH 620676	A	19801215	CH 1976-10362	19760813
	JP 52025011	A2	19770224	JP 1976-97655	19760816
	JP 61010465	B4	19860329		
	JP 52025034	A2	19770224	JP 1976-97654	19760816
	JP 61014122	B4	19860417		
	FR 2321301	A1	19770318	FR 1976-24850	19760816
	FR 2321301	B1	19800430		
	FR 2321302	A1	19770318	FR 1976-24851	19760816
	FR 2321302	B1	19790504		

PRAI DE 1975-2536617 19750816
DE 1975-2536618 19750816
DE 1976-2616049 19760412

AB Storage-stable concs. for oxidn. of org. materials (i.e, hair, textiles) and for disinfection of machinery such as that used in the food industry comprise a C2-3 peracid and/or aliph. monocarboxylic acid 0.5-20, **H2O2** 25-40, a **phosphoric acid** or a H2O-sol. **phosphonic acid** salt 0.25-10, an anionic alkylbenzenesulfonate, alkyl **sulfate** or alkylsulfonate wetting agent 0.05-5% and H2O. These concs. also do not show the strong odor and irritancy of mixts. contg. only peracetic acid and AcOH [64-19-7]. For example, a conc. for prepn. (by diln.) of microbicidal agents comprised AcOH 5, **H2O2** 27.6, hydroxyethane-1.1-**diphosphonic acid** [2809-21-4] 1.5, a C12 alkylbenzenesulfonate 1.5, and H2O 64.4%. The conc. contained 28.3% H2O2 and 2.3% peracetic acid just after prepn., 27.5 and 2.3%, resp., after 1 mo, and 26.5 and 2.3%, resp., after 6 mo. Similar storage stability was also obsd. at 40.degree..

AB Storage-stable concs. for oxidn. of org. materials (i.e, hair, textiles) and for disinfection of machinery such as that used in the food industry comprise a C2-3 peracid and/or aliph. monocarboxylic acid 0.5-20, **H2O2** 25-40, a **phosphoric acid** or a H2O-sol. **phosphonic acid** salt 0.25-10, an anionic alkylbenzenesulfonate, alkyl **sulfate** or alkylsulfonate wetting agent 0.05-5% and H2O. These concs. also do not show the strong odor and irritancy of mixts. contg. only peracetic acid and AcOH [64-19-7]. For example, a conc. for prepn. (by diln.) of microbicidal agents comprised AcOH 5, **H2O2** 27.6, hydroxyethane-1.1-**diphosphonic acid** [2809-21-4] 1.5, a C12 alkylbenzenesulfonate 1.5, and H2O 64.4%. The conc. contained 28.3% H2O2 and 2.3% peracetic acid just after prepn., 27.5 and 2.3%, resp., after 1 mo, and 26.5 and 2.3%, resp., after 6 mo. Similar storage stability was also obsd. at 40.degree..

IT Carboxylic acids, biological studies
(aliph. peroxy, in disinfectant and oxidizing concs. with **hydrogen peroxide** and **phosphonates** and acylbenzenesulfonates)

IT 64-19-7, biological studies
RL: BIOL (Biological study)
(in disinfectant and oxidizing concn. with **hydrogen peroxide** and **phosphonates** and acylbenzenesulfonates)

L3 ANSWER 146 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1976:7905 CAPLUS

DN 84:7905

TI Stimulation of recovery from underground deposits

IN Sareen, Sarvajit S.; Girard, Lucien, III; Hard, Robert A.

PA Kennecott Copper Corp., USA

SO U.S., 4 pp. Division of U.S. 3,865,435.

CODEN: USXXAM

DT Patent

LA English

FAN.CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 3896879	A	19750729	US 1974-517677	19741024
	US 3865435	A	19750211	US 1973-422233	19731206
PRAI	US 1973-422233		19731206		

AB Metal recovery by in-situ leaching from underground formations is stimulated by injection of stabilized H2O2 [7722-84-1]. The stabilizer is pptd. by reaction with metals in the formation, and the unstabilized H2O2 decomposes and acts as a fracturing agent. Cu [7440-50-8] is leached from chalcopyrite formations by injection of 0.4M Fe2(SO4)3 [10028-22-5]. The wells are periodically shut down and treated with a 75% aq. **H2O2** (pH 4.0) contg. aminotrimethylene **phosphonic acid** [6419-19-8] as the stabilizing agent.

AB Metal recovery by in-situ leaching from underground formations is stimulated by injection of stabilized H₂O₂ [7722-84-1]. The stabilizer is pptd. by reaction with metals in the formation, and the unstabilized H₂O₂ decomposes and acts as a fracturing agent. Cu [7440-50-8] is leached from chalcopyrite formations by injection of 0.4M Fe₂(SO₄)₃ [10028-22-5]. The wells are periodically shut down and treated with a 75% aq. H₂O₂ (pH 4.0) contg. aminotrimethylene **phosphonic acid** [6419-19-8] as the stabilizing agent.

IT 7440-50-8P, preparation

RL: PREP (Preparation)

(by leaching of chalcopyrite under ground, by iron **sulfate** solns., **hydrogen peroxide** in)

IT 6419-19-8

RL: PROC (Process)

(**hydrogen peroxide** contg., in leaching of copper by iron **sulfate** under ground)

L3 ANSWER 147 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1975:437276 CAPLUS

DN 83:37276

TI Detecting hydrogen peroxide

IN Nakagiri, Yoshitaka; Kihara, Riichiro

PA Toyobo Co., Ltd.

SO Jpn. Tokkyo Koho, 3 pp.

CODEN: JAXXAD

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 49035715	B4	19740925	JP 1970-125534	19701229
PRAI	JP 1970-125534		19701229		

AB A reagent mixt. for detecting H₂O₂ by means of the color change in an indicator owing to oxidn. contains an anionic surfactant, preferably a carboxylate, a higher alc. **sulfate**, an alkyl benzenesulfonate, an alkyl naphthalenesulfonate, an alkyl phosphate, or a dialkyl sulfosuccinate, to improve color stability. The preferred concn. of anionic surfactant is 0.001% when testing liqs. or >0.01% in a soln. for prepg. a testing paper. Thus, 0.2 ml of 20 mg uric acid/10 ml, 1.0 ml Na azide, and 1.0 ml uricase were mixed, and 1.0 ml citric acid buffer, 0.5 ml anionic surfactant, 1.0 ml o-toluidine hydrochloride, and 0.5 ml peroxidase were added to the mixt.

AB A reagent mixt. for detecting H₂O₂ by means of the color change in an indicator owing to oxidn. contains an anionic surfactant, preferably a carboxylate, a higher alc. **sulfate**, an alkyl benzenesulfonate, an alkyl naphthalenesulfonate, an alkyl phosphate, or a dialkyl sulfosuccinate, to improve color stability. The preferred concn. of anionic surfactant is 0.001% when testing liqs. or >0.01% in a soln. for prepg. a testing paper. Thus, 0.2 ml of 20 mg uric acid/10 ml, 1.0 ml Na azide, and 1.0 ml uricase were mixed, and 1.0 ml citric acid buffer, 0.5 ml anionic surfactant, 1.0 ml o-toluidine hydrochloride, and 0.5 ml peroxidase were added to the mixt.

IT Benzenesulfonic acid, alkyl ester

Butanedioic acid, sulfo-, dialkyl ester

Naphthalenesulfonic acid, alkyl ester

Phosphoric acid, alkyl ester

Sulfuric acid, ester

RL: ANST (Analytical study)

(stabilizer, for color formed in hydrogen peroxide detection)

L3 ANSWER 148 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1975:549525 CAPLUS

DN 83:149525

TI Bleaching with hydrogen peroxide in organic solvent

IN Cormany, Charles P.
PA PPG Industries, Inc., USA
SO Fr. Demande, 51 pp.
CODEN: FRXXBL
DT Patent
LA French
FAN.CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	FR 2225567	A1	19741108	FR 1974-12612	19740410
	FR 2225567	B1	19771014		
	CA 1021659	A1	19771129	CA 1974-194044	19740305
	SE 7404454	A	19741017	SE 1974-4454	19740402
	JP 50040874	A2	19750414	JP 1974-42671	19740416
	GB 1470332	A	19770414	GB 1974-16681	19740416
PRAI	US 1973-351730		19730416		
	US 1973-413410		19731106		

AB Cotton and cotton-polyester textiles were bleached during dry cleaning with H2O2 at .gtoreq.45.degree. in the presence of Wayfos DNP-10K (I) [39464-64-7] and a small amt. of water [7732-18-5], and the H2O2 was removed by rinsing with the dry cleaning solvent contg. a detergent, iso-PrOH [67-63-0], methyl cellulose [9004-67-5], or water. Thus, cotton greige, prewashed with 3% aq. NaOH and 0.1% Triton X 100, was treated 15 min at 57.degree. with perchloroethylene contg. 35% H2O2 0.4 (based on fabric wt.), water 15.0 (based on fabric wt.), and I 0.75% (based on solvent vol.) and then rinsed 3 times with II contg. 0.75% Wayfos 6TD [55200-06-1] to give fabric contg. 0.085% residual H2O2.

IT Polyethers

RL: USES (Uses)

(alkylarylated ethers of, detergents, contg. org. **sulfonates**, in **hydrogen peroxide** removal in bleaching of fabrics in org. solvents)

IT **Sulfonates**

RL: USES (Uses)

(detergents, contg. alkaryl ethers of polyethers, in **hydrogen peroxide** removal in bleaching of fabrics in org. solvents)

IT Benzenesulfonic acid, dodecyl-, amine salts

Phosphoric acid, esters, salts

Poly(oxy-1,2-ethanediyl), .alpha.-hydro-.omega.-hydroxy-, ethers, alcs.

RL: USES (Uses)

(detergents, in hydrogen peroxide removal in bleaching of fabrics in org. solvents)

L3 ANSWER 149 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1974:479689 CAPLUS

DN 81:79689

TI Stable bleaching agents

IN Stalter, Neil J.

PA du Pont de Nemours, E. I., and Co.

SO Ger. Offen., 11 pp.

CODEN: GWXXBX

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 2333201	A1	19740124	DE 1973-2333201	19730629
	US 3811833	A	19740521	US 1972-268051	19720630
	CA 1003605	A1	19770118	CA 1973-173982	19730613
	IT 990818	A	19750710	IT 1973-26029	19730628
	BE 801681	A1	19731015	BE 1973-132921	19730629
	FR 2190912	A1	19740201	FR 1973-23905	19730629
	JP 49052784	A2	19740522	JP 1973-73343	19730630
	GB 1419184	A	19751224	GB 1973-31493	19730702

PRAI US 1972-268051 19720630

AB Stable aq. bleaching compns. of pH .sim.0.5-7.0 and used for bleaching detergent (Tide)-contg. laundry baths contained **hydrogen peroxide** [7722-84-1], Na stannate (I) [12773-27-2], ammonium **sulfate** (II) [7783-20-2], and **Dequest** 2010 (III, contg. alkylidenediphosphonic acid) [51888-66-5]. Thus, an aq. bleaching compn. of pH 2.0 (adjusted with HNO₃) contained H₂O₂ 35, III 0.1, I 0.01, and II 30% and lost 1.8% of active O on heating .sim.15 hr at 100.deg..

AB Stable aq. bleaching compns. of pH .sim.0.5-7.0 and used for bleaching detergent (Tide)-contg. laundry baths contained **hydrogen peroxide** [7722-84-1], Na stannate (I) [12773-27-2], ammonium **sulfate** (II) [7783-20-2], and **Dequest** 2010 (III, contg. alkylidenediphosphonic acid) [51888-66-5]. Thus, an aq. bleaching compn. of pH 2.0 (adjusted with HNO₃) contained H₂O₂ 35, III 0.1, I 0.01, and II 30% and lost 1.8% of active O on heating .sim.15 hr at 100.deg..

L3 ANSWER 150 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1973:85853 CAPLUS

DN 78:85853

TI Bleach-stabilizing agents

IN Kowalski, Xavier

PA Monsanto Co.

SO Ger. Offen., 28 pp.

CODEN: GWXXBX

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
	-----	----	-----	-----	-----
PI	DE 2226784	A	19721214	DE 1972-2226784	19720602
	DE 2226784	C3	19860417		
	US 3740187	A	19730619	US 1971-149856	19710603
	US 3766078	A	19731016	US 1971-149854	19710603
	US 3795625	A	19740305	US 1971-149855	19710603
	IT 956067	A	19731010	IT 1972-25149	19720531
	FR 2140213	A1	19730112	FR 1972-19960	19720602
	FR 2140213	B1	19771223		
	GB 1354989	A	19740530	GB 1972-25941	19720602
	ES 403436	A1	19751201	ES 1972-403436	19720602
	JP 56052080	B4	19811209	JP 1972-54398	19720602
PRAI	US 1971-149854		19710603		
	US 1971-149855		19710603		
	US 1971-149856		19710603		

AB Synergistic stabilizers for alk. peroxy bleaching solns. consisted of Na salts of 1-hydroxyethylidene-1,1-diphosphonic acid and nitrilotriacetic, and Ca and Mg salts. Thus, a mixt. of disodium 1-hydroxyethylidene-1,1-diphosphonate [7414-83-7], nitrilotriacetic acid trisodium salt [5064-31-3], and magnesium **sulfate** [7487-88-9] was an effective stabilizer for a H₂O₂ bleach contg. NaOH, Na silicate, and 2ppm Fe²⁺. Cotton fabrics bleached with the stabilized bleaching solns. were whiter than fabrics bleached with bleaches contg. conventional stabilizers.

AB Synergistic stabilizers for alk. peroxy bleaching solns. consisted of Na salts of 1-hydroxyethylidene-1,1-diphosphonic acid and nitrilotriacetic, and Ca and Mg salts. Thus, a mixt. of disodium 1-hydroxyethylidene-1,1-diphosphonate [7414-83-7], nitrilotriacetic acid trisodium salt [5064-31-3], and magnesium **sulfate** [7487-88-9] was an effective stabilizer for a H₂O₂ bleach contg. NaOH, Na silicate, and 2ppm Fe²⁺. Cotton fabrics bleached with the stabilized bleaching solns. were whiter than fabrics bleached with bleaches contg. conventional stabilizers.

L3 ANSWER 151 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1966:50704 CAPLUS

DN 64:50704
OREF 64:9435f-h
TI Water purification by liquid-liquid extraction
PA Canadian Industries Ltd.
SO 10 pp.
DT Patent
LA Unavailable

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	NL 65004197		19651004	NL	
PRAI	CA		19640402		

AB Fe, Cu, and Mn ions are extd. from H₂O by an org. acid dissolved in an org. solvent, not miscible with H₂O. The org. acid is regenerated by extg. the metal ions with an inorg. acid, such as 2-20 wt. %, H₂SO₄, HNO₃, H₃PO₄, or HCl. Org. acids with >8 C atoms per mol., e.g. aromatic and aliphatic fatty acids, esp. tertiary monobasic acids with 9-11 C atoms in the main chain and aliphatic and aromatic phosphonic and sulfonic acids are useful. Org. solvents, not miscible with H₂O, that can be used are aromatic liquid hydrocarbons, such as benzene, toluene, and xylene, and cyclic and acyclic aliphatic hydrocarbons, such as hexane, cyclohexane, and kerosine, and alcs. such as n-hexanol. The concn. of the org. acids are 10-50 wt. %. The invention is applied particularly to aq. solns. for the prepn. of H₂O₂ and to the process of Can. 673.355 (CA 60, 5110b). This method uses less inorg. acid than the process with ion exchange resins. In an example 25% di(2-ethylhexyl)phosphoric acid in xylene was used to decrease the Fe content from 1.15 to 0.2 ppm. and Cu from 1.18 0.31 ppm.

IT **Sulfonic acids**
(metal extn. from H₂O₂ soln. and water by immiscible soln. of)

IT 13598-36-2, **Phosphonic acid**
(derivs., metal extn. from H₂O₂ soln. and water by immiscible soln. of)

L3 ANSWER 152 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1965:401793 CAPLUS

DN 63:1793

OREF 63:278c-e

TI Hydrogen peroxide

PA Imperial Chemical Industries Ltd.

SO 11 pp.

DT Patent

LA Unavailable

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	NL 64008112		19650125	NL	
PRAI	GB		19630723		

AB The H₂O₂ was prepd. by treating H with O in a liquid phase, consisting of water, a stabilizer, a catalyst, an acid, and an org. compd. Suitable stabilizers are sequestering agents, such as aluminates, silicates, titanates, stannates, antimonates, salts of oxyacids of P, phenolic compds. (thymol), salicylates, citrates, hydroxyquinolines, carboxypyridines, carboxyquinolines, hexamethylenetetramine, nitriloacetic acids, .alpha.-amino carboxylic acids, acetanilide, phenacetin, BzOH, glycerol Ph ether, thiourea, uric acid, and glycol monoacetate. Colloids, such as gum arabic, gum tragacanth, albumin, gelatin, agar, and phosphatides, and cation-exchange resins are likewise useful as stabilizers. The quantity present can be 20-200 ppm. In some cases it is sufficient to wash the reaction vessel with a satd. soln. of stabilizer. Elements of Groups I and VIII of the periodic system can be used as catalysts. Preferred is Pd with silica gel as a carrier. The acid used must be stronger than AcOH. The acidity of the phase must be > 10⁻⁵N,

e.g. 10-4N-N. An inorg. acid is preferred, such as HNO₃, HF, HBr, H₃PO₄, and sulfonic acids. It is useful if the chloride ion is used. The concn. should be 10-4-0.1 N. The org. compd. must be water-sol. It can be an alc., aldehyde, ketone, ether, ester, amide, or an O-contg. amine of a carboxylic acid. Acetone (40-90 vol. %) is preferred. Example: A Pyrex vessel is washed with a satd. soln. of Na₄P₂O₇, 100 ml. of 10-2N HCl and 10-1N H₂SO₄ in a mixt. of acetone (75 vol. %) and water is filled in the vessel. One g. Pd powder (5%) on silica gel is added. The mixt. is cooled to 0.degree. and agitated. A mixt. of 1 part H and 4 parts of air is led in at normal pressure for 5 hrs. (25 l./hr.). H₂O₂ (2.85 g.) is formed. When the vessel was not washed with Na₄P₂O₇, 0.83 g. of H₂O₂ is formed. Meth. Appl. 6,408,187; 7 pp. H₃BO₃ and a 2nd acid, stronger than AcOH, are used.

IT Sulfonic acids
(in hydrogen peroxide manuf. from H and O)

IT Boric acid
Phosphoric acid, hexose esters
(in hydrogen peroxide manuf. from H and O)

L3 ANSWER 153 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1956:92550 CAPLUS

DN 50:92550

OREF 50:17354a-b

TI Hydrogen peroxide

IN Rust, Frederick F.; Porter, Lee M.; Vaughan, Wm. E.

PA N. V. de Bataafsche Petroleum Maatschappij

DT Patent

LA Unavailable

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	GB 751508		19560627	GB	
AB	A faster conversion and higher yield are obtained in the H ₂ O ₂ synthesis described earlier (C.A. 48, 11017a) if the alc. contains small amts. of both an alc.-sol. heavy metal salt (I) and a H ₂ O ₂ stabilizer (II). Suitable I-II combinations are Co (as Co isovalerylacetone) 0.00009 and H ₃ PO ₄ 0.0035; Fe (as Fe ₂ (SO ₄) ₃) 0.0001 and H ₃ PO ₄ 0.0003%. Other I are Ni, Mn, or Cr salts. Other II are oxy acids of Sn and alkali metal salts of oxyacids of Sn or P.				
IT	Iron sulfates, Fe ₂ (SO ₄) ₃ (secondary alc. stabilization for H ₂ O ₂ manuf. with H ₃ PO ₄ or Na ₄ P ₂ O ₇ and)				
IT	7664-38-2, Phosphoric acid (mixts. of, with heavy-metal salts, iso-PrOH stabilization with, for H ₂ O ₂ manuf.)				

L3 ANSWER 154 OF 158 CAPLUS COPYRIGHT 2003 ACS

AN 1955:34951 CAPLUS

DN 49:34951

OREF 49:6705f-i,6706a

TI Stabilization of hydrogen peroxide

AU Seshadri, K.

CS Indian Inst. Sci., Bangalore

SO Journal of Scientific & Industrial Research (1954), 13B, 575-80

CODEN: JSIRAC; ISSN: 0022-4456

DT Journal

LA Unavailable

AB cf. Isemura, C.A. 45, 5500h. The decompn. of 3-30% H₂O₂ solns. in the presence of various stabilizers was investigated. The stability of H₂O₂ solns. was detd by storing the solns. for 2 yrs. in the presence of various stabilizers and then testing for remaining H₂O₂. Thirteen solns. were further tested by boiling under reflux the H₂O₂ soln. with added stabilizer and detn. of residual H₂O₂ after 1, 3, and 6-hr. intervals. The decompn. of H₂O₂ soln: at b.p. for 1 hr. was approx. equal to the

decompn. of the soln. stored for 1 yr. at room temp. Among the stabilizers examd., in the order of their decreasing effectiveness were: tartaric acid and Na salicylate mixt., sulfanilamide, 8-hydroxyquinoline, nipagin (I), acetanilide (II), and phenacetin (III). The activity of these stabilizers was attributed to their ability to deactivate the decompn. of catalysts of H₂O₂ by complex formation or chelation. Other compds. tested and found to be less effective were: AcOH, AcNH₂, aspirin, adipic acid, barbituric acid, barbitol, BzOH, BzNHPh, BzNH₂, Ca(H₂PO₄)₂, CaHPO₄, Ca₃(PO₄)₂, EtOH, (CH₂OH)₂, Et₂O, glycerol, hydroquinone, malic acid, mepacrine, malonic acid, 1-naphthylamine, HPO₃, H₃PO₄, phenylbarbital, phenyl cincholine, phenyl salicylate, quinine sulfate, quinine glycerophosphate, quinine salicylate, salicylic acid, Na stannate, Na₄P₂O₇, Na glycerophosphate, Na₃PO₄, H₂SO₄, succinic acid, uric acid, and urea. The results obtained with I, II, and III indicated that high concn. of stabilizer was worse than no stabilizer at all, and there seemed to be no simple relation between concn. of stabilizer and the rate of retardation of decompn. of H₂O₂. The optimum concns. of I, II, and III are 0.10, 0.02, and 0.025%, resp. A mixed stabilizer of Deutsche Gold und Silber Scheide-Anstalt was superior to all other mixts. of stabilizers tried.

- IT Quinine, glycerophosphate, salicylate
Quinine, glycerophosphate, **sulfate**
(hydrogen peroxide stabilization by)
- IT 50-06-6, Phenobarbital 50-78-2, Acetylsalicylic acid 54-21-7, Sodium salicylate 55-21-0, Benzamide 56-81-5, Glycerol 57-13-6, Urea 60-29-7, Ethyl ether 60-35-5, Acetamide 62-44-2, Phenacetin 63-74-1, Sulfanilamide 64-17-5, Ethyl alcohol 64-19-7, Acetic acid 65-85-0, Benzoic acid 67-52-7, Barbituric acid 69-93-2, Uric acid 83-89-6, Quinacrine 87-69-4, Tartaric acid 93-98-1, Benzanilide 99-76-3, Benzoic acid, p-hydroxy-, methyl ester 103-84-4, Acetanilide 107-21-1, Ethylene glycol 110-15-6, Succinic acid 118-55-8, Salol 123-31-9, Hydroquinone 124-04-9, Adipic acid 134-32-7, 1-Naphthylamine 141-82-2, Malonic acid 6915-15-7, Malic acid 7601-54-9, Sodium phosphate, Na₃PO₄ 7664-38-2, **Phosphoric acid**
(as **hydrogen peroxide** stabilizer)
- L3 ANSWER 155 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 1954:31454 CAPLUS
DN 48:31454
OREF 48:5624g-h
TI Delaying action of distilled water on catalytic decomposition of hydrogen peroxide
AU Krause, Alfons; Olejnik, Gwidon
CS Univ. Poznan, Pol.
SO Roczniki Chemii (1953), 27, 17-22
CODEN: ROCHAC; ISSN: 0035-7677
DT Journal
LA German
AB The decompn. of H₂O₂ by Ag₂O at 37.degree. was delayed by distd. water contg. H₂CO₃. Small amts. of HCl, H₂SO₄, Na₂SO₄, H₃PO₄, and CH₃COOH have similar action, attributed to formation of complex compds. with Ag₂O.
- IT 64-19-7, Acetic acid 463-79-6, Carbonic acid 7647-01-0, Hydrochloric acid 7664-38-2, **Phosphoric acid** 7757-82-6, Sodium **sulfate**, Na₂SO₄
(effect on H₂O₂ decompn. by Ag₂O)
- L3 ANSWER 156 OF 158 CAPLUS COPYRIGHT 2003 ACS
AN 1952:10789 CAPLUS
DN 46:10789
OREF 46:1891e-f
TI Activated carbon electrodes for air-depolarized wet cells. I. The decomposition of hydrogen peroxide on account of activated charcoals
AU Watanabe, Jun; Shiramoto, Tomizo
CS Matsushita Elec. Ind. Co., Osaka

SO Denki Kagaku (1951), 19, 274-9
 CODEN: DNKKA2; ISSN: 0366-9440

DT Journal

LA Unavailable

AB The activity of charcoals activated by many inorg. reagents was studied by measuring the decompn. velocity of aq. H₂O₂ in the presence of the sample charcoal. The charcoals activated with NaOH, KOH, KMnO₄, H₂SO₄, H₃PO₄, and other metallic sulfates gave the lowest decompn. velocity values; while HClO₄, HNO₃, NH₄Cl, (NH₄)₂SO₄, MgCl₂, FeCl₂, Co(NO₃)₂, SnCl₂, etc., gave medium values; and MnCl₂, CoCl₂, PdCl₂, H₂PtCl₆, etc., gave considerably higher values.

IT **Sulfates**
 (charcoal activated with, H₂O₂ decompn. on)

IT 7664-38-2, **Phosphoric acid**
 (charcoal activated with, H₂O₂ decompn. in)

IT 1310-58-3, Potassium hydroxide 1310-73-2, Sodium hydroxide 7601-90-3, Perchloric acid 7647-10-1, Palladium chloride, PdCl₂ 7664-93-9, Sulfuric acid 7697-37-2, Nitric acid 7722-64-7, Potassium permanganate 7758-94-3, Iron chloride, FeCl₂ 7772-99-8, Tin chloride, SnCl₂ 7773-01-5, Manganese chloride, MnCl₂ 7783-20-2, Ammonium **sulfate** 10141-05-6, Cobalt nitrate, Co(NO₃)₂ 12125-02-9, Ammonium chloride 16941-12-1, Chloroplatinic(IV) acid (H₂PtCl₆)
 (charcoal activated with, H₂O₂ decompn. on)

L3 ANSWER 157 OF 158 JAPIO COPYRIGHT 2003 JPO

AN 2000-197899 JAPIO

TI AGENT AND METHOD FOR PURIFICATION TREATMENT OF LAKE, MARSH, AND THE LIKE

IN TANIMOTO KAYOHIKO; YARA HIDEO; FURUYA MASAAKI; GIBO MASAMI

PA KANKYO BIKEN:KK

PI JP 2000197899 A 20000718 Heisei

AI JP 1999-1797 (JP11001797 Heisei) 19990107

PRAI JP 1999-1797 19990107

SO PATENT ABSTRACTS OF JAPAN (CD-ROM), Unexamined Applications, Vol. 2000

AB PROBLEM TO BE SOLVED: To restore a natural ecological system by converting nutritious components eluted from bottom sludge in a reduced state into inactive substances by using a purifying agent prepared by mixing an oxidation promoter and sodium silicate in an appropriate ratio.
 SOLUTION: A purifying agent for lakes, marshes, and others comprises an oxidation promoter and sodium silicate as an additive, the mixing ratio of which is about 50%. In the promoter, 500-1,000 g of pure water, 25-50 g of ammonium persulfate, 25-50 g of calcium nitrate, 2-4 g of **phosphonic acid**, and 15-30 g of dilute sulfuric acid are mixed. In place of ammonium **sulfate** which is used as a component of the promoter, sodium persulfate and **hydrogen peroxide** can be used. By injecting an appropriate amount of the purifying agent directly at each time into lake bottom sludge, the sludge is reformed fundamentally so that lake water can be clarified while the original self-purifying capacity of a lake is being restored.
 COPYRIGHT: (C)2000,JPO

AB . . . the promoter, 500-1,000 g of pure water, 25-50 g of ammonium persulfate, 25-50 g of calcium nitrate, 2-4 g of **phosphonic acid**, and 15-30 g of dilute sulfuric acid are mixed. In place of ammonium **sulfate** which is used as a component of the promoter, sodium persulfate and **hydrogen peroxide** can be used. By injecting an appropriate amount of the purifying agent directly at each time into lake bottom sludge, . . .

L3 ANSWER 158 OF 158 JAPIO COPYRIGHT 2003 JPO

AN 2000-191328 JAPIO

TI PRODUCTION OF FERRIC SULFATE SOLUTION

IN TSUKAMOTO TATSUO; KAKIO TOSHIHIKO

PA TAKI CHEM CO LTD

PI JP 2000191328 A 20000711 Heisei

AI JP 1998-368298 (JP10368298 Heisei) 19981225

PRAI JP 1998-368298 19981225

SO PATENT ABSTRACTS OF JAPAN (CD-ROM), Unexamined Applications, Vol. 2000

AB PROBLEM TO BE SOLVED: To obtain an iron type inorganic flocculant having a turbidity removing effect comparable to or higher than that of an aluminum type flocculant by oxidizing a ferrous sulfate solution having a specified concentration of total iron, containing phosphate or polyphosphate ions in a specified P to Fe molar ratio and having a specified Fe^{3+} to Fe^{2+} molar ratio.

SOLUTION: A ferrous sulfate solution having ≥ 10 wt.% (expressed in terms of Fe) concentration of total iron, containing phosphate or polyphosphate ions in a P to Fe molar ratio of 0.1-0.45 and having an Fe^{3+} to Fe^{2+} molar ratio of ≤ 1 is oxidized. The ferrous sulfate solution is not particularly limited, a solution discharged from a step for pickling a steel sheet, a steel wire or the like may be used and powdery or liquid ferrous sulfate discharged from the titanium industry may also be used. **Phosphoric acid** or its water-soluble salt may be used as a phosphate or polyphosphate ion donor. The oxidation is carried out by direct oxidation with oxygen or the like or with an oxidizing agent such as **hydrogen peroxide** to such a degree that all of iron (II) in the ferrous **sulfate** solution is oxidized to iron (III).

COPYRIGHT: (C) 2000, JPO

AB . . . the like may be used and powdery or liquid ferrous sulfate discharged from the titanium industry may also be used. **Phosphoric acid** or its water-soluble salt may be used as a phosphate or polyphosphate ion donor. The oxidation is carried out by direct oxidation with oxygen or the like or with an oxidizing agent such as **hydrogen peroxide** to such a degree that all of iron (II) in the ferrous **sulfate** solution is oxidized to iron (III).

COPYRIGHT: (C) 2000, JPO